

DMSO hydration redefined: unraveling the hydrophobic hydration of solutes with a mixed hydrophilic–hydrophobic characteristic

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Abstract

Hydrophobic hydration of solutes with a mixed hydrophilic–hydrophobic characteristics is still poorly understood. This is because both experimental and theoretical methods find it difficult to see the ice-like water structure around the nonpolar solute groups, unlike hydrogen bonds with the hydrophilic groups. In order to unravel this problem, we have investigated DMSO hydration by means of infrared spectroscopy and theoretical methods, namely DFT, ONIOM calculations and AIMD simulations, which allowed us to redefine its hydration. In dilute DMSO solutions the clathrate-like water is formed around the DMSO molecule, supported by interactions of water molecules with the methyl hydrogens (the blue-shifted hydrogen bonds). The cage is constructed by water molecules that form hydrogen bonds of the comparable energy and length with the SO group and between water molecules. When the construction of the cage is completed, DMSO molecule partially regains its rotational freedom inside. Strong hydrogen bonds within the frame

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are masked by the relatively small population of weakened hydrogen bonds of water molecules in the vicinity of the SO group, due to the improper fit to the bulk water of water molecules hydrogen bonded to the oxygen atom of DMSO. We also propose a new explanation of the highly non-ideal mixing behavior of aqueous DMSO solutions at the eutectic point, as the positive excess entropy of the equimolar amounts of molecular complexes distinguished in the system.

Keywords: Water structure, Hydrophobic hydration, DMSO-water complexes, FTIR spectroscopy, DFT calculations, AIMD simulations

1. Introduction

“Water is an active matrix of life for cell and molecular biology” [1]. In this sense, the role of water relies on its diverse structural and dynamic characteristics in such systems. One of the very important specific structural forms which water adopts near non-polar surfaces of biomolecules is known as the hydrophobic hydration which contributes to many chemical and biochemical processes, such as protein folding, protein–protein, and protein–co-solute or self–assembly of lipid membranes [2, 3, 4, 5, 6, 7, 8, 9, 10]. However, since the formulation of the “iceberg formation” hypothesis by Frank and Evans [11], opinions on the water structure near the non-polar fragments of solutes have been inconsistent (see for example ref. [12] for a review). Recently, their model has been confirmed by Grdadolnik *et al.* [13] by means of the high-pressure infrared spectroscopy for small purely hydrophobic solutes (methane, ethane, krypton, and xenon). The strengthened water structure in such cases is similar to the one of ice or solid clathrates. Other direct



16 experimental evidence has also been demonstrated [14, 15], however, studies
17 of soluble solutes with both hydrophilic and hydrophobic moieties elude any
18 simple explanation of their complex hydration [16, 17]. Dimethyl sulfoxide
19 (DMSO) is a simple molecule which can serve as a model for studying such
20 a type of mixed hydration.

21 DMSO is an aprotic solvent and its aqueous solutions have many inter-
22 esting physicochemical and biological properties. DMSO is miscible with
23 water in all proportions [18]. Dielectric spectroscopy studies of liquid DMSO
24 have proven the existence of its dimers and longer forms (“polymers”) with
25 anti-parallel ordering of molecular dipoles [19, 20]. Numerous studies have re-
26 vealed uncommon physicochemical properties of DMSO–water system, man-
27 ifested by strong deviations of their thermodynamic properties from ide-
28 ality [21]. Many reports indicate that water–DMSO hydrogen bonds are
29 stronger than those between water molecules [19, 20, 22]. As a result of
30 these interactions, molecular complexes of water and DMSO are relatively
31 stable, regardless of composition [23]. It is assumed that the unusual features
32 of DMSO solutions are due to the various species water–DMSO complexes
33 [24, 25]. These clusters at various DMSO concentrations have been exten-
34 sively studied both experimentally [20, 21, 26, 24, 25, 27] and theoretically
35 [23, 24, 28, 29, 30, 31]. Those which consist of $3\text{DMSO}\cdot\text{1H}_2\text{O}$, $2\text{DMSO}\cdot\text{1H}_2\text{O}$,
36 and $1\text{DMSO}\cdot\text{1H}_2\text{O}$ have been found at mole fractions of water (x_w) lower than
37 0.5 [21, 23, 25]. The strongest deviations from ideality and the strongest hy-
38 drogen bonds occur at x_w between 0.6 and 0.7 [26, 32, 17]. The presence
39 of stable $1\text{DMSO}\cdot\text{2H}_2\text{O}$ aggregate has been observed in this range [20, 26].
40 It is also the eutectic composition with freezing temperature of ca. -70°C



41 [33], while freezing points of water and DMSO are 0 °C and 18.6 °C, respec-
42 tively. But according to the hypothesis of Kirchner and Reiher [31], many
43 energetically similar but structurally different complexes exist in the DMSO–
44 water mixture near the eutectic point. They proposed the mechanism of the
45 clusters’ influence on the non–ideal mixing behavior of these systems.

46 There is no general agreement on the influence of DMSO on water struc-
47 ture. The results of experimental studies [34, 35, 36, 37] and computer sim-
48 ulations [28, 35, 38, 39] indicate that water structure is enhanced in dilute
49 DMSO aqueous solutions. On the other hand, some experimental studies
50 [20, 40, 41] lead to the conclusion that water structure is weakened at a low
51 concentration of DMSO. The experimental techniques, including IR spec-
52 troscopy and MD simulations demonstrate that DMSO acts as a “structure
53 breaker” at high concentrations of DMSO [28, 37].

54 In this paper, we examine the hydration of dimethyl sulfoxide (DMSO)
55 in the whole mole fraction range by means of the FT-IR spectroscopy and
56 computational methods. We demonstrate the existence of the “ice-like” water
57 cage around DMSO molecules in diluted solutions, which enables partial
58 rotational freedom of the guest molecule inside the water cage. We propose
59 also a novel explanation of the strong deviation from ideality of aqueous
60 solutions of DMSO at concentration corresponding to the eutectic point of
61 the system.



62 2. Materials and methods

63 2.1. Chemicals and solutions

64 Dimethyl sulfoxide (99.9%, Alfa Aesar) and D₂O (isotopic purity 99.9%,
65 Aldrich) were used as supplied to prepare solutions without purification.
66 Water used to prepare these solution was deionized (<0.01 S·cm⁻¹). All
67 solution have been prepared by weight and their densities were determined
68 with Anton Paar DMA 5000 densitometer at 25.000 ± 0.001 °C. The solution
69 preparation procedure for FTIR measurements of HDO spectra has been
70 described in Supporting Material in ref. [42].

71 2.2. FTIR spectroscopy

72 All FTIR spectra of aqueous solutions of DMSO were recorded on the
73 IFS 66 Bruker spectrometer. The spectrometer was purged with dry nitro-
74 gen during the measurement. A liquid transmission cell (model A145, Bruker
75 Optics) was equipped with two CaF₂ windows separated by teflon spacers.
76 For H₂O transmission spectra of DMSO-water mixtures for high concentra-
77 tion of DMSO (in the range of $\nu_{S=O}$ vibration), 512 independent scans were
78 taken with resolution of 2 cm⁻¹. The path length was equal to 0.0053 mm, as
79 determined interferometrically. In the case of HDO spectra, in the range of
80 ν_{OD} vibration 256 independent scans were taken with the resolution of 4 cm⁻¹.
81 The path length was equal to 0.0306 mm, as determined interferometrically.
82 The same path length was used also in the case of spectra for DMSO·H₂O
83 mixtures for very low DMSO concentration, in the range of the $\nu_{S=O}$ vibra-
84 tion band. The temperature was kept at 25.0 ± 0.1 °C and monitored using
85 thermocouples placed in the sample cell.

86 The spectra have been analyzed using the following commercial software:
87 GRAMS/32 (Galactic Industries Corp.) and RAZOR (Spectrum Square As-
88 sociates, Inc.) run under GRAMS/32.

89 *2.3. Analysis of vibrational spectra*

90 The difference spectra method was applied to extract the DMSO-affected
91 HDO spectrum on the basis of spectra series measured for different molalities
92 of aqueous solutions. An assumption was made that the water in solution can
93 be divided into two additive contributions: the bulk water (pure water) and
94 the “solute-affected” water (modified by interactions with the solute). The
95 method of analysis of the spectral data towards extraction of solute-affected
96 water spectrum was described in details in refs. [43, 44, 45] and some of
97 the most basic information are included in the Supplementary Material. The
98 difference spectra method was also applied to the DMSO spectra in the range
99 of the $\nu(\text{S}=\text{O})$ vibrations to separate specific states of the S=O oscillators.

100 The factor analysis, in the version written by Malinowski [46], was per-
101 formed using the commercial computer program Factor Analysis Toolbox for
102 MATLAB (Applied Chemometrics Inc., Sharon). The spectral data were
103 assembled into matrices of absorbances at given wavenumbers and concen-
104 trations of DMSO–water mixtures. The main application of this method
105 was window factor analysis (WFA) [47], which involves the extraction of the
106 concentration profiles of the individual chemical species contributing to the
107 spectral data (i.e. factors). The correct number of these factors was obtained
108 by using the basic principal factor analysis algorithm (PFA).



109 *2.4. Theoretical calculations*

110 *2.4.1. DFT and ONIOM calculations*

111 All calculations were performed with the GAUSSIAN 09 v.D1 software
112 [48] available at the Academic Computer Center in Gdansk (TASK). The
113 program Avogadro was used to prepare of input data and for visualization
114 of computed results. The analysis of resulting wavefunction files, involving
115 the reduced density gradient (RDG) method [49], was performed with the
116 Multiwfn software v.3.3.9 [50]. The RDG method allowed to visualize weak
117 interaction sites (hydrogen bonds, van der Waals interactions, steric clashes)
118 in molecular complexes and to classify them according to their strengths.
119 All RDG-related figures are presented in the stereo view for better clarity.
120 The D3 version of Grimme's empirical dispersion correction, including the
121 Becke–Johnson damping (GD3-BJ), was applied [51].

122 Structures of medium-sized complexes: 1DMSO·nH₂O (where n = 1÷15),
123 3DMSO·2H₂O, and 3DMSO were optimized using the density functional
124 theory (DFT) level with the B3LYP hybrid exchange-correlation functional
125 [52, 53] and 6-311++G(d,p) basis set [54]. The conductor-like polarizable
126 continuum model (CPCM) of the self-consistent reaction field theory (SCRF)
127 was used to simulate the solvent environment [55, 56]. In the case of three
128 different types of the considered complexes: 1DMSO·nH₂O, 3DMSO·2H₂O,
129 and 3DMSO, the following solvents were used: water, solvent mixture of
130 DMSO:water with a mixing ratio of 3:2, and DMSO, respectively.

131 For more complex structures of one DMSO molecule with up to 100 hy-
132 dration water molecules the ONIOM approach was applied [57]. 5 to 90 water
133 molecules were added to the DMSO molecule with a step of 5 molecules. Ad-

ditional complexes with 2 to 4 molecules were prepared, however, a more reliable results for such complexes were obtained with the method presented in previous paragraph thanks to the larger basis set and the use of the CPCM solvent model. Initial complexes were optimized with a simple MMFF94a force field [58, 59]. Next, the central DMSO molecule was selected for the high level of ONIOM calculations – B3LYP/aug-cc-pVQZ, and all water molecules for the low level – B3LYP/cc-pVDZ, all performed *in vacuo*. Such a selection of basis sets allowed to obtain satisfactory structures within a reasonable calculation time.

Three different water clusters have been used to obtain the reference parameters for hydrogen bonds energies between solvent molecules: 1) a triangle prism, or nano-drop of 6 molecules for medium-sized complexes, 2) an unstructured complex of 100 molecules for ONIOM-based calculations.

2.4.2. AIMD simulations

AIMD simulations [60] were performed using the DFT-based QUICKSTEP electronic structure module [61] of the CP2K 6.0 computational suite [62, 63]. We applied the BLYP functional [52, 64] together with the DFT-D3 empirical dispersion correction [65]. The cutoff for the latter was set to 16 Å. QUICKSTEP defines a mixed Gaussian type orbitals plus plane waves (GPW) basis set scheme [66], and we used a TZV2P basis set for atomic orbitals and a 500 Ry cutoff for the plane wave expansion of the electron density. Only valence electrons were treated explicitly, while the core electrons were represented by GTH pseudo potentials [67].

