



Azo group(s) in selected macrocyclic compounds

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Abstract

Azobenzene derivatives due to their photo- and electroactive properties are an important group of compounds finding applications in diverse fields. Due to the possibility of controlling the *trans*–*cis* isomerization, azo-bearing structures are ideal building blocks for development of e.g. nanomaterials, smart polymers, molecular containers, photoswitches, and sensors. Important role play also macrocyclic compounds well known for their interesting binding properties. In this article selected macrocyclic compounds bearing azo group(s) are comprehensively described. Here, the relationship between compounds' structure and their properties (as e.g. ability to guest complexation, supramolecular structure formation, switching and motion) is reviewed.

Keywords Macrocyclic compounds · Azo group · *Trans*–*cis* isomerization · Host–guest interactions · Molecular switches

Introduction

The year 2017 appears to be a very special for supramolecular chemistry. 50 years ago Charles Pedersen [1] published papers describing the syntheses and completely untypical and unknown until that time intriguing complexing properties of macrocyclic polyethers, i.e. crown ethers [2, 3]. The discovery turned out to be a milestone in chemistry that changed the whole chemical world, gave new fascination and opened up new perspectives for science and technology. Crown ethers are excellent example of unexpected discovery that gained worldwide fame. Since discovery of crown ethers, many their applications have been developed,

for example in chromatography [4, 5], sample preparations [6], catalysis [7–9], and chemical sensing [10].

Macrocyclic compounds had entered the laboratories all over the world, in particular after the discovery of macrocycles containing oxygen and nitrogen electron donors, being the base for three dimensional cryptands, synthesized and studied by Lehn [11, 12] and spherands, obtained and investigated by Cram [13–16]. All these discoveries initiated host–guest [17, 18] and supramolecular chemistry [19–22]. For their achievements Pedersen [23], Cram [18] and Lehn [20] were honored in 1987 with a Nobel Prize. The Nobel Prize in Chemistry 2016 was awarded to: Jean-Pierre Sauvage, Sir J. Fraser Stoddart and Ben L. Feringa “for the design and synthesis of molecular machines” [24–26], which have close relationship with the above mentioned branches of chemistry.

A year after Pedersen's publication on crown ethers and their unique metal cation binding abilities, Park and Simmons published work on macrobicyclic amines i.e. catapinands, the first anion receptors [27–29]. Since that time supramolecular chemistry of anions for many years seemed to be almost forgotten, but last two decades were a renaissance of anion recognition studies [30–33].

Within the last 50-years a lot of macrocyclic compounds of sophisticated structures have been synthesized and investigated [34]. The skeleton of the vast majority of macrocycles can be more or less easily modified by introducing functional groups, which bring about additional chemical or physical features in comparison to the parent compounds as well as

Ewa Wagner-Wysiecka, Natalia Łukasik and Elżbieta Luboch dedicate this article to Professor Jan F. Biernat on occasion of His 80th birthday.

The original version of this article was revised: Scheme 12 was updated.

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to the respective supramolecular species. Functionalized supramolecular systems can be applied in many branches of science and life [35], including e.g. the development of new analytical [36–41] and therapeutic [42–44] systems, modern, intelligent (nano)materials [45–50] and molecular devices and machines [51–57].

One of the most convenient and useful functionalization of macrocyclic compounds is the introduction of azo group(s): incorporated in the ring or on its periphery. Azo moiety due to its ability to alter the geometry upon photochemical or thermal *trans*–*cis* isomerization can be utilized as a light triggered switch in vast variety of functional materials such as for example molecular containers, polymers, supramolecular protein channels, and sensors. As upon photoisomerization process of azo bearing molecules electromagnetic radiation is converted to mechanical work, those compounds can be used in light-driven molecular machines. Here, we present an extensive review of selected azomacrocyclic compounds with the special focus on supramolecular interactions (host–guest complex formation, self-assembly) and *trans*–*cis* isomerization of azo group.

Azobenzene and its derivatives

The properties and functions of the supramolecular systems can be controlled by external stimuli such as changing of pH, temperature, irradiation with the selected wavelength, action with electric or magnetic field. For specific goals, moieties sensitive to one or more of the above factors must be present or introduced to macrocycle structure upon its functionalization. The synthetic routes leading to macrocyclic compounds are often laborious, hence additional functionalization preferably needs relatively simple procedures. A nice example of relatively easy-to-implement functional unit with photo- and redox active properties is the azo group $-\bar{N}=\bar{N}-$, which is also pH sensitive.

Azo compounds are one of the oldest synthesized organic compounds, being produced till now on a large scale in dye industry [58]. The main synthetic approach is based on

diazotization reaction discovered by Peter Griess in nineteenth century. The most common methods of azo group incorporation are schematically shown in Fig. 1. Nowadays, diverse modifications of the original process of diazocoupling are available; also new, synthetic procedures are proposed for preparation of azo compounds for varied purposes [59–66], including methods identified as environmentally friendly [67–69].

Colored azobenzenes and their more sophisticated derivatives, among others, can undergo light-driven reversible *trans*–*cis* ($E \rightleftharpoons Z$) isomerization. The reversible $E \rightleftharpoons Z$ photoisomerization of azobenzene presents well-understood process widely used for construction of light-driven functional molecules for energy storage or conversion of light energy into mechanical motion, exemplified by molecular devices and machines [70]. *Cis* isomer of azobenzene was discovered in 1937 by Hartley [71]. *Trans* (*E*) and *cis* (*Z*) azobenzene isomers are shown in Fig. 2 that also illustrates the reversible isomerization.

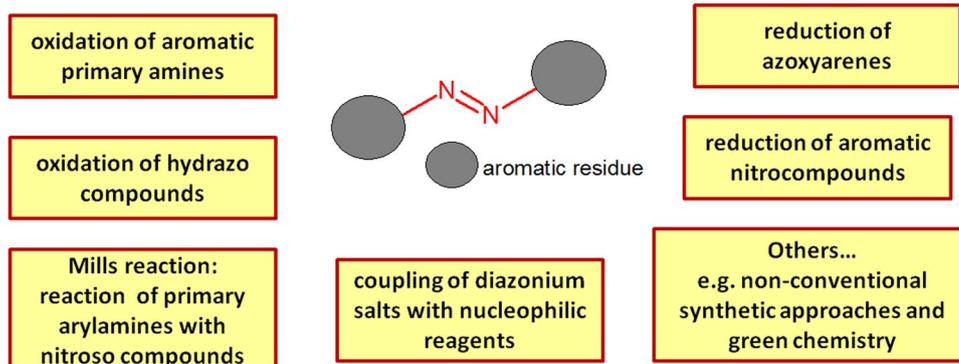
Trans isomer of azobenzene is thermodynamically more stable than the *cis* isomer. In most cases *trans*→*cis* isomerization occurs upon irradiation with UV light (Fig. 2a). However, azobenzene derivatives undergoing reversible *trans*↔*cis* isomerization upon visible light illumination have been also reported [72–74]. Such molecular switches are more applicable and safer for biological uses where harmful ultraviolet light should be avoided.

The *cis*↔*trans* isomerization may occur by spontaneous thermal back reaction or reverse photoisomerization cycle.

The light-driven reversible $E \rightleftharpoons Z$ isomerization of azobenzene is associated with substantial changes of structure, size, geometry and physical properties. Structural changes of azobenzene moiety inbuilt into a larger or more complicated compound affect also the behavior and properties of the azo-functionalized molecular systems like it is for example in photoswitches.

Dipole moment of *trans* isomer of azobenzene is near zero. *Cis* isomer of azobenzene has dipole moment 3.1 D, what determines hydrophobic/hydrophilic character of isomers. *Trans* (*E*) azobenzene is almost planar, opposite to

Fig. 1 Schematically: the main methods for the synthesis of azo compounds



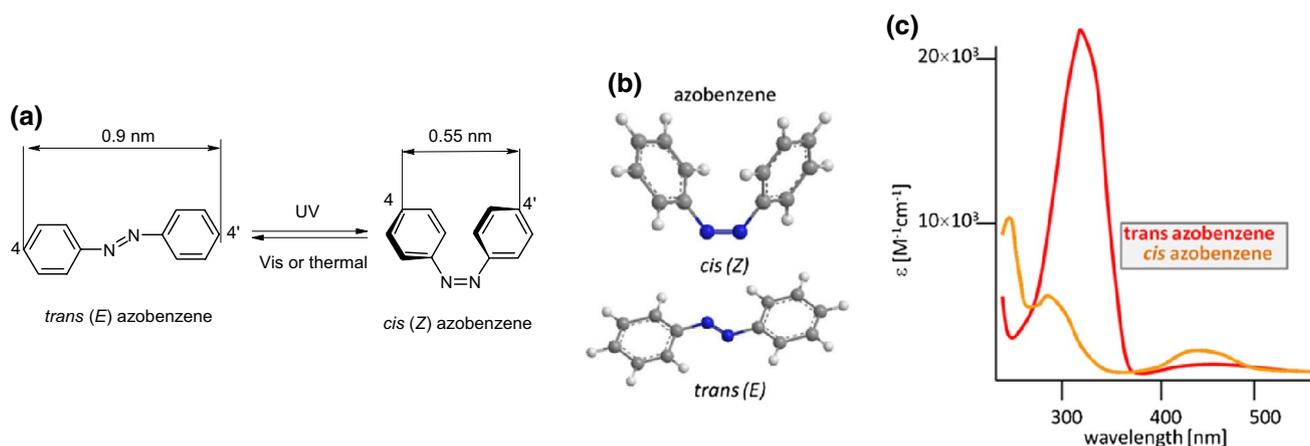


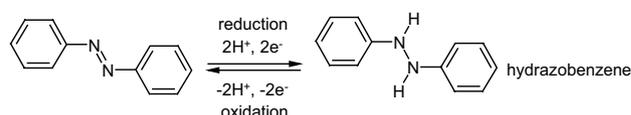
Fig. 2 **a** *Trans* and *cis* isomers of azobenzene and reversible photoisomerization process, **b** models of *cis* and *trans* isomers of azobenzene, **c** schematic UV–Vis spectra of *trans* and *cis* isomers of azobenzene

cis (*Z*) isomer. In solid state in *cis* azobenzene the parallel phenyl rings are twisted 56° out of the plane of the azo group (Fig. 2). The different geometry of *trans* and *cis* isomers of azobenzene affects their UV–Vis spectra. The spectra (Fig. 2c) of *trans* and *cis* isomers are overlapping, but differ significantly. Band at ~ 440 nm originating from $n \rightarrow \pi^*$ transition is more distinct for *cis* isomer. Strong absorption band at ~ 320 nm for *trans* isomer can be attributed to $\pi \rightarrow \pi^*$ transition. In a spectrum of *cis*-azobenzene less intensive $\pi \rightarrow \pi^*$ transition bands are observed at lower wavelength. The spectral differences cause different colors of both isomers, what makes the observation of isomerization process possible also in non-instrumental manner (by naked eye). Spectral properties of azobenzene derivatives are strongly dependent on the substituents in phenyl rings.

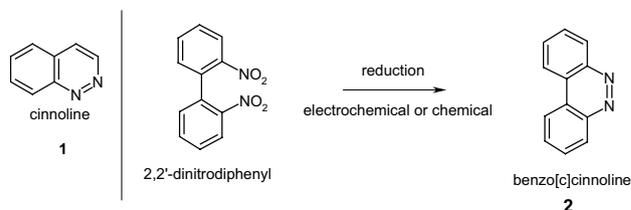
Azobenzene can also act as an important functional unit if incorporated into electrochromic materials (ECMs), which properties can be stimulated by applied potential. Such substances are outstanding candidates for materials used for production of electronic paper [75–78] or dual-stimuli-responsive systems [79, 80].

The electrochemistry of azobenzene and its derivatives in different solvents was studied exhaustively in details for both *trans* and *cis* isomers [81–85]. It was found that the electrochemical reduction of azobenzene is strongly dependent on conditions, such as type of the solvent, pH or reagent concentrations. However, in general it can be summarized that the reduction of azobenzene occurs in a single two electrons, two protons process with a final formation of hydrazobenzene. The simplified way of the electrochemical reduction of azobenzene is shown in Scheme 1.

The properties of self-assembled monolayers of azobenzene derivatives—also macrocyclic—on different surfaces [86–92] showed, that such materials are promising



Scheme 1 The electrochemical reduction–oxidation of azobenzene



Scheme 2 Cinnoline and reductive cyclization of 2,2'-dinitrophenyl as preparation method of benzo[*c*]cinnoline

candidates for molecular devices for energy storage and conversion.

Cyclic and macrocyclic derivatives of azobenzene(s)

Small rings

Derivatives of cinnoline **1**, e.g. benzo[*c*]cinnoline **2** (Scheme 2) can be considered as structural, cyclic analogs of azobenzene. These compounds are used in manufacturing of dyes, electrochromic polymers, coloured polyamide fibers and have microbial and herbicidal activities [93, 94]. Cinnolines were also studied as potential anticancer agents [95, 96]. The reduction of 2,2'-dinitrophenyl to

3,4-benzocinnoline (benzo[*c*]cinnoline) **2** (Scheme 2) was first described by Wohlfart [97] and later by other groups [98–107].

The crystal structure of benzo[*c*]cinnoline complex with ytterbium Yb(BC)₃(thf)₂ (BC = benzo[*c*]cinnoline) was described [108]. Fe₂(BC)(CO)₆ complex was examined as a candidate for a new structural and functional model for [FeFe]-hydrogenases [109, 110].

Modified with benzo[*c*]cinnoline or its derivatives surfaces of e.g. glassy carbon [111, 112], gold [113] or platinum [114, 115] are often used in organic, inorganic, and biochemical catalytic transformations.

Öztürk et al. [116] reported an amperometric lactate biosensor based on a carbon paste electrode modified with benzo[*c*]cinnoline and multiwalled carbon nanotubes. Its characteristics showed, that it can be used for determination of lactate in human serum. Incorporation of benzo[*c*]cinnoline moieties into poly[2-methoxy-5-(2-ethylhexyloxy)-1,4-phenylenevinylene] (MEH-PPV) indicated that p-type semiconductors based on the above polymer can be transformed into n-type materials [117].

Larger analog of cinnoline, (5,6-dihydrodibenzo[*c,g*] [1,2]diazocine) (Fig. 3, compound **3**) comprising azobenzene moiety joined by ethylene bridge at 2,2'-positions was identified as a molecular switch with interesting photochemical characteristics [118–120].

Interestingly, in this case *cis* isomer is thermodynamically more stable than *trans* isomer. Remarkably, both isomers of **3** have well pronounced n→π* bands in UV–Vis absorption spectra. Reversible *trans* to *cis* photoisomerization occurs with efficiency close to 100% under the illumination with visible light of 480–550 nm. The back process of rapid kinetics can be achieved at near ultraviolet at 380–400 nm.

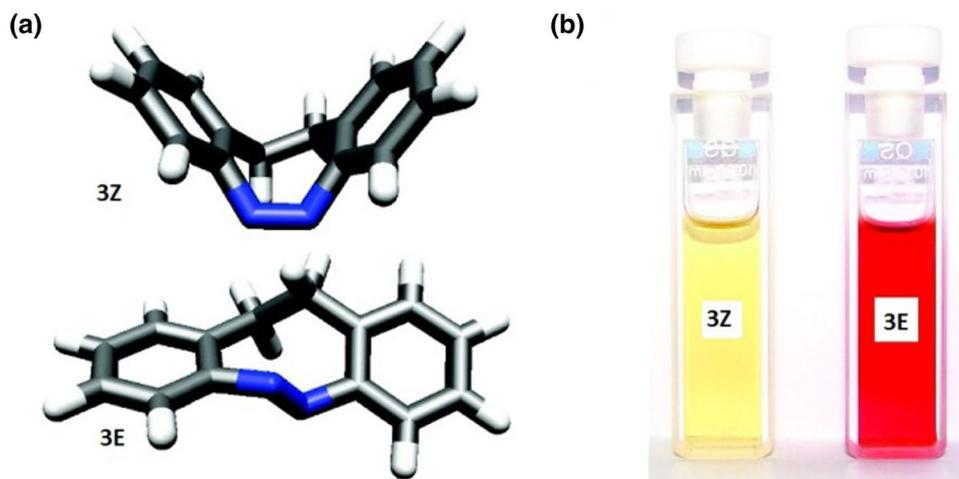
Cyclic oligomers of azobenzene (oligoazobenzophanes)

Oligoazobenzophanes are compounds consisting of at least two or more azobenzene units forming macrocycles. Azobenzene moieties can be joined in *para*-, *meta*- or *ortho*-positions by sp³ hybridized spacers with or without heteroatoms forming relatively flexible, non-conjugated azobenzophanes (Fig. 4a). More rigid, conjugated azobenzophanes are obtained by joining azobenzene units without sp³ tether. As an example of conjugated azobenzophane the simplest cyclotrisazobenzene is shown in Fig. 4b.

Synthetic procedures leading to azobenzophanes involve also approaches typical for macrocyclization, e.g. high dilution technique or template synthesis. The formation of azo group can be the final step of ring closure or can be achieved from substrate(s) bearing functional group(s) by substitution or condensation reactions. An exhaustive review on synthetic protocols was published by Reuter and Wegner [121] that shows preparation of vast varieties of azobenzophane skeletons by cyclizations based on nucleophilic reactions, Schiff bases condensations, reductive or oxidative azocouplings, palladium catalyzed *N*-arylations, and electrophilic aromatic substitutions of diazonium salts.

The utility of azobenzophanes lies in reversible photoisomerization. Opposite to azobenzene for which only two possible states *Z* or *E* can be achieved by photoisomerization, macrocyclic azobenzophanes offer multiple molecular states, depending on the number of azo units. For example, for azobenzophane composed of two azobenzene fragments three states can be considered: *E,E*, *E,Z* and *Z,Z* (Fig. 4c) with the ratio of the isomers depending on e.g. the structure of macrocycle and photoisomerization conditions. The simplest conjugated azobenzophane cyclotrisazobenzene **4** (Fig. 4b) exists only in all-*E* form and has no tendency to be converted into *Z* form under illumination [121]. Unusual behavior of cyclotrisazobenzene was exhaustively

Fig. 3 **a** 5,6-dihydrodibenzo[*c,g*][1,2]diazocine (**3**): equilibrium structures of **3Z** and **3E** in the electronic ground states from quantum chemical calculations at the B3LYP/def2-TZVP level of theory using the TURBOMOLE program and **b** colors of **3Z** before irradiation and color **3E** upon irradiation in *n*-hexane. Adapted with permission from [119]. Copyright 2009 American Chemical Society



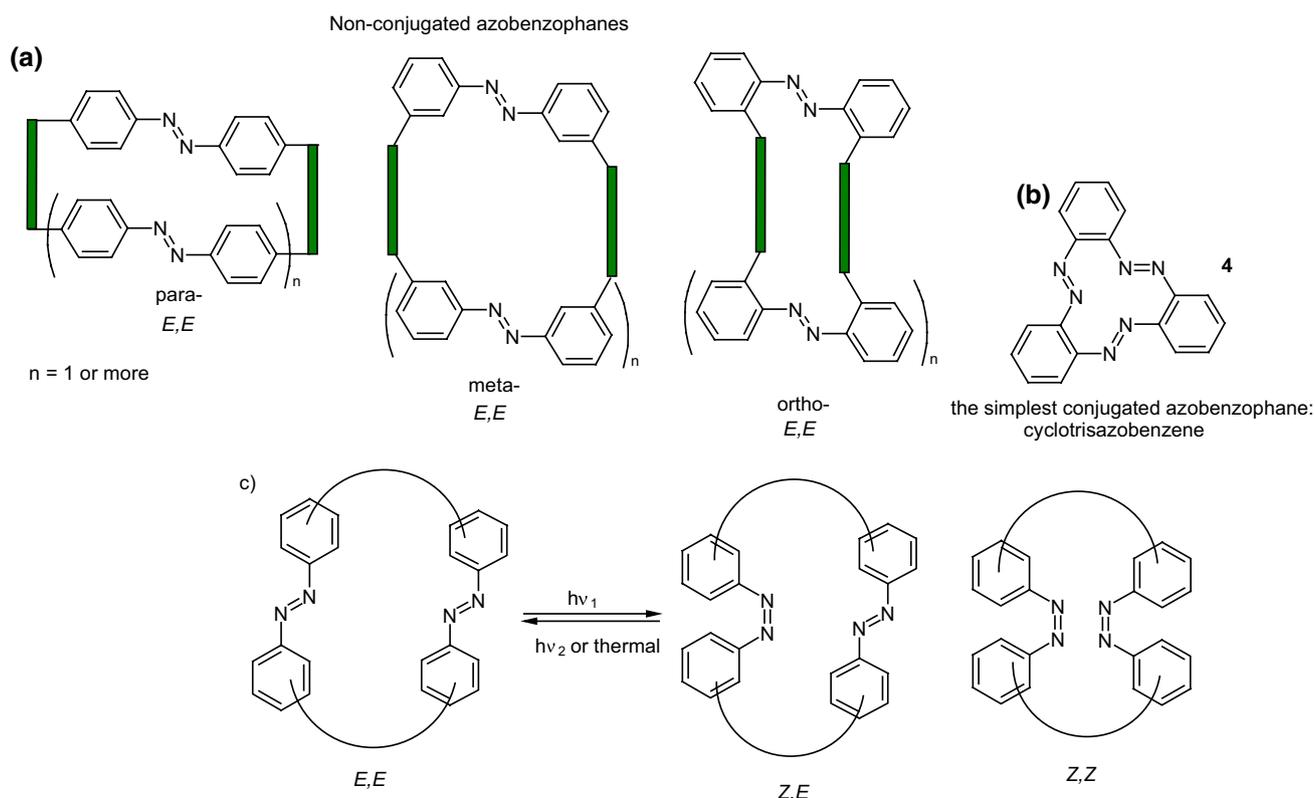


Fig. 4 Azobenzophanes, schematic view: **a** joined in *para*-, *meta*- and *ortho*- positions (*E,E* isomers), **b** the simplest conjugated azobenzophane: cyclotrisazobenzene, **c** reversible isomerization of azobenzophanes

investigated by Dreuw and Wachtveitl [122]. According to experimental and theoretical studies on ultrafast dynamics of this macrocycle the authors stated that the structural constraints prevent isomerization of azo units. The azo bonds respond elastically to the motion along the isomerization coordinates leading to complete and ultrafast dissipation of the UV excitation as heat. It was proposed that the molecules of this type can be used as UV absorbers e.g. in sunscreens.

Azobenzophanes of various structures are studied inter alia as metal cation complexing reagents. Tamaoki and co-workers [123] have obtained a series of azobenzophanes **6–8** (Fig. 5) by reductive macrocyclization of bis(3-nitrophenyl) methane under high dilution conditions. Macrocycles constructed of two, three or four azobenzene units with methylene linkers were isolated as all-*E* isomers. For comparative studies *trans*-3,3'-dimethylazobenzene **5** was prepared (Fig. 5).

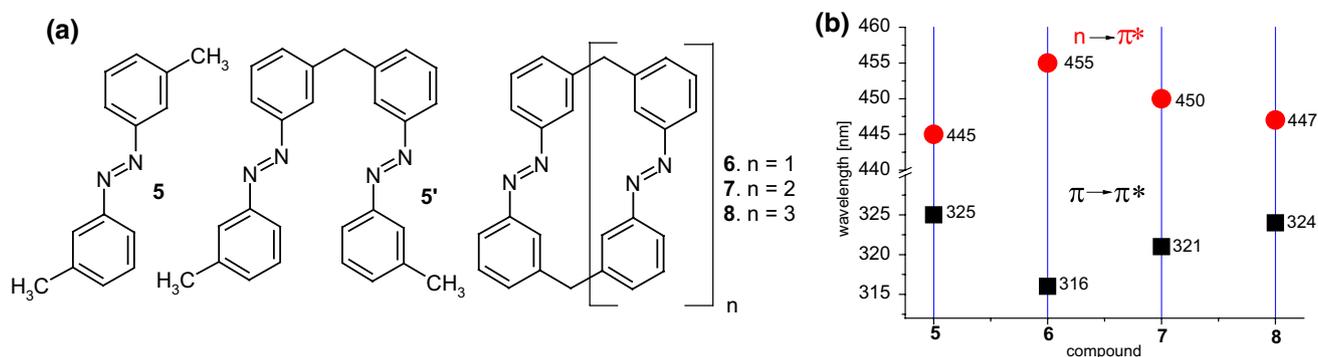


Fig. 5 **a** Azobenzophanes **6–8** and acyclic compounds **5** and **5'** obtained by Tamaoki and co-workers [123, 124], **b** the position of $\pi \rightarrow \pi^*$ and $n \rightarrow \pi^*$ absorption bands in UV–Vis spectra (benzene) of macrocyclic compounds **6–8** and acyclic analogs **5** [123]

The position of UV–Vis absorption maxima for compounds **6–8** and **5** registered in benzene are ring size dependent. The shift of the main band ($\pi \rightarrow \pi^*$) towards longer wavelength can be ordered as follows: $8 > 7 > 6$ and reverse order for $n \rightarrow \pi^*$ (cf. Fig. 5b) and can be associated with steric distortion of the azobenzene moieties. The position of the main absorption band of the largest compound **8** is comparable to a spectrum of open chain analog *trans*-3,3'-dimethylazobenzene.

The UV–Vis spectra of all-*trans* **6** macrocycle and all-*trans* isomer of acyclic dimer **5'** (Fig. 5a) in acetonitrile are compared in Fig. 6 [124]. The same Figure shows also changes upon irradiation with 313 nm wavelength light.

Photoisomerization of macrocycles **6–8** (Fig. 5) and acyclic compound **5'** from all-*trans*(*E*) to all-*cis*(*Z*) isomers proceeds gradually *via* respective *trans*(*E*)/*cis*(*Z*) isomers (depending on the number of azo groups). Comparison of UV–Vis spectra of macrocycle **6** *trans/trans*, and its acyclic analog **5'** is shown in Fig. 6 (left). Photoisomerization studies of all-*trans* isomers of compounds **6–8** showed that the ratio of all-*cis* isomers is irradiation wavelength dependent. The increase in quantity of *cis* azobenzene units upon

irradiation at 366 nm and decrease at 436 nm was observed for **7** and **8** in chloroform. Similar behaviour was found for photoisomerization of **6** in acetonitrile. The ratios of isomers at the photostationary state for compounds **6–8** are schematically shown in Fig. 7a–c [123].

It was found that macrocyclic compounds **6–8** form complexes with alkali metal cations in methanol (determined by mass spectrometry, MS ESI). For all-*trans* isomers the highest peak in mass spectra was observed for cesium complexes; peak intensities for particular macrocycles can be ordered as: $6 > 7 > 8$. The observed trend was disturbed upon irradiation when *cis* isomers also participate in complex formation. It was explained by the softer character of *trans* isomers. However, the clear relationship: the intensity of peaks versus ion diameter in correlation with the size of macrocycle ring was not defined. It was concluded that other factors than only host–guest geometrical complementarity affect the binding strength of metal cations by azobenzophanes **6–8** [123].