The two studied systems consisted of bulk H₂O (80 molecules) and a DMSO(H₂O)₈₀ solution, contained in cubic supercells with applied periodic

159 boundary conditions. In order to compare directly to the FTIR HDO/H₂O
160 spectra, all water hydrogen atoms were given the mass of deuterium, i.e.,
161 we effectively simulate D₂O. Initial volumes of the systems were chosen to
162 reflect the experimental density of heavy water [68] combined the apparent
163 molar volume of the DMSO solution [69]. While the experimental data were
164 measured at 298 K, we apply a slight temperature overscaling, typical for
165 AIMD simulations neglecting nuclear quantum effects, in order to recover
166 the proper diffusional behavior of individual molecules in our systems and to
167 avoid the ‘glassy dynamics’ regime [70]. The extent of this scaling is slight
168 for D₂O ($T = 323.15$ K)[71].

169 Both systems were first equilibrated for at least 20 ps with a time step of
170 0.5 fs in the NVT ensemble using massive Nosé-Hoover chain thermostating
171 [72]. After the equilibration period, 20 initial conditions were sampled every
172 3 ps from a further NVT simulation to initialize microcanonical (NVE) tra-
173 jectories of 20 ps length each. During these runs the centers of maximally
174 localized Wannier functions (MLWFs) [73] were computed every 2 fs. All an-
175 alyzed observables were averaged over the NVE trajectories yielding proper
176 canonical averages.

177 Molecular dipole moments were obtained classically by summing over
178 positive nuclei and negative MLWF centers. The IR spectra were calculated
179 as Fourier transforms of time correlation functions of dipole moment finite
180 differences [74, 75] using various recently introduced dipolar decomposition
181 schemes for solute–solvent systems, see refs. [75, 76] for details. The spectral
182 resolution was set to 1 cm^{-1} by setting the upper limit of the correlation
183 time to ~ 16.66 ps and the final spectra were smoothed by passing through

184 a 20 cm^{-1} Gaussian filter. Numerical Kramers–Krönig transform was used
185 to remove the refractive index contributions to the IR spectra [75] using the
186 experimental refractive index of D_2O , $n_{\text{D}} = 1.328$ [77].

187 3. Results and discussion

188 3.1. FTIR spectra of HDO in DMSO-water mixtures

189 HDO spectra in the range of the $\nu(\text{OD})$ vibrations are shown for all com-
190 positions of the DMSO–water mixture in Fig. 1a. It is clear that the spectral
191 series exhibits two isosbestic points: the first one at ca. 2600 cm^{-1} corre-
192 sponding to the spectra from pure water to $x_w = 0.83 \pm 0.07$, and the second
193 one at ca. 2530 cm^{-1} corresponding to spectra from the smallest dilutions of
194 water in DMSO to $x_w = 0.41 \pm 0.01$.

195 The main parameters of the HDO spectra as a function of x_w are presented
196 in Fig. 1b (and also in the Supplementary Material, see Fig. S2). Particularly
197 noteworthy is the dependence for $\nu(\text{OD})$ band position at maximum and at
198 the center of gravity of the band. The first one serves as a measure of
199 the most probable energy of hydrogen bonds of water (as well as the most
200 probable $\text{O}\cdots\text{O}$ distance), and the latter as a measure of the average energy
201 of hydrogen bonds (as well as the average $\text{O}\cdots\text{O}$ distance).

202 —Figure 1—

203 In simple terms, DMSO dissolved in water in larger quantities should be
204 classified as a water “structure-breaking” solute. Only in the range of high
205 dilutions, above ca. 0.8 mole fraction of water, DMSO can be considered
206 as a “structure-making” agent. This may be inferred from the position of
207 the center of gravity of the $\nu(\text{OD})$ band, which is red-shifted relative to the

208 pure water, Fig. 1b. This conclusion is in agreement with other IR studies
209 [37]. The DMSO–water interaction is stronger than water–water interaction:
210 the energy of vaporization of water from DMSO (infinitely diluted solution
211 of water) equals to $46.87 \text{ kJ}\cdot\text{mol}^{-1}$ vs. $41.53 \text{ kJ}\cdot\text{mol}^{-1}$, which corresponds
212 to the energy of vaporization of pure water [78]. The results of the DFT
213 calculations confirm this result: interactions of water molecules with the
214 oxygen atom of S=O group are stronger than those between water molecules
215 in the pure water (see Fig. S7 in the Supplementary Material). This is also
216 in agreement with the structural data obtained from the AIMD simulations:
217 the O \cdots O distance in pure D₂O is 0.02 Å longer than the O_{DMSO} \cdots O_{D₂O}
218 distance; see Fig. S10 and Table S2 in the Supplementary Material. On the
219 other hand, this is a typical situation for the aprotic solvent–water systems for
220 which solvent–water interactions are stronger than water–water interactions
221 [78, 79]. Further explanation can be found in the Supplementary Material
222 (section S1.3).

223 3.1.1. DMSO-affected HDO spectra at high water concentration

224 Fig. 2a shows DMSO-affected HDO spectra for high water content in the
225 mixtures, along with the affected spectrum obtained for solution of DMSO
226 infinitely diluted in water, the bulk HDO and the dependence of the affected
227 number (i.e. number of moles of water molecules affected by one mole of
228 DMSO, N), on the water mole fraction (see the inset). It should be noted that
229 this parameter generally does not correspond to a hydration number in strict
230 sense, it rather shows the number of solvent molecules statistically influenced
231 by a solute. Because of lability of the hydration sphere for the most of solutes,
232 only a few water molecules appear to be influenced in their proximity, the rest

233 resembling bulk water. This way, the solute-affected spectrum represents the
234 water status in the “concentrated form”. As it can be seen in Fig. 2a, DMSO-
235 affected $\nu(\text{OD})$ band positions are blue-shifted relative to the bulk water and,
236 on visual inspection, do not show complex structure. The most interesting
237 seems to be the variability of the derivatives $(d\varepsilon(\nu)/dm)_{m=m_i}$ obtained for the
238 molalities of DMSO (m_i) corresponding to the water mole fractions shown in
239 the inset in Fig. 2a. The derivative corresponding to the infinite dilution of
240 DMSO in water ($m_i = 0$) is characterized by two occurrences of the maximum
241 increase of absorption: (1) at 2538 cm^{-1} , which belongs to water molecules
242 with weaker hydrogen bonds than in the bulk water, and (2) at 2438 cm^{-1} ,
243 which corresponds to the strong hydrogen bands of water, as in the case of
244 ice [13]. The last component quickly loses its intensity with the increase of
245 DMSO content in solution and virtually disappears at $x_w \simeq 0.85$, while the
246 first component remains at its place.

247 —Figure 2—

248 For further discussion of the observed spectral effects, we will use the
249 transformation of the DMSO-affected water band-shape to the interatomic
250 oxygen–oxygen ($\text{O} \cdots \text{O}$) distance distribution function between water molecules,
251 $P(R_{OO})$, as described in ref. [45]. The “O” symbol can also denote the oxy-
252 gen atom of the $\text{S}=\text{O}$ group, when the oxygen atom of this group is hydrogen
253 bonded with water molecules. Fig. 2b shows the differences in interatomic
254 $\text{O} \cdots \text{O}$ distance distribution functions, $\Delta P(R_{OO})$, between DMSO-affected
255 water (for water at infinite dilution or for water at $x_w \simeq 0.85$) and the bulk
256 water. These differences qualitatively resemble the features of the evolution
257 of the derivative $(d\varepsilon(\nu)/dm)_{m=m_i}$ in Fig. 2a. However, the difference deter-



258 mined for DMSO-affected water at $x_w \simeq 0.85$, $\Delta P(R_{OO})_{x=0.85}$ (Fig. 2b), still
259 shows an increase of the population of water at short O \cdots O distances, corre-
260 sponding to the ice-like structure, but this population is much smaller than
261 for the solution at infinite dilution.

262 As the water concentration increases ($x_w > 0.85$), the water cage around
263 DMSO molecule develops, thus increasing the population of strong hydrogen
264 bonds between the water molecules surrounding the methyl groups. This
265 increase is illustrated by the derivative $(d\varepsilon(\nu)/dm)_{m=m_i}$ (Fig. 2a) and the
266 $\Delta\Delta P(R_{OO})$ function (Fig. 2b). The structural and energetic state of the
267 DMSO hydration water at the infinitely diluted solution is characterized by
268 the $\Delta P(R_{OO})_{x=1}$ function in Fig. 2b. Accordingly, it should be assumed that
269 for $x_w > 0.85$ strong hydrogen bonds are formed by water molecules both
270 (1) interacting with the S=O group of DMSO, as well as (2) participating
271 in the hydrogen bond network around methyl groups. These two groups
272 of strongly bound water molecules were confirmed by DFT calculation (see
273 Fig.S7 in Supplementary Material). As it can be seen, in terms of the O \cdots O
274 distance both types of hydrogen bonds are poorly distinguishable for small
275 hydration complexes. However, a tendency can also be observed for hydrogen
276 bonds with the S=O group to become weaker as the water cage develops.
277 Only a small population of the water molecules in the hydration shell can
278 be regarded as weakened, as evidenced by the presence of a maximum at
279 ca. 2.90 Å in the $\Delta P(R_{OO})_{x=1}$ function. This population is also visible on
280 the corresponding difference obtained from AIMD simulation (Fig. 3b) and
281 reflects the water–water hydrogen bonds around the DMSO oxygen atom
282 (explanation in section 3.1.2).



283 3.1.2. *Confrontation of experimental results with those obtained from AIMD*
284 *simulations*

285 It is most instructive to confront the experimental HDO spectra in the
286 $\nu(\text{OD})$ range with the computational IR spectra of D_2O obtained from AIMD
287 simulations (see section S4.4 in the Supplementary Material for implementa-
288 tion details). The representative distance-dependent IR spectra that selec-
289 tively capture the absorption of the solute–water complex up to a specified
290 cutoff radius R_c are shown in Fig. 3a. Simultaneously, the hydrogen bond
291 definitions (see section S4.1 in the Supplementary Material) have been used
292 to obtain interatomic oxygen–oxygen distance distribution functions corre-
293 sponding to the experimental ones.

294 —Figure 3—

295 Compared to the experimental liquid D_2O spectrum [80], the respective
296 IR spectrum computed on the basis of AIMD simulations is red-shifted by
297 $\sim 90 \text{ cm}^{-1}$. This magnitude of the red shift is typical of the liquid wa-
298 ter simulations using the generalized gradient approximation (GGA)-based
299 functionals [76, 81] and in explicit inclusion of nuclear quantum effects (NQE)
300 is required to obtain a better agreement with experiment [82, 83]. However,
301 even in the absence of NQE the band shifts obtained from the distance-
302 dependent spectra can be meaningfully compared to experiment [76, 84].

303 We first focus on the spectrum of DMSO molecule extracted from the
304 aqueous solution. As demonstrated previously, such spectra record important
305 information about the hydration shell due to the dipole moment induced
306 by the solvent’s local electric field and the appearance of the thus allowed
307 vibrational transitions in the IR spectrum [76, 84]. In the case of DMSO,

308 the band position at maximum is blue-shifted by 44 cm^{-1} . Coincidentally,
309 this is in perfect agreement with the experimental $\nu(\text{OD})$ band shift for HDO
310 infinitely diluted in DMSO (section S1.3 in Supplementary Material). Since
311 for other hitherto studied solutes similar effects can be observed (i.e., the
312 band shift in the distance-dependent IR spectra is the most pronounced at
313 the vanishing cutoff radius), we hypothesize now that the solute molecule
314 extracted from the solution encodes in a unique way the energetic state of
315 the water molecules at their most perturbed state in the solution, such as at
316 the limit of the infinite dilution in the given solvent.

317 With the increasing cutoff radius, the distance-dependent IR spectra un-
318 dergo major changes in the position and the intensity of the water stretching
319 band [76]. These changes can be monitored selectively at a chosen probe
320 wave number, here taken as the position at maximum of the band at $R_c \rightarrow 0$,
321 $\nu^\circ = 2461 \text{ cm}^{-1}$; see inset in Fig. 3a. With increasing R_c , the intensity of
322 the $\nu(\text{OD})$ band also increases and reaches a local maximum at $R_c = 3.8 \text{ \AA}$.
323 This particular cutoff radius is then selected as the most representative of
324 the DMSO–water complex. The $\nu(\text{OD})$ band position at maximum at this
325 R_c is blue-shifted by 16 cm^{-1} , in excellent agreement with the experimentally
326 observed shift by 11 cm^{-1} ($2519 \text{ cm}^{-1}(\nu^\circ_{\text{DMSO}}) - 2508 \text{ cm}^{-1}(\nu^\circ_{\text{bulk}})$, values for the
327 infinite dilution solution, see Table S1 in Supplementary Material). Simulta-
328 neously, we can obtain the normalization factor for the distance-dependent
329 IR spectra and thus determine the fractional number of water molecules in
330 the considered solute-centered sphere. It is found to be equal to 3.2 ± 0.6
331 (the uncertainty is based on the uncertainty of the R_c value, $\pm 0.1 \text{ \AA}$), also
332 in perfect correspondence with the experimental affected number $N = 2.8$



333 (value for the infinite dilution solution, see inset in Fig. 2a and Table S1 in
334 Supplementary Material). Thus, our distance-dependent IR spectra firmly
335 support the results of the analysis of the experimental data with the affected
336 spectra method.