Norikane et al. [125] obtained azobenzophanes **9** and **10** (Fig. 8), having structures similar to **6–8** (Fig. 5). The modification of the macrocyclic skeleton by attaching

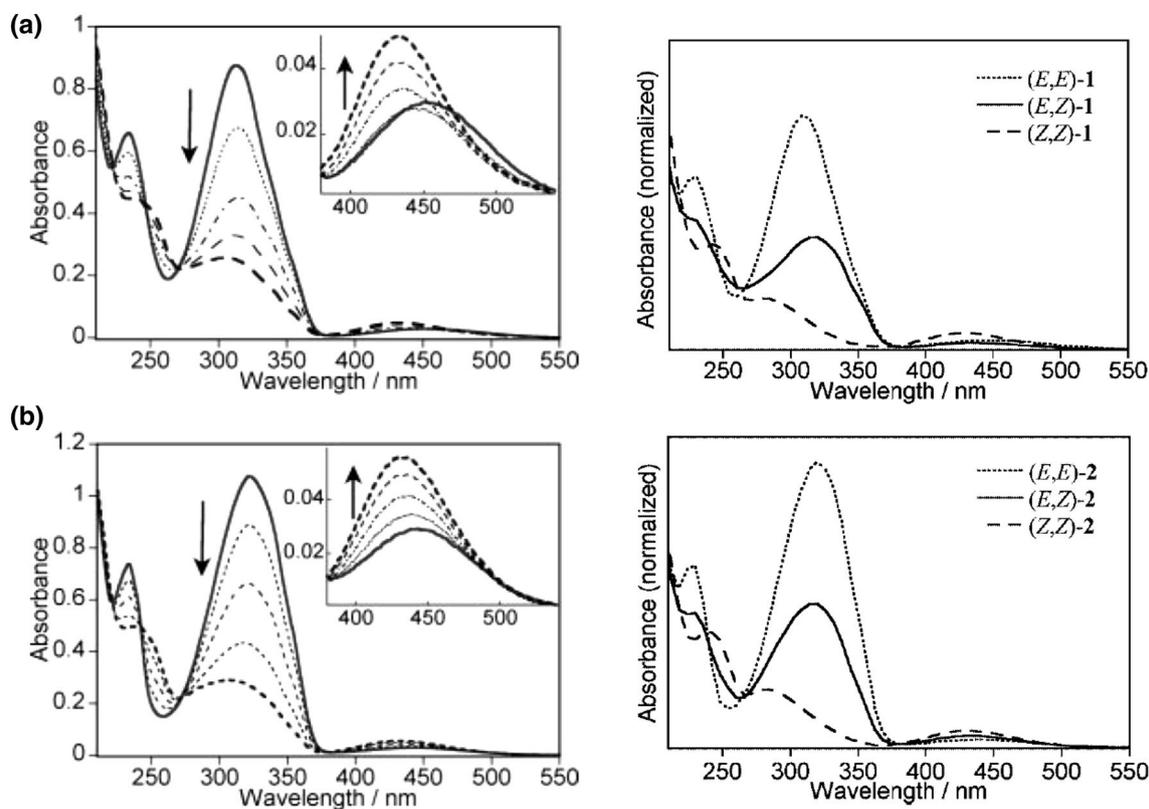


Fig. 6 Left: changes in the absorption spectra of **a 6** and **b 5'** (Fig. 5) in acetonitrile upon irradiation at 313 nm. The insets show the n, π^* band, spectral range (370–550 nm). Bold lines are the initial (solid) and final (dash) traces. Right: absorption spectra of each isomer of **a 6** and **b 5'** measured by use of a photodiode array detector attached to

an HPLC system. Spectra are normalized at the isosbestic points (269 and 272 nm for **6** and **5'**, respectively). Numbers for compounds in reproduced material correspond to following numbers of compounds in this work: 1 = **6**, 2 = **5'**. Reprinted from [124]. Copyright 2006 with permission from John Wiley and Sons

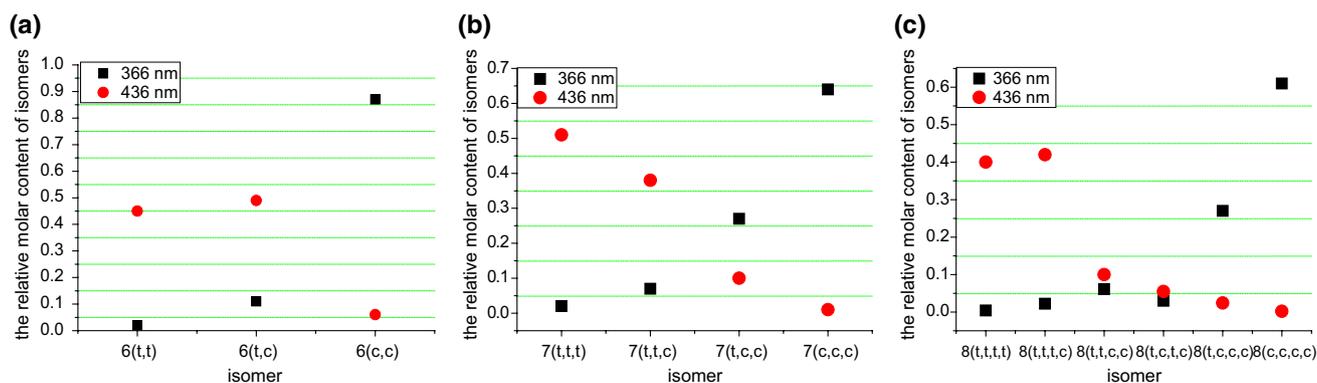


Fig. 7 The ratios of isomers at the photostationary state (PPS) at various wavelength irradiation for **a** **6** in acetonitrile, for **b** **7** in chloroform and **c** **8** in chloroform [123]

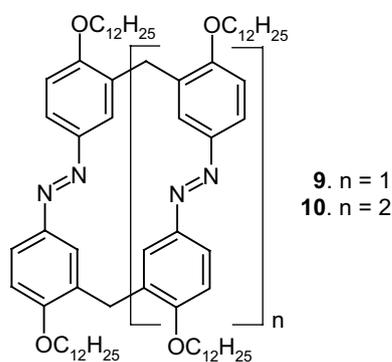


Fig. 8 Azobenzophanes obtained by Norikane et al. [125]

long alkoxy chains resulted in photoresponsive liquid crystallinity.

The effect of different bulky substituents on the properties of azobenzophanes having the same macroring size was investigated by Mayor and co-workers [126]. Four *m*-terphenyl compounds **11–14** (Fig. 9) comprising different peripheral substituents were synthesized by multistep reactions, and different strategies, with the final step of reductive macrocyclization (LiAlH_4 , THF, r.t.) of the respective nitro compounds. Compounds **11–13** are symmetric opposite to derivative **14** with two different substituents at peripheral positions.

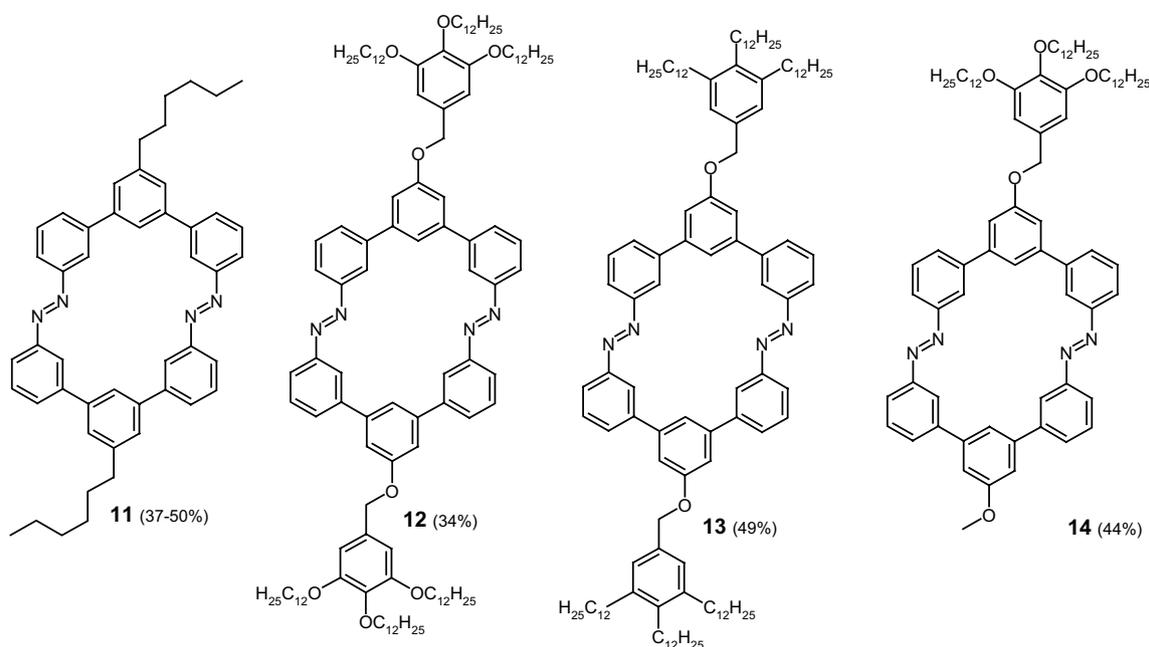


Fig. 9 Bulky azobenzophanes synthesized by Mayor and co-workers [126]

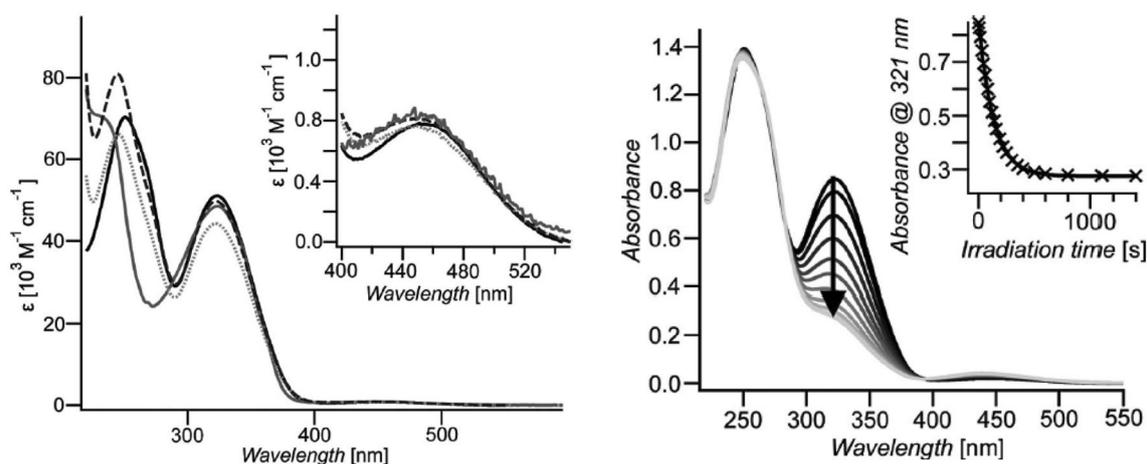


Fig. 10 Left: absorption spectra of azobenzophanes **11–14**: **11**—solid black line, **12**—dashed line, **13**—solid grey line, **14**—dotted line, in THF. Right: changes in absorption spectra of macrocycle **11** in THF upon

irradiation at 313 nm. Inset: absorbance change versus irradiation time. Reprinted from [126]. Copyright 2009 with permission from John Wiley and Sons

The structures of obtained macrocycles were confirmed spectroscopically and the molecular weights of oligomers were determined by vapor pressure osmometry. UV–Vis spectra registered for macrocycles **11–14** in THF are shown in Fig. 10 (left) [126].

Similar shape of spectra, i.e. $\pi \rightarrow \pi^*$ ~ 350 nm and $n \rightarrow \pi^*$ ~ 450 nm was found for all compounds, although below 300 nm in UV–Vis spectra of macrocycles **11–14** the differences are more pronounced. The blue shift of absorption bands for compounds **12** and **13** can be attributed to the effect of substituents on the central phenyl rings. UV–Vis spectra of **11–14** undergo changes upon illumination (313 nm, in THF). For all compounds comparable changes were observed, what is exemplified for **11** in Fig. 10

(right). The photostationary state was reached within 8 min. Photoisomerization is observable in UV–Vis spectra by the decrease of the $\pi \rightarrow \pi^*$ and the increase in the $n \rightarrow \pi^*$ absorption bands upon irradiation over time. These changes are associated with the formation of *Z* isomer. Photoisomerization was monitored by ^1H NMR measurements along with UV–Vis experiments (Fig. 11). By integration of the corresponding ^1H NMR signals the amounts of *E* isomer at the photostationary state was determined to be 15%. In all cases no intermediate *E,Z* isomer was observed as it was in the case of similar systems studied by Tamaoki [123]. This property can be attributed to the extremely rigid structure of macrocycles **11–14**. The back *Z* \rightarrow *E* isomerization proceeds upon illumination or thermally. The *Z* isomers of the

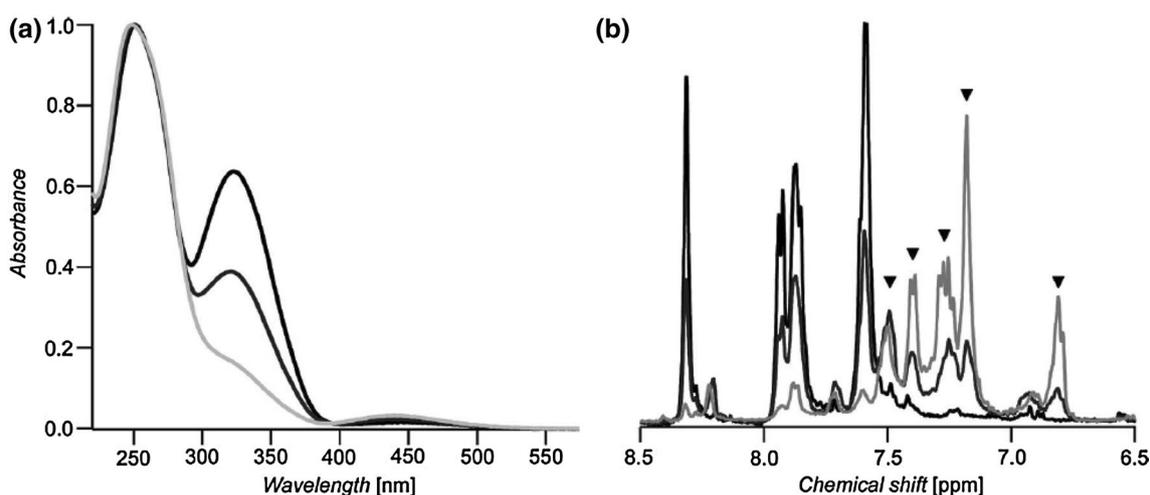


Fig. 11 **a** UV–Vis spectra of macrocycle **11** showing the corresponding *E/Z* ratio (black: thermally stable state; dark grey: 50% isomerized; light grey: photostationary state). **b** Corresponding ^1H NMR

spectra (markers indicate the peaks corresponding to the *Z* isomer). Reprinted from [126]. Copyright 2009 with permission from John Wiley and Sons

macrocycles **11–14** are stable pointing to very slow thermal back-reaction (the rate constant $1.15 \times 10^{-6} \text{ s}^{-1}$). The reversibility of the photoisomerization was investigated under illumination (450 nm). Contrary to the thermal-back process, under which macrocycles were fully converted back to *E* isomers, upon light stimuli ~ 15% remain in their *Z* form. However, this process seems to be reversible what was confirmed by experiments performed in several cycles.

The effect of the strain in azobenzophanes on the photoisomerization of azobenzene unit is well seen in cyclotrisazobenzenes, a special class of azobenzophanes, in which all azobenzene units are conjugated. The simplest compound of this class has been already shown in Fig. 4b. Wegner and co-workers [127] prepared bromo- and *t*-butyl derivatives of the simplest cyclotrisazobenzene **4** using *o*-phenylenediamine as a substrate (Scheme 3).

Irradiation of **4**, **15** and **16** (Scheme 3) showed no isomerization under various conditions. The unfavorable change of geometry upon possible photoisomerization should result in extreme strain in the macrocyclic skeleton, thus **4**, **15** and **16** exist only as all-*E* isomers [cf. 122]

Light controlled sol–gel transition of azobenzene bismacrocycle **17** (Fig. 12) was described by Reuter and Wegner [128].

Due to significant π – π -stacking interactions macrocycle **17** forms 3D networks. Its gelation was observed in aromatic solvents, attributable to the incorporation of the solvent molecule inside the 3D π -stacking network. After UV irradiation at 365 nm the gel in *o*-xylene slowly liquefies as a result of dissociation of 3D network. The gel–liquid conversion of **17** upon irradiation till now is the first example of switchable 3D system which was controlled by two factors: incorporation of azobenzene units and non-covalent interaction, namely π -stacking of the azobenzene macrocycles. The proposed system can be potentially used in process where small molecules are released from the 3D network upon light stimulation.

In photoswitchable cyclic azobenzenes several factors such as ring strain and the number of azo units are crucial for photochemical properties. These features depend also on rigidity and the position of linkers connecting the

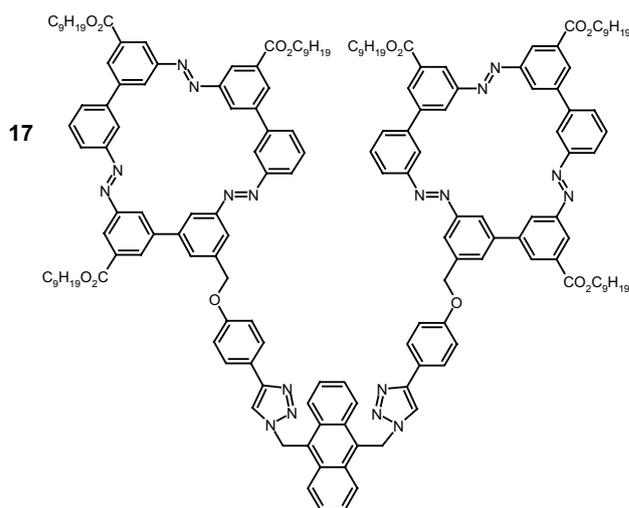
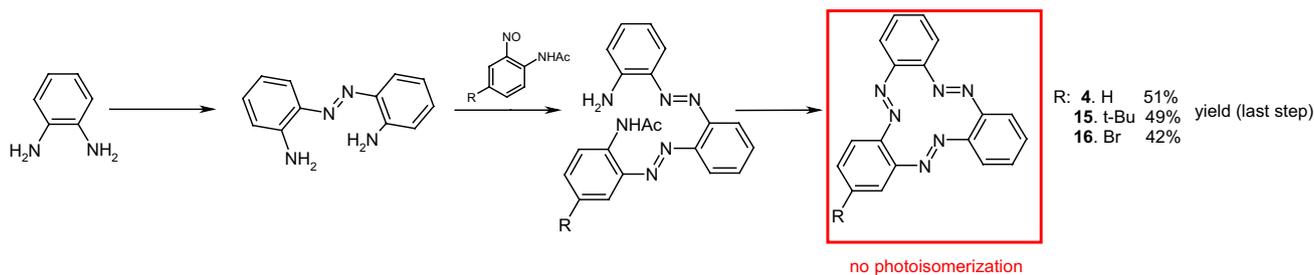


Fig. 12 Macrocycle **17** described by Reuter and Wegner [128]

azobenzene units, the symmetry of the macrocycle and the degree of bonds conjugation. If at the beginning, i.e. before the illumination, a compound with several azo units is fully symmetrical in all-*E* configuration, the change of one of the azo groups into *Z* isomer affects the geometry of the macrocycle. The more azo units in macrocycle the more configuration variations (number of isomers) and geometrical changes can be expected. Wegner and co-workers [129] investigated the effect of symmetry changes on the photostationary state upon *E*→*Z* isomerization stimulated by both light and temperature. For this purpose they used macrocycle **18** with four azo moieties shown in Fig. 13. The isomerization of **18** was monitored by UV–Vis measurements and ^1H NMR spectroscopy with in situ light irradiation. **18** in THF exists in the form of all-*E* isomer. Upon irradiation of this solution (125 μM) at 424 nm for 73 min. a mixture of five among six possible isomers was detected: the starting all-*E* (21%), *E,E,E,Z* (49%), *E,E,Z,Z* (19%), *E,Z,E,Z* (7%), and *E,Z,Z,Z* (4%). Under elevated (50 °C) temperature, at photostationary state, much higher ratio of the all-*E* isomer (55%) was detected, but lower quantities of *E,E,E,Z* (32%), *E,E,Z,Z* (4%) and *E,Z,Z,Z* (1.7%) isomers and almost unchanged



Scheme 3 The general synthetic route for preparation of cyclotrisazobenzenes **4**, **15** and **16** reported by Wegner and co-workers [127]

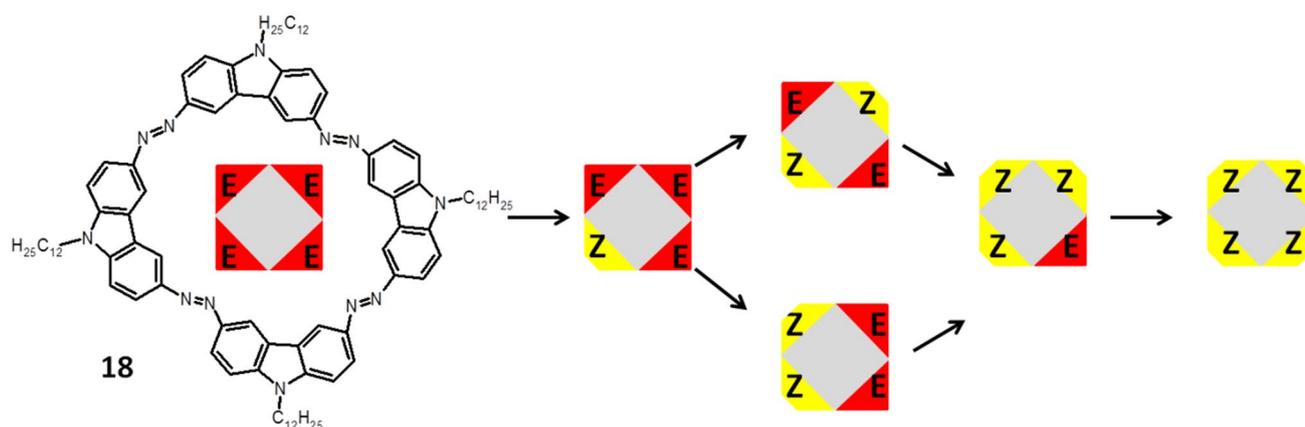


Fig. 13 Macrocycle **18** with four azo moieties studied by Wegner and co-workers and combination of its possible *E-Z* isomers [129]. (Color figure online)

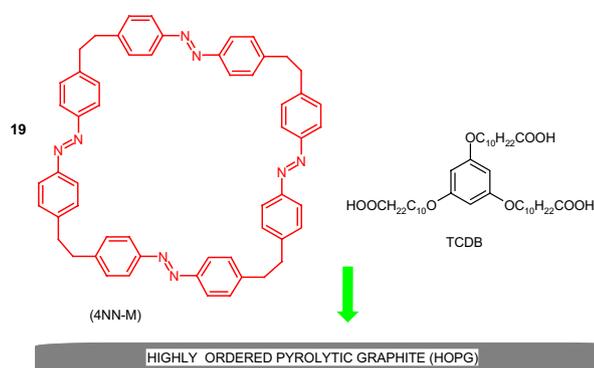


Fig. 14 Photoresponsive system based on macrocycle **19** (4NN-M) and TCDB on HOPG surface [130]

amount of *E,Z,E,Z* (6.8%) isomer. It was concluded that at photostationary state the *E,E,Z,Z* isomer is favored over the *E,Z,E,Z* isomer. Comparison of the rates of thermal back isomerization reveals that the *E,Z,E,Z* isomer has the highest and the *E,E,Z,Z* isomer the lowest thermal stability. This can be ascribed to the ring strain of the particular forms. Different states can be achieved by the arrangement of the azo groups in macroring reflecting the overall symmetry of the molecule without introduction of additional substituents or applying different wavelength of the light used for illumination.

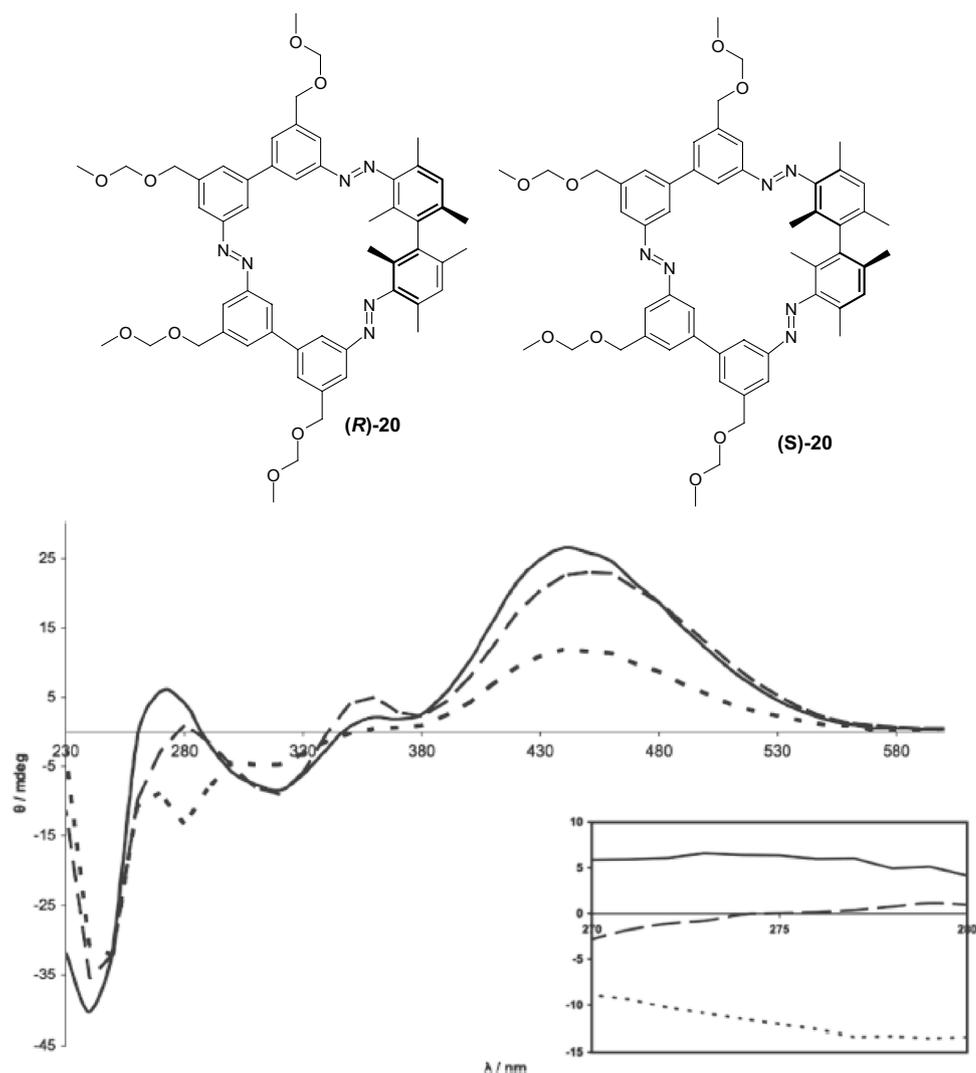
Interesting, well organized system utilizing highly ordered pyrolytic graphite (HOPG) based on the photosensitive macrocycle bearing four azobenzene units **19** (4NN-M, Fig. 14) immobilized in the TCDB network was obtained and investigated by Wang and co-workers [130]. Upon UV illumination of the prepared material *E,E,E,Z* and *E,Z,E,Z* isomers are present at photostationary state. The proposed methodology was found to be useful for fabrication of

nanostructures and can be valuable for production of photo-sensitive nanodevices.

A ternary switch utilizing chiral macrocycle was presented by Reuter and Wegner [131]. They obtained both enantiomers *R* and *S* and racemic form of chiral bis-mesitylcyclotrisazobiphenyl compound **20** (Fig. 15) in about 40% yield.

Three different photostationary states were gained by irradiation of **20** with different UV (302 and 365 nm) and visible light. The photoisomerization was investigated by CD spectroscopy. A large increase in the optical rotation angle for (*S*)-**20**: $[\alpha]_D^{20} = 2128^\circ$ and for (*R*)-**20**: $[\alpha]_D^{20} = -2077^\circ$ in comparison with acyclic 3,3'-diaminobismesityls comes from the helical shape of macrocyclic compounds. All *E*-isomer was obtained by heating samples of (*S*)-**20** and (*R*)-**20** at 45 °C overnight. CD spectra of two all-*E* enantiomers are mirror images with four different absorption maxima. Upon irradiation of all *E*-isomer of (*S*)-**20** with three different wavelength the photostationary state was reached after ~15 min. For (*S*)-**20** seven different isomers were detected by ^1H NMR measurements: six species being different *E/Z* isomers (one (*E,E,E*), two (*E,E,Z*), two (*E,Z,Z*), and one (*Z,Z,Z*)). The seventh one was described as a stable conformer of the (*E,E,Z*)-isomer with azo bond next to the bimesityl unit in *Z*-form. The different ratio of these isomers at particular photostationary state is manifested in CD spectra that varied mostly in intensities, but with preserved similar overall shape. However, a difference can be observed at 275 nm, when irradiating sample with mentioned above three different wavelengths: at 302 nm—positive value, visible light—no dichroism, at 365 nm negative value what is promising for ternary switch with +, – and 0 output (Fig. 15, bottom).

Fig. 15 Top: (*R*)-**20** and (*S*)-**20** enantiomer of bis(mesityl)clotrisbiphenyl macrocycle obtained by Reuter and Wegner [131]. Bottom: CD spectra at different photostationary states at 302 nm (solid line), visible light (dashed line), and 365 nm (dotted line) (5.9×10^{-5} M), with an enlarged graph for the region 270–280 nm (inset). Reprinted with the permission from [131]. Copyright 2011 American Chemical Society



Crown ethers with azobenzene moiety(-ies)

Azobenzene unit incorporated into crown ethers skeleton was first reported by Takagi and co-workers almost 40 years ago [132, 133]. Azo bearing crowns **21–24** (Fig. 16a) were obtained by Williamson reaction from dihydroxyazobenzene and alkylating agents. The synthesis of this type of compounds (**21–23**, Fig. 16a) was also elaborated in details by Biernat and co-workers [134–140]. Reductive macrocyclization of dinitropodands allowed the preparation of vast number of macrocyclic compounds showing diverse properties. By this method azoxycompounds are formed next to azocompounds. They were studied e.g. as ionophores in ion-selective membrane electrodes and chromogenic agents for metal cation complexation. At first glance—these simple compounds bring a great potential in supramolecular chemistry not only as metal cation complexing properties, but also due to photosensitivity. There are also known crown ethers with azo group located at the periphery of the molecule

with brilliant example of so called “butterfly crown ethers” obtained and investigated by Shinkai et al. [141, 142]. These photo-switchable compounds were used for light-driven transport of potassium and sodium. Figure 16b shows the scheme of light-driven transport of potassium cations across organic bulk membrane with the use of photoresponsive azobis(benzo-15-crown-5) **25**.