337 To further explore the parallel character of experimental and computa-
338 tional IR spectra, we again turn to the interatomic oxygen–oxygen ($\text{O}\cdots\text{O}$)
339 distance distributions; see Fig. 3b. We monitor the changes in the geom-
340 etry of the hydrogen bond network with respect to bulk water by applying
341 the distance distribution differences, $\Delta P(R_{OO})$, by analogy to the analysis
342 of experimental spectra (section 3.1.1). By selectively computing $P(R_{OO})$
343 functions around specific sites of DMSO, the local state of the hydrogen
344 bond network can be revealed. As seen in Fig. 3b, the distance distributions
345 around DMSO oxygen undergo much more pronounced changes than around
346 the methyl groups. Interestingly, the $\text{O}_{\text{DMSO}}\cdots\text{O}_{\text{water}}$ hydrogen bonds show
347 a completely different trend than the water–water hydrogen bonds in which
348 water molecules located in the first hydration shell around DMSO oxygen
349 are involved. While the former are found to be strengthened with respect
350 to the bulk water, in accordance with the previously discussed DFT studies
351 of static aqueous clusters (see Fig. S7 in the Supplementary Material), the
352 latter are weakened, possibly because the water molecules involved in short
353 hydrogen bonds with O_{DMSO} have trouble with incorporating into the ex-
354 tended hydrogen bond network. The weakening of hydrogen bonds of water
355 is also illustrated in the appropriate distance distribution difference obtained
356 from the spectral data Fig. 2b. On the other hand, the water–water hydrogen
357 bonds around the hydrophobic groups show only a slight reduction of weak

358 elongated hydrogen bonds with the corresponding appearance of medium-
359 and short-length bonds, but this behavior is qualitatively in agreement with
360 the discussion of the experimental results presented above. We note that
361 similar enhancement of the hydrogen bond network around the hydrophobic
362 groups of TMAO has been recently found [85].

363 3.2. FTIR spectra of DMSO in aqueous solutions

364 3.2.1. Spectra in the range of $\nu(\text{S}=\text{O})$ vibrations

365 The results of FTIR studies of DMSO–water mixtures in the entire com-
366 position range for the $\nu(\text{S}=\text{O})$ band are summarized in Fig. 5. Spectra shown
367 in Fig. 5a correspond to the $\nu(\text{S}=\text{O})$ vibrations in the 975–1150 cm^{-1} range,
368 while the band at ca. 950 cm^{-1} is due to the rocking vibrations of DMSO
369 methyl groups [86]. The assignment of the two low-wavenumber components
370 of the complex band of the $\nu(\text{S}=\text{O})$ vibrations solely to the methyl rocking
371 modes, as in refs [25, 87, 88], is not justified in view of the most recent reas-
372 signment which concludes that the CH_3 rocking modes are actually tightly
373 coupled with the $\nu(\text{S}=\text{O})$ band [89].

374 —Figure 5—

375 As it can be seen in Fig. 5a, the addition of water to DMSO signifi-
376 cantly changes the shape and the position of the $\nu(\text{S}=\text{O})$ band. Generally,
377 the increase of water mole fraction (x_w) results in the decrease of the high-
378 wavenumber component of the SO band (band fitting procedure indicates
379 that this component disappears at $x_w \simeq 0.6$), while the other two compo-
380 nents are red-shifted and their total intensity increases. It is difficult to
381 identify their individual behavior, because most probably these component
382 bands exchange positions at $x_w \simeq 0.4$. The overall spectral changes observed

383 in the case of these two component bands are typical for the proton accep-
384 tor group forming a hydrogen bond, in this case with water molecules. Our
385 AIMD simulations show that (at $x_w \simeq 0.988$) on average 2.5 hydrogen bonds
386 are formed between the DMSO oxygen and the water molecules; see Figs. S8
387 and S9 in the Supplementary Material. It is worth noting that the band at
388 ca. 950 cm^{-1} exhibits a similar behavior. When the mole fraction of water
389 exceeds ca. 0.98 (water to DMSO molar ratio higher than 50) unexpected
390 spectral changes take place: the intensity of the SO band decreases, accom-
391 panied with a small blue-shift, but at the same time its high-wavenumber
392 part, which corresponds to the vibration of the SO group of DMSO in the
393 gas phase [87, 90] increases. This absorption band is broad and flat and its
394 intensity increases with the DMSO dilution. In another work [89] this band
395 was not analyzed and was probably treated as a background.

396 *3.2.2. Molecular complexes of water and DMSO – analysis of the $\nu(S=O)$*
397 *vibration bands*

398 It can be expected that fragments of DMSO linear clusters may also exist
399 in aqueous solutions, resulting mainly from dipole–dipole interactions and
400 the $\text{CH}\cdots\text{O}=\text{S}$ blue-shifted hydrogen bonds. [91]. To check this hypothesis
401 we have removed the contribution of the liquid DMSO spectrum from the
402 measured spectra of DMSO–water mixtures shown in Fig.5a. The results are
403 shown in Fig. 5d, which presents the corresponding dependence of the relative
404 contribution of DMSO not forming hydrogen bonds with water. Residual
405 spectra after this treatment (shown in Fig.S1 in the Supplementary Material)
406 should correspond to the spectra of DMSO forming hydrogen bonds with
407 water. Their position vs. x_w is shown in Fig. 5b. A straight line connecting

408 the extreme values of the band position in Fig. 5b determines the hypothetical
409 ideal behavior of the DMSO–water solutions. It is obvious that the system
410 is non-ideal. The inflection points of the discussed relationship suggest the
411 existence of various molecular complexes. The first pronounced inflection
412 point occurs at $x_w \simeq 0.4$, which corresponds to the $3\text{DMSO}\cdot 2\text{H}_2\text{O}$ complex.
413 However, such complexes are highly labile and can form many structures in
414 the solution (see Fig.S3 in the Supplementary Material). The representative
415 $3\text{DMSO}\cdot 2\text{H}_2\text{O}$ complex is indicated in Fig. 4b. The next point is less marked
416 and occurs at $x_w \simeq 0.85$. In this case, the stoichiometry of the complex is
417 poorly defined and can be approximated by the set of complexes of the type
418 $1\text{DMSO}\cdot n\text{H}_2\text{O}$, where $n = 4 \div 6$. To simplify further discussion, we will
419 use the formula $1\text{DMSO}\cdot 4\text{H}_2\text{O}$, which corresponds to the core structure of
420 the nascent hydration shell around the hydrophobic groups of DMSO. The
421 structure of such a complex is presented in Fig. 4c. The discussed complex
422 still lacks hydrogen bonds between water molecules around the non-polar
423 parts of DMSO molecule. Therefore, the strong hydrogen bonds in this
424 complex, can be attributed to the water molecules interacting with the S=O
425 group and correspond to the population of the $\text{O}\cdots\text{O}$ distances characterized
426 by the maximum at ca. 2.75 \AA in $\Delta P(R_{OO})_{x=0.85}$, as shown in Fig. 2b. In
427 the discussed complex a water cage is developed around the DMSO molecule.
428 This expansion is confirmed by DFT calculations and presented in the form of
429 various structures in Fig. S6 in the Supplementary Material. The formation
430 of the water cage is also evidenced from AIMD simulations which confirm (by
431 spatial distribution functions of water oxygens around DMSO, see Fig. S11)
432 that the particular static structures from DFT calculations indeed correlate



433 with the average solvation structure of DMSO. The last point of inflection is
434 very sharp and corresponds to $x_w \simeq 0.98$.

435 It is worth noting that the distinguished solution compositions (i.e. $x_w \simeq$
436 0.4 and $x_w \simeq 0.85$) correspond well to those observed for the HDO spectra
437 in the range of the $\nu(\text{OD})$ vibrations (section 3.1). Thus, changes in DMSO
438 vibrational structure is highly correlated with changes in the arrangement of
439 solvating water.

440 The results of the principal factor analysis (PFA algorithm) applied to
441 the spectra in Fig. 5a lead to the conclusion that they can be adequately
442 reproduced using four principal factors. This is also the number of absorbing
443 species present in the solutions. By using the window factor analysis, the
444 relative concentration profiles of these species were obtained and presented
445 in Fig. 5c. It should be stressed that factor analysis confirmed the presence of
446 molecular complexes in the tested solutions with compositions corresponding
447 to the inflection points observed in Fig. 5b (black and blue). The first fac-
448 tor corresponding to $\nu(\text{S=O})$ band (red) shows maximum intensity for pure
449 DMSO and can be assigned to DMSO molecules that do not form hydrogen
450 bonds with water. As it can be seen, the dependence of the relative contri-
451 bution of this factor vs. x_w is qualitatively similar to the one obtained in a
452 different way and presented in Fig. 5d (red). On the other hand, the assign-
453 ment of the factor dominant in water-rich solutions (green) requires careful
454 consideration. Its relative contribution can be qualitatively reconstructed by
455 a similar procedure as in the case of the first factor. For this purpose, we
456 analyzed the participation of the boundary spectrum at $x_w \simeq 0.98$ (with the
457 highest intensity in Fig. 5a) in the spectra of water-rich mixtures. It must be

458 stressed, however, that this spectrum does not correspond to the infinitely
459 diluted DMSO. It is only the closest approximation we were able to get,
460 but the obtained contribution profile is reasonable and closely resembles the
461 one calculated with chemometric algorithm. We attribute the fourth fac-
462 tor (green) to DMSO molecules that are able to embed via hydrogen bonds
463 into a clathrate-like cage of water. The corresponding relationship in Fig. 5c
464 shows the variability of the relative proportion of DMSO molecules in this
465 state. It should be added here that the spectra that correspond to $x_w \gtrsim 0.98$
466 were assigned to DMSO molecules that are partly free from hydrogen bonds
467 within the water cage and have gained some degree of rotational freedom.

468 Molecular complexes $3\text{DMSO}\cdot 2\text{H}_2\text{O}$ and $1\text{DMSO}\cdot 4\text{H}_2\text{O}$ are composed of
469 the same molecules, so it can be expected that their enthalpy of mixing will
470 be almost ideal or close to zero. However, they differ in size and, most impor-
471 tantly, in structure. Therefore, the non-ideality of the solution resulting from
472 the mixing of such complexes will be reflected in the positive excess entropy,
473 in accordance with the theory of athermal solutions developed by Huggins
474 [92], Flory [93] and Miller [94]. Hence, a solution with a composition cor-
475 responding to equal relative contributions of both complexes, corresponding
476 to $x_w = 0.67$ (marked with the red dashed line in Fig. 5c) should have the
477 maximum entropy. This, in turn, translates into the minimum of the Gibbs
478 free energy of such solution. Finally, it should consequently have a mini-
479 mum temperature of solidification. In our opinion, the presented reasoning
480 explains the occurrence of an eutectic point for the DMSO–water system at
481 $x_w = 0.67$. Its appearance is thus of entropic nature and corresponds to the
482 largest number of configurations responsible for the intermolecular interac-



483 tions between the highlighted molecular complexes. Previously, Kircher and
484 Reiher [31] proposed a somewhat similar concept. Accordingly, in the eutec-
485 tic mixture many energetically similar, but structurally different molecular
486 complexes exist (proposed by the authors on the basis of *ab initio* calcula-
487 tions) that cannot easily transform their conformations into one another. It
488 leads to the hindered formation of hydrogen bonds between clusters. We
489 want to emphasize that on the basis of our results we see no premise for
490 the existence of the 1DMSO·2H₂O molecular complex, often postulated by
491 other authors (ref. [25] and references cited therein), which would justify the
492 depression of melting point of the solutions at the discussed composition.