These early works on azo group bearing crown ethers inspired further development of synthetic methods, challenging functionalization, and studies (both experimental and theoretical) of properties and finally applications of macrocyclic polyethers.

Crown ethers with inherent azobenzene group(s)

Among the first synthesized crown ethers with azo unit incorporated into the macrocycle were so called “all or nothing” crown ethers exemplified by **26** (Fig. 17) obtained by Shinkai et al. [141, 143]. These photoswitchable

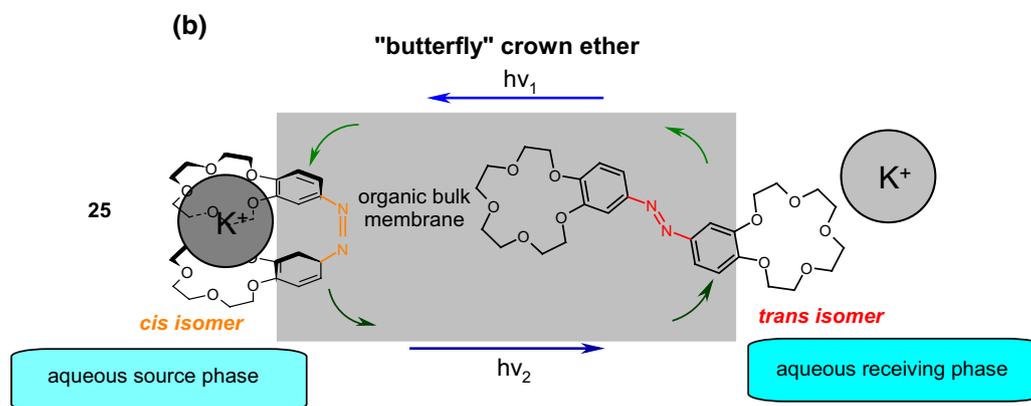
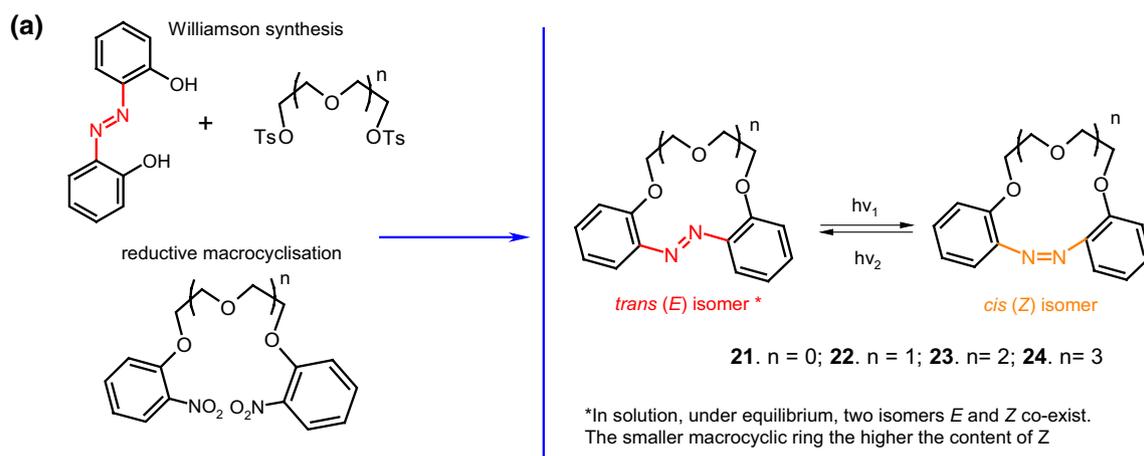


Fig. 16 a Crown ethers **21–24** with inherent azo group (Takagi's and Biernat's group) [132–140] (b) **25** - azobis(benzo-15-crown-5) an example of butterfly crown ethers obtained and studied by Shinkai [141, 142]

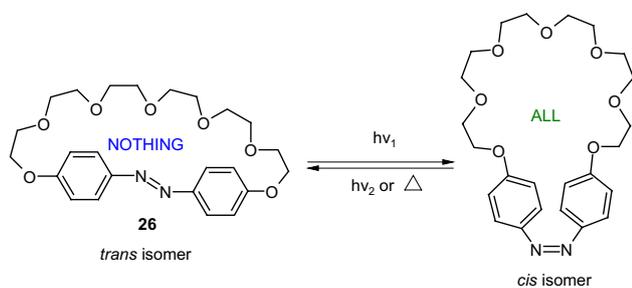


Fig. 17 Example of "all-or nothing" photoswitchable crown ether **26** [141, 143]

compounds form complexes with metal cations with affinity that depends on the geometry of azo group. The *cis* isomer obtained by illumination binds cations, whereas in the dark the cation is released due to decreasing the cavity size being a consequence of isomerization to *trans* form. The spectral behavior of "all or nothing" crowns of different size of the macrocycle and their ability to form complexes with alkali metal cations was later studied theoretically

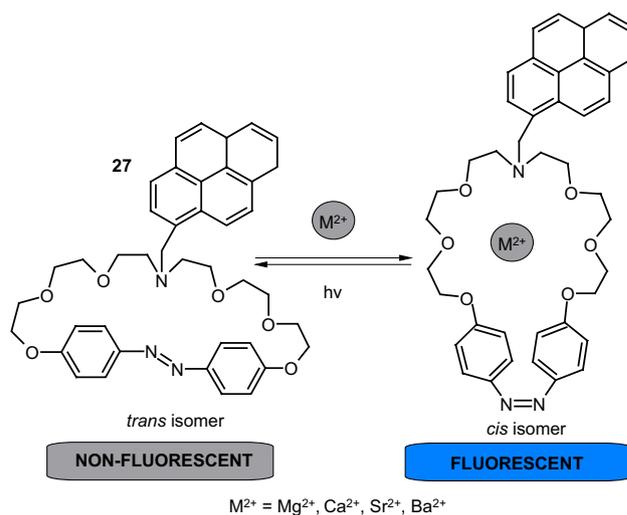


Fig. 18 Azobenzene embedded *N*-(11-pyrenyl methyl)aza-21-crown-7, **27** studied by DFT by Wang and co-workers [145]

using density functional theory (DFT) [144]. The results showed good agreement between experimental and computational attempts.

Computational methods were also used by Wang and co-workers [145] to study *trans*-azobenzene embedded *N*-(11-pyrenyl methyl)aza-21-crown-7 **27** (Fig. 18) as a fluorogenic receptor for alkaline-earth metal cations.

According to density functional theory using B3LYP/6-31G(d) it was determined that the ether chain of *trans* isomer of the compound becomes almost a straight line forming a strip crown ring. Calculated structure of *cis* isomer shows cavity enables coordination of metal cation inside the macrocycle. The optimized structures of complexes of the host molecule and alkaline earth metal cations (Mg^{2+} , Ca^{2+} , Sr^{2+} , and Ba^{2+}) indicate that the ligand binds calcium cations the strongest due to the best match of ion radius to the cavity size. These results showed, that proposed system can act as molecular device of double function.

Tamaoki and co-workers [146] studied the effect of *trans*–*cis* isomerization of [5.5](4,4′)azobenzene(1,5)naphthalenophane **28** (Fig. 19) on silver(I) complexation. The resolved crystal structure of 1:1 complex showed that two silver cations are complexed to form dimeric structure with azobenzenonaphthalenophane in *trans* form (Fig. 19 left).

^1H NMR studies showed that complexation of silver cation is controlled by reversible *trans*–*cis* isomerisation of azo moiety; photoisomerization of *trans* to *cis* isomer causes the cleavage of the π –cation interaction. The opposite change was found under reverse isomeriation (Fig. 19, right).

Kirichenko and co-workers [147] described synthesis and complexing properties of four crownphanes **29**–**32** (Fig. 20) containing 2,7-dioxyfluorenone and 4,4′-azobiphenoxy groups joined with di-, tri-, tetra-, and pentaethylene glycol moieties. Based on NMR, UV–Vis, and X-ray data it was concluded that all macrocycles exist in solution and in solid state in *trans*-configuration of azobenzene unit. The *trans* to *cis* isomerization of **30** can be achieved by UV-light (365 nm) irradiation.

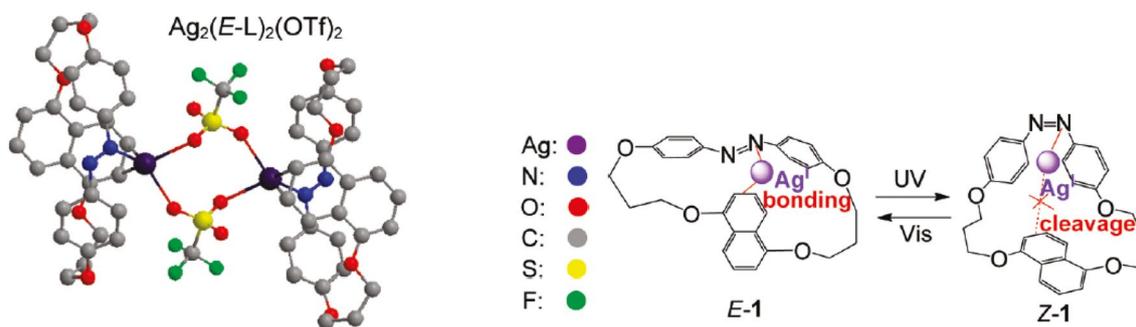


Fig. 19 Left: crystal structure of dimeric Ag^{I} complex *E*-**28**. Right: schematic illustration of photoresponsive cleavage/binding of cation– π bond. Numbering of compound in the reproduced material

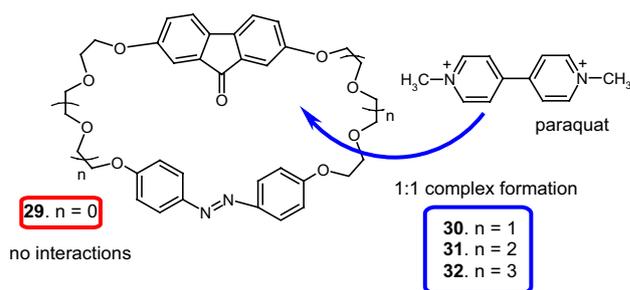


Fig. 20 Crownphanes **29**–**32** bearing 2,7-dioxyfluorenone and 4,4′-azobiphenoxy groups synthesized and studied by Kirichenko and co-workers [147] showing binding ability of dimethylbipyridinium (paraquat) bis(hexafluorophosphate)

Macrocycles **30**–**32** bind 4,4′-dimethylbipyridinium (paraquat) bis(hexafluorophosphate), an electron-deficient model compound. The derivatives of this compound are used in synthetic procedures leading to interpenetrating complexes (pseudorotaxanes). Complex formation of paraquat with macrocycles is based on π – π interactions between π -donor aromatic moieties of cyclophanes and π -acceptor dipyridinium core of the guest. ^1H NMR and MS measurements showed the formation of 1:1 inclusion complexes of pseudorotaxane type. The stability of complexes changes in the order: **31** > **30** > **32**. The smallest macrocycle **29** does not complex the guest due to lack of complementarity between size of the guest and cavity of the host.

Described by Takagi’s and Biernat’s groups 13- and 16-membered crown ethers **22** and **23**, as it was stated earlier, form complexes with metal cations. The X-ray structure of complexes of 13-membered crown with lithium bromide [148] and sodium iodide [149] were described. Metal cation complexes of larger, 16-membered crowns were also obtained. In solid state 16-membered crown forms sodium complex of 1:1 stoichiometry [150] while with potassium salt sandwich type 2:1 (crown:ion) complex [151] is created. In all cases the azo group is in *trans* configuration. It was

corresponds to number of compound in this work: 1 = **28**. Adapted with permission from [146]. Copyright 2010 American Chemical Society. (Color figure online)

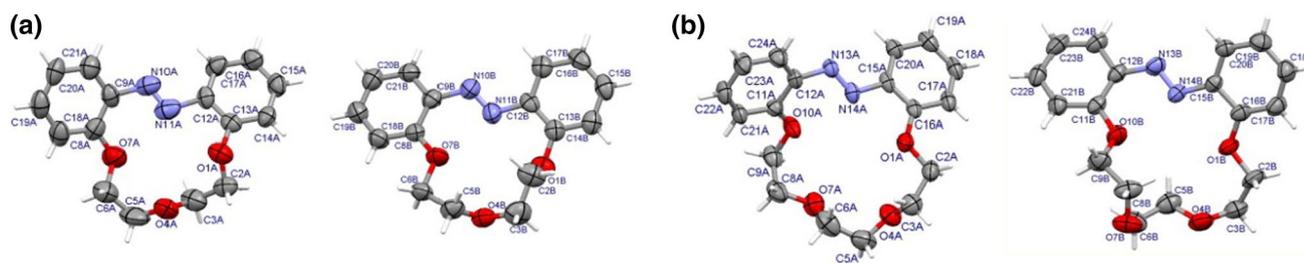
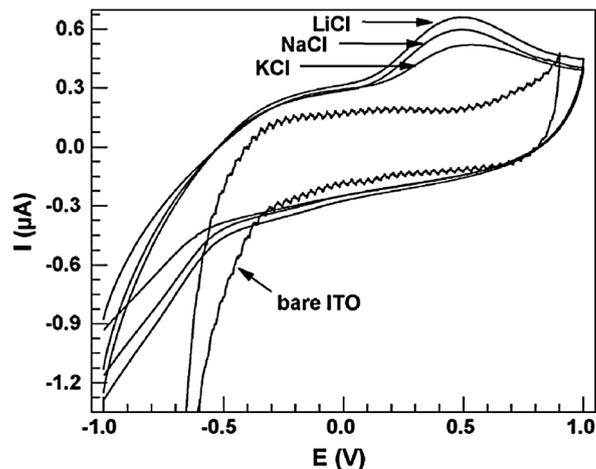
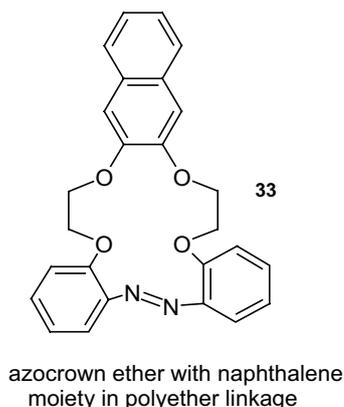


Fig. 21 ORTEP view of **a** crown **22** (molecules **22A** and **22B**), **b** crown **23** (molecules **23A** and **23B**). In both cases the thermal ellipsoids are drawn at the 50% probability level. Reprinted from [153]. Copyright 2008 with permission from Elsevier. (Color figure online)

Fig. 22 16-membered crown **33** and voltammograms for 22-monolayers LB films on ITO, based on this macrocyclic compound, electrode in solutions of KCl, NaCl and LiCl (0.1 M) registered with a sweeping rate of 50 mVs⁻¹. Reprinted from [155]. Copyright 2009 with permission from Elsevier



also shown that the analysis of crystal structures of complexes of azobenzocrown ethers with alkali metal cations can be helpful in interpretation of the selectivity of ion-selective electrodes doped with particular macrocycle [152].

The X-ray structures of uncomplexed *trans* isomers of crown **22** and **23** were also investigated [153]. In the unit cells there are two independent molecules **22A** and **22B** or **23A** and **23B** (Fig. 21).

The kinetics of the buildup and decay of photoinduced birefringence of crown ethers with inherent azo groups **21–24** (Fig. 16a) of different size of the macrocycle was investigated in poly(methyl methacrylate) matrix [154]. For all cases it was found that the kinetics of the buildup of the birefringence was suitably described by a sum of two exponential functions, the time constants (being function of the pumping light characteristic) and sample thickness. The dark decays were described the best by the stretched exponential function, with the characteristic parameters (time constant and stretch coefficient) being practically independent of the type of crown ether. The time constants of the signal decay were orders of magnitude shorter than the respective constants of the dark isomerization of the azo crown ethers. Thus it indicates that the process controlling the decay was a relaxation of the polymer matrix and/or a rearrangement of the flexible parts of the crowns.

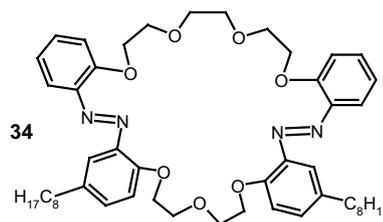


Fig. 23 29-Membered diazocrown **34** showing electrochemical response towards potassium cations [156]

The introduction of the azo group into compounds results not only in photoresponsive but also redox active properties. An example can serve 16-membered crown **33** (Fig. 22) [155] with naphthalene joined by two oxyethylene chains. This macrocycle was used for the preparation of Langmuir–Blodgett (LB) film deposited onto solid ITO substrate. The complexation of metal cations on these electrodes can be successfully observed by cyclic voltammetry (CV). Figure 22 shows CV obtained for a 22-monolayers LB film on an ITO electrode in solutions of KCl, NaCl and LiCl (0.1 M). Bare ITO shows no redox peaks in the presence of K⁺, Na⁺ or Li⁺ ions. For the LB film based on crown **33** film, an electrochemical response in the presence of metal salts was observed. The change of observed signal was attributed

to the specific interactions between the film and the metal ions. The peaks in voltammograms can be ascribed to the electro-reduction of the azo moiety to the hydrazo group, which consumes two electrons and two protons according to the overall reaction. The strongest effect was observed in the presence of lithium cation, showing the possibilities of its electrochemical sensing.

Similar experiments were performed for a number of macrocyclic compounds, e.g. larger 29-membered macrocycle **34** (Fig. 23), bearing two *n*-octyl substituents in benzene rings and two azo groups as a part of macrocycle [156]. Langmuir–Blodgett (LB) and physical vapor deposition (PVD) films on ITO showed electrochemical response towards metal cations. Cyclovoltamperometric curves registered for LB films of 29-membered compound **34** point out that among alkali metal cations Li^+ , Na^+ , K^+ , potassium ion was preferentially complexed under applied conditions suggesting the best host and guest size complementarity.

The selectivity of crown ethers and other host molecules towards metal cation can be controlled also by changing the type of donor atoms. 16- and 18-membered azo- and azoxythiacrown (forming next to azo compounds) ethers **35–40** (Fig. 24, right) were obtained in satisfactory yields by Kertmen and Szczygelska-Tao [157] using reductive macrocyclization procedure. Thiacrowns were tested as ionophores in ion-selective, graphite screen printed electrodes. Opposite to

their oxygen analogs, sulfur containing compounds preferentially supposed to form complexes with softer metal cations. All electrodes doped both with azo- and azoxythiacrowns **35–40** (Fig. 24) showed high sensitivity towards heavy metal cations. The effect of softer sulfur donor atom in the skeleton of macrocycles on the response of ISE with membrane doped with **35–40** can be visualized by comparison of the order of potentiometric selectivity of thia-crown and its oxoanalog [137], shown in Fig. 24 (right, in a frame).

Potassium selectivity of electrodes based on derivatives of 16-membered crown ether **23** was well-proved over years of working with ISEs. 13-membered azobenzocrowns, derivatives of compound **22** (Fig. 16a) are sodium ionophores [134, 136–138, 158–160]. To improve the characteristic of the sodium and potassium sensors, important for clinical analyses, new derivatives of both 13- and 16-membered crowns were prepared and at the same time new technical solutions, including miniaturization of the sensors, were applied. Recently, a series of bis-(azobenzocrown)s (compounds **41–48**, Scheme 4) based on the skeleton of parent 13- and 16-membered crowns **22** and **23** (Fig. 16a) linked by α,ω -dioxalkane chains between two macrocycles have been obtained [162]. Bis-crowns were synthesized from the respective hydroxyazobenzocrowns obtained in reaction analogous to Wallach rearrangement elaborated by Luboch [161].

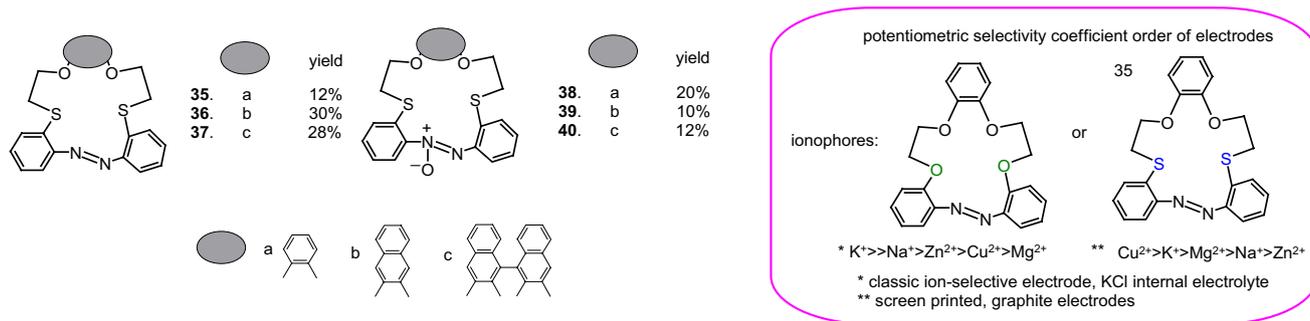
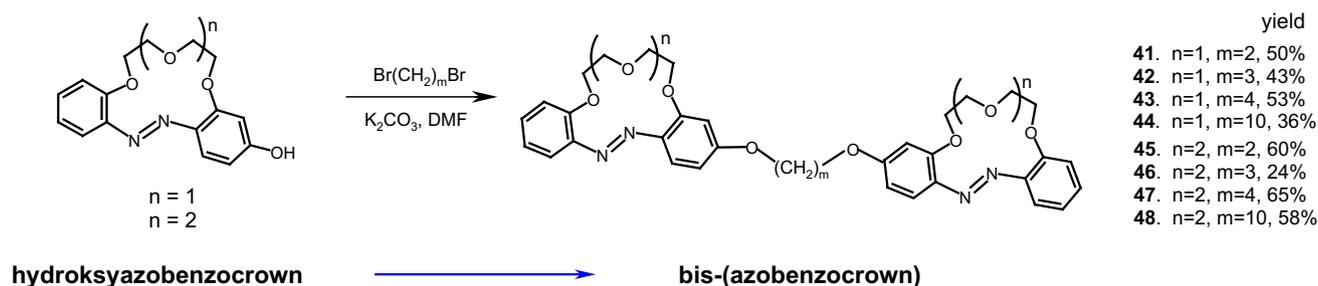
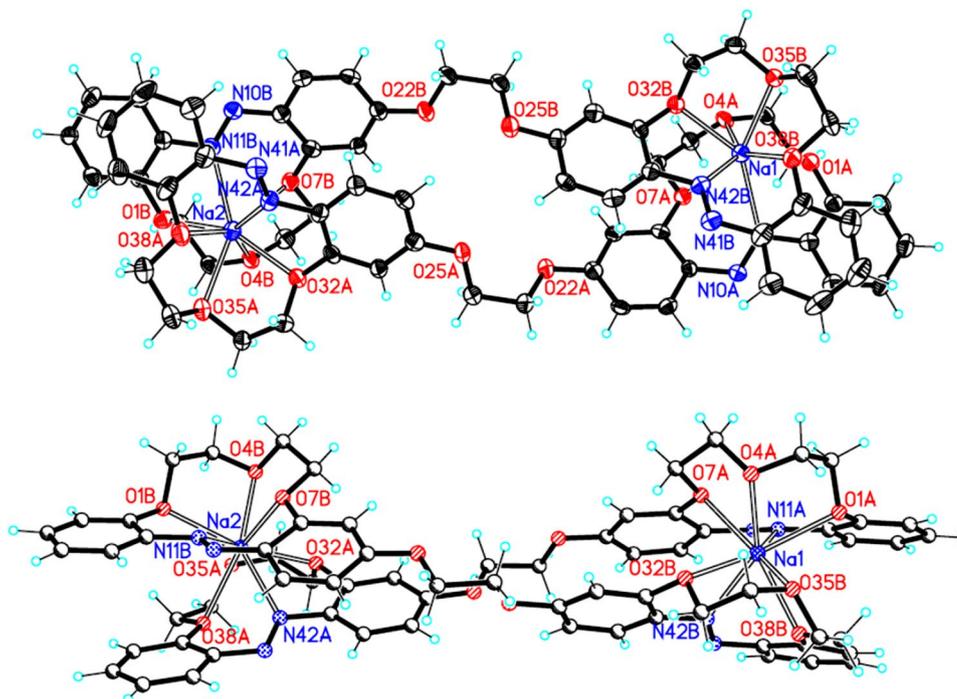


Fig. 24 Left: Thiaazo- (**35–37**) and thiaazoxy (**38–40**) crown ethers obtained by Kertmen and Szczygelska-Tao. Right: comparison of the trend of potentiometric selectivity coefficients of electrodes with crown **36** and its oxygen analog shown is in a frame [157]



Scheme 4 Synthetic route for preparation of bis-(azobenzocrown)s **41–48** from hydroxyazobenzocrowns as substrates [162]

Fig. 25 Two projections of macrocyclic cation $[\text{Na}_2(\text{trans-41})_2]^{2+}$ in **41**-NaI complex with a partial labeling scheme. Reprinted from [162]. Copyright 2012 with permission from Elsevier



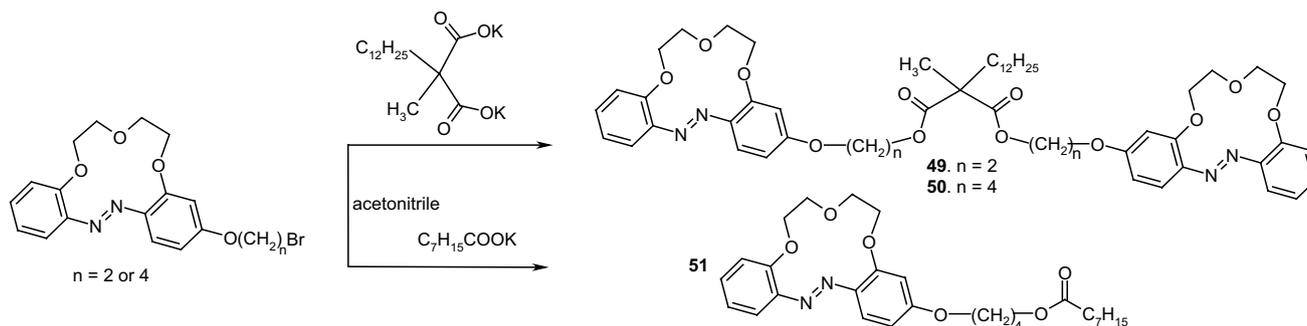
The unique structure of intermolecular of 2:2 stoichiometry sandwich-type complex of bis-(azobenzocrown) **41** with sodium iodide was obtained [162]. It is presented in Fig. 25.

Bis-(azobenzocrown)s **41–48** were used as ionophores both in classic and miniature, all-solid state, screen-printed, graphite ion-selective electrodes. New sodium and potassium sensors feature by short response times, stable potential and high selectivity, in particular high K/Na selectivity.

Bis-(azobenzocrown)s **41–48** form complexes with metal cations also in acetonitrile. The increase of stability constant values comparing analogous monocrown bearing alkoxy substituent proves beneficial effect of the presence of two binding sites in one molecule.

Another example of biscrowns are diester derivatives of dodecylmethylmalonic acid joining two 13-membered azobenzocrown moieties obtained in Luboch group [163] (compounds **49** and **50**, Scheme 5). Biscrowns were obtained using bromoalkoxy derivatives of azobenzocrowns [164] and potassium salt of dodecylmethylmalonic acid in ~40% yield. For comparative studies monoester derivative **51** was synthesized.

For biscrowns **49** and **50** three isomers *trans-trans*, *trans-cis* and *cis-cis* can be considered. From ^1H NMR spectra registered in d-acetone it was found that in solutions of **49** and **50** *trans-trans* and *trans-cis* isomers dominate representing altogether ~90% of the total amount of compounds. The presence of *cis-cis* isomer of **49** was observed upon irradiation with UV light. For monoester



Scheme 5 Synthesis of bis-(azobenzocrown)s **49** and **50**, diesters of dodecylmethylmalonic acid and monoazobenzocrown **51** [163]

derivative **51** the ratio of *trans* to *cis* isomer was evaluated as 6:4. *Trans-trans* and *trans-cis* isomers of **49** and especially of **50**, differ significantly in TLC properties. This can be associated with different complexation properties of both isomers [166]. *Trans* isomers of azobenzocrowns show higher affinity towards metal cations than *cis* forms. Thus *trans-trans* isomer is probable able to form intramolecular sandwich type complexes (Fig. 26) with metal cations whereas for *trans-cis* isomer rather intermolecular complexes are expected. This hypothesis finds confirmation in previously published works of the above authors and in articles published by other groups [149, 166, 167].

Formation of sodium complex by *trans-trans* isomer of **49** was confirmed also by ^1H NMR measurements. Stability constant value of (1:1) complex of **49** in acetone was estimated as $\log K \sim 3.0$ from UV-Vis titrations.