493 3.2.3. Spectra in the range of $\nu(\text{C-H})$ vibrations

494 The asymmetric and symmetric $\nu(\text{C-H})$ bands of DMSO methyl groups
495 (located at 2996 ± 1 and $2912 \pm 1 \text{ cm}^{-1}$, respectively, in the case of pure
496 DMSO; data not shown) demonstrate a clear blue-shift of 22 and 12 cm^{-1} ,
497 respectively, as the water content in mixtures with DMSO increases. Such a
498 shift corresponds to the formation of hydrogen bonds of the C-H···OH₂ type
499 between methyl groups of DMSO and lone electron pairs of water molecules
500 [31, 95, 96, 97, 98, 99, 100]. The rate of the shift with increasing x_w clearly
501 accelerates after exceeding $x_w \simeq 0.5$. The corresponding dependencies of the
502 band shift on x_w can be found in previous works [95, 100]. It is important
503 to note the change of the blue-shift into the red-shift for $x_w > 0.96$. We
504 can associate this unexpected change in the direction of the band shift with
505 the already discussed change of the behavior of the $\nu(\text{S=O})$ band in a very
506 similar range of compositions. Clearly, relatively weak hydrogen bonds of
507 the C-H···OH₂ type, which are two to four times weaker than water–water

508 hydrogen bonds [31, 97, 99], play a special or even a key role in the in-
509 termolecular interactions between DMSO and water. The high importance
510 of such hydrogen bonds, disregarding their relatively weak character, stems
511 from their non-negligible number: the analysis of AIMD simulations shows
512 that almost 0.5 hydrogen bonds per C-H bond, and simultaneously 2.5 such
513 bonds per DMSO molecule, are formed on average (see Figs. S8 and S9 in
514 the Supplementary Material).

515 The role of C-H \cdots OH₂ “improper” hydrogen bonds becomes clear when
516 the RDG method is applied to the electron densities calculated for DFT-
517 optimized complexes of DMSO with different number of surrounding water
518 molecules (including both low—classical DFT—and high—ONIOM-based—
519 water content complexes). Such bonds are weaker than other hydrogen
520 bonds, yet their orientation and localization between the hydrogen bond
521 donor and acceptor are as good (Fig. 4). These bonds are present even in pure
522 DMSO complexes and facilitate the self-association of molecules (Fig. 4a).
523 When water molecules are introduced into the complex, such interactions be-
524 come possible also between the oxygen atom of water molecules and methyl
525 groups of DMSO (Fig. 4b). However, the most important role of such bond-
526 ing is clearly visible when the number of water molecules is high enough to
527 form an initial core structure of the future hydration shell (Fig. 4c). These
528 hydrogen bonds impose the position of water molecules and stabilize the
529 shell formation. The core is stabilized solely by proper and improper hydro-
530 gen bonds, and the van der Waals interactions are almost absent, even with
531 10 water molecules per one DMSO (Fig. 6a). These interactions, however,
532 tend to substitute C-H \cdots OH₂ bonds when the number of water molecules is



533 high enough to almost cover one of the methyl groups, as in Fig. 6b. In this
534 figure, two “improper” bonds are visible around the left CH₃ group, while
535 the right one interacts with its hydration shell through van der Waals in-
536 teractions marked with green/olive flat surfaces. In complexes with higher
537 water content the, “improper” hydrogen bonds are almost absent and the
538 hydrophobic part of the DMSO molecule is covered with patches of van der
539 Waals interactions (Fig. 6c).

540 —Figure 6—

541 The change in the character of DMSO–water interactions is reflected also
542 in the simple geometry of ONIOM-calculated complexes. When the hydra-
543 tion shell starts to cover the central DMSO molecule, more and more C–H
544 bonds are released from the “improper” hydrogen bonds. Consequently, the
545 mean C–H bond length is shortened (Fig. S4 in Supplementary Material),
546 which translates into the blue-shift observed in the FTIR spectra. The pro-
547 cess is stopped when the hydration shell is complete and such hydrogen bonds
548 are no more present in the complexes. We note here that the complete hy-
549 dration shell of DMSO at $x_w \simeq 0.988$ (the AIMD-studied system) contains
550 on average 29.2 water molecules (see Fig. S10 and Table S2 in the Supple-
551 mentary Material). The C–H bond length reaches minimum at ca. 30 water
552 molecules per 1 DMSO and increases at higher water contents. We note
553 here that this closely corresponds to the number of water molecules in the
554 full hydration shell of DMSO, as evidenced by DMSO···D₂O radial distribu-
555 tion function obtained from AIMD simulations (see Fig. S10 and Table S2
556 in Supplementary Material). Moreover, the mean $C_{\text{DMSO}} \cdots O_{\text{water}}$ distance
557 increases with the number of water molecules in the shell (Fig. S5 in Supple-



558 mentary Material). Such geometry change of the water cage around DMSO
559 favors partial rotational freedom of the guest molecule inside, as suggested by
560 FTIR spectra of DMSO in the range of $\nu(\text{S}=\text{O})$ vibrations (see also section
561 3.2.1).

562 3.2.4. Increase of the rotational freedom of DMSO molecule in its water cage

563 It seems very likely that hydrogen bonding of the red-shifted and the blue-
564 shifted types are cooperative according to refs [98, 99, 100]. It is also possible
565 that DMSO is not an exception and other solutes, including biomolecules,
566 can undergo a similar process. The way of formation of the hydration cage
567 is generally consistent with the progressive hydration model proposed by
568 Mrázková and Hobza [97]. The only caveat resulting from our research
569 concerns the fact that the blue-shifted hydrogen bonds become less important
570 when the construction of the water cage becomes fully complete and the cage
571 itself becomes more relaxed.

572 The aforementioned weakening of the water hydrogen bonds with the
573 $\text{S}=\text{O}$ group of DMSO (section 3.2.1), progressing with the increasing size
574 of the hydration complex, together with changes in the interaction of the
575 water cage with DMSO molecule and the modification of the size of the cage
576 itself (inferred from the DFT calculations presented in section 3.2.3) create
577 conditions for the partial release of the DMSO molecule from the network
578 of hydrogen bonds with water to a degree that allows a certain rotational
579 freedom. This is supported by the spectral effects observed in Fig. 5a (blue
580 lines) and the change of the blue-shift into the red-shift of the $\nu(\text{C}-\text{H})$ DMSO
581 bands [95, 100] taking place at the lowest DMSO concentrations. The driving
582 force behind the observed phenomenon is the entropy gain associated with



583 the recovery of rotational freedom of the guest molecule.

584 We also note here that the formation of the water cage is connected with
585 the orientational retardation of water molecules in the DMSO hydration shell,
586 as clearly evidenced by the orientational relaxation time of the O–D bond of
587 D₂O increasing from 2.3 ps in bulk water to 3.1 ps in the hydration shell, see
588 the Supplementary Material for details.

589 *3.3. Hydration of DMSO in the light of other water-soluble solutes regarded*
590 *as “hydrophobic”*

591 Recently, Grdadolnik *et al.* [13] have demonstrated, the presence of
592 the ice-like hydration water around small purely hydrophobic solutes (i.e.,
593 methane, ethane, krypton, and xenon). Ice (Ih), solid clathrates of these
594 substances and their solutions gave a single $\nu(\text{OD})$ band of HDO at the same
595 position, ca. $2440 \pm 10 \text{ cm}^{-1}$, which corresponds to $R_{\text{OO}} = 2.76 \pm 0.01 \text{ \AA}$.

596 —Figure 7—

597 Substances included in Fig. 7 allow us to better justify the thesis that the
598 hydration of DMSO in water-rich solutions is hydrophobic in nature, leading
599 to the formation of a clathrate-like structure. Both the tetrabutylammonium
600 cation (Bu_4N^+) and tetrahydrofuran (THF) form solid clathrates [103, 104,
601 105] and are commonly regarded as model hydrophobic molecules [105, 106,
602 44, 101]. The hydration sphere in both cases is composed of ice-like water
603 molecules and the accompanying population of water molecules with $\text{O}\cdots\text{O}$
604 distances longer than in the bulk water. The latter often dominates and
605 obscures the effect of the presence of “ice-like” water molecules [79]. However,
606 among the solutes listed in Fig. 7, DMSO is characterized by a hydration
607 sphere in which water molecules with the “ice-like” structure are dominant

608 ($R_{OO} \simeq 2.75 \text{ \AA}$).

609 The use of differences between interatomic O \cdots O distance distribution
610 functions (Fig. 7b) is probably the most appropriate method for uncovering
611 the presence of subtle and labile structures in solution. Simulation methods
612 can also sometimes fail in this respect if we take into account the information
613 resulting from the analysis of the affected spectra, regarding the affected
614 number, N (note, however, that the normalization factor for the distance-
615 dependent IR spectra derived from AIMD simulations often remains in close
616 correspondence to the experimental N value, as found also in this work).

617 It can be estimated on the basis of Fig. 7b that within the hydration
618 sphere of DMSO the contribution of “ice-like” water ($R_{OO} \simeq 2.75 \text{ \AA}$) relative
619 to the water with a weakened structure ($R_{OO} \simeq 2.90 \text{ \AA}$) is close to 2/3. Tak-
620 ing into account the N value for the infinite dilution solution (2.8, see inset
621 in Fig. 2a and Table S1 in Supplementary Material), we obtain ca. 2 wa-
622 ter molecules which statistically form the ice-like structure in the hydration
623 sphere of DMSO. However, we must take into account that about half of this
624 population are water molecules that form hydrogen bonds with the DMSO
625 oxygen atom (Fig. 2b). On the other hand, for Bu_4N^+ it can be estimated,
626 basing on data in ref. [44], that N equals to ca. 3 and that the contribution
627 of the “ice-like” water in the cation hydration sphere is close to 1/3 (Fig. 7),
628 which indicates that statistically only 1 water molecule of the “ice-like” struc-
629 ture can be found in its hydration sphere. As it can be seen in Fig. 7b, THF
630 ($N = 3.2$ [101]) lies between DMSO and Bu_4N^+ in respect of proportion
631 of the “ice-like” water and the water with weakened structure. Therefore,
632 for typical water-soluble small solutes that contain non-polar groups one can



633 expect on average one water molecule forming the “ice-like” structure.

634 In the light of the above-mentioned Grdadolnik’s findings [13], the ques-
635 tion arises about the origin of the accompanying population of water molecules
636 with weakened hydrogen bonding in the case of hydrophobic water-soluble
637 molecules. The formal answer to this question arises from simple thermody-
638 namic predictions. Because the formation of the “iceberg”, triggered by the
639 presence of the solute, is very entropy consuming and leads to a low solubility
640 of such solute, any other structural opportunity for water molecules which
641 is less “structured” will be more thermodynamically favorable and improve
642 the solubility in water. An explanation on the molecular level should take
643 into account also molecular premises. It should be noted that in the case of
644 Bu_4N^+ cation the solvent accessible surface consists of convex and concave
645 areas. Water molecules in the surroundings of the convex area have a chance
646 to build a network of ice-type hydrogen bonds with an approximately parallel
647 molecule orientation relative to the surface. On the other hand, in the con-
648 cave areas the probability of perpendicular orientation of water molecules,
649 i.e. with their lone electron pairs pointing towards the hydrophobic surface,
650 is growing. As has been discussed in the Supplementary Material (section
651 S1.3), water molecules which do not engage their lone electron pairs in clas-
652 sical hydrogen bonds are less polarized and exhibit a blue-shifted $\nu(\text{OD}/\text{OH})$
653 band. According to this argumentation, those water molecules which are in
654 the concave areas of the Bu_4N^+ cation should contribute to the population of
655 molecules with “weakened hydrogen bonds”. Those molecules which are in
656 the vicinity of convex areas should have contribute to the population of the
657 ice-like water molecules. In the case of DMSO, the accompanying population



658 of molecules with weakened hydrogen bonds originate from improper fit of
659 water molecules hydrogen bonded to the oxygen atom of DMSO, to the bulk
660 water, as recognized in this work from AIMD simulations. In the case of
661 THF hydration, it can be presumed that the source of water molecules with
662 weakened hydrogen bonds is similar in nature to the DMSO case.

663 4. Conclusions

664 DMSO-water mixtures have been studied by means of the FTIR spec-
665 troscopy in the range of DMSO vibration and HDO stretching vibrations
666 (as a probe of hydration water) in the whole range of mixture compositions.
667 Theoretical calculations facilitated the interpretation of experimental results.
668 This helped us to redefine the view on the hydration of DMSO.

669 It has been established that in diluted solutions of DMSO the clathrate-
670 like water cage is created around the DMSO molecule. The formation of
671 this cage is facilitated by interactions between water molecules and methyl
672 groups of the guest molecule (the blue-shifted hydrogen bonds) and it is
673 particularly evidenced by the spatial distribution function of water around
674 DMSO that clearly displays the emergence of cage structures around methyl
675 groups. For even more diluted solutions, when the water cage is fully com-
676 pleted, the DMSO molecule partly regains its rotational freedom inside. The
677 driving force behind this effect is the entropy gain. Weakened hydrogen
678 bonds of water molecules population arise from those water molecules which
679 are hydrogen bonded to the ones already interacting with the oxygen atom of
680 DMSO. Most probably, it happens because strong hydrogen bonds of water
681 molecules directly bonded to the hydrophilic group of the solute poorly fit



682 to the structure of the bulk water. This population with weakened hydrogen
683 bonds efficiently obscures the presence of the population with strong hydro-
684 gen bonds formed with the SO group of DMSO and between water molecules
685 around the methyl groups in the overall spectral effect. We also propose
686 a novel explanation of the strong deviation from ideality of DMSO-water
687 mixtures at composition corresponding to the eutectic point of the system.
688 Namely, it stems from an equimolar mixture of the molecular complexes of
689 the $3\text{DMSO}\cdot 2\text{H}_2\text{O}$ and of $1\text{DMSO}\cdot n\text{H}_2\text{O}$ ($n = 4 \div 6$) type. For such a mixing
690 the, maximum excess entropy is expected, which justifies the depression of
691 the melting point of the solution at $x_w = 0.67$.

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696 **6. Declaration of interest**

697 The authors declare that they have no conflict of interest.

698 [1] P. Ball, Water is an active matrix of life for cell and molecular biology.,
699 Proceedings of the National Academy of Sciences of the United States
700 of America 114 (2017) 13327–13335.

701 [2] W. Kauzmann, Some Factors in the Interpretation of Protein Denatu-
702 ration, Advances in Protein Chemistry 14 (1959) 1–63.

- 703 [3] C. Tanford, The hydrophobic effect and living matter, *Science* 200
704 (1978) 1012–1018.
- 705 [4] M. Charton, B. I. Charton, The structural dependence of amino acid
706 hydrophobicity parameters, *Journal of Theoretical Biology* 99 (1982)
707 629–644.
- 708 [5] R. L. Baldwin, Temperature dependence of the hydrophobic interaction
709 in protein folding., *Proceedings of the National Academy of Sciences*
710 83 (1986) 8069–8072.
- 711 [6] K. A. Dill, Dominant forces in protein folding, *Biochemistry* 29 (1990)
712 7133–7155.
- 713 [7] W. Blokzijl, J. B. Engberts, *Hydrophobic Effects. Opinions and Facts,*
714 *Angewandte Chemie International Edition in English* 32 (1993) 1545–
715 1579.
- 716 [8] V. R. Agashe, M. C. R. Shastri, J. B. Udgaonkar, Initial hydrophobic
717 collapse in the folding of barstar, *Nature* 377 (1995) 754–757.
- 718 [9] H. J. Dyson, P. E. Wright, H. A. Scheraga, The role of hydrophobic in-
719 teractions in initiation and propagation of protein folding, *Proceedings*
720 *of the National Academy of Sciences* 103 (2006) 13057–13061.
- 721 [10] P. Ball, Water as an Active Constituent in Cell Biology, *Chemical*
722 *Reviews* 108 (2008) 74–108.
- 723 [11] H. S. Frank, M. W. Evans, Free Volume and Entropy in Condensed
724 Systems III. Entropy in Binary Liquid Mixtures; Partial Molal En-



- 725 tropy in Dilute Solutions; Structure and Thermodynamics in Aqueous
726 Electrolytes, *The Journal of Chemical Physics* 13 (1945) 507–532.
- 727 [12] T. Hajari, S. Bandyopadhyay, Water structure around hydrophobic
728 amino acid side chain analogs using different water models, *The Journal*
729 of *Chemical Physics* 146 (2017) 225104.
- 730 [13] J. Grdadolnik, F. Merzel, F. Avbelj, Origin of hydrophobicity and en-
731 hanced water hydrogen bond strength near purely hydrophobic solutes,
732 *Proceedings of the National Academy of Sciences* 114 (2017) 322–327.
- 733 [14] D. T. Bowron, A. Filipponi, M. A. Roberts, J. L. Finney, Hydrophobic
734 Hydration and the Formation of a Clathrate Hydrate, *Physical Review*
735 *Letters* 81 (1998) 4164–4167.
- 736 [15] N. Uras-Aytemiz, I. Abrey Monreal, J. P. Devlin, Communication:
737 Quantitative Fourier-transform infrared data for competitive loading
738 of small cages during all-vapor instantaneous formation of gas-hydrate
739 aerosols, *Journal of Chemical Physics* 135 (2011) 2–5.
- 740 [16] J. L. Finney, Overview lecture. Hydration processes in biological and
741 macromolecular systems, *Faraday Discussions* 103 (1996) 1.
- 742 [17] A. K. Soper, Neutron scattering studies of solvent structure in systems
743 of chemical and biological importance, *Faraday Discussions* 103 (1996)
744 41.
- 745 [18] D. Martin, A. Weise, H.-J. Niclas, *The Solvent Dimethyl Sulfoxide,*
746 *Angewandte Chemie International Edition in English* 6 (1967) 318–
747 334.

- 748 [19] U. Kaatze, R. Pottel, M. Schaefer, Dielectric spectrum of dimethyl
749 sulfoxide/water mixtures as a function of composition, *The Journal of*
750 *Physical Chemistry* 93 (1989) 5623–5627.
- 751 [20] Z. Lu, E. Manias, M. Lanagan, D. Macdonald, Dielectric Relaxation
752 in Dimethyl Sulfoxide/Water Mixtures, *ECS Transactions* 28 (2010)
753 11–21.
- 754 [21] S. Y. Lam, R. L. Benoit, Some Thermodynamic Properties of the
755 Dimethylsulfoxide–Water and Propylene Carbonate–Water Systems at
756 25°C, *Canadian Journal of Chemistry* 52 (1974) 718–722.
- 757 [22] J. T. Cabral, A. Luzar, J. Teixeira, M.-C. Bellissent-Funel, Single-
758 particle dynamics in dimethyl–sulfoxide/water eutectic mixture by neu-
759 tron scattering, *The Journal of Chemical Physics* 113 (2000) 8736.
- 760 [23] I. A. Borin, M. S. Skaf, Molecular association between water and
761 dimethyl sulfoxide in solution: A molecular dynamics simulation study,
762 *The Journal of Chemical Physics* 110 (1999) 6412–6420.
- 763 [24] P. P. Wiewiór, H. Shirota, E. W. Castner, Aqueous dimethyl sulfox-
764 ide solutions: Inter- and intra-molecular dynamics, *The Journal of*
765 *Chemical Physics* 116 (2002) 4643–4654.
- 766 [25] V. M. Wallace, N. R. Dhumal, F. M. Zehentbauer, H. J. Kim,
767 J. Kiefer, Revisiting the Aqueous Solutions of Dimethyl Sulfoxide
768 by Spectroscopy in the Mid- and Near-Infrared: Experiments and
769 Car–Parrinello Simulations, *The Journal of Physical Chemistry B* 119
770 (2015) 14780–14789.



- 771 [26] J. M. G. Cowie, P. M. Toporowski, Association in the binary liquid
772 system dimethyl sulphoxide-water, *Canadian Journal of Chemistry* 39
773 (1961) 2240–2243.
- 774 [27] S. S. N. Murthy, Phase Behavior of the Supercooled Aqueous Solu-
775 tions of Dimethyl Sulfoxide, Ethylene Glycol, and Methanol As Seen
776 by Dielectric Spectroscopy, *The Journal of Physical Chemistry B* 101
777 (1997) 6043–6049.
- 778 [28] I. I. Vaisman, M. L. Berkowitz, Local structural order and molecu-
779 lar associations in water-DMSO mixtures. Molecular dynamics study,
780 *Journal of the American Chemical Society* 114 (1992) 7889–7896.
- 781 [29] I. A. Borin, M. S. Skaf, Molecular association between water and
782 dimethyl sulfoxide in solution: the librational dynamics of water,
783 *Chemical Physics Letters* 296 (1998) 125–130.
- 784 [30] A. Vishnyakov, A. P. Lyubartsev, A. Laaksonen, Molecular Dynam-
785 ics Simulations of Dimethyl Sulfoxide and Dimethyl Sulfoxide-Water
786 Mixture, *The Journal of Physical Chemistry A* 105 (2001) 1702–1710.
- 787 [31] B. Kirchner, M. Reiher, The Secret of Dimethyl Sulfoxide-Water Mix-
788 tures. A Quantum Chemical Study of 1DMSO- n Water Clusters, *Jour-
789 nal of the American Chemical Society* 124 (2002) 6206–6215.
- 790 [32] G. J. Safford, P. C. Schaffer, P. S. Leung, G. F. Doebbler, G. W. Brady,
791 E. F. X. Lyden, Neutron Inelastic Scattering and X-Ray Studies of
792 Aqueous Solutions of Dimethylsulphoxide and Dimethylsulphone, *The
793 Journal of Chemical Physics* 50 (1969) 2140–2159.

- 794 [33] D. H. Rasmussen, A. P. MacKenzie, Phase Diagram for the System
795 Water–Dimethylsulphoxide, *Nature* 220 (1968) 1315–1317.
- 796 [34] E. S. Baker, J. Jonas, Transport and relaxation properties of dimethyl
797 sulfoxide–water mixtures at high pressure, *The Journal of Physical*
798 *Chemistry* 89 (1985) 1730–1735.
- 799 [35] A. K. Soper, A. Luzar, Orientation of Water Molecules around Small
800 Polar and Nonpolar Groups in Solution: A Neutron Diffraction and
801 Computer Simulation Study, *The Journal of Physical Chemistry* 100
802 (1996) 1357–1367.
- 803 [36] Y. Koga, Y. Kasahara, K. Yoshino, K. Nishikawa, Mixing Schemes
804 for Aqueous Dimethyl Sulfoxide: Support by X-ray Diffraction Data,
805 *Journal of Solution Chemistry* 30 (2001) 885–893.
- 806 [37] Z. S. Klemenkova, E. G. Kononova, Elucidation of the Water–DMSO
807 Mixing Process Based on an IR Study, *Journal of Solution Chemistry*
808 44 (2015) 280–292.
- 809 [38] R. L. Mancera, M. Chalaris, K. Refson, J. Samios, Molecular dy-
810 namics simulation of dilute aqueous DMSO solutions. A temperature-
811 dependence study of the hydrophobic and hydrophilic behaviour
812 around DMSO, *Physical Chemistry Chemical Physics* 6 (2004) 94.
- 813 [39] R. L. Mancera, M. Chalaris, J. Samios, The concentration effect on
814 the ‘hydrophobic’ and ‘hydrophilic’ behaviour around DMSO in dilute
815 aqueous DMSO solutions. A computer simulation study, *Journal of*
816 *Molecular Liquids* 110 (2004) 147–153.



- 817 [40] K. Subbarangaiyah, N. M. Murthy, S. V. Subrahmanyam, Ultra-
818 sonic Investigation on the Structure of Aqueous Solutions of *N,N*-
819 Dimethylformamide and Dimethyl Sulfoxide, Bulletin of the Chemical
820 Society of Japan 54 (1981) 2200–2204.
- 821 [41] G. Petrella, M. Petrella, M. Castagnolo, A. Dell’Atti, A. De Giglio,
822 Solute-Solvent interactions in water-rich mixtures. II. Ionic conduc-
823 tances in water-dimethylsulfoxide mixtures at 25°C, Journal of Solu-
824 tion Chemistry 10 (1981) 129–138.
- 825 [42] A. Panuszko, P. Bruździak, E. Kaczkowska, J. Stangret, General Mech-
826 anism of Osmolytes’ Influence on Protein Stability Irrespective of the
827 Type of Osmolyte Cosolvent, The Journal of Physical Chemistry B
828 120 (2016) 11159–11169.
- 829 [43] J. Stangret, Solute-Affected Vibrational Spectra of Water in $\text{Ca}(\text{ClO}_4)_2$
830 Aqueous Solutions, Spectroscopy Letters 21 (1988) 369–381.
- 831 [44] J. Stangret, T. Gampe, Hydration Sphere of Tetrabutylammonium
832 Cation. FTIR Studies of HDO Spectra, The Journal of Physical Chem-
833 istry B 103 (1999) 3778–3783.
- 834 [45] M. Śmiechowski, J. Stangret, Vibrational spectroscopy of semiheavy
835 water (HDO) as a probe of solute hydration, Pure and Applied Chem-
836 istry 82 (2010) 1869–1887.
- 837 [46] E. Malinowski, Factor Analysis in Chemistry, Wiley, New York, 2002.
- 838 [47] E. R. Malinowski, Window factor analysis: Theoretical derivation and



839 application to flow injection analysis data, *Journal of Chemometrics* 6
840 (1992) 29–40.