Bis-crowns **49** and **50** based on 13-membered rings, were tested as sodium ionophores in classic and miniature,

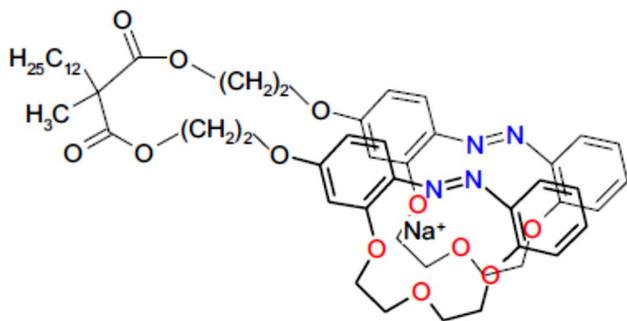


Fig. 26 Proposed organization of bis-crown **49** sodium cation complex. Reprinted without changes from [163]. Copyright 2016 with permission from Springer Publishing Company (<http://creativecommons.org/licenses/by/4.0/>)

solid contact: screen-printed and particularly glassy carbon membrane ion-selective electrodes. Plasticizers 2-nitrophenyl octyl ether (*o*-NPOE) and more lipophilic di(2-ethylhexyl) sebacate (DOS) can be successfully used for bis(azobenzocrown) containing membranes. It was proved that possible isomerization under usual conditions does not significantly affect the characteristics of the prepared electrodes. The influence of UV irradiation on the properties of glassy carbon electrode with ionophore **49** is shown in Fig. 27. After exposition to UV light (1 h, 365 nm), the electrode regains its properties practically after 2 h conditioning in NaCl solution.

Electrodes with the tested biscrowns **49** and **50** were found to have better selectivity coefficients $K_{\text{Na/K}}$ than the electrodes with the monocrown **51**. The best selectivity coefficient Na/K was achieved for the screen printed graphite electrode with the addition of carbon nanotubes into the membrane (**50** as the ionophore, $\log K_{\text{Na,K}} = -2.6$). No significant differences were also observed between the selectivities of the classic and solid contact electrodes. In the last case lower detection limits (LDL) may be obtained. The membrane doped with carbon nanotubes deposited onto graphite screen-printed electrodes results in the better potential stability, detection limit and selectivity of bis-crown-based electrodes. The electro-conductive material was introduced directly into the membrane in a manner analogous to that proposed by Ivaska and co-workers [168]. For glassy carbon electrodes to improve the conductivity, between the membrane and glassy carbon the conductive PEDOT/PSS polymer blend was introduced by electropolymerization. Such electrodes have better (lower) LDL than plain glassy carbon electrode. Electrodes with ionophores **49** and **50** characterize with response times not longer than 10 s, illustrated in Fig. 28 for membrane electrode doped with **49**.

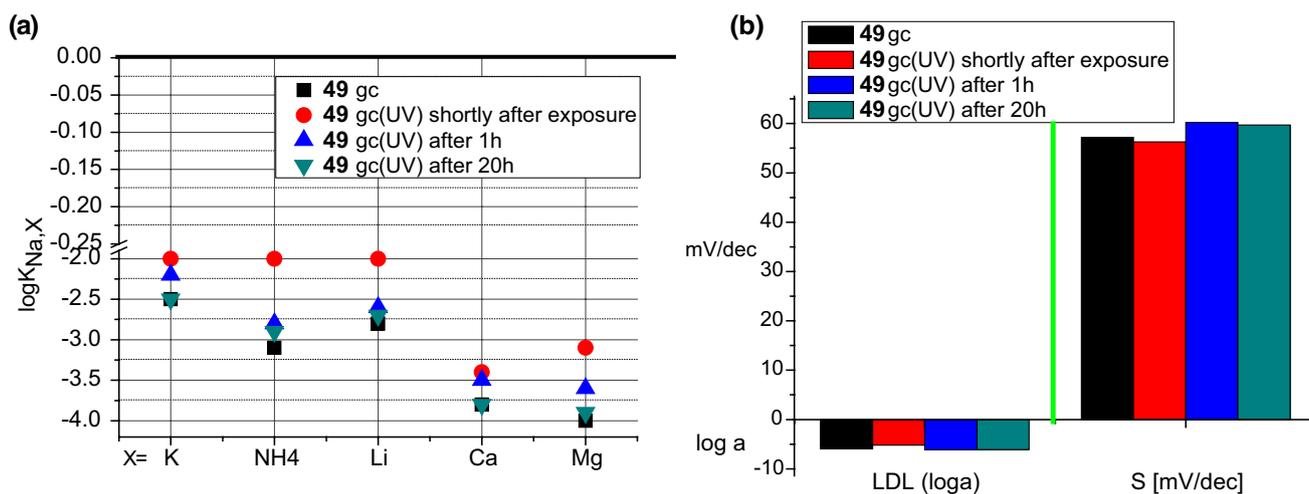


Fig. 27 **a** Selectivity coefficients (SSM, 0.1M) and **b** potentiometric response characteristics: LDL [log a] and slope [mV/dec] for glassy carbon sodium selective electrodes based on **49** as ionophore. UV-electrode upon irradiation with UV light (9W) [163]. (Color figure online)

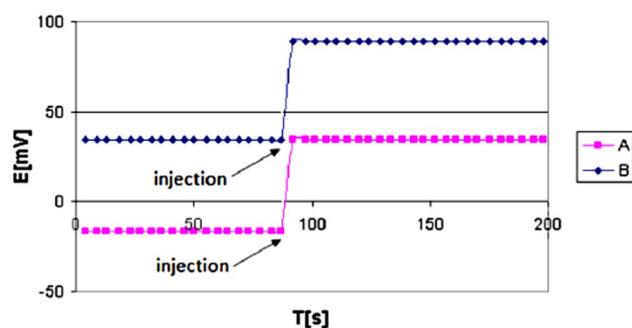


Fig. 28 Response time of glassy carbon electrode with membrane with ionophore **49** (*o*-NPOE as plasticizer) A 0.9 mL of NaCl solution (0.1M) was injected to 100 mL of NaCl solution (10^{-4} M), B 0.9 mL of NaCl solution (1M) was injected to 100 mL of NaCl solution (10^{-3} M). Reprinted without changes from [163]. Copyright 2016 with permission from Springer Publishing Company (<http://creativecommons.org/licenses/by/4.0/>). (Color figure online)

Electrodes based on **49–51** characterize by stable potential in a wide range of pH, depending on the type of the used plasticizer, e.g. electrodes with compound **49** and DOS show stable potential in the pH range 2–10 (0.1M NaCl). Proposed sodium sensor (based on **50**) fulfills requirements for electrodes used in clinical analysis [169]. The response of electrodes based on **50** for sodium in the presence of interfering metal cations corresponding to their blood plasma levels are shown in Fig. 29.

The electrodes were tested for sodium in blood plasma giving consistent results with independent measurements carried out in clinical analytical laboratory.

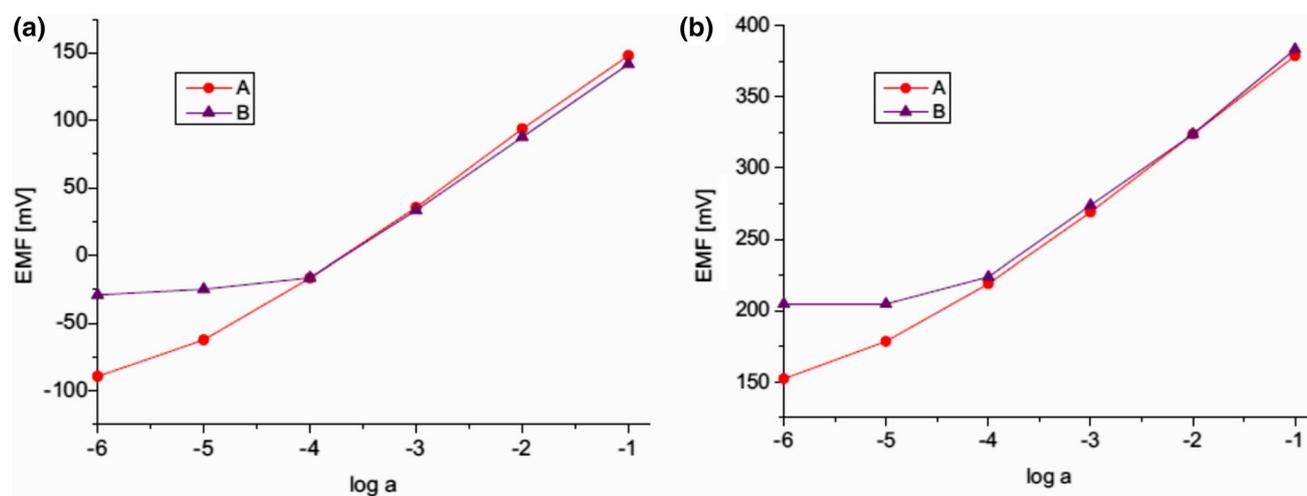


Fig. 29 Response curves for Na^+ obtained with ISEs based on ionophore **50** **a** graphite screen-printed electrode **b** glassy carbon electrode. Curve A indicates the response for Na^+ without and curve B Na^+ in the presence of interfering ions (4.2 mM K^+ , 1.1 mM Ca^{2+} , 0.6 mM

Crown ethers with peripheral azo group

The interactions between photoswitchable azobis-(benzo-18-crown-6) and alkaline earth metal cations were studied by DFT and reactive molecular dynamics (reactive MD) by Pang et al. [170]. Optimized structures of complexes revealed that in the case of Ba^{2+} complex the distance between two cations is the largest among tested complexes in their *trans* form, and the shortest among *cis* complexes. Macrocycles become face-to-face when complexing Ba^{2+} ions. Small energy difference between Ba^{2+} complex in its *trans* and *cis* form indicates facile *cis* to *trans* thermal conversion. Calculation the Ba^{2+} complex allows to conclude that it is a suitable candidate for photocontrolled catalysis.

To mimick the structure and function of biological ion channels the light-regulated transmembrane system was proposed by using tris(macrocycle) system based on diaza-18-crown-6 joined by azobenzene photoswitchable moieties **52** (hydrophile 1, Fig. 30) [171]. The liposome-based ion transport assays revealed that compound **52** displays an efficient transmembrane activity with Y_{max} around 0.7 at 40 $\mu\text{mol/L}$ of **52** in DMSO. Due to the presence of azobenzene moieties the potassium ion transport by the molecule across bilayer membranes can be regulated by applying of external source of light. The photoisomerization of azo groups induces changes of transmembrane length of the ion channel and this way regulating the efficiency of the ion transport.

In many chemical and photochemical processes donor–acceptor complexes (D–A complexes) play an important role. Such systems are also investigated as organic conductors and photoconductors that find applications in nonlinear optics. D–A complexes of a series of bis(crown

Mg^{2+}). Reprinted without changes from [163]. Copyright 2016 with permission from Springer Publishing Company (<http://creativecommons.org/licenses/by/4.0/>)

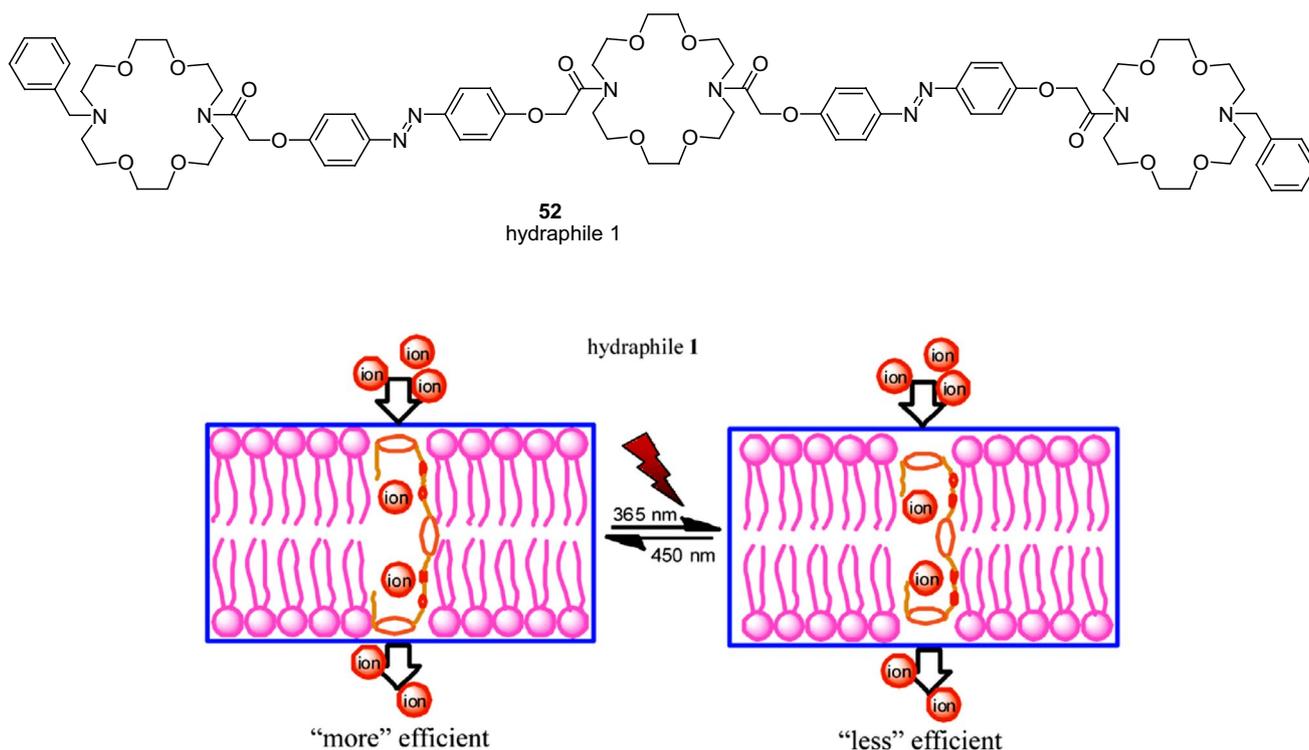


Fig. 30 Tris(macrocycle), amphiphilic azobenzene moiety bearing compound **52** - hydraphile 1 (top) and schematic presentation of transmembrane ion transport in photoswitchable system based on hydraphile. Reprinted from [171]. Copyright 2015 with permission from Elsevier

stilbenes, and also of bis(crown)azobenzene with salts of alkylammonium viologen derivatives were studied in solution and in a solid state by Gromov and co-workers [172]. X-ray structure of complex of bis(18-crown-6)azobenzene **53** (Fig. 31) with viologen derivative **54** showed that the central parts of donor and acceptor molecules feature planar geometry. The proposed systems can be used for the design of optical sensors and molecular devices.

Colorimetric and spectrophotometric ion receptors

Molecular recognition can be utilized in many branches of science and technique if the information about host–guest interaction could be converted into analytically useful signal, e.g. optical or electrochemical. Optical signaling in the visible range of the electromagnetic spectrum draws special attention because it enables non-instrumental sensing of various chemical species such as ions or neutral molecules, e.g. for monitoring of ions of biological or/and environmental importance. The receptor molecule besides binding site

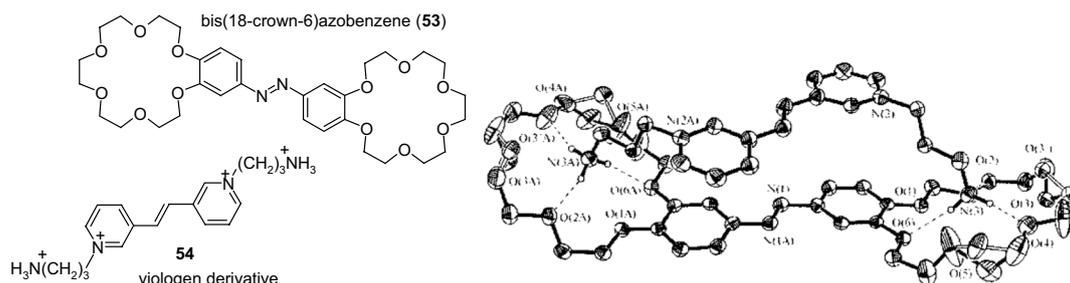


Fig. 31 Structure of bimolecular complex of bis(18-crown-6)azobenzene **53** with viologen derivative **54**. Reproduced from [172]. Copyright 2008 with permission from Springer Publishing Company

should be equipped with additional signaling unit, a functional group joined via linker or chromophoric/fluorophoric moiety forming an integral part of the molecule. Schematically, the idea of chromo- and fluorogenic molecular receptors is shown in Scheme 6. The mechanism of sensing depends on the nature of both the host and the guest. The binding mode, selectivity and sensitivity can be also influenced or controlled by the effect of the solvent and/or receptor immobilization on solid surfaces of various properties.

Inter alia functionalized macrocyclic compounds bearing azo moiety belong to this relatively popular group of sensing materials.

In the case of *para*- and *ortho*-hydroxyderivatives of azo-compounds the color signaling mechanism may be associated with the change in the tautomeric equilibrium upon complexation. This is well illustrated by tautomeric switch based on functionalized azacrown ether **55** (Fig. 32) synthesized and investigated by Antonov and co-workers [173]. Uncomplexed ligand in acetonitrile exists in azophenol form stabilized by intramolecular hydrogen bond between phenolic OH group and nitrogen atom of crown ether residue. In the presence of alkali and alkaline earth metal cations—the color of the solution turns from yellow to orange–red, what is a result of bathochromic and hyperchromic effects

in UV–Vis spectra. The complex formation is connected with the shift of the tautomeric equilibrium towards ketone (quinone-hydrazone) form. Metal cations are complexed by ether oxygen donor atoms and by carbonyl oxygen atom of ketone form.

Lithium and sodium cations form complexes of 1:1 stoichiometry with azacrown **55** (Fig. 33). For magnesium and calcium initially 1:1 complex is formed. Under an excess of a metal salt 2:2 complex dominates. Direct 2:2 complex formation was found for barium perchlorate. Absorption spectra of azacrown registered in the presence of metal perchlorates are shown in Fig. 33a. In Fig. 33b the values of the stability constants of 1:1 and 2:2 metal complexes with discussed azacrown **55** are presented.

Aza-15-crown-5 **56** (Fig. 34) skeleton is a hopeful building block for colorimetric sensors. Lincoln and Sumbly [174] used this macrocycle to synthesize *N*-[4-(phenyldiazo)benzenesulfonyl]-aza-15-crown-5 **57** (Fig. 34). This chromogenic compound was obtained in 55% yield by treating commercially available 4-phenyldiazobenzenesulfonyl chloride with aza-15-crown-5 in DMF in the presence of triethylamine. The synthesized lariat ether was studied as metal cation reagent in ethanol–water (75:25 v/v, pH 6.66) mixture. The stability constant values

Scheme 6 Schematic: the idea of chromo- and fluorogenic molecular receptors

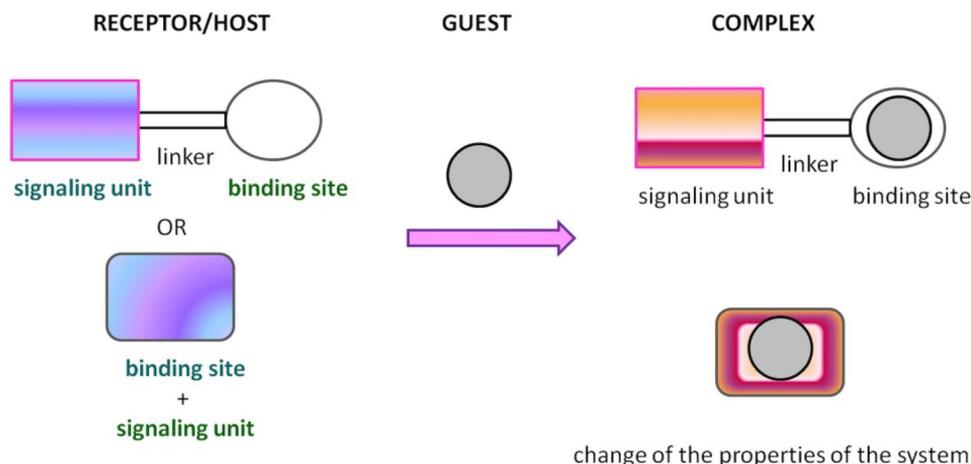


Fig. 32 The mechanism of color change of azacrown ether modified with 4-(phenyldiazenyl)naphthalen-1-ol **55** synthesized by Antonov et al. exemplified by sodium complexation [173]

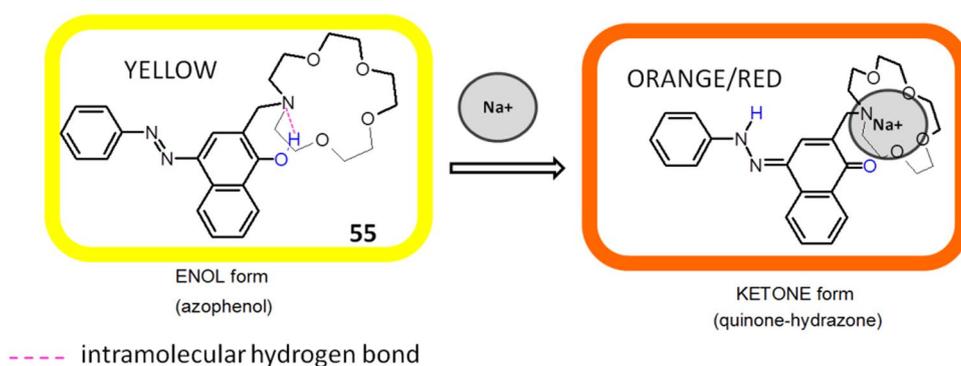


Fig. 33 **a** Normalized absorption spectra of azacrown **55** (—) in CH₃CN and its final complexes with alkali and alkaline earth metal ions. Reprinted from [173]. Copyright 2010 with permission from Elsevier. **b** the values of the stability constants of azacrown with metal cations and the position of the absorption maxima for the respective complexes

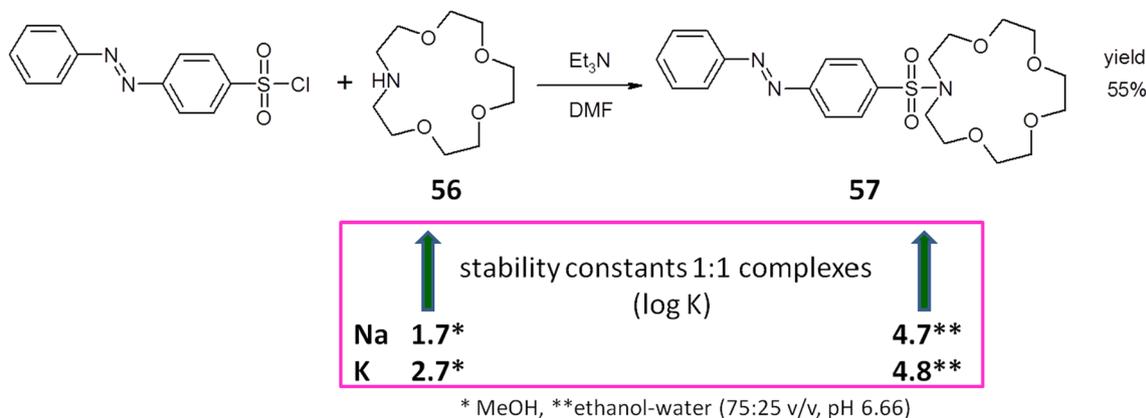
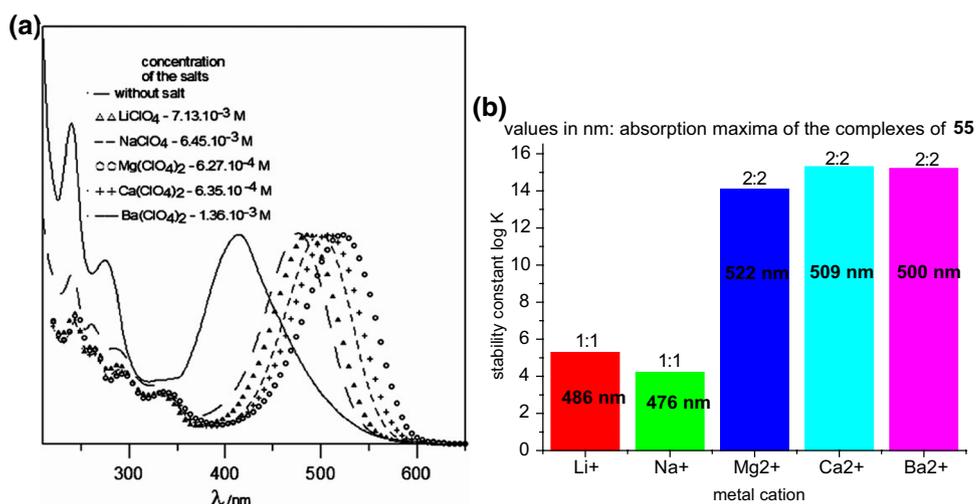


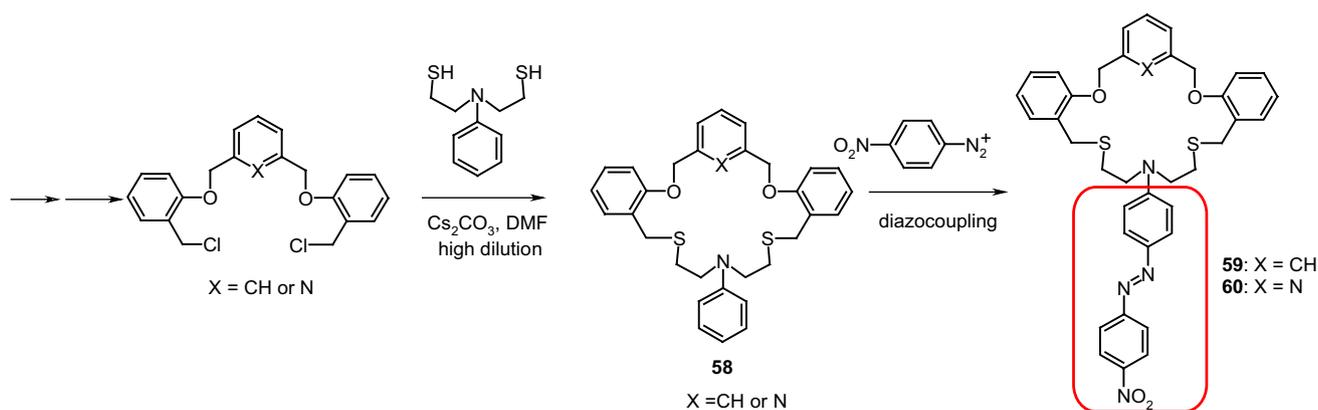
Fig. 34 Lariat ether: *N*-[4-(phenyldiazo)benzenesulfonyl]-aza-15-crown-5 **57**—synthesis and comparison of sodium and potassium complexes stability constant values with parent aza-15-crown-5 **56** [174–178]

of 1:1 complexes of sodium and potassium cations with **57** are higher than for the parent aza-15-crown-5 **56** (Fig. 34) and its derivatives [175–178]. The solved X-ray structure of [Na(**57**)(H₂O)]₂(ClO₄)₂ complex showed that it is a dimer with the sulfonamide oxygen atom engaged in cation complexation. This indicates the cooperation of sulfonamide side arm and crown ether moiety in ion binding and explains the higher values of the stability constant compared with data for unsubstituted aza-15-crown-5.

The selectivity of metal cation binding can be controlled by using macrocycles with softer, sulfur donor atoms. Lee and Lee [179] synthesized, under high dilution conditions, macrocyclic derivatives incorporating aromatic moiety, i.e. benzene **59** or pyridine **60** (Scheme 7) within the macroring. Chromogenic character of macrocycles was achieved by extending the structure by diazocoupling of the obtained in the first step *N*-phenylated macrocycles **58** with *p*-diazonium salt.

Both compounds **59** and **60** selectively bind mercury(II) in acetonitrile forming 1:1 complexes. Complexation of Hg²⁺ causes hypsochromic shift of absorption bands from 480 to 339 and 378 nm for **59** and **60**, respectively. Among other investigated metals only copper(II) cations cause bathochromic shift of absorption band of **59**, whereas spectral behavior of **60** remains intact. Color and spectral changes of **59** and **60** in acetonitrile solutions in the presence of metal salts are shown in Fig. 35. The crystal structure of **60** complex with mercury(II) ion showed metal cation located inside the macrocycle cavity. The difference in selectivity towards mercury ions versus other metal cations was explained by the engagement of the pyridine nitrogen atom in complex formation in case of **60**.

Spectral and color changes in the presence of Hg²⁺ were found to be anion dependent (Fig. 36). Addition of perchlorates or nitrates to the acetonitrile solution of mercury(II) complexes of **59** and **60** causes spectral and color changes,



Scheme 7 Synthetic route for preparation chromogenic macrocycles **59** and **60** [179]

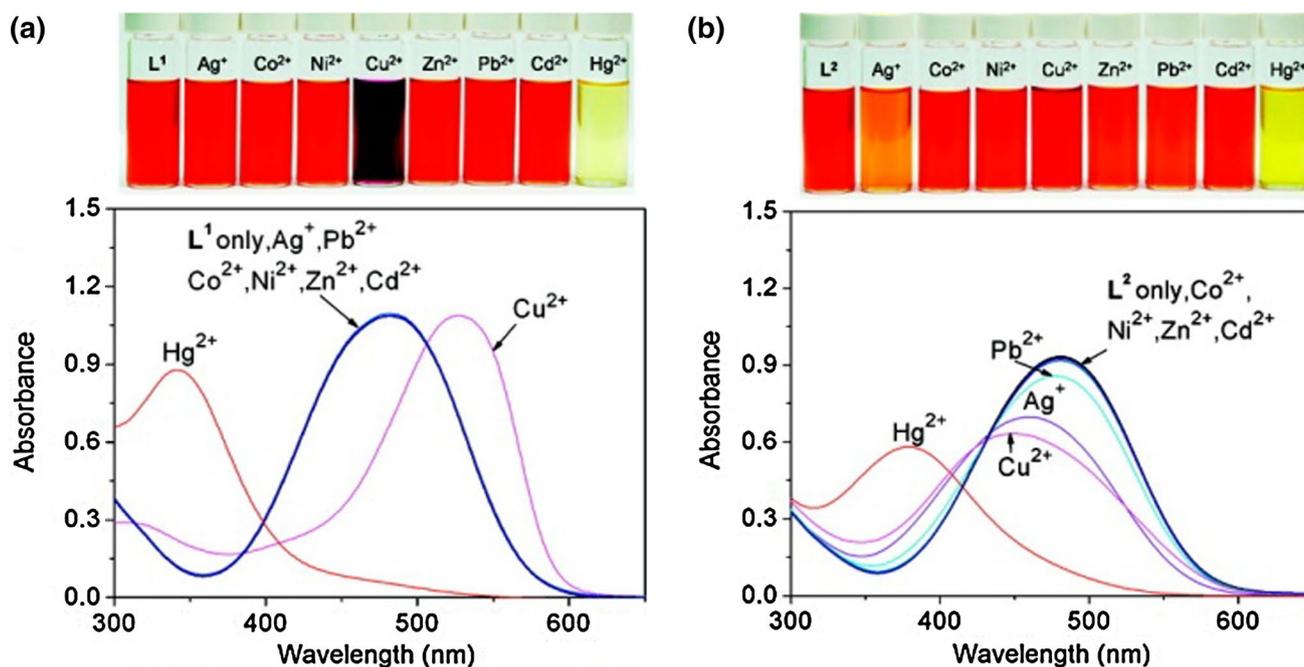


Fig. 35 UV-Vis spectra of **a 59** and **b 60**—(40 μ M) in the presence of metal perchlorates (5.0 equiv) in acetonitrile. Numbers of compounds in reproduced material correspond to following numbers of

compounds in this work: $L^1 = 59$; $L^2 = 60$. Reprinted with permission from [179]. Copyright 2009 American Chemical Society

which can be attributed to the ability of mercury to coordinate these anions. The obtained results indicate that proposed macrocycles can be used not only as mercury, but also as anion sensing molecules.

Spectral and color changes caused by complexation of heavy metal cations were also found for macrocycle **61** bearing as chromogenic substituent *p*-nitroazobenzene [180] that was obtained in multistep reaction shown in Scheme 8.

Red acetonitrile solution of **61** changes color to yellow upon addition of metal salt, which is a result of metal cation induced hypsochromic shift of absorption band. The largest spectral and color changes among investigated metal cations

causes copper(II) ($\Delta\lambda_{\max} = 174$ nm). The spectral and color changes of **61** in the presence of metal nitrates are shown in Fig. 37.

Compound **61** forms two types of solid state complexes, which differ in color: $[\text{Cu}(\mathbf{61})\text{NO}_3]\text{NO}_3 \cdot \text{CH}_2\text{Cl}_2$, a pale-yellow and dark red $\{[\text{Cu}(\mathbf{61})]2(\mu\text{OH})_2\}(\text{ClO}_4)_2 \cdot 2\text{CH}_2\text{Cl}_2 \cdot 2\text{H}_2\text{O}$. The effect of counter ion on spectral changes upon copper(II) complexation was investigated using chloride, nitrate, perchlorate, acetate, and sulfate salts. A blue shift was observed and the influence of anion can be set in the following order: NO_3^- , $\text{ClO}_4^- > \text{Cl}^-$,

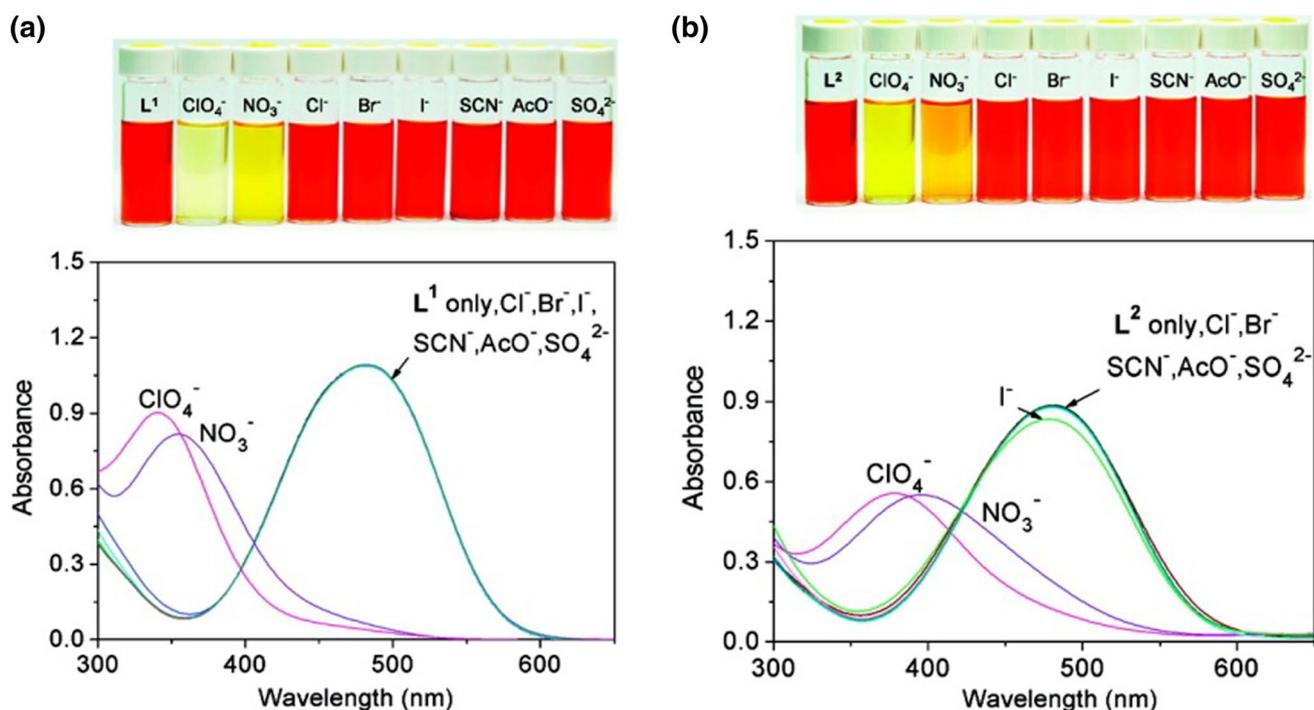
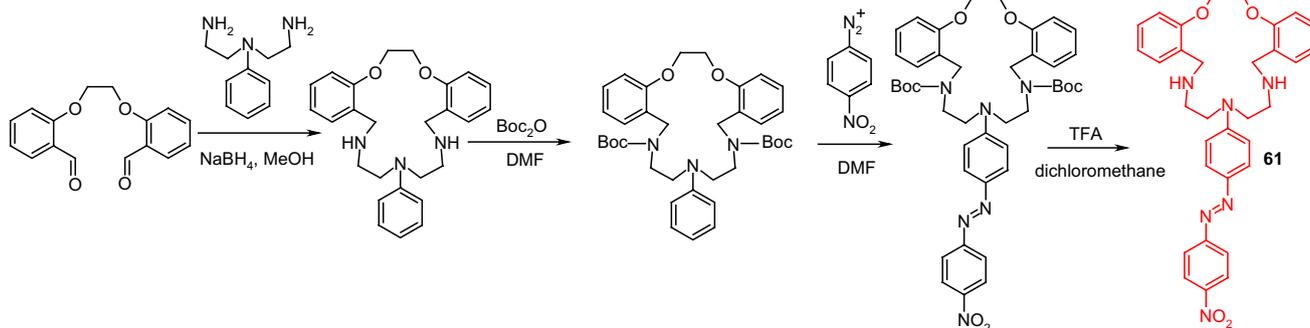


Fig. 36 UV-Vis spectra of **a** **59** and **b** **60** - ($40 \mu\text{M}$) in the presence of Hg^{2+} (5.0 equiv) upon addition of anion salts in acetonitrile. Numbers of compounds in reproduced material correspond to following

numbers of compounds in this work: $\text{L}^1 = \mathbf{59}$; $\text{L}^2 = \mathbf{60}$. Reprinted with permission from [179]. Copyright 2009 American Chemical Society



Scheme 8 Synthesis of chromogenic macrocycle **61** [180]

$\text{AcO}^- > \text{SO}_4^{2-}$, which is in accordance with the Hofmeister series of relative anion lipophilicities.

Colored systems can be also used for preparation of sensing materials by immobilization of the respective receptor(s) on a chosen solid surface. For example, macrocycle **62** (Scheme 9) bearing azo unit, was immobilized on a silica nanotubes (SNT) using sol-gel method [181]. The described system (SNT-**62**) was presented as a heterogenous “naked-eye” and spectrophotometric metal cation sensor.

Inorganic-organic nanomaterial (SNT-**62**) shows in water selective response by color change from yellow to violet towards Hg^{2+} among all other investigated metal cations:

Ag^+ , Co^{2+} , Cd^{2+} , Pb^{2+} , Zn^{2+} , Fe^{3+} , Cu^{2+} . The color of suspensions also changes in the presence of nitrate and perchlorate anions from yellow to pink and violet, respectively. The addition of Cl^- , Br^- , I^- , SCN^- , or SO_4^{2-} salts does not cause color change. It was also shown that modified silica nanotubes can act not only as colorimetric sensor for mercury(II) cation, but also for preparation of stationary phases for ion chromatography. The use of suspensions can be sometimes troublesome, thus a portable chemosensor kit was prepared by modification of the glass surface with SNT-**62**. The material also in this form exhibits selective response towards Hg^{2+} with color change from yellow to violet upon dipping

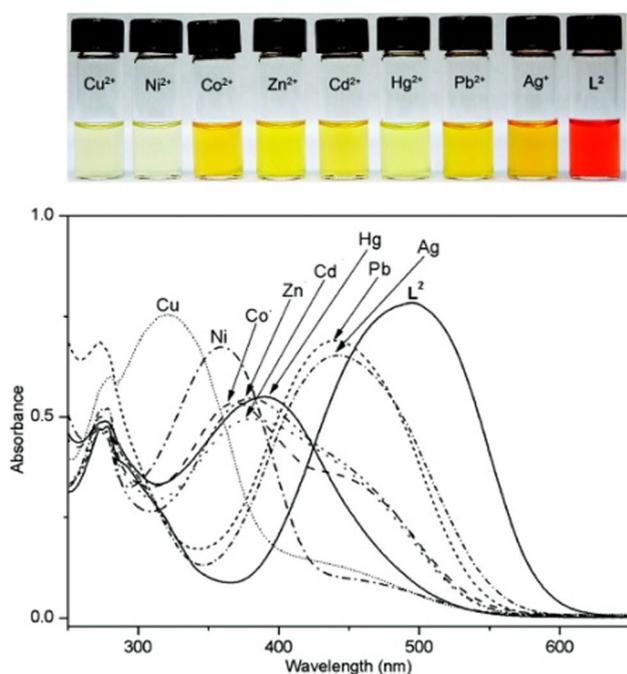
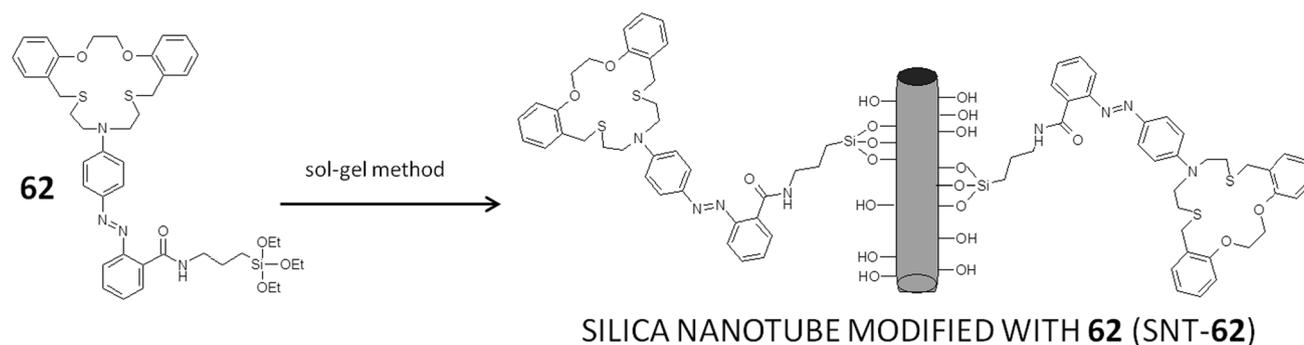


Fig. 37 Changes in the UV–Vis spectrum of **61** on addition of metal nitrates in acetonitrile (ligand concentration, 5.0×10^{-5} M; and added metal ion, 3.0 equiv). Number of compound in reproduced material corresponds to following number of compound in this work: $L^2 = 61$. Reprinted with permission from [180]. Copyright 2009 American Chemical Society

in solution of mercury(II) salt. Color changes of the water suspensions of SNT-**62** upon addition of mercury(II) nitrate at different concentrations are shown in Fig. 38 (left). The color change of the glass sensor modified with SNT-**62** upon immersion into mercury(II) and for comparison copper(II) aqueous solutions is shown in Fig. 38 (right).

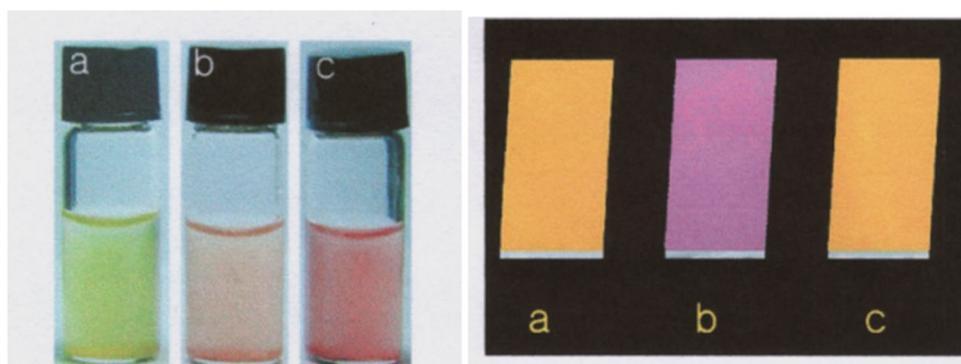
Environmentally hazardous mercury(II) sensing based on dithiaazadioxo crown ether system with peripheral azo unit was described by Ha and co-workers [182]. Compounds **63** and **64** (Fig. 39) were investigated as Hg^{2+} receptors in solvents of diverse polarity (acetonitrile, its mixture with water and in chloroform). It was found that host–guest interaction strongly depends on the solvent nature. According to 1H NMR and spectrophotometric measurements it was stated that both ligands in acetonitrile form 1:1 complexes, if Hg^{2+} is coordinated inside the macrocyclic cavity (Fig. 39). As a consequence of molecular recognition solutions of both ligands undergo discoloration in the presence of Hg^{2+} ions. In less polar chloroform, different mechanism of ligand-ion interaction was proposed. Two molecules of **63** probably bind one mercury(II) cation forming sandwich complex. This is manifested by color change from yellow to pink. In the case of macrocycle **64** in chloroform complexes of 2:2 stoichiometry are formed.

Ha and Jeon continued the work on selective mercury(II) sensing using compound **63** (Fig. 39) [183]. The colored



Scheme 9 Schematically: modification of silica nanotubes with chromogenic macrocyclic derivative **62** (SNT-**62**) [181]

Fig. 38 Left: pictures of the suspensions: **a** SNT-**62**, **b** SNT-**62** + 0.01 mM $Hg(NO_3)_2$, **c** SNT-**62** + 1.0 mM $Hg(NO_3)_2$. Right color changes of glass plates coated with SNT-**62**: **a** before immersion and after immersion in **b** Hg^{2+} (0.01 mM) and **c** Cu^{2+} (0.01 mM) solution in water. Reprinted from [181]. Copyright 2007 with permission from John Wiley and Sons



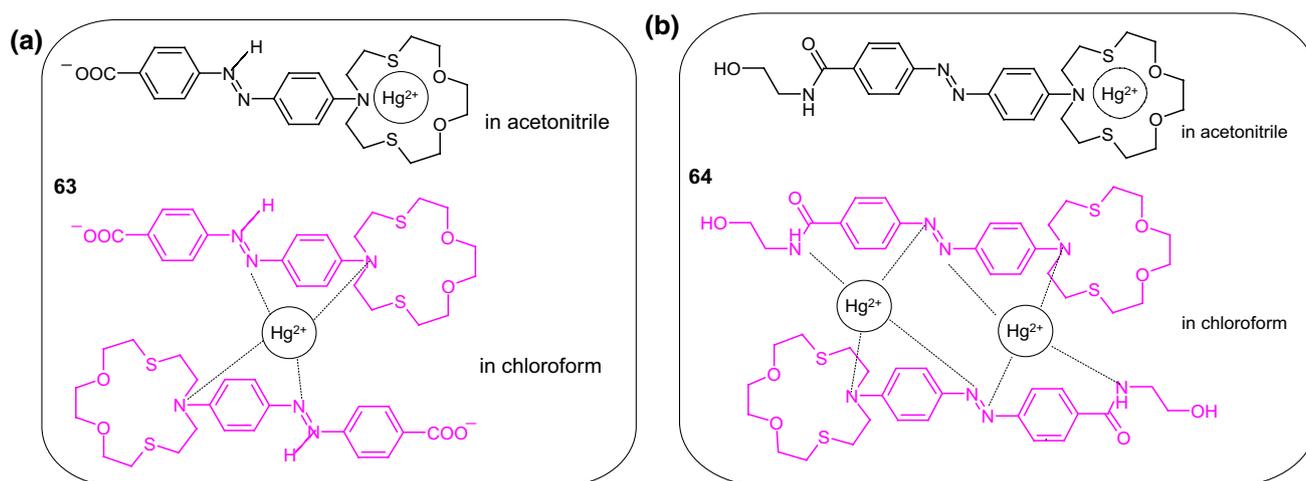


Fig. 39 Proposed mechanism of mercury(II) complexation by macrocycles **63** and **64** depending on the solvent type [182]

macrocycle was applied for recognition of Hg^{2+} ions in aqueous solution. The effect of two surfactants cetyltrimethylammonium bromide (CTAB) and sodium dodecylsulfate (SDS) on spectral and color behavior of **63** was investigated. In the presence of CTAB the solution of **63** is yellow, while pink color is observed in the presence of SDS (Fig. 40). In the presence of Hg^{2+} the pink solution of **63**-SDS system becomes colorless enabling naked-eye ion recognition with detection limit 1.6 μM . The **63**-SDS based system was also used for preparation of the mercury sensitive cellulose test strips.

Functionalized azobenzocrowns (azo moiety as a part of the macrocycle)

This chapter highlights the preparation and properties of azobenzocrowns, of different size of the macrocycle, equipped with additional functional groups in benzene rings: hydroxyl, amino, and dimethylamino, as well as *pull-push* type azobenzocrowns with nitro and dimethylamino groups.

13-, 16- and 19-membered crowns bearing electron donating (dimethylamino) **65–67** or two different electron donating/accepting groups (dimethylamino and nitro) **68–70**

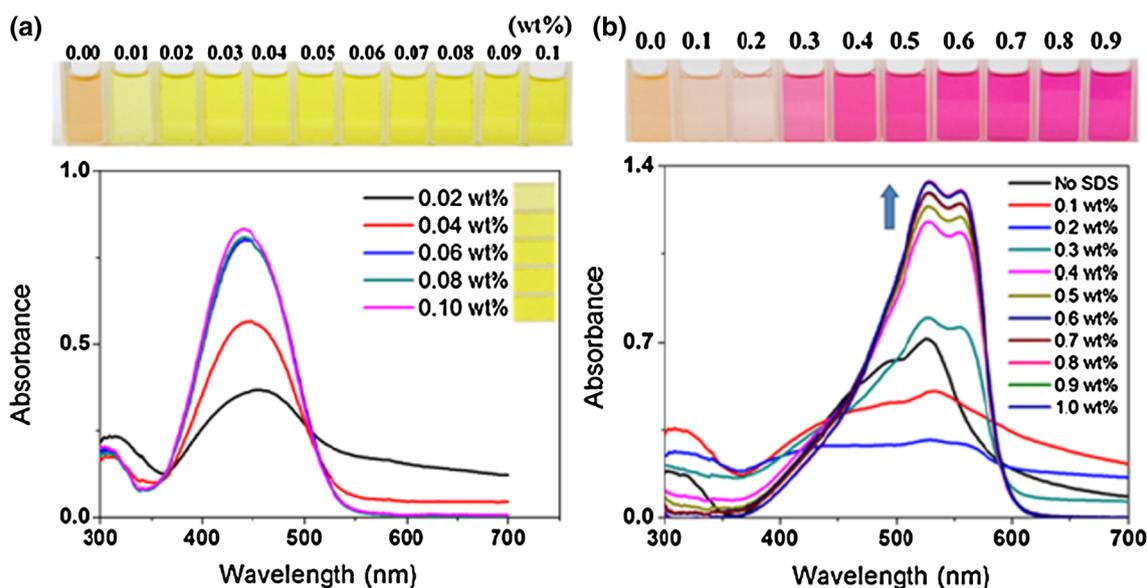
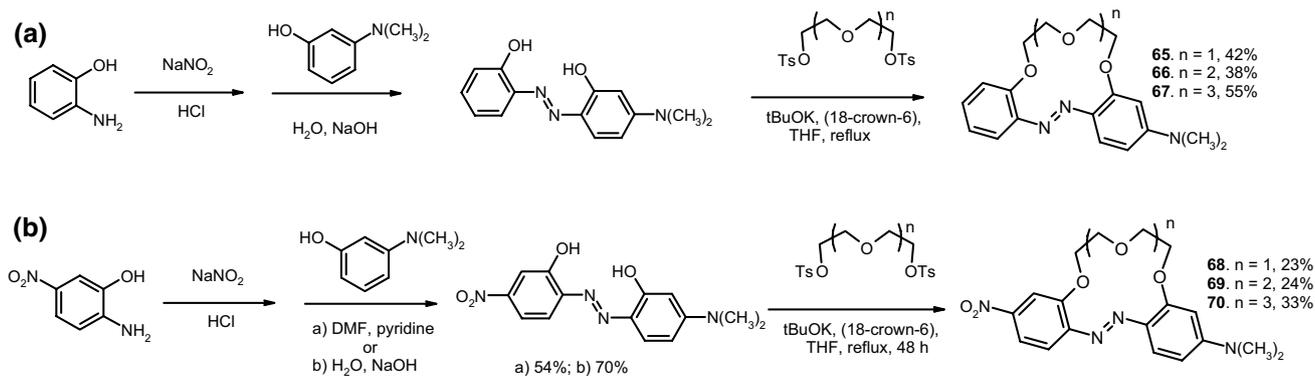


Fig. 40 Color and spectral changes of **63** in various concentrations of **a** CTAB and **b** SDS. Reprinted from [183] Copyright 2015 with permission from John Wiley and Sons. (Color figure online)



Scheme 10 Synthesis of functionalized azobenzocrowns **65–70** [159]

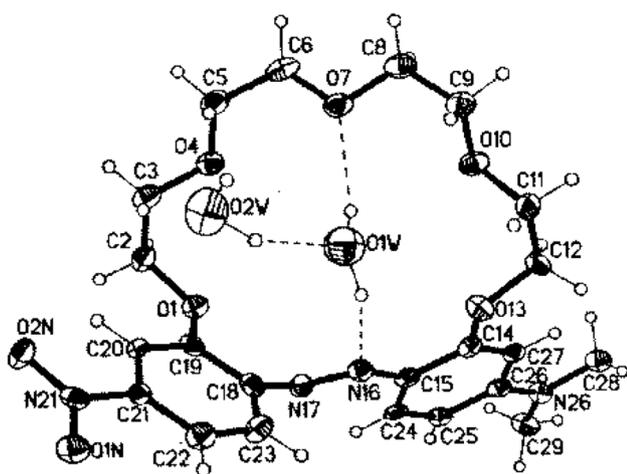


Fig. 41 Molecular structure of **70·2H₂O** with atom labeling scheme; ellipsoids are drawn at 50% probability level. Reprinted from [159]. Copyright 2005 with permission from Elsevier

in the azobenzene fragment (Scheme 10) were obtained with the yields up to 55% [159].

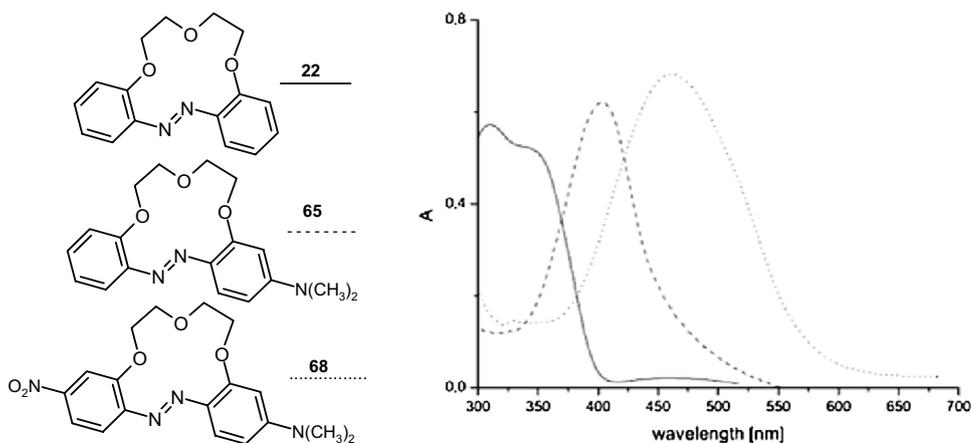
Compounds **65–70** exist only in *E* form, both in the solid state and in solution. In Fig. 41 X-ray structure of **70·2H₂O** is presented, showing the *E* geometry of the azo unit with aromatic moieties in the *trans*-positions and proved the existence of a molecular diaqua-complex [159].

The absorption spectra of **65–70**, opposite to the parent azobenzocrowns **22–24** (Fig. 16a) have sharp and well pronounced maxima (Fig. 42).

The UV–Vis studies of alkali and alkaline metal cation complexation by compounds **65–70** showed magnesium selectivity of 19-membered azobenzocrown **67** in acetonitrile. Only in this case the complexation is characterized by significant spectral shift (Fig. 43) and by distinctive color change from orange to pink.

Another set of synthesized and investigated functionalized azobenzocrowns consists of derivatives with a hydroxyl substituent. Azobenzocrowns with hydroxyl group located in one of the benzene rings, in the *para* position to the azo group, have been synthesized prior to 2002 [158] and are also a part of current works carried in Luboch's group.

Fig. 42 Comparison of absorption spectra of azobenzocrowns: parent—**22** (solid line), and functionalized with: dimethylamino—**65** (dashed line) and dimethylamino- and nitro—**68** (dotted line); ($c = 7.0 \times 10^{-5}$ M) in acetonitrile. Spectra reproduced from [159]. Copyright 2005 with permission from Elsevier



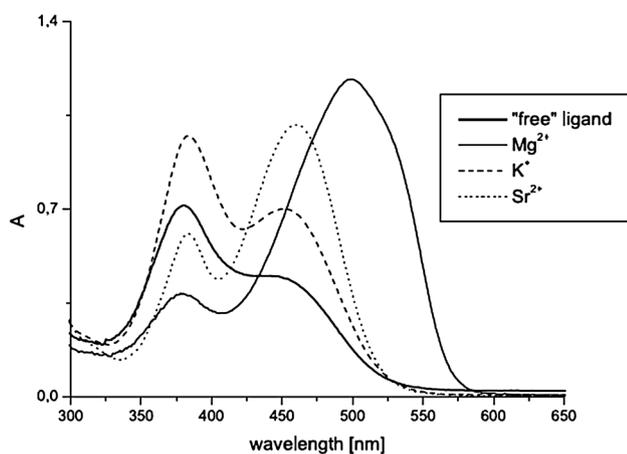


Fig. 43 The comparison of absorption spectra of **67** (solid line, $c=5.7 \times 10^{-5}$ M) and limiting spectra in the presence of: potassium (dashed line, $c=2.6 \times 10^{-3}$ M), strontium (dotted line, $c=4.4 \times 10^{-5}$ M) and magnesium (solid line with maximum at ~ 520 nm, $c=1.9 \times 10^{-3}$ M) perchlorates in acetonitrile. Reprinted from [159]. Copyright 2005 with permission from Elsevier

A simple method for the synthesis of 13- and 16-membered azobenzocrown ethers, derivatives 4-hexylresorcinol **71–73** with two peripheral groups, i.e. nitro and hydroxyl groups at two opposite sides of the conjugated chromophoric system has been described by Luboch et al. (Scheme 11) [160].