841 [48] M. J. Frisch, G. W. Trucks, H. B. Schlegel, G. E. Scuseria, M. A.
842 Robb, J. R. Cheeseman, G. Scalmani, V. Barone, B. Mennucci, G. A.
843 Petersson, H. Nakatsuji, M. Caricato, X. Li, H. P. Hratchian, A. F.
844 Izmaylov, J. Bloino, G. Zheng, J. L. Sonnenberg, M. Hada, M. Ehara,
845 K. Toyota, R. Fukuda, J. Hasegawa, M. Ishida, T. Nakajima, Y. Honda,
846 O. Kitao, H. Nakai, T. Vreven, J. A. Montgomery Jr., J. E. Peralta,
847 F. Ogliaro, M. Bearpark, J. J. Heyd, E. Brothers, K. N. Kudin, V. N.
848 Staroverov, R. Kobayashi, J. Normand, K. Raghavachari, A. Rendell,
849 J. C. Burant, S. S. Iyengar, J. Tomasi, M. Cossi, N. Rega, J. M. Millam,
850 M. Klene, J. E. Knox, J. B. Cross, V. Bakken, C. Adamo, J. Jaramillo,
851 R. Gomperts, R. E. Stratmann, O. Yazyev, A. J. Austin, R. Cammi,
852 C. Pomelli, J. W. Ochterski, R. L. Martin, K. Morokuma, V. G. Za-
853 krzewski, G. A. Voth, P. Salvador, J. J. Dannenberg, S. Dapprich, A. D.
854 Daniels, Ö. Farkas, J. B. Foresman, J. V. Ortiz, J. Cioslowski, D. J.
855 Fox, *Gaussian 09*, Revision D.01, 2009.

856 [49] E. R. Johnson, S. Keinan, P. Mori-Sanchez, J. Contreras-Garcia, A. J.
857 Cohen, W. Yang, Revealing noncovalent interactions, *Journal of the*
858 *American Chemical Society* 132 (2010) 6498–6506.

859 [50] T. Lu, F. Chen, Multiwfn: A multifunctional wavefunction analyzer,
860 *Journal of Computational Chemistry* 33 (2012) 580–592.

861 [51] S. Grimme, S. Ehrlich, L. Goerigk, Effect of the damping function in

- 862 dispersion corrected density functional theory, *Journal of Computa-*
863 *tional Chemistry* 32 (2011) 1456–1465.
- 864 [52] C. Lee, W. Yang, R. G. Parr, Development of the Colle-Salvetti
865 correlation-energy formula into a functional of the electron density,
866 *Phys. Rev. B* 37 (1988) 785–789.
- 867 [53] A. D. Becke, Density-functional thermochemistry. III. the role of exact
868 exchange, *The Journal of Chemical Physics* 98 (1993) 5648–5652.
- 869 [54] R. Ditchfield, W. J. Hehre, J. A. Pople, Self-consistent molecular-
870 orbital methods. IX. an extended gaussian-type basis for molecular-
871 orbital studies of organic molecules, *The Journal of Chemical Physics*
872 54 (1971) 724–728.
- 873 [55] V. Barone, M. Cossi, Quantum calculation of molecular energies and
874 energy gradients in solution by a conductor solvent model, *Journal of*
875 *Physical Chemistry A* 102 (1998) 1995–2001.
- 876 [56] M. Cossi, N. Rega, G. Scalmani, V. Barone, Energies, structures, and
877 electronic properties of molecules in solution with the C-PCM solvation
878 model, *Journal of Computational Chemistry* 24 (2003) 669–681.
- 879 [57] V. P. Ananikov, D. G. Musaev, K. Morokuma, Real size of ligands,
880 reactants and catalysts: Studies of structure, reactivity and selectivity
881 by ONIOM and other hybrid computational approaches, *Journal of*
882 *Molecular Catalysis A: Chemical* 324 (2010) 104–119.



- 883 [58] T. A. Halgren, Merck molecular force field. I. Basis, form, scope, pa-
884 rameterization, and performance of MMFF94, *J. Comput. Chem.* 17
885 (1996) 490–519.
- 886 [59] T. A. Halgren, Merck molecular force field. II. MMFF94 van der Waals
887 and electrostatic parameters for intermolecular interactions, *J. Com-
888 put. Chem.* 17 (1996) 520–552.
- 889 [60] D. Marx, J. Hutter, *Ab Initio Molecular Dynamics*, Cambridge Uni-
890 versity Press, Cambridge, 2009.
- 891 [61] J. VandeVondele, M. Krack, F. Mohamed, M. Parrinello, T. Chassaing,
892 J. Hutter, Quickstep: Fast and accurate density functional calculations
893 using a mixed Gaussian and plane waves approach, *Computer Physics
894 Communications* 167 (2005) 103–128.
- 895 [62] The cp2k Developers Group, cp2k v. 6.0, 2001–2018.
896 <http://www.cp2k.org/>.
- 897 [63] J. Hutter, M. Iannuzzi, F. Schiffmann, J. VandeVondele, cp2k: atom-
898 istic simulations of condensed matter systems, *WIREs Comput. Mol.
899 Sci.* 4 (2014) 15–25.
- 900 [64] A. D. Becke, Density-functional exchange-energy approximation with
901 correct asymptotic behavior, *Phys. Rev. A* 38 (1988) 3098–3100.
- 902 [65] S. Grimme, J. Antony, S. Ehrlich, H. Krieg, A consistent and accurate
903 *ab initio* parametrization of density functional dispersion correction
904 (DFT-D) for the 94 elements H-Pu, *J. Chem. Phys.* 132 (2010) 154104.

- 905 [66] G. Lippert, J. Hutter, M. Parrinello, A hybrid Gaussian and plane
906 wave density functional scheme, *Mol. Phys.* 92 (1997) 477–487.
- 907 [67] S. Goedecker, M. Teter, J. Hutter, Separable dual-space Gaussian
908 pseudopotentials, *Phys. Rev. B* 54 (1996) 1703–1710.
- 909 [68] M. Nakamura, K. Tamura, S. Murakami, Isotope effects on thermody-
910 namic properties: mixtures of $x(\text{D}_2\text{O or H}_2\text{O}) + (1-x)\text{CH}_3\text{CN}$ at 298.15
911 K, *Thermochim. Acta* 253 (1995) 127–136.
- 912 [69] R. B. Torres, P. L. O. V. A. C. M. Marchiore, Volumetric properties
913 of binary mixtures of (water + organic solvents) at temperatures be-
914 tween $T = 288.15$ K and $T = 303.15$ K at $p = 0.1$ MPa, *J. Chem.*
915 *Thermodynamics* 38 (2006) 526–541.
- 916 [70] I.-F. W. Kuo, C. J. Mundy, M. J. McGrath, J. I. Siepmann, J. Van-
917 deVondele, M. Sprik, J. Hutter, B. Chen, M. L. Klein, F. Mohamed,
918 M. Krack, M. Parrinello, Liquid Water from First Principles: Investi-
919 gation of Different Sampling Approaches, *J. Phys. Chem. B* 108 (2004)
920 12990–12998.
- 921 [71] T. Bryk, A. P. Seitsonen, Ab initio molecular dynamics study of collec-
922 tive excitations in liquid H_2O and D_2O : Effect of dispersion corrections,
923 *Condens. Matter Phys.* 19 (2016) 1–14.
- 924 [72] G. J. Martyna, M. L. Klein, M. Tuckerman, Nosé–Hoover chains: The
925 canonical ensemble via continuous dynamics, *J. Chem. Phys.* 97 (1992)
926 2635–2643.



- 927 [73] N. Marzari, D. Vanderbilt, Maximally localized generalized Wannier
928 functions for composite energy bands, *Phys. Rev. B* 56 (1997) 12847–
929 12865.
- 930 [74] M. Thomas, M. Brehm, R. Fligg, P. Vöringer, B. Kirchner, Computing
931 vibrational spectra from ab initio molecular dynamics, *Phys. Chem.*
932 *Chem. Phys.* 15 (2013) 6608–6622.
- 933 [75] M. Śmiechowski, J. Sun, H. Forbert, D. Marx, Solvation shell resolved
934 THz spectra of simple aqua ions – distinct distance- and frequency-
935 dependent contributions of solvation shells, *Phys. Chem. Chem. Phys.*
936 17 (2015) 8323–8329.
- 937 [76] M. Śmiechowski, Unusual Influence of the Fluorinated Anions on the
938 Stretching Vibrations of Liquid Water, *J. Phys. Chem. B* 122 (2018)
939 3141–3152.
- 940 [77] H. Odhner, D. T. Jacobs, Refractive Index of Liquid D₂O for Visible
941 Wavelengths, *J. Chem. Eng. Data* 57 (2012) 166–168.
- 942 [78] J. Stangret, Donor properties of water in organic solvents derived from
943 infrared spectra of HDO, *Journal of Molecular Structure* 643 (2002)
944 29–35.
- 945 [79] E. Gojło, T. Gampe, J. Krakowiak, J. Stangret, Hydration of Aprotic
946 Donor Solvents Studied by Means of FTIR Spectroscopy, *The Journal*
947 *of Physical Chemistry A* 111 (2007) 1827–1834.
- 948 [80] J. E. Bertie, M. K. Ahmed, H. H. Eysel, Infrared intensities of liquids.
949 5. Optical and dielectric constants, integrated intensities, and dipole



- 950 moment derivatives of H₂O and D₂O at 22 °C, *J. Phys. Chem.* 93
951 (1989) 2210–2218.
- 952 [81] M. Heyden, J. Sun, S. Funkner, G. Mathias, H. Forbert, M. Havenith,
953 D. Marx, Dissecting the THz spectrum of liquid water from first prin-
954 ciples via correlations in time and space, *Proc. Natl. Acad. Sci. USA*
955 107 (2010) 12068–12073.
- 956 [82] L. R. Pestana, N. Mardirossian, M. Head-Gordon, T. Head-Gordon,
957 Ab initio molecular dynamics simulations of liquid water using high
958 quality meta-GGA functionals, *Chem. Sci.* 8 (2017) 3554–3565.
- 959 [83] O. Marsalek, T. E. Markland, Quantum Dynamics and Spectroscopy
960 of Ab Initio Liquid Water: The Interplay of Nuclear and Electronic
961 Quantum Effects, *J. Phys. Chem. Lett.* 8 (2017) 1545–1551.
- 962 [84] M. Leśniewski, M. Śmiechowski, Communication: Inside the water
963 wheel: Intrinsic differences between hydrated tetraphenylphosphonium
964 and tetraphenylborate ions, *J. Chem. Phys.* 149 (2018) 171101.
- 965 [85] S. Imoto, H. Forbert, D. Marx, Aqueous TMAO solutions as seen by
966 theoretical THz spectroscopy: hydrophilic versus hydrophobic water,
967 *Phys. Chem. Chem. Phys.* 20 (2018) 6146–6158.
- 968 [86] W. N. Martens, R. L. Frost, J. Kristof, J. Theo Klopogge, Raman
969 spectroscopy of dimethyl sulphoxide and deuterated dimethyl sulphox-
970 ide at 298 and 77 K, *Journal of Raman Spectroscopy* 33 (2002) 84–91.
- 971 [87] M.-T. Forel, M. Tranquille, Spectres de vibration du diméthylsulfoxyde

- 972 et dn diméthylsulfoxyde-d6, *Spectrochimica Acta Part A: Molecular*
973 *Spectroscopy* 26 (1970) 1023–1034.
- 974 [88] W. R. Fawcett, A. A. Kloss, Solvent-Induced Frequency Shifts in the
975 Infrared Spectrum of Dimethyl Sulfoxide in Organic Solvents, *The*
976 *Journal of Physical Chemistry* 100 (1996) 2019–2024.
- 977 [89] K.-I. Oh, K. Rajesh, J. F. Stanton, C. R. Baiz, Quantifying Hydrogen-
978 Bond Populations in Dimethyl Sulfoxide/Water Mixtures, *Angew.*
979 *Chem. Int. Ed.* 56 (2017) 11375–11379.
- 980 [90] M. Y. Skripkin, P. Lindqvist-Reis, A. Abbasi, J. Mink, I. Pers-
981 son, M. Sandström, Vibrational spectroscopic force field studies of
982 dimethyl sulfoxide and hexakis(dimethyl sulfoxide)scandium(III) io-
983 dide, and crystal and solution structure of the hexakis(dimethyl sul-
984 foxide)scandium(III) ion, *Dalton Trans.* 0 (2004) 4038–4049.
- 985 [91] N. S. Venkataramanan, A. Suvitha, Nature of bonding and cooperativ-
986 ity in linear DMSO clusters: A DFT, AIM and NCI analysis, *Journal*
987 *of Molecular Graphics and Modelling* 81 (2018) 50–59.
- 988 [92] M. L. Huggins, Thermodynamic Properties of Solutions of Long Chain
989 Compounds, *Annals of the New York Academy of Sciences* 43 (1942)
990 1–32.
- 991 [93] P. J. Flory, Thermodynamics of High Polymer Solutions, *The Journal*
992 *of Chemical Physics* 10 (1942) 51–61.