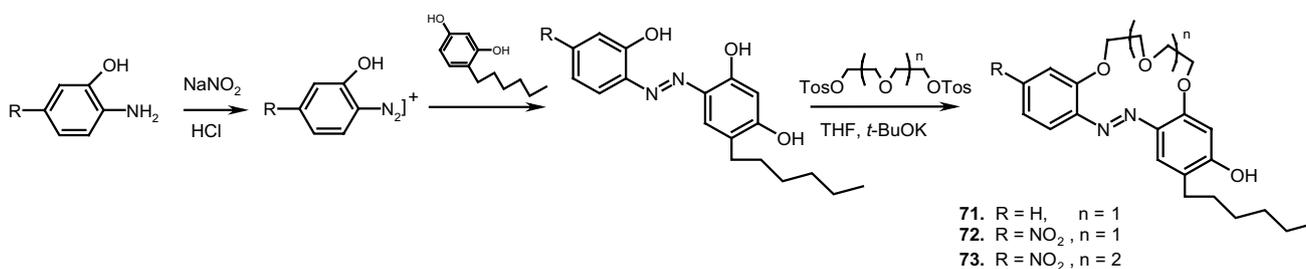
Typical for 13-membered azobenzocrowns, including compounds **71** and **72** is selective binding of lithium cations. The most significant, among all investigated so far compounds of this type, is the spectral shift of 95 nm and

color change from yellow to pink found for **72** (Fig. 44) in basic acetonitrile (Et_3N) solution. Crown **73** is more lithium sensitive, but less selective, versus sodium and potassium (Fig. 44) giving the color change. The chromoionophoric behavior of the compounds potentially allows their application, under selected conditions, for construction of optical sensors.

As shown in Scheme 12 the hydroxyazobenzocrowns undergo tautomeric equilibrium to quinone-hydrazone. The tautomeric equilibrium of hydroxyazobenzocrowns is affected by the size of the macrocycle. The larger the cavity size the lower the tendency to occur in the quinone-hydrazone form. This can be explained by weaker hydrogen bonds in macrocycles of larger cavity. 13-membered macrocyclic *p*-hydroxyazobenzene derivative—compound **74** (Scheme 11), in the solid state and in solvents of different polarity (chloroform, acetonitrile, acetone or methanol) exists in the quinone-hydrazone form. The azophenol form was observed ($\sim 30\%$) in DMSO.

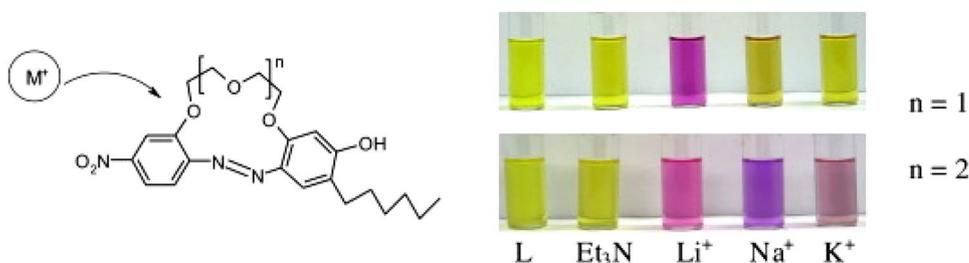
16-Membered crown **75** in chloroform and in acetonitrile exists, like **74**, in the quinone-hydrazone form, but in DMSO only the azophenol form was found. Compound **76** of 19-membered ring entirely exists in the azophenol form in DMSO and chloroform. In acetonitrile no less than 75% of this form was detected [159], but in acetone both forms exist in comparable amounts.

p-Hydroxyazobenzocrown ethers can be obtained from *O*-protected podands by reduction [158] or directly from dihydroxyazocompounds as shown for sterically hindered crowns **71–73** [160]. The reaction analogous to the Wallach rearrangement was proposed as a method for preparation of



Scheme 11 The synthesis of azobenzocrown ethers with peripheral hydroxyl group **71–73**—derivatives of 4-hexylresorcinol [160]

Fig. 44 Color changes of azobenzocrown solutions with peripheral hydroxyl group **72** ($n=1$) and **73** ($n=2$) in the presence of metal perchlorates in acetonitrile. Reprinted from [160]. Copyright 2009 with permission from Elsevier



p-hydroxyazobenzocrowns using azoxybenzocrowns as substrates [161]. However, the reaction carried out in the mixture of concentrated sulfuric acid and ethanol suffers from the formation of side products and large amounts of used reagents [162]. Exhaustive synthetic research on the applicability of Wallach rearrangement allowed to conclude that the decrease of the side-products formation, lower amounts of reagents and finally, the most importantly, significant yield increase is obtained by carrying out the Wallach rearrangement in a mixture of concentrated sulfuric acid and dimethylformamide [165]. Under elaborated reaction and isolation conditions a series of hydroxyazobenzocrowns **74**, **75** and **77–83** were successfully obtained (Scheme 13).

In contrast to the Wallach rearrangement conducted under strongly acidic conditions where mostly *p*-hydroxyazo compounds are formed, the photochemical rearrangement leads also to *ortho*-hydroxyazo compounds **84**, **85** (Scheme 14) [162]. Under fixed conditions the ratio of *para* to *ortho* hydroxyazobenzocrown isomers was dependent on the

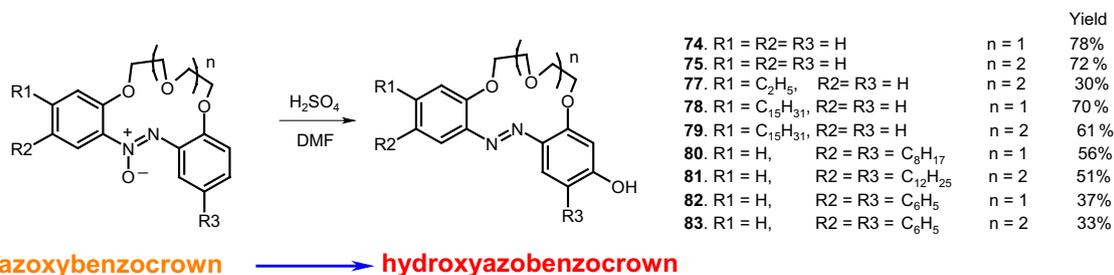
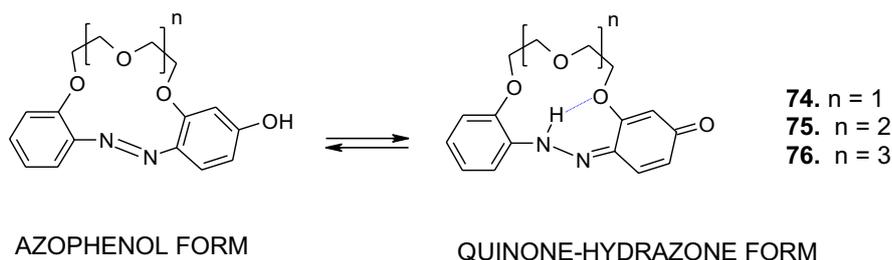
solvent. In toluene *o*-substituted compounds were dominating, *p*-substituted crowns were the main product in ethanol, whereas in DMF a mixture of comparable amounts of both isomers were obtained.

o-Hydroxyazobenzocrowns opposite to *p*-substituted analogs, exist mainly in azophenol form.

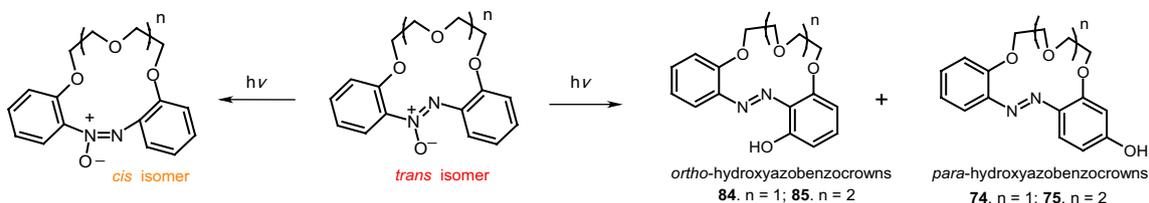
The spectral properties of *o*-hydroxyazobenzocrowns **84** and **85** were compared with **74** and **75**, and unsubstituted crowns **22** and **23** (Fig. 16a) [165]. Their normalized UV–Vis spectra (acetonitrile) (solid lines) and the corresponding protonated forms (dashed lines) are shown in Fig. 45a, b. Protonation constants in acetonitrile solutions are compared in Fig. 45 (right). The protonation constants can be ordered: *p*-hydroxyazobenzocrowns > *o*-hydroxyazobenzocrowns > unsubstituted azobenzocrowns ~ acyclic analog of azobenzocrowns **86**.

p-Hydroxyazobenzocrowns were used as substrates in the synthesis of bisazobenzocrowns (Scheme 15) of different lipophilicity, where two macrocyclic residues are

Scheme 12 Tautomeric equilibrium for hydroxyazobenzocrown ethers **74–76** showing hydrogen bond inside the cavity of quinone-hydrazone form [158]



Scheme 13 Rearrangement of azoxybenzocrowns in the presence of concentrated sulfuric acid and DMF [165]



Scheme 14 Azoxybenzocrowns: *trans*–*cis* photoisomerisation and photochemical rearrangement leading to *ortho*- (**84**, **85**) and *para*-hydroxyazobenzocrowns (**74**, **75**) [165]

joined via dioxymethylene group. Biscrowns (**87–93**) were obtained in yields up to 72%.

Bisazobenzocrowns were used as ionophores in classic and miniature (screen-printed) ion-selective electrodes. A selectivity coefficient $\log K_{Na,K} = -2.5$ (SSM, 10^{-1} M) for electrode with crown **87** as ionophore was one of the best result obtained for the whole group of the electrodes based on the 13-membered azobenzocrowns. Within the investigated 16-membered bisazocrowns the best potassium over sodium selectivity coefficient for potassium electrodes was $\log K_{K,Na} = -3.5$ (SSM, 10^{-1} M) found for compound **88**.

13- and 16-Membered azobenzocrowns (Scheme 16) with aromatic amino (**94, 95**), amide (**96, 97**), ether–ester (**98–103**) or ether–amide (**104–107**) residue in *para* position to an azo moiety were synthesized and investigated [184].

The studies of tautomeric equilibrium of aminocrowns **94** and **95** showed that in majority of solvents, similarly to open chain aminoazocompounds [185] they exist in aminoazofrom (Scheme 17). It is opposite to discussed above

hydroxyazobenzocrowns for which tautomeric equilibrium was found to be more solvent dependent [158, 159, 186].

The protonation of aminoazobenzocrowns shifts the tautomeric equilibrium towards protonated iminohydrazone form.

13-Membered crown **94**, as expected, in acetonitrile preferentially complexes lithium ions. Stability constant of this 1:1 complex is $\log K = 4.0$. This value is comparable with the value for unsubstituted **22** ($\log K = 4.1$), but it is higher than for the corresponding 13-membered hydroxyazobenzocrown **74**. The stability constant obtained for magnesium complex, $\log K = 6.43$, is the highest value for magnesium complex among all studied so far azobenzocrowns. Changes in the absorption spectra upon spectrophotometric titration of a solution of **94** with lithium and magnesium perchlorates in acetonitrile are illustrated in Fig. 46a, c. Fig. 46b shows limiting spectra for **94** upon titration with alkaline earth metal perchlorates.

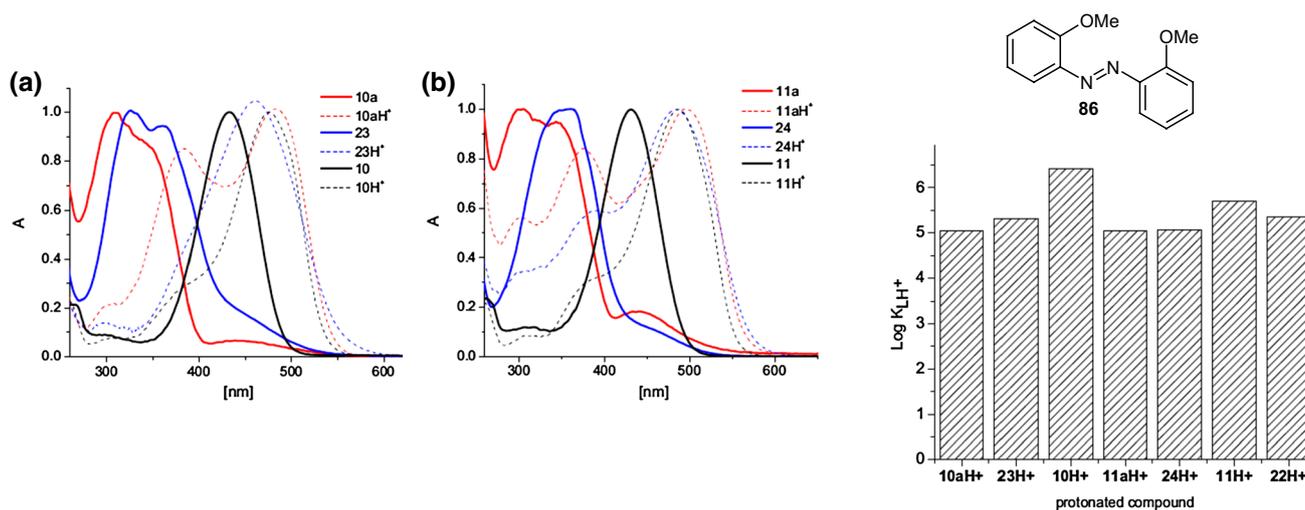
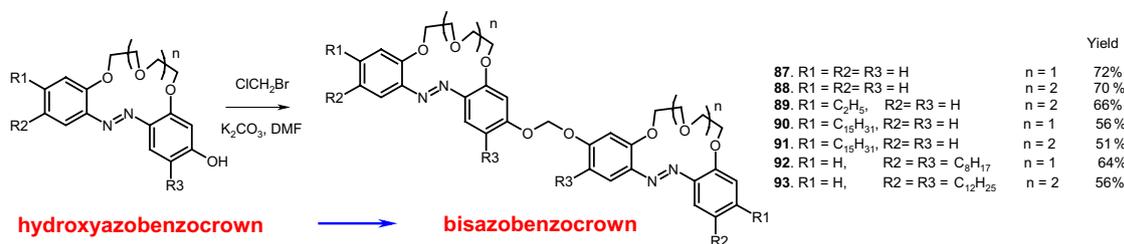
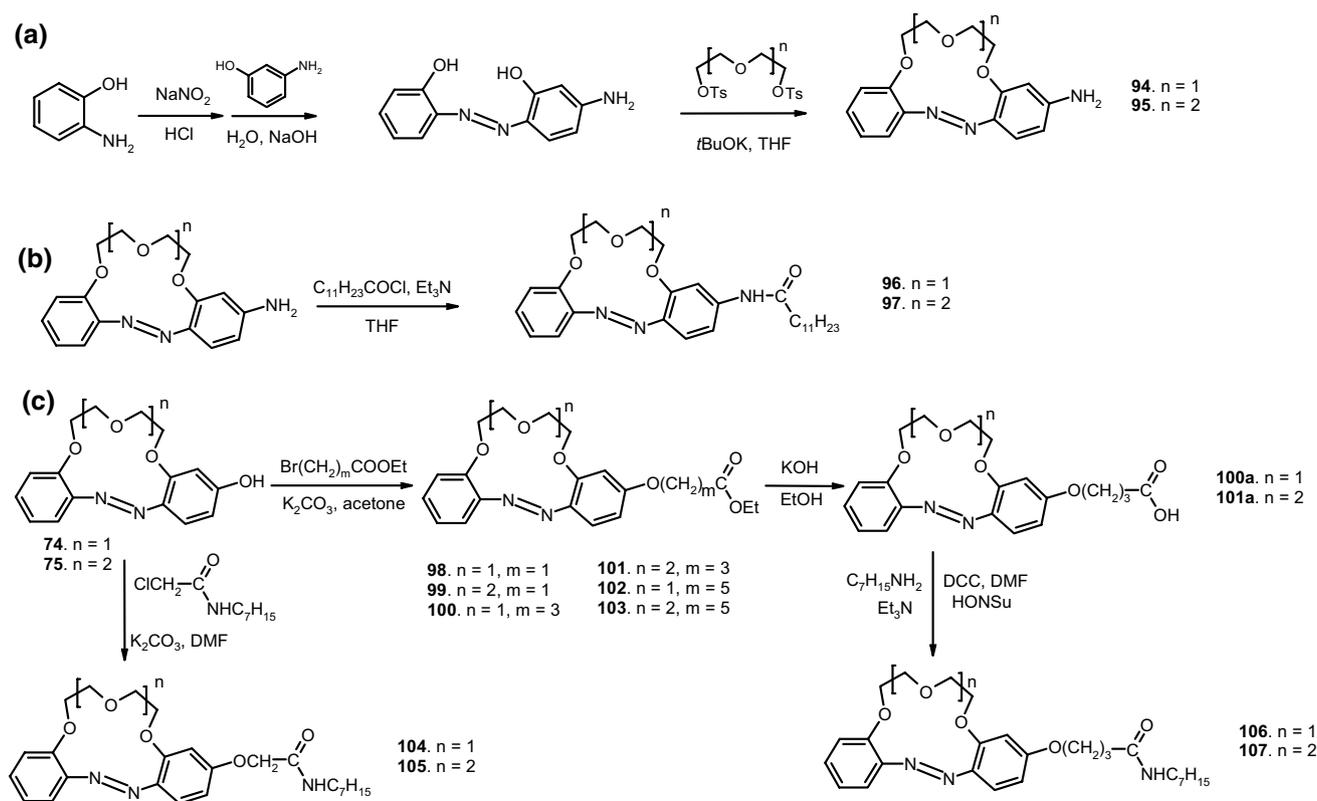


Fig. 45 Left: comparison of normalized UV-Vis spectra of (a) 13-membered (b) 16-membered azobenzocrowns (solid) and their protonated forms (dashed lines) Right: proton binding constants for **22, 84, 74** and **23, 85, 75** azobenzocrowns and for acyclic analog **86** in acetonitrile. Number of compounds in reproduced material cor-

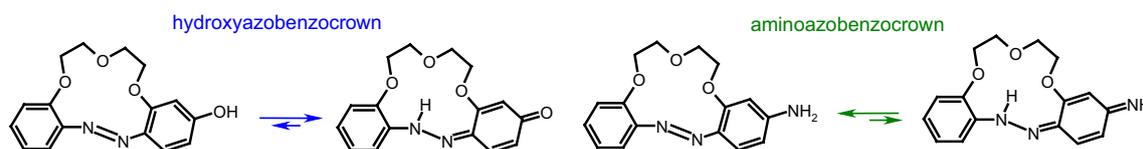
respond to following numbers of compounds in this work: 10=**74**; 10a=**22**; 23=**84**; 11=**75**; 11a=**23**; 24=**85**; 22=**86**. Reprinted from [165]. Copyright 2013 with permission from Elsevier. (Color figure online)



Scheme 15 Synthesis of bisazobenzocrowns **87–93** with dioxymethylene spacer [165]



Scheme 16 Synthesis of **a** amino (**94**, **95**), **b** amide (**96**, **97**), **c** ether–ester (**98–103**) or ether–amide (**104–107**) derivatives of azobenzocrowns [184]



Scheme 17 Comparison of tautomerism of 13-membered hydroxy- and aminoazobenzocrowns [184]

16-Membered aminoazobenzocrown **95** forms 1:1 complexes with alkali and alkaline earth metal cations. In all cases, with exception for potassium, the values of the corresponding stability constants are higher than for parent azobenzocrown **23**, which are in turn higher than for complexes of hydroxyazobenzocrown **75**. The introduction of electron-donor amino group into the benzene ring in *para* position to azo moiety enhanced binding properties of azobenzocrowns.

Lithium binding was investigated for a series of 13-membered azobenzocrown with oxyalkylcarbonester moiety as side chain **98**, **100**, **102** (Scheme 16) and was compared with properties of **22** and its alkoxy derivative **108** (Fig. 47).

The general trend of spectral changes upon lithium complexation for oxyalkylcarbonester derivatives is similar as for **108**. The length of aliphatic acid chain has some effect on the binding strength of the lithium ions, however it cannot be the complexation of the same type as for lariat type crowns. The side chain seems to be too short to participate in complex formation. This is to some extent confirmed by the crystal structure of sodium iodide complex **101** (Fig. 48).

Vast majority of azo compounds, with few exceptions [105, 187–190] show no fluorescence. Protonated azobenzocrowns exhibit orange-red fluorescence. The position of emission band, and the value of the Stoke's shift is dependent on the presence and nature of the substituent in the *para* position to the azo group [184]. Comparison of normalized absorption and the corresponding emission

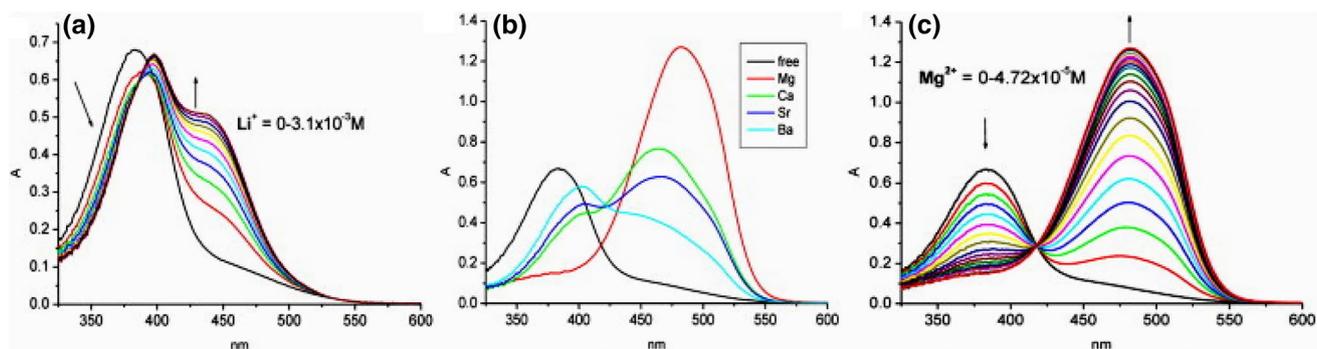


Fig. 46 Changes in absorption spectra upon titration of solution of **94** ($c = 3.27 \times 10^{-5}$ M) with perchlorates: **a** lithium; **c** magnesium; **b** the limiting spectra obtained during spectrophotometric titration of

solution of **94** with alkaline earth metal perchlorates in acetonitrile. Reprinted from [184]. Copyright 2013 with permission from Elsevier. (Color figure online)

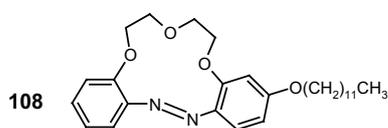


Fig. 47 Alkoxy azobenzocrown **108** [184]

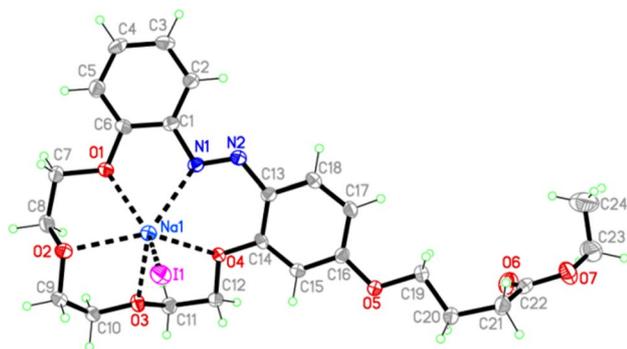


Fig. 48 The crystal structure of sodium iodide complex of compound **101**. Reprinted from [184]. Copyright 2013 with permission from Elsevier

spectra for protonated forms of 13-membered azobenzocrowns **22**, **94** and **96** are shown in the Fig. 49a, b.

Changes in the UV–Vis and emission spectra of compound **96** solution upon titration with solution of *p*-toluenesulfonic acid in acetonitrile are shown in Fig. 50a, b. Photos show a color change and red fluorescence of **96** caused by protonation.

Selected functionalized azobenzocrowns (Scheme 16) were tested as ionophores in ion-selective membrane electrodes. Ether-ester **100–103** and ether-amide **104–107**, similarly to described earlier alkyl and dialkyl derivatives [136, 137, 158] are good ionophores in membrane electrodes. The

mechanism of ion selectivity can be explained by formation of “sandwich” type complexes with the main ions [149, 151].

The ^1H NMR studies of tautomeric equilibrium of hydroxyazobenzocrowns with phenyl substituents in benzene rings (Fig. 51) showed that 13-membered crown **109** in acetonitrile exists in quinone-hydrazone form [191]. 10% of azophenol form was detected in DMSO. For 16-membered crown **110** the presence of quinone-hydrazone form was stated in acetone and chloroform [191] and also in DMSO (50%).

For 13-membered hydroxyazobenzocrown **109** spectral response with spectral shift ~ 40 nm (Fig. 52a) and color change from yellow to orange (Fig. 52, top) caused by the presence of lithium salt was observed only in basic solution (Et_3N) of acetonitrile. This corresponds to lithium complex formation by ionized hydrazone form of **109**. Lithium response is observed also in the presence of the excess of sodium salt (Fig. 52b).

16-Membered crown without phenyl substituents **75** (Scheme 12) forms complexes both in neutral and basic acetonitrile solution. In pure acetonitrile the presence of alkali and alkaline earth metal cations causes a blue spectral shift corresponding to shift of tautomeric equilibrium and formation of complex in azophenol form. Under the same conditions no spectral changes were observed in the presence of lithium and sodium for **110**. In basic acetonitrile (Et_3N) for both **75** and **110** the presence of lithium and sodium salts (no changes in the presence of potassium) causes red shift of absorption band. This corresponds to formation of complexes by quinone-hydrazone forms. The comparison of the stability constants ($\log K$), determined by UV–Vis titrations, of lithium and sodium complexes of **75** in neutral and basic acetonitrile and **110** in basic acetonitrile are shown in Fig. 53. Stability constant values of complexes of **23** in neutral acetonitrile were also included for comparison.

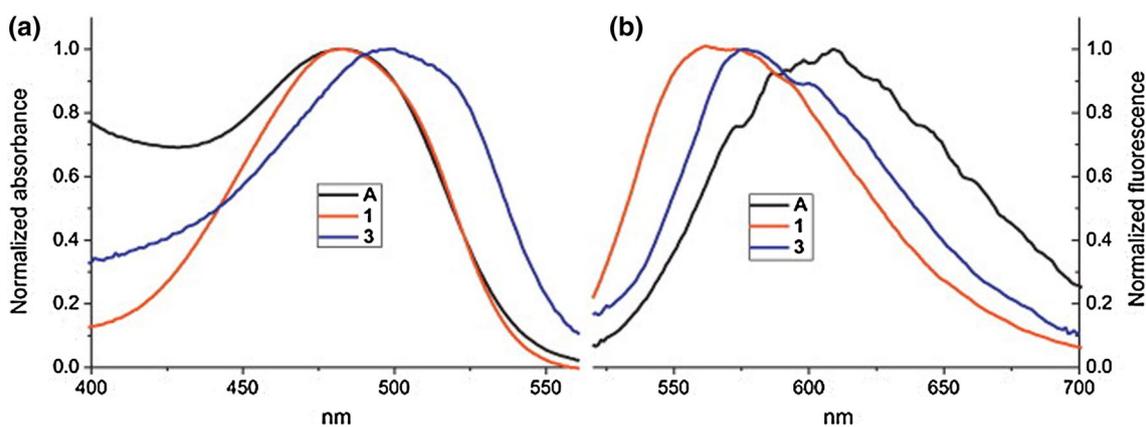


Fig. 49 Normalized **a** UV and **b** fluorescence spectra of protonated 13-membered azobenzocrowns **22** ($\lambda_{\text{ex}}=487$ nm, $\lambda_{\text{em}}=608$ nm), **94** ($\lambda_{\text{ex}}=482$ nm, $\lambda_{\text{em}}=568$ nm) and **96** ($\lambda_{\text{ex}}=490$ nm, $\lambda_{\text{em}}=576$ nm) in acetonitrile. Numbers of compounds in reproduced material cor-

respond to following numbers of compounds in this work: A=**22**; 1=**94**; 3=**96**. Reprinted from [184]. Copyright 2013 with permission from Elsevier. (Color figure online)

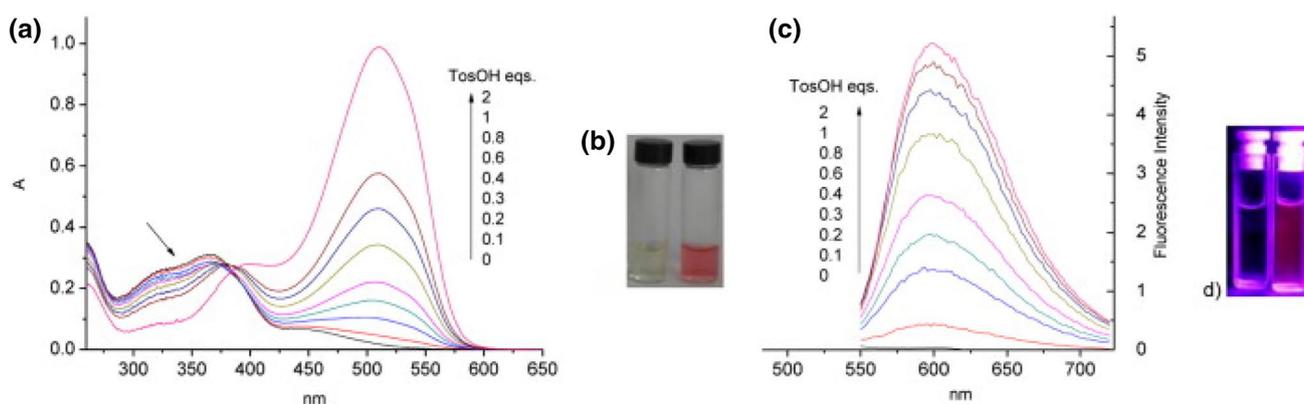


Fig. 50 a Changes in UV-Vis and **c** in emission spectra of **96** ($\lambda_{\text{ex}}=510$ nm, $\lambda_{\text{em}}=598$ nm) (3.73×10^{-5} M) upon titration its solution with *p*-toluenesulfonic acid solution (TosOH); **b** color change

and **d** red fluorescence of **96** in the presence of two-fold excess of TosOH in acetonitrile. Reprinted from [184]. Copyright 2013 with permission from Elsevier. (Color figure online)

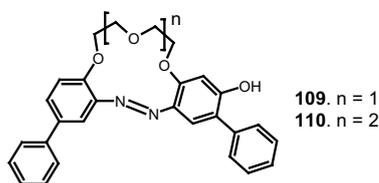


Fig. 51 Hydroxyazobenzocrowns **109** and **110** with phenyl substituents in benzene rings [191]

Lithium cation is stronger complexed by the ionized quinone-hydrazone form than by azophenol tautomer. For **109** and **110** the presence of phenyl rings appears to be important factor affecting tautomeric equilibrium driven by metal cation complexation and the strength of ion–azobenzocrown interactions.

Spectrofluorimetric titrations of **109** with lithium and sodium perchlorates causes the decrease of fluorescence intensity, especially in the presence of organic base. The value of Stern–Volmer (K_{SV}) constant was the highest for lithium in the presence of triethylamine 222 M^{-1} . For comparison this value for **74** is 80 M^{-1} . For 16-membered diphenyl derivative **110** Stern–Volmer constants in the presence of triethylamine were also higher than for compound without phenyl substituents in benzene rings **75**, both for lithium and sodium. In pure acetonitrile quenching constants caused by metal perchlorates were higher for **75** than for **110**, what can be connected with complex formation by this compound in azophenol form in neutral solvent.

Acid–base properties of hydroxyazobenzocrowns were compared with their analogs without hydroxyl substituent by determination of proton binding constants in acetonitrile

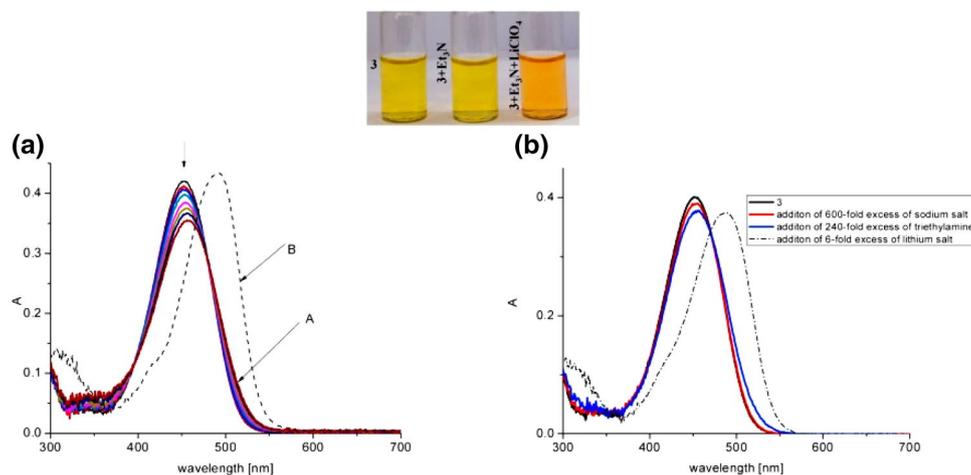


Fig. 52 Top: comparison of color acetonitrile solution of **109** in the presence of Et₃N and both Et₃N and LiClO₄. Bottom: **a** UV-Vis titration of **109** (7.2×10^{-5} M) with LiClO₄ ($0\text{--}5.5 \times 10^{-2}$ M) in pure acetonitrile. Dashed lines are spectra registered upon addition to the titrated system solution of Et₃N (2.8×10^{-2} M); **b** UV-Vis spectra

showing the competitive binding of lithium by **109** in the presence of sodium salt and triethylamine in acetonitrile. Number of compound in reproduced material correspond to following number of compound in this work: 3 = **109**. Reprinted from [191]. Copyright 2017 with permission from Elsevier

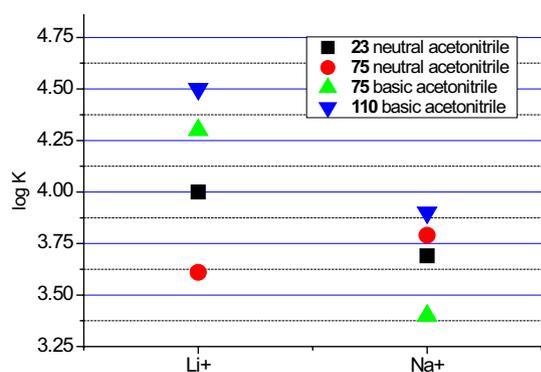
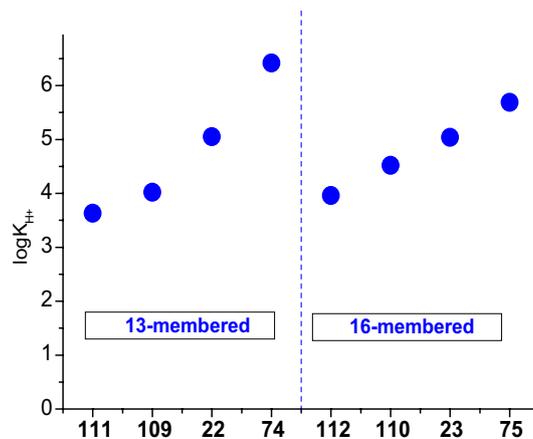
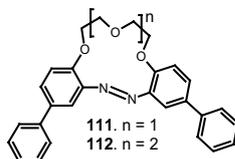


Fig. 53 Comparison of stability constant values of complexes of **23**, **75** and **110** in neutral and basic acetonitrile calculated from UV-Vis titrations [133, 166, 191]

Fig. 54 Comparison of proton binding constants for 13- (**111**, **109**, **22**, **74**) and 16-membered (**112**, **110**, **23**, **75**) azobenzocrowns in acetonitrile [191]



(Fig. 54). Protonation constants of hydroxyazobenzocrowns **109**, **110** and **74**, **75** are higher than for compounds without hydroxyl group (**111**, **22**, **112**, **23**) and among hydroxyazobenzocrowns are higher for macrocycles without phenyl substituents in benzene rings **74** and **75**.

X-ray structure of **109** was solved confirming the existence of this compound in quinone-hydrazone form in solid state (Fig. 55).

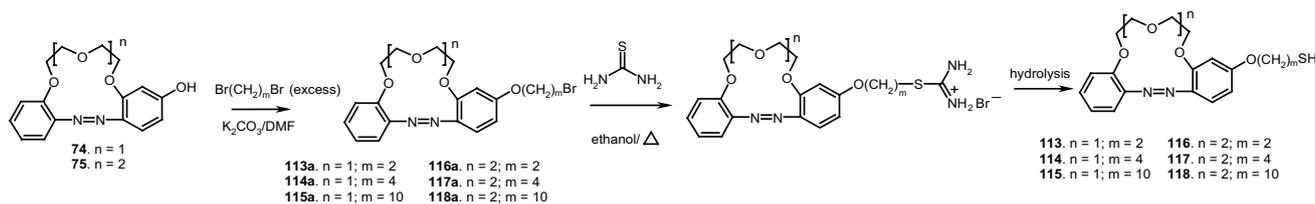
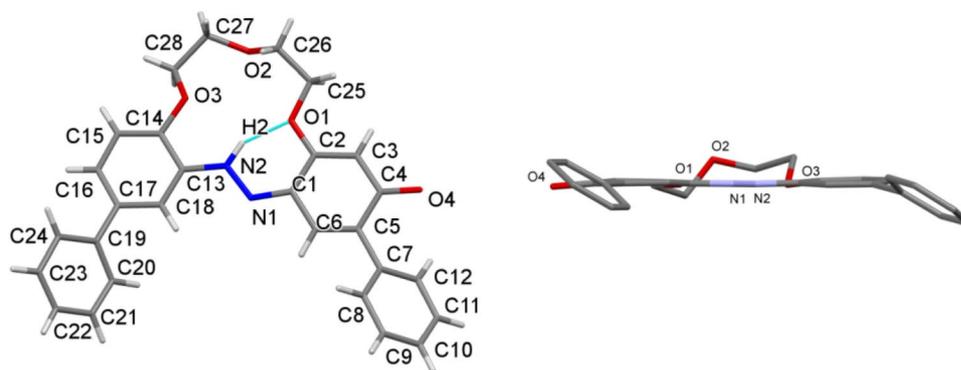
p-Hydroxyazobenzocrowns can be considered as universal substrates for azobenzocrowns skeleton modifications. 13- and 16-membered azobenzocrowns **113**–**118** with peripheral thiol moieties (Scheme 18) were obtained from hydroxyazobenzocrown *via* the respective bromoderivatives (**113a**–**118a**) as substrates under optimized reaction conditions [164]. The most effective preparation method of thiol

derivatives turned out to be a relatively simple thiourea route giving functionalized crowns in yields 27–46%.

Selected thiol derivatives were anchored onto the surfaces of Gold NanoParticles (GNPs) and studied as plasmonic sensors. The stable systems, which can be stored without change for several month were obtained by bifunctionalization of GNPs with thiol and liponic acid. GNPs modified

with 16-membered thiol derivatives and liponic acid, showed spectral and visual response towards potassium cations in aqueous environment. In Fig. 56 the comparison of the spectrophotometric response of GNPs of different concentrations modified with thiol derivative **116** and liponic acid (**Au-116**) towards potassium chloride is shown.

Fig. 55 Left: molecular structure and labeling scheme for **109**; Right: view of **109** along the mean plane defined by the azobenzene moiety. Reprinted from [191]. Copyright 2017 with permission from Elsevier



Scheme 18 Synthesis of thiol derivatives of azobenzocrown ethers **113–118** [164]

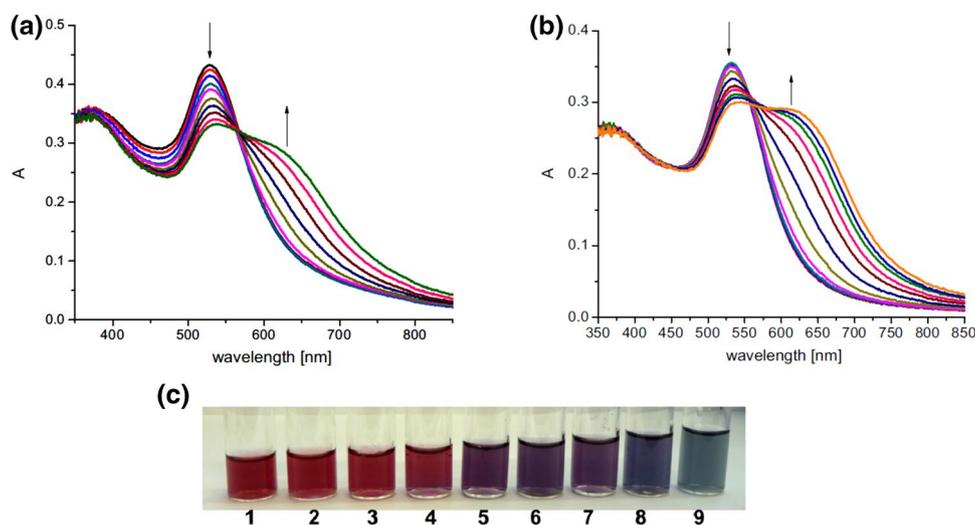


Fig. 56 Changes in UV–Vis spectrum of: Left: Au-**116** (2.3 mL, 0.095 mM) upon titration with KCl solution (0.05 M, titration step: 0.15 mL), Right: Au-**116** (2.2 mL, 0.022 mM) upon titration with KCl solution (0.05 M, titration step: 0.1 mL). Bottom: Au-**116** solutions (1.5 mL) at concentration of 0.36 mM containing increasing ali-

quots of 0.05 M KCl: (1) 0 mL; (2) 0.1 mL; (3) 0.2 mL; (4) 0.3 mL; (5) 0.4 mL; (6) 0.5 mL; (7) 0.6 mL; (8) 0.7 mL; (9) 0.8 mL. Photograph taken 5 min after addition of KCl solution. Reprinted from [164]. Copyright 2015 with permission from Springer Publishing Company

The appearance of new absorption band ~ 600 nm in the presence of potassium is more distinct for diluted solutions GNPs. Color changes of **Au-116** system upon addition of potassium salt are shown in Fig. 56 (bottom). The presence of magnesium and calcium salts do not cause spectral changes, moreover, selective potassium response was found in the presence of these two ions (Fig. 57, left). Color of **Au-116** system in the presence of sodium, magnesium and calcium and color change upon addition of potassium salt is shown in Fig. 57 (right). This behavior is important from practical point of view when considering possible applications of modified GNPs in biomedical analysis.

Among investigated 16-membered crowns the most promising results as potential potassium plasmonic sensor—regarding the stability of the system and fast response—were

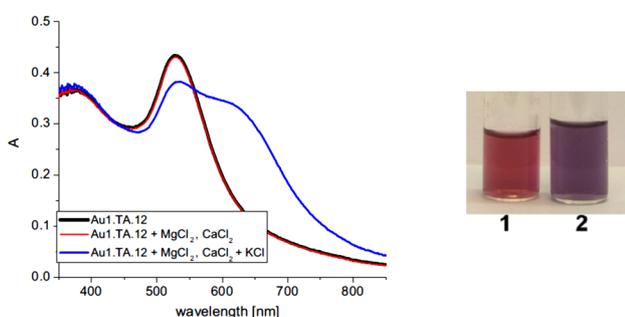


Fig. 57 Left: Comparison of UV–Vis spectra of **Au-116** (2.3 mL, 0.095 mM) registered upon addition of mixture of magnesium and calcium chlorides (0.3 mL, 8×10^{-4} M) and the effect of addition of potassium chloride solution (0.6 mL, 0.05 M). Right: **Au-116** (1.5 mL) at concentration of 0.36 mM: (1) after addition of 0.3 mL of solution containing 8×10^{-4} M CaCl_2 , 8×10^{-4} M MgCl_2 , 1×10^{-3} M NaCl ; (2) after addition of 0.3 mL of solution containing 8×10^{-4} M CaCl_2 , 8×10^{-4} M MgCl_2 and 0.4 ml of 0.05 M KCl . Number in reproduced material corresponds to following number in this work: Au1.TA.12=**Au-116**. Reprinted from [164]. Copyright 2015 with permission from Springer Publishing Company. (Color figure online)

obtained for GNPs modified with crown **117** (**Au-117**). In Fig. 58 changes in absorption spectrum for the highest, among all tested, concentration of colloid is shown.

Azo group(s) in macrocyclic compounds bearing azole rings

Azo group(s) as integral part of the macrocyclic ring

The first series of crown analogs, mainly 18-membered, consisting of two azo groups and heterocyclic (pyrrole or imidazole) residues (compounds **119–121**, **126** and **127** Fig. 59) were obtained and studied since 2003 [192].

The further studies including 21-membered pyrrole derivatives (compounds **122–125** Fig. 59) showed the lead(II) selectivity of this class of compounds [193]. Complexation of lead(II) in acetonitrile solution was connected with large, bathochromic shift resulting in color change from orange to blue. High lead(II) selectivity of these compounds, over other studied metal cations, was also confirmed by experiments on transport through the liquid membranes [193]. Membrane electrodes doped with 21-membered pyrrole derived crowns **123–125** (Fig. 59) also exhibited lead(II) sensitivity with the best parameters for sensor in which lipophilic derivative **123** was used as ionophore.

The relatively high yields of macrocyclization and attractive properties of the initially synthesized azomacrocycles with heterocyclic unit(s) motivated to expanded investigation of this class of compounds. The modification of the skeleton of macroring in different parts of the molecule (schematically drawn in Fig. 60) allowed to obtain macrocycles featuring with many interesting properties.

The structure of 23- and 21-membered macrocyclic pyrrole derivatives **128–131** shown in Fig. 61a comprise chromogenic and fluorogenic character [194]. Pyrrole containing

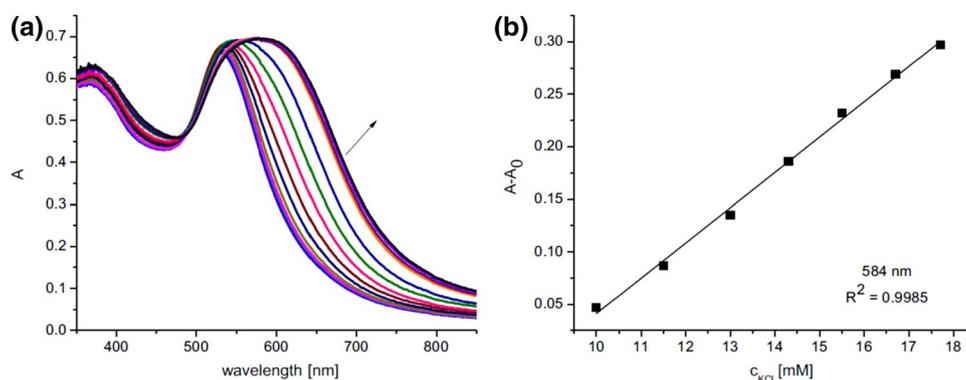


Fig. 58 **a** Changes in UV–Vis spectrum of **Au-117** (2.0 mL, 0.25 mM) upon titration with KCl solution (0.05 M, titration step: 0.1 mL); **b** the range of linear relationship between absorbance and potassium chloride concentration (10–17.7 mM) given as a difference

$A - A_0$, where A is absorbance upon addition of salt and A_0 is absorbance of **Au-117** before salt addition, at 584 nm. Reprinted from [164]. Copyright 2015 with permission from Springer Publishing Company. (Color figure online)

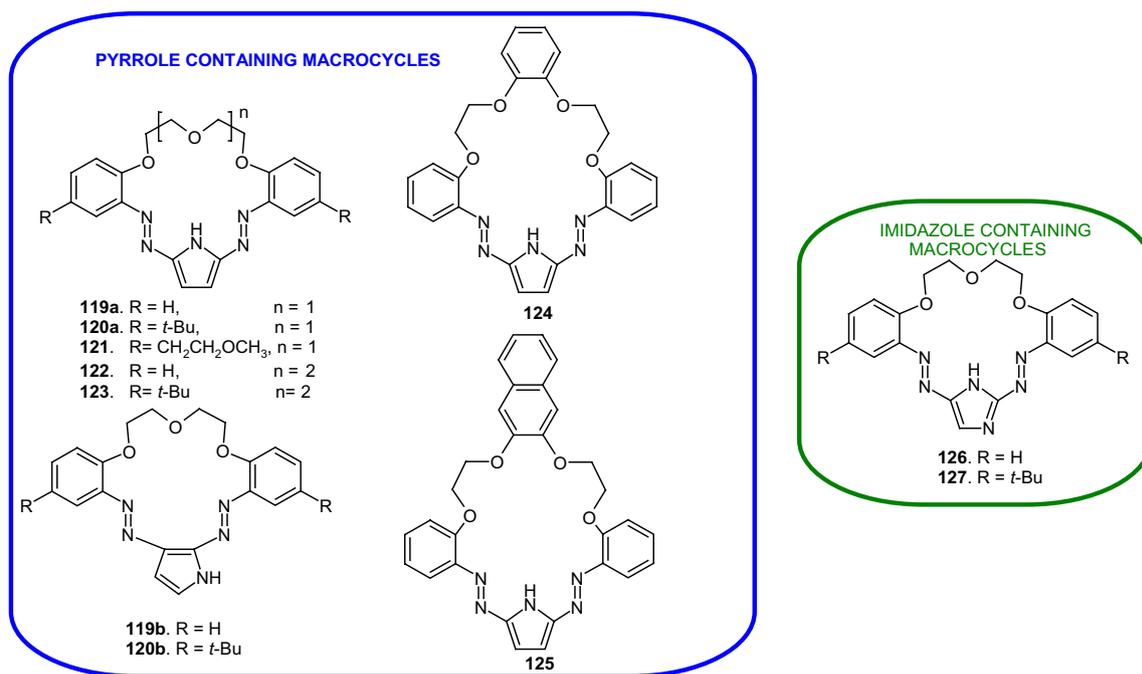


Fig. 59 Chromogenic crown ethers with pyrrole (**119–125**) and with imidazole (**126, 127**) heterocyclic residue and with two azo groups in macrocoring [192, 193]

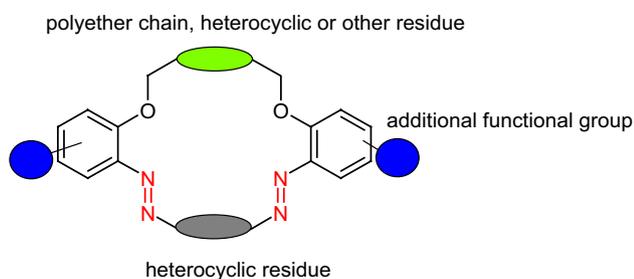


Fig. 60 Schematically: the general formula of macrocyclic compounds with two azo groups and heterocyclic residue(s) as inherent part of the macrocoring

macrocycles were prepared in multistep reactions; the final step was the azocoupling of the respective bisamines with pyrrole that proceeds with moderate yields.

Azocrowns bearing pyrrole residue **128–131** are lead(II) selective with the highest values of the stability constants and the most significant spectral shifts. The comparison of the color change in the presence of lead(II) and other metal cations is shown in Fig. 61b. The comparison of the stability constants of 2:1 (L:Pb) complexes of crowns **128–131** with lead(II) perchlorate in acetonitrile [195] is presented in Fig. 61c.

The incorporation of the fluorescent moieties into the structure of azomacrocycles **125** (Fig. 59) and **128–131** (Fig. 61a) results in compounds for which lead(II)

complexation is associated with quenching or increase of the fluorescence intensity, depending on the type of substituent. The changes of fluorescence intensity for crowns **125** and **128–131** in the presence of 100-fold excess of metal perchlorates in acetonitrile are shown in Fig. 62 (left).

The binding properties and associated changes in fluorescence spectra of crowns **125** and **128–131** can be to some extent explained by analyzing their X-ray structures (Fig. 62, right). Naphthalene fragment(s) influences the geometry of this part of the molecule leading to the significant distortion of its shape. On the other hand the arrangement of the two naphthalene moieties in compound **130** causes molecule to be more flat, resembling a flying butterfly.

For the above-described compounds **119–125** (Fig. 59) the characteristic constituent of the structure is polyether chain coexisting with heterocyclic residue and two azo groups. More “soft” nitrogen atom donor, instead of the polyether fragment pyridine was inserted and the effect of such structure modification on metal cation binding was investigated [196]. Crown **132** (Scheme 19) constitutes of two heterocyclic residues: pyrrole and pyridine, and two azo groups forming inherent part of the macrocoring.

Metal cation complexation by **132** was studied using UV–Vis spectroscopy in acetonitrile and in its mixture with water. In acetonitrile, spectral and color changes were not observed in the presence of alkali metal ions and

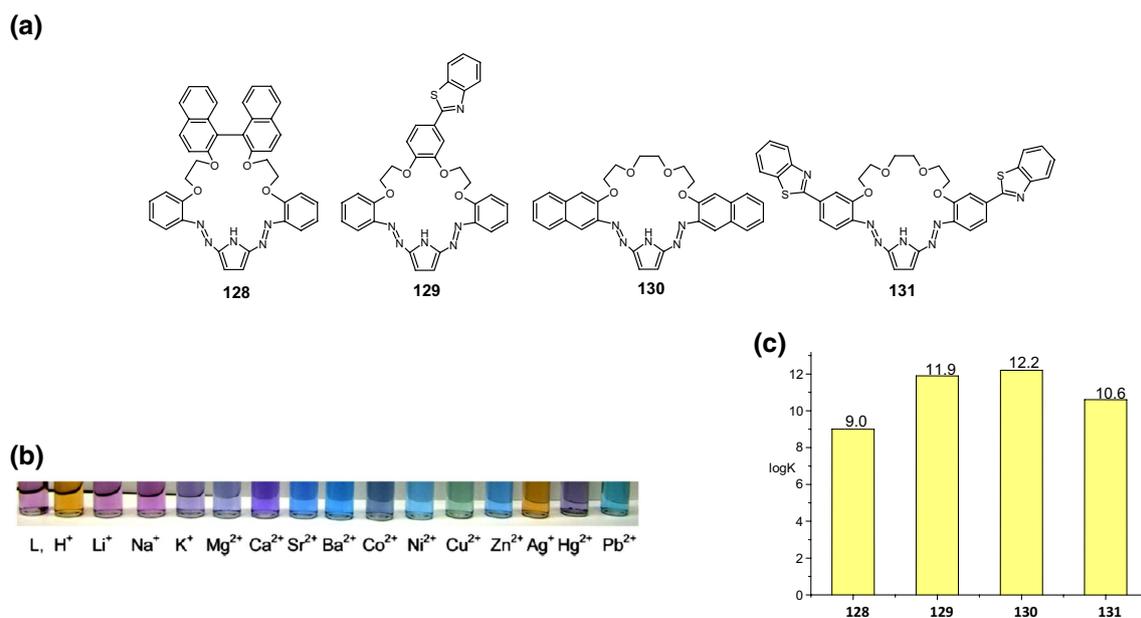


Fig. 61 **a** Chromo- and fluorogenic macrocyclic derivatives with pyrrole residue in the macroring [194]; **b** Color changes of acetonitrile solution of **131** in the presence of the excess of metal perchlorates. Reprinted from [194]. Copyright 2011 with permission from Else-

vier; **c** the comparison of the stability constants of 2:1 (L:Pb) complexes of compounds **128–131** in acetonitrile [195]. (Color figure online)

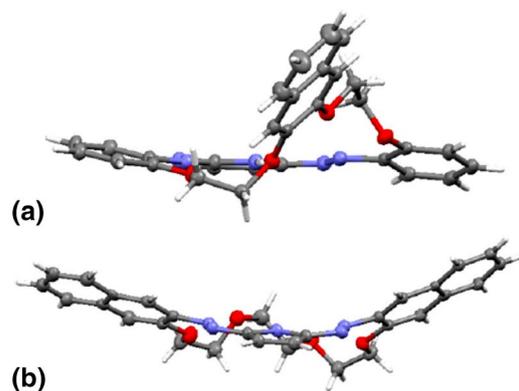
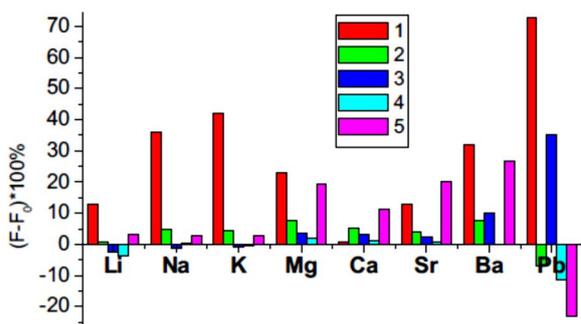
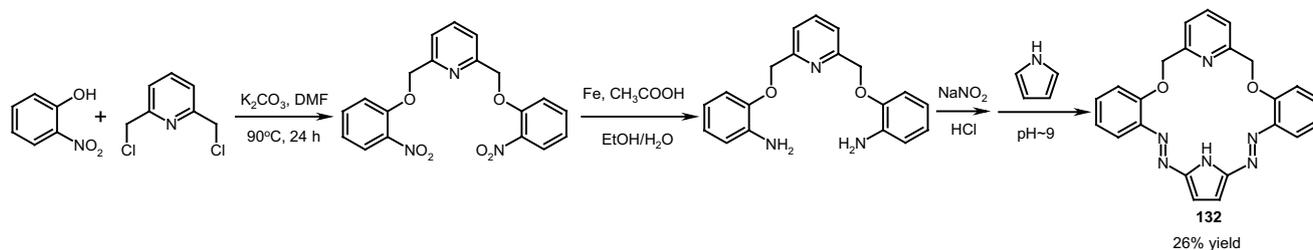


Fig. 62 Left: the changes of fluorescence intensity for crowns **125** and **128–131** in the presence of 100-fold excess of metal perchlorates in acetonitrile. Right: the side view of the X-ray structure of the macrocycles **a** **125** and **b** **130**. Numbers of compounds in reproduced

figure correspond to following numbers of compounds in this manuscript: 1 = **125**, 2 = **128**, 3 = **129**, 4 = **130**, 5 = **131**. Reprinted from [194]. Copyright 2011 with permission from Elsevier. (Color figure online)



Scheme 19 The synthesis of crown **132** with two heterocyclic residues: pyrrole and pyridine as inherent part of the macrocycle [196]

magnesium, but the presence of calcium, strontium and barium affected the absorption spectra of **132**. Among heavy metal cations in the case of nickel(II) and cobalt(II) slow kinetics of complex formation was observed. The presence of zinc(II) and lead(II) caused the most significant color changes of acetonitrile solution of **132** from red to violet. In Fig. 63 the values of the stability constants of 2:1 (L:M) complexes of **132** with metal cations in acetonitrile are presented with the highest value for zinc(II) complex. In water containing system (acetonitrile:water, 9:1), the spectral changes were observed only in the presence of lead(II) and zinc(II) salts.

Besides pyrrole, also imidazole and its derivatives were used as substrates in diazocoupling reaction (reactions carried similarly to that shown in Scheme 19, under high dilution conditions) resulting in 17-(compound **133**), 18- (**136**, **138**, **140**), 20- (**134**) and 21-membered (**135**, **137**, **139**, **141**) macrocycles differing in lipophilicity and substitution of the imidazole moiety (Fig. 64) in the yield up to 55% [197].

The synthesis of macrocyclic imidazole derivatives **126**, and **133–141** was also carried out in the presence of α -, β - and γ -cyclodextrins (CD) [198]. The most spectacular

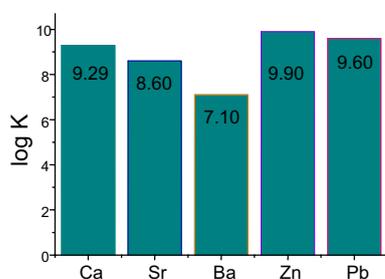
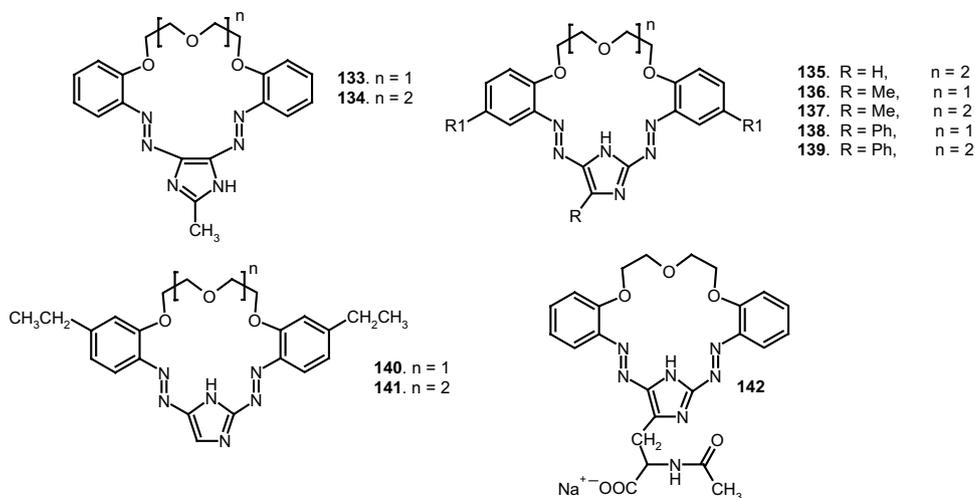


Fig. 63 The comparison of the values of the stability constants of 2:1 (L:M) complexes of compound **132** with metal perchlorates in acetonitrile [196]

Fig. 64 Imidazole bearing macrocyclic azo compounds [197]



results were found for γ -CD in the case of diazocoupling of 2- and 4-methylimidazole with bisdiazonium salt. The yield of products **133** and **136** was as high as 89 and 75%, respectively. In reaction carried out without γ -CD compounds **136** and **133** were obtained in $\sim 45\%$ yield.

Besides imidazole, also 18- (**119a**, **119b**) and 21-membered (**122**) pyrrole derivatives were obtained by cyclodextrin assisted synthesis. The size of used cyclodextrin affects the total yield and the ratio of isomeric crowns **119a** and **119b**. The overall yield of the two macrocyclic compounds formed simultaneously reached 87% in the presence of β -CD. In the absence of β -CD the cumulative yield of both isomers is $\sim 40\%$. Cyclodextrins assisted syntheses of azomacrocyles bearing pyrrole or imidazole confirmed the role of CDs as a molecular reactors.

Metal cations binding by 21-membered imidazole crown **137** was investigated as a model compound by UV–Vis spectroscopy in acetonitrile, methanol, and methanol–water (4:1) mixture. In acetonitrile, only calcium, strontium and barium among alkali and alkaline earth metal cations cause spectral change, i.e. hypsochromic shift, with simultaneous color changes from orange–red to yellow. In the presence of heavy metal salts the color of acetonitrile solution of **137** turns to pink-purple. Schematically, these color changes are illustrated in Fig. 65a. From spectrophotometric measurements it was concluded that complexes of **137** have different stoichiometry depending on the complexed metal cation. Alkaline earth metal ion complexes have 1:1 stoichiometry; in the case of cobalt(II) and nickel(II) 2:1 (L:M) complexes dominate. In systems containing **137** and zinc(II), lead(II) and copper(II) salts two 1:1 and the 2:1 (L:M) complexes exist under equilibrium. The values of stability constants ($\log K$) and the most probable stoichiometry of the formed complexes are shown in Fig. 65b. It was found (1H NMR) that polyether chain plays a bigger role in the case of barium ion complexation

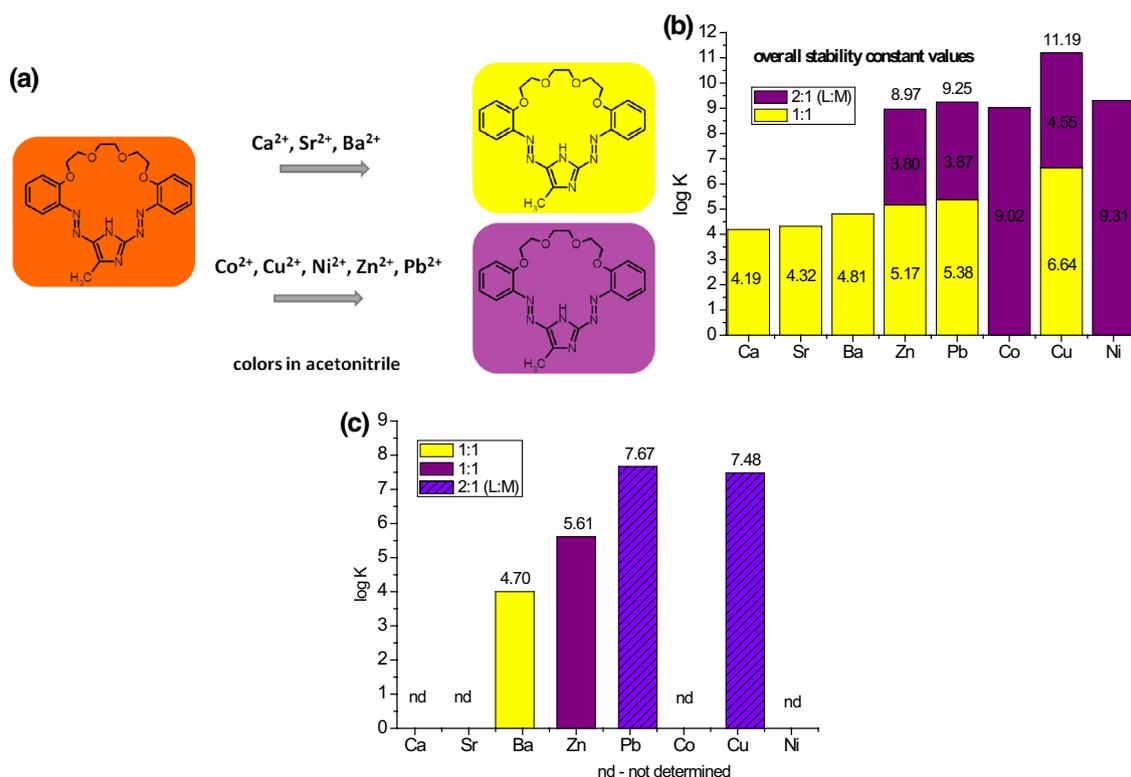


Fig. 65 a Color changes of acetonitrile solution of **137** in the presence of metal perchlorates b the values of the stability constants of **137** complexes with metal perchlorates in: b acetonitrile c methanol [197]

than in the case of lead(II) complex, where imidazole residue was supposed to be engaged. In methanol, only barium, among alkaline earth metal cations, and zinc(II), lead(II) and copper(II), among heavy metal cations, caused significant changes in UV–Vis spectra. For barium and zinc(II) 1:1 complexes dominate, whereas lead(II) and copper(II) form complexes of 2:1 (L:M) stoichiometry. The values of the complex stability constants of **137** with metal perchlorates in methanol are shown in Fig. 65c. Comparing with acetonitrile, in methanol larger spectral shift, particularly for lead(II) complex was observed. The further increase of solvent polarity by water addition to methanol (mixture 4:1, methanol:water) resulted in increase of molar absorption coefficient of **137** and further increase of the selectivity. Only copper(II) complex with spectral shift of 55 nm and lead(II) with spectral shift of 61 nm caused spectral changes in water containing system. The values of the stability constants ($\log K$) 4.75 and 5.51 for lead(II) and copper(II) 1:1 complexes are lower than in acetonitrile and methanol.

Macrocyclic imidazole derivatives (Fig. 64) were also tested as ionophores in membrane ion-selective electrodes. The obtained potentiometric sensors were lead(II) sensitive.

Selected macrocyclic derivatives of imidazole crowns **126** (Fig. 59), **135** and **137** (Fig. 64) were entrapped in silica

xerogel matrix and tested as optical recognition elements for metal cations in water [199]. Among prepared materials, elements based on 21-membered derivatives of imidazole **135** and **137** showed spectral response (change of reflectance) in the presence of lithium salt in aqueous solution.

Structurally similar to the macrocyclic derivatives of imidazole discussed above is compound **142** (Fig. 64), *N*-acetylhistidine based macrocycle with two azo groups [200]. This is an example of crown ether containing azo group and $-\text{COOH}$ mobile protons of unique pH controlled properties. The electrochemical behavior of imidazole derivative **126** (Fig. 59) and *N*-acetylhistidine azocrown ether **142** adsorbed on the electrode surface was investigated under different pH conditions. At pH ~ 12 similar voltammograms for **126** and **142** were obtained, but different changes were noticed under acidic conditions. It was concluded, that at pH higher than 4 the adsorbed compound **142** is present in electrochemically reducible azo form on the electrode surface. In voltammograms the first step of reduction for *N*-acetylhistidine azomacrocyclic **142** was not observed below pH 4.0. This suggests that transformation to the hydrazine form occurs in chemical pathway, not by means of an electrode process. This is a unique example of pH dependent on/off process due to the presence of azo group.

Tetrapyrrolic macrocycles with azo moiety as a part of macroring, compounds **143** and **144**, (Fig. 66) were obtained by Tsuda and co-workers [201].

^1H NMR experiments performed in chloroform-*d* showed that after exposure of **143** solution onto UV-light (350 nm) a mixture of isomers was present. In the photostationary state the *trans* to *cis* ratio was 6:4. Higher conversion to the *cis* isomer was obtained for compound **144** with the *trans* to *cis* isomers ratio 4:6. According to theoretical calculations for *cis* isomer of **143** the photoisomerization results in a larger structural distortion of the macrocycle. Reverse *cis* to *trans* reaction for both compounds occurs thermally and can be accelerated upon irradiation with visible light. The half-life at 25 °C of thermal back reaction for **143** was estimated to be 8 days in toluene solution, whereas for **144** the time is much shorter (14 h). Taking into account the values of activation energy for **143** and **144** obtained from van't Hoff plots it was stated that the thermal stability of *cis*-**144** is lower than *cis*-**143** having less overcrowded structure. The

obtained tetrapyrrolic derivatives were suggested to be used for photoswitchable molecular devices.

The reversible *trans*–*cis* isomerization of host molecules, as it was shown in relatively early works of Shinkai [141] can be used for light-driven transport of guest molecules. Scherman and co-workers [202] described the synthesis and complexing properties of macrocycle **145** (Fig. 67) containing bis(imidazolium)-azobenzene motifs.

According to ^1H NMR analysis, compound **145** exists as all-*trans* isomer (*E,E*-**145**). UV-light-promoted isomerization generates a mixture of stereoisomers: *E,E*-**145** (18%), *E,Z*-**145** (38%) and *Z,Z*-**145** (44%). Visible light illumination causes the reversible process; the isomer *E,E*-**145** (64%) prevails in the photostationary state. Macrocycle **145** in all-*trans* form interacts with 4,4'-dipyridyl-*N,N'*-dioxide (4DPDO). According to quantum mechanical calculations, in the complex the ligand adopts cage-like conformation and the guest is encapsulated inside the host cavity. Each of the oxygen atoms of 4DPDO are hydrogen-bonded by two of

Fig. 66 Tetrapyrrolic macrocycles **143** and **144** with azo group in macroring obtained by Tsuda and co-workers [201]

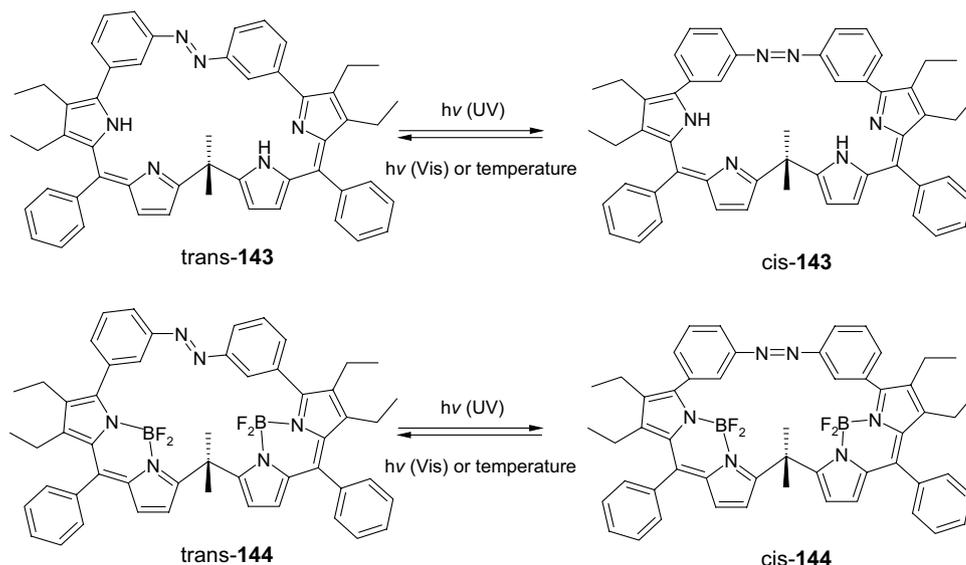
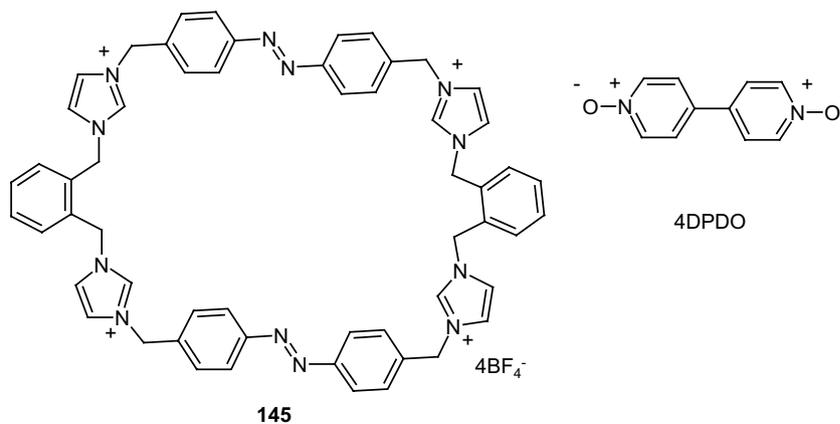


Fig. 67 Molecular container **145** and its complementary guest—4,4'-dipyridyl-*N,N'*-dioxide (4DPDO) [202]



the four acidic protons from imidazole rings. Exposure of the complex to UV light induces *trans* to *cis* isomerization of the macrocycle azo groups resulting in the guest release from the cavity, what was detected by ^1H NMR spectroscopy. The studies revealed that the *Z*-predominant isomeric mixture of **145** interacts with the guest very weakly due to significant decrease of the **145** cavity size and change of the ligand conformation resulted from *trans* to *cis* isomerization. Irradiation of the mixture with visible light increases the ligand affinity to 4DPDO and ensures the guest encapsulation inside the macrocycle cavity. Thanks to these properties ligand **145** can be considered as photoswitchable molecular container.

Azo group(s) as peripheral part of the macrocycle

Thanks incorporation of azo group “spectroscopically silent” in the visible range compounds gain the chromogenic character and photoactivity. Ballester and co-workers [203] reported dimerization of urea-based calix[4]pyrrole with four appended azobenzene units **146** (Fig. 68) templated by the encapsulation of 4,4'-dipyridyl-*N,N'*-dioxide (4DPDO). The assembly can be observed only when all azo groups are in *trans* form. Light-induced *trans* to *cis* isomerization of a single azo moiety is probably sufficient to trigger the disassembly of the capsule. The subsequent back isomerization conducted in the dark or upon visible light irradiation results in the quantitative recovery of the complex.

Molecular switches or molecular devices can be based on porphyrins and phthalocyanines skeletons. An azobenzene moiety was joined at the meso position of expanded

porphyrins, called also smaragdyrins [204]. Azobenzene-smaragdyrin conjugates were obtained using dipyrromethane bearing azo unit as precursor. Smaragdyrins **147–154** (Fig. 69) were synthesized by acid-catalyzed [3 + 2] oxidative coupling of dipyrromethane precursors with para-substituted 5,10-diphenyl-16-oxatripyranes in ~25% yields. The spectroscopic and electrochemical measurements showed that the azobenzene residue covalently linked to the meso carbon atom leads to interaction between the azobenzene residue and the macrocyclic π -system. In solid state compound **151** azo group is in *trans* form. The reversible *trans* \rightleftharpoons *cis* isomerization for smaragdyrin–azobenzene conjugate was studied for compound **151** in toluene by irradiation (360 nm) and monitored by UV–Vis spectroscopy. In UV–Vis spectrum upon illumination the following changes were observed: a decrease of absorption intensity of band at 350 nm and some increase in the absorption band at 450 nm with a sharp decrease of the band intensity in the Soret band region. These changes were attributed to the *trans* to *cis* isomer conversion. Upon continuous irradiation within about 5 min a green color characteristic for smaragdyrin disappeared indicating decomposition of the system.

The linkage of porphyrin—both free and in form of zinc(II) complex—and fullerene C_{60} moieties (Scheme 20) with a central azobenzene moiety gave dyads **155** and **156** which properties can be controlled by photoinduced changes only of the azo moiety [205]. The properties of obtained macrocycles were exhaustively studied by spectroscopic and electrochemical methods. Unfortunately, no evidence of photoinduced *cis*–*trans* isomerization was found; moreover the photochemical decomposition of the azo dyads was noted upon continuous irradiation.

Fig. 68 Calix[4]pyrrole **146** bearing azobenzene moieties described by Ballester and co-workers [203]

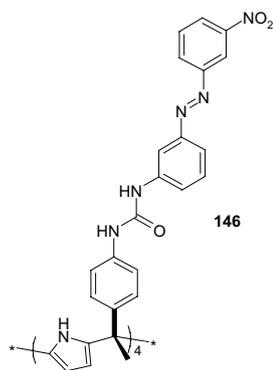
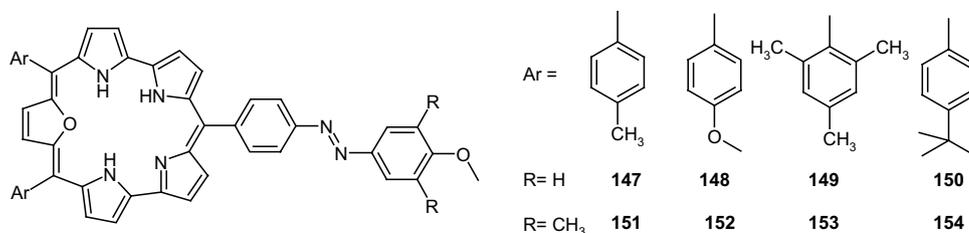
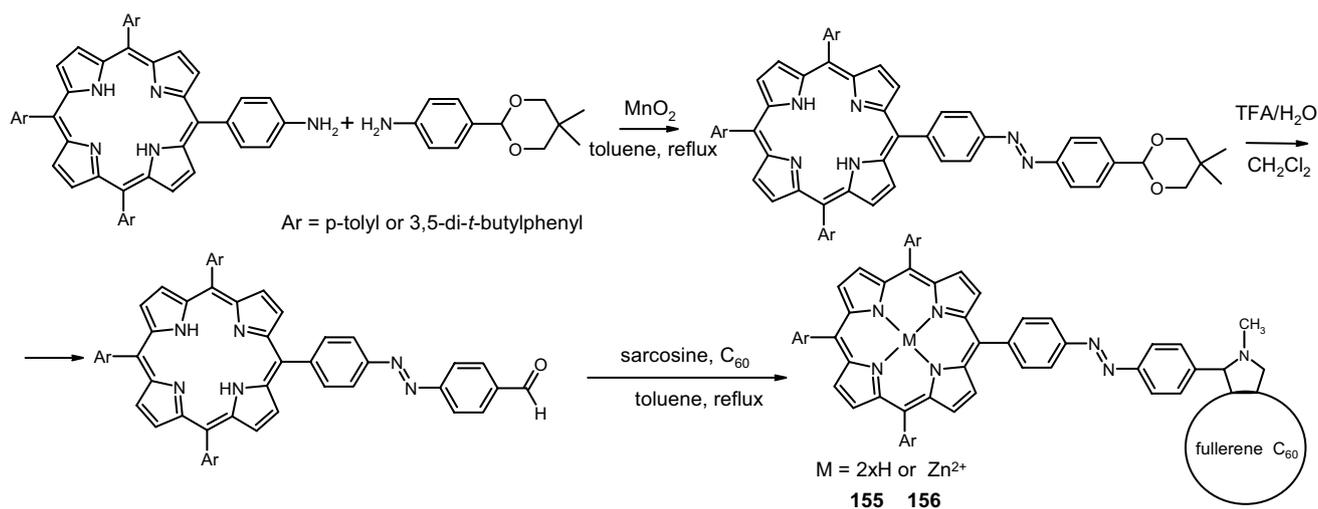


Fig. 69 Azobenzene-smaragdyrin conjugates **147–154** [204]





Scheme 20 The synthesis of porphyrin-azobenzene-fullerene dyads **155** and **156** [205]

molecules, which properties can be modified towards chromogenic, photoswitchable and redox active compounds by introducing an azo moiety. Modification of calixarenes skeleton can lead to conformational and structural changes and can affect the binding properties of these macrocyclic compounds.

The conformation of calix[4]arene, i.e. *cone*, *partial cone* or *1,3-alternate* depends inter alia on the number and type of substituents of the macroring. Calix[4]arenes in the *cone* conformation with azo moieties at the *meta* position (in relation to the alkoxy groups) (Fig. 70) were obtained via nitroso derivatives obtained by preceded mercuration [215]. The effect of the unique for calixarenes substitution pattern and introduction of two bulky residues is rigidified molecular structure. In solution *pinched cone*–*pinched cone* interconversion is stopped (found by temperature dependent ¹H NMR in dichloromethane), meaning that modified calix[4]arene adopts only one of the two possible pinched cone conformations. The resulting *meta* position modified calix[4]arenes can be considered as molecular receptors exhibiting cation- π binding properties.

A series of azocalix[4]arenes **161**–**164** with one, two, three, and four free phenolic groups were synthesized in reaction of 4-nitro- and 2,4-dinitrophenylhydrazines with calix[4]arene diquinones and also in diazocoupling reactions of calix[4]arenes [216]. As an example, the synthesis of dialkyl azocalix[4]arenes with two free phenolic groups by diquinone route is shown in Scheme 21.

The relationship between the conformation of azocalix[4]arene and the number of free phenolic OH groups was established by X-ray studies. It was concluded that azocalix[4]arene adopts a *cone* conformation if it contains at least one free phenolic OH group. *Partial cone* or *1,3-alternate*

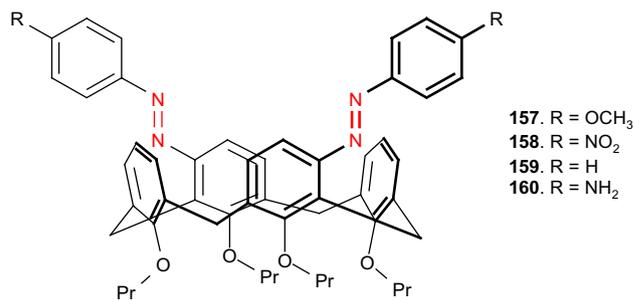
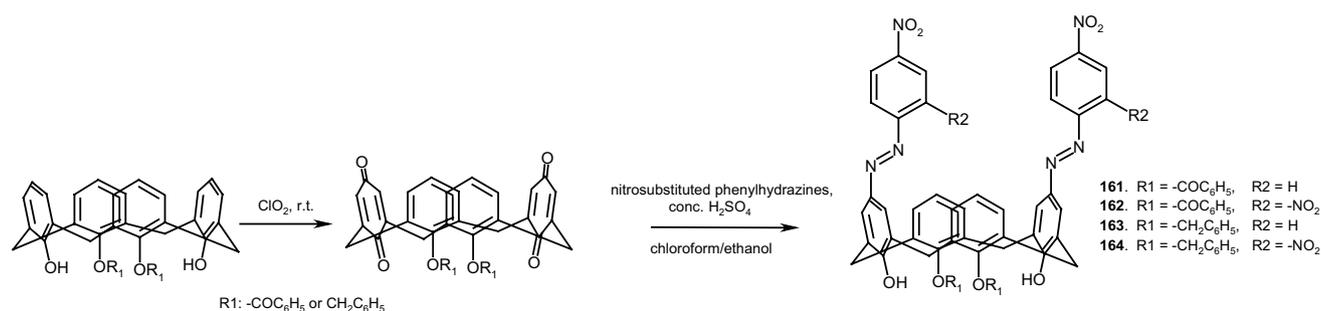


Fig. 70 *Meta*-substituted calix[4]arenes **157**–**160** of unusual rigidity [215]

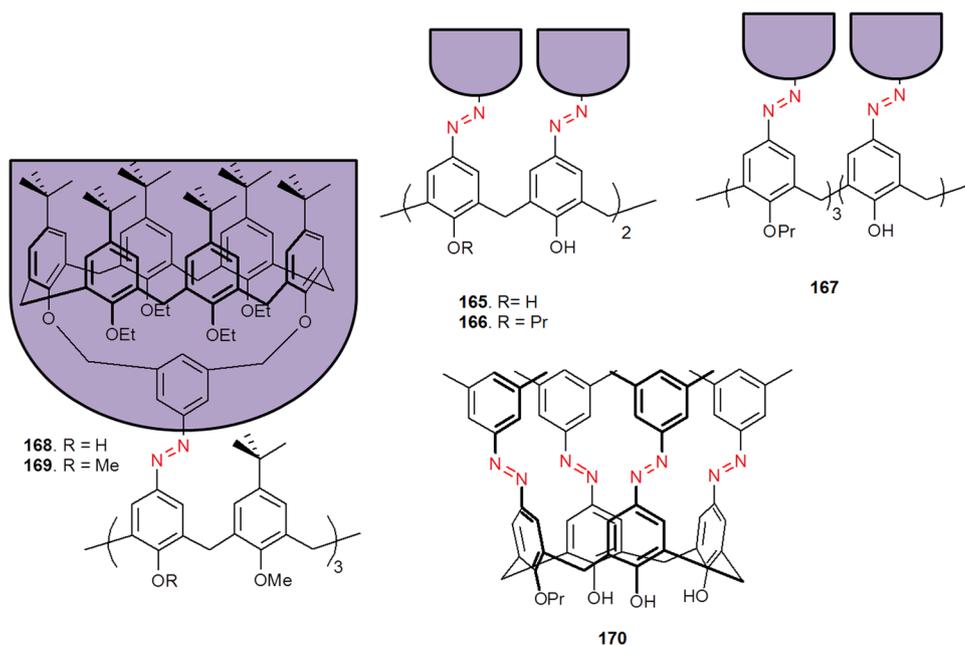
conformers of azocalix[4]arenes were found in macrocycles with all substituted phenolic groups. The possibility of the controlling the conformation of calix[4]arenes is very important in design of the complementary hosts to a particular guest molecules. Additional benefit is the introduction of chromogenic groups useful in preparation of spectrophotometric and naked-eye sensing reagents.

Dendricalixarenes