- 993 [94] A. R. Miller, The Vapour-Pressure Equations of Solutions and the Os-
994 motic Pressure of Rubber, *Mathematical Proceedings of the Cambridge*
995 *Philosophical Society* 39 (1943) 54.
- 996 [95] K. Mizuno, S. Imafuji, T. Ochi, T. Ohta, S. Maeda, Hydration of
997 the CH Groups in Dimethyl Sulfoxide Probed by NMR and IR, *The*
998 *Journal of Physical Chemistry B* 104 (2000) 11001–11005.
- 999 [96] H.-C. Chang, J.-C. Jiang, C.-M. Feng, Y.-C. Yang, C.-C. Su, P.-J.
1000 Chang, S. H. Lin, High-pressure spectroscopic probe of hydrophobic
1001 hydration of the methyl groups in dimethyl sulfoxide, *The Journal of*
1002 *Chemical Physics* 118 (2003) 1802–1807.
- 1003 [97] E. Mrázková, P. Hobza, Hydration of Sulfo and Methyl Groups in
1004 Dimethyl Sulfoxide Is Accompanied by the Formation of Red-Shifted
1005 Hydrogen Bonds and Improper Blue-Shifted Hydrogen Bonds: An ab
1006 Initio Quantum Chemical Study, *The Journal of Physical Chemistry*
1007 *A* 107 (2003) 1032–1039.
- 1008 [98] L. Qingzhong, N. Wang, Y. Zhiwu, Effect of hydration on the C-H...O
1009 hydrogen bond: A theoretical study, *Journal of Molecular Structure:*
1010 *THEOCHEM* 847 (2007) 68–74.
- 1011 [99] Q. Li, X. An, B. Gong, J. Cheng, Spectroscopic and theoretical evi-
1012 dence for the cooperativity between red-shift hydrogen bond and blue-
1013 shift hydrogen bond in DMSO aqueous solutions, *Spectrochimica Acta*
1014 *Part A: Molecular and Biomolecular Spectroscopy* 69 (2008) 211–215.



- 1015 [100] Q. Li, X. An, B. Gong, J. Cheng, Comparison of contribution
1016 of OH \cdots OS hydrogen bond and CH \cdots O_w interaction to the methyl
1017 blueshift in hydration of dimethyl sulfoxide, *Vibrational Spectroscopy*
1018 46 (2008) 28–33.
- 1019 [101] J. Stangret, T. Gampe, Hydration of tetrahydrofuran derived from
1020 FTIR spectroscopy, *Journal of Molecular Structure* 734 (2005) 183–
1021 190.
- 1022 [102] A. Panuszko, E. Gojło, J. Zielkiewicz, M. Śmiechowski, J. Krakowiak,
1023 J. Stangret, Hydration of Simple Amides. FTIR Spectra of HDO and
1024 Theoretical Studies, *The Journal of Physical Chemistry B* 112 (2008)
1025 2483–2493.
- 1026 [103] R. McMullan, G. A. Jeffrey, Hydrates of the Tetra *n*-butyl and Tetra
1027 *i*-amyl Quaternary Ammonium Salts, *The Journal of Chemical Physics*
1028 31 (1959) 1231–1234.
- 1029 [104] R. K. McMullan, M. Bonamico, G. A. Jeffrey, Polyhedral Clathrate
1030 Hydrates. V. Structure of the Tetra-*n*-butyl Ammonium Fluoride Hy-
1031 drate, *The Journal of Chemical Physics* 39 (1963) 3295–3310.
- 1032 [105] M. Bach-Vergés, S. J. Kitchin, K. D. M. Harris, M. Zugic, C. A. Koh,
1033 Dynamic Properties of the Tetrahydrofuran Clathrate Hydrate, Inves-
1034 tigated by Solid State 2 H NMR Spectroscopy, *The Journal of Physical*
1035 *Chemistry B* 105 (2001) 2699–2706.
- 1036 [106] W.-Y. Wen, Aqueous solutions of symmetrical tetraalkylammonium

1037 salts, in: R. A. Horne (Ed.), Water and Aqueous Solutions, Wiley-
1038 Interscience, New York, 1972, pp. 613–661.

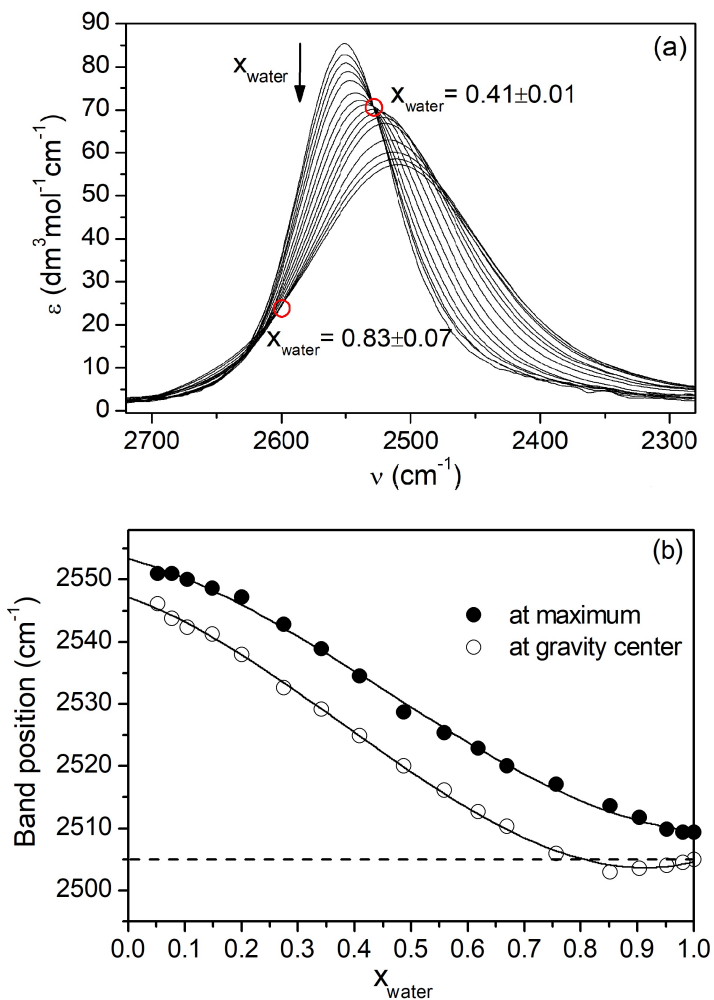


Figure 1: (a) Measured FTIR spectra of DMSO–water solutions at different mole fractions of water in the $\nu(\text{OD})$ stretching vibrations range. The arrow shows increasing water mole fraction in the mixture. Red circles represent isosbestic points together with the corresponding water mole fractions: $x_w = 0.41 \pm 0.01$ (the maximum mole fraction to which the spectra intersect) and $x_w = 0.83 \pm 0.07$ (the minimum mole fraction from which the spectra begin to intersect). (b) The dependence of HDO band parameters on the mole fraction of water in the DMSO–water solutions. The dashed horizontal line denotes the band position at gravity center for pure water spectrum. Solid lines suggest the approximate relationship.

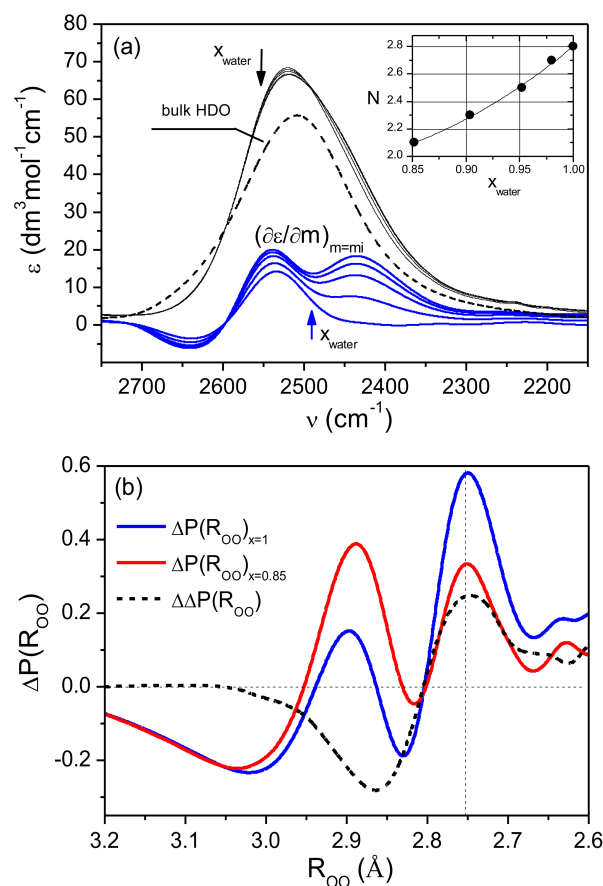


Figure 2: (a) DMSO-affected HDO spectra for high water content in mixtures, including the DMSO-affected HDO spectrum at the infinite dilution, for the water mole fractions as indicated in the inset (solid lines), along with the bulk HDO spectrum (black, dashed). Blue lines represent the derivatives $(d\varepsilon(\nu)/dm)_{m=m_i}$ for aqueous solutions of DMSO with molalities corresponding to the water mole fractions as indicated in the inset. The arrows indicate the direction of change of x_w in the mixture. Insert: The dependence of affected number, N , on the water mole fraction in the mixtures with the high water content. (b) Difference between interatomic oxygen-oxygen distance distribution function of DMSO-affected water at infinite dilution and the bulk water, (obtained on the basis of the spectra shown in Fig. 2a), $\Delta P(R_{OO})_{x=1}$, and the corresponding distance difference for the DMSO-affected water for high water content ($x_w \simeq 0.85$) in the mixture (obtained on the basis of spectra in Fig. 2a), $\Delta P(R_{OO})_{x=0.85}$. The black dashed line shows the difference: $\Delta\Delta P(R_{OO}) = \Delta P(R_{OO})_{x=1} - \Delta P(R_{OO})_{x=0.85}$. The vertical dashed line corresponds to the oxygen-oxygen distance in ice-like water ($2.76 \pm 0.01 \text{ \AA}$).

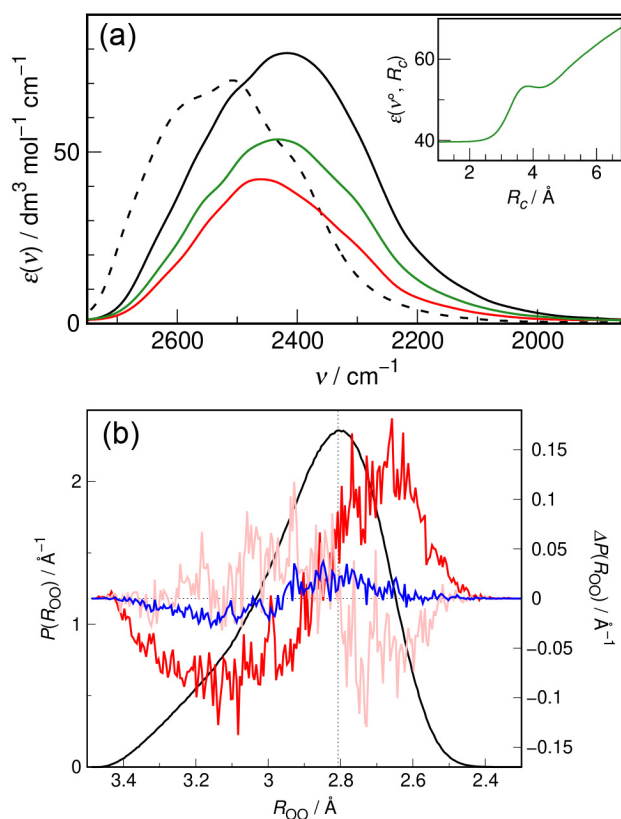


Figure 3: (a) Distance-dependent IR spectra from AIMD simulations at the cutoff radius $R_c \rightarrow 0$ (i.e., DMSO molecule extracted from the solution, red) and $R_c = 3.8 \text{ \AA}$ (green), together with the bulk D_2O spectrum (black) and the experimental IR spectrum of pure liquid D_2O (ref [80], black dashed). The inset shows the dependence of the intensity of the distance-dependent IR spectrum on R_c at the probing wavenumber $\nu^\circ = 2461 \text{ cm}^{-1}$ (the position of the maximum at $R_c \rightarrow 0$). (b) The interatomic oxygen–oxygen distance distribution function for hydrogen bonded water molecules in bulk D_2O (black, left ordinate axis), $P(R_{OO})$, together with the distance differences with respect to bulk D_2O (right ordinate axis), $\Delta P(R_{OO})$, for: water–water hydrogen bonds in the hydration shell of methyl groups of DMSO (up to 4.7 \AA , blue, cf. Fig. S10 and Table S2 in the Supplementary Material), water–water hydrogen bonds to water hydrogen bonded to the DMSO oxygen (up to 3.5 \AA , pink), as well as water–DMSO oxygen hydrogen bonds (red). See section S4.1 in the Supplementary Material for hydrogen bond definitions. The vertical dashed line corresponds to the oxygen–oxygen distance in bulk D_2O ($R_{OO} \simeq 2.8 \text{ \AA}$).



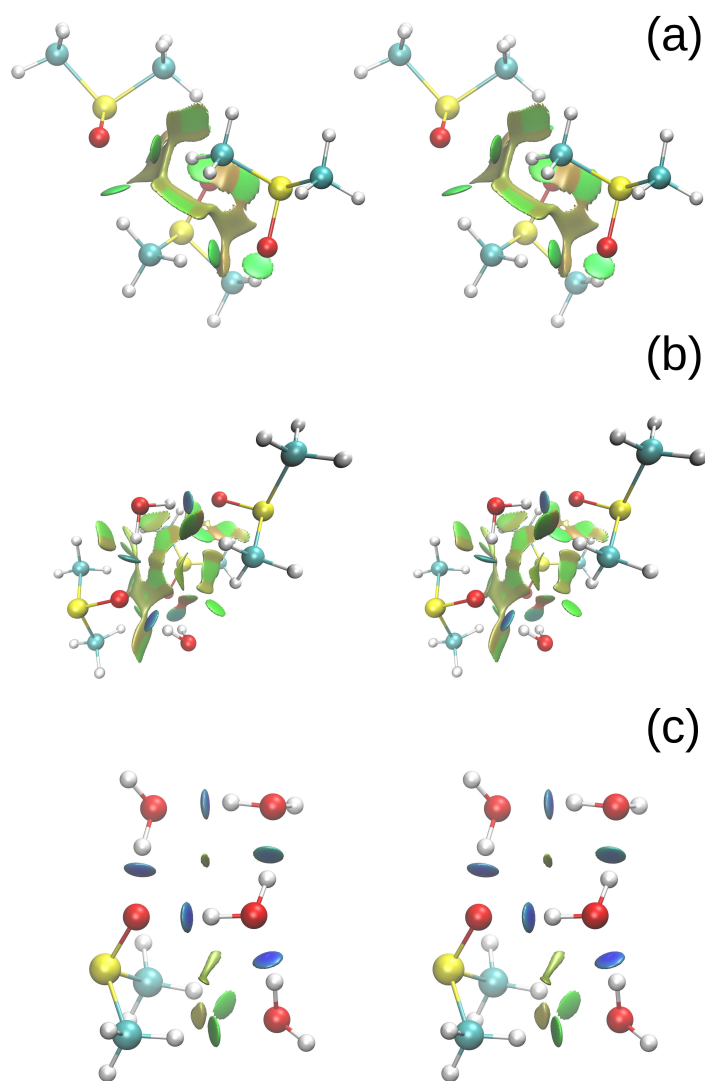


Figure 4: Stereo images of the results of weak interaction analysis of RDG function for (a) 3DMSO and small DMSO–water complexes: (b) 3DMSO·2H₂O, chosen on the basis of the course of changes in the shape of SO band and results of WFA analysis, and (c) 1DMSO·4H₂O, the first step of hydration sphere formation (the core type complex). Brown/olive flat, shapeless or elongated patches indicate van der Waals interactions, blue/green disks correspond to the hydrogen bonds (light green – weak HB, blue – strong HB). RDG calculated on the basis of electron densities obtained for DFT-optimized complexes in the CPCM model (B3LYP/6-311++G(d,p)). See explanation in text.

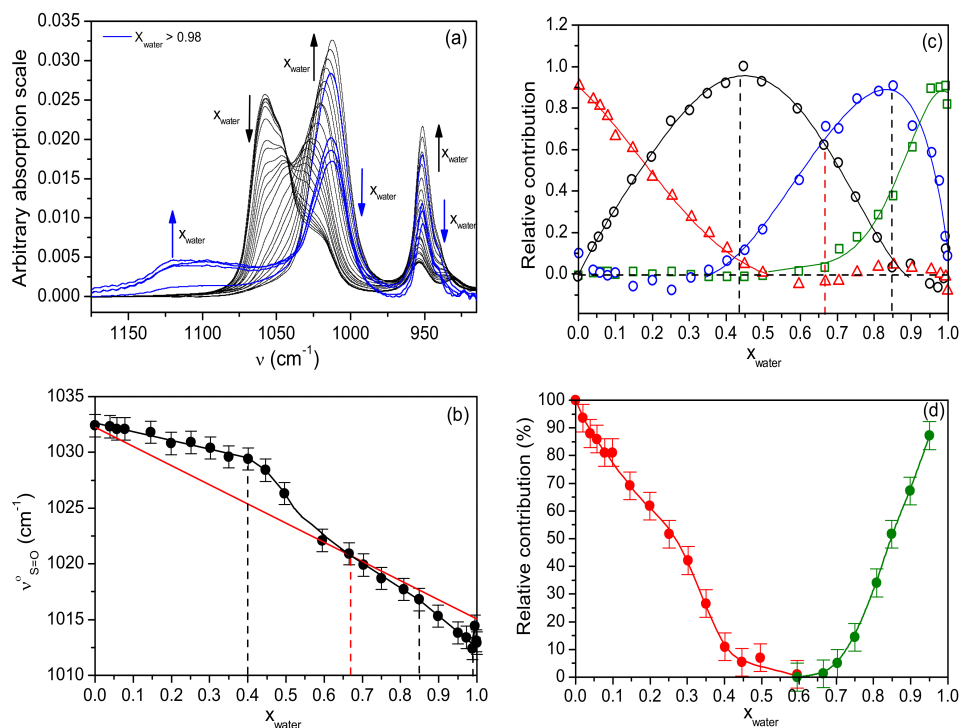


Figure 5: (a) FTIR spectra of DMSO–water mixtures in the range of the $\nu(\text{S}=\text{O})$ and the rocking vibrations of the methyl groups depending on the solution composition. Blue lines represent the spectra for high water content in the mixtures ($x_w > 0.98$). Increasing/decreasing water mole fraction is shown by arrows. (b) The dependence of the $\nu(\text{S}=\text{O})$ band position in maximum for DMSO forming hydrogen bonds with water (based on the spectra shown in Fig. S1 in the Supplementary Material) on the mole fraction of water in DMSO–water solutions. In the set of these residual spectra for $x_w > 0.1$, both low-wavenumber component bands determine the maximum of the complex band of $\nu(\text{S}=\text{O})$. For $x_w < 0.1$ the band position of the low-wavenumber component has been taken into account. Solid red line indicates hypothetical ideal solutions. Vertical dashed lines indicate the composition of complexes: red line denotes the eutectic composition [33], while the black lines correspond to the observed complexes at their maximum concentrations. (c) Relative contributions of different forms of DMSO as a function of water mole fraction from the factor analysis of DMSO–water spectra: DMSO non-bonded to water (red Δ); molecular complex $3\text{DMSO}\cdot 2\text{H}_2\text{O}$ (black \circ); molecular complex corresponding to the core type complex: $1\text{DMSO}\cdot n\text{H}_2\text{O}$, where $n = 4 \frac{51}{6}$, see text (blue \circ); DMSO involved in the creation of the water cage (green \square). (d) Relative contribution of DMSO not forming hydrogen bonds with water (red \bullet), and DMSO involved in the creation of the water cage (green \bullet) as a function of the mole fraction of water in DMSO–water solutions, obtained from the difference spectra method.

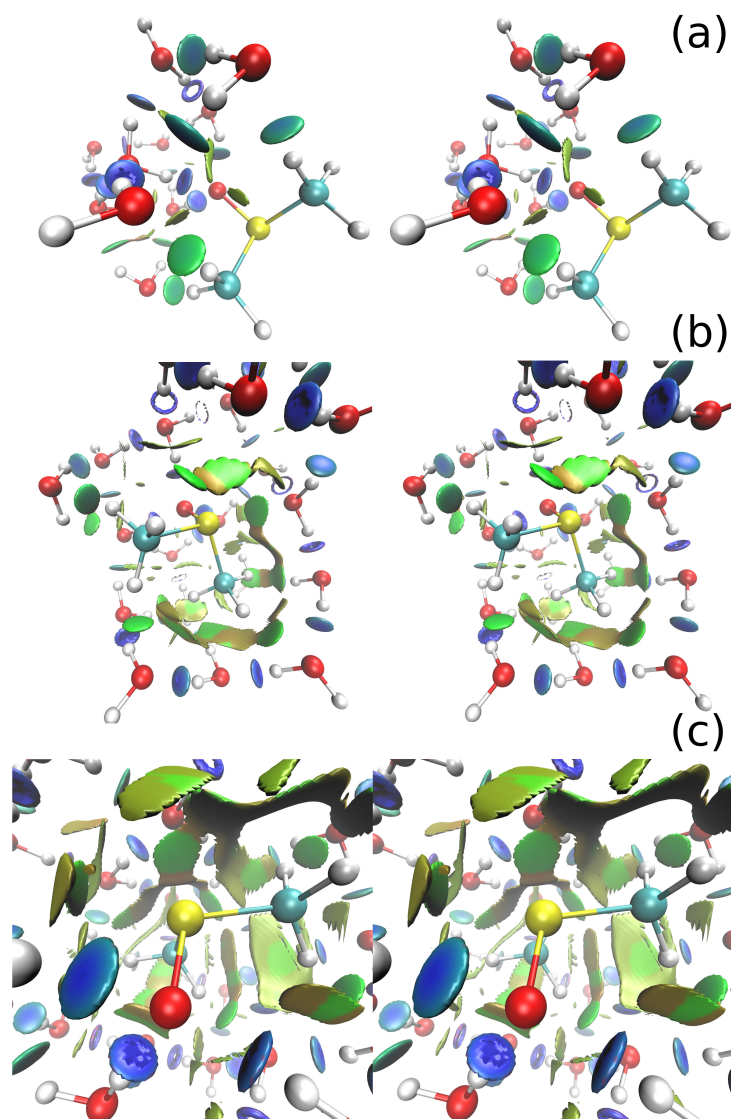


Figure 6: Stereo images of the results of weak interaction analysis by RDG function for larger DMSO–water complexes: a) 1DMSO·10H₂O, b) 1DMSO·20H₂O, c) 1DMSO·30H₂O. The meaning of various structures and colours is the same as in Figure 4. RDG calculated on the basis of electron densities obtained for ONIOM-optimized (B3LYP/aug-cc-pVQZ and B3LYP/cc-pVDZ) systems for high and low level calculation, respectively. See explanation in text.

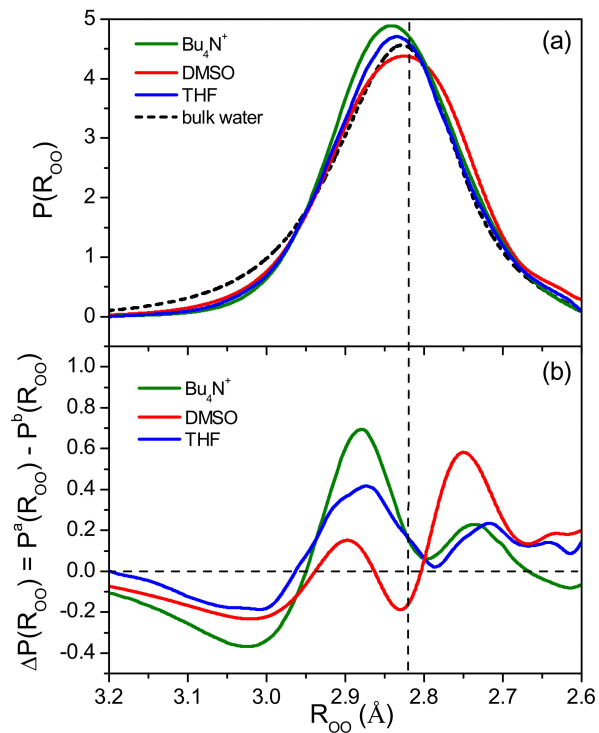


Figure 7: a) Interatomic oxygen–oxygen distance distribution function derived from the HDO spectra affected by DMSO, Bu_4N^+ [44], and THF [101], together with the bulk HDO distance distribution curve. (b) Differences between interatomic oxygen–oxygen distance distribution function for solute-affected water, $P^a(R_{OO})$, and the bulk water, $P^b(R_{OO})$, for DMSO, Bu_4N^+ [102], and THF. The vertical dashed line corresponds to the value of the most probable oxygen–oxygen distance in bulk water (2.83 Å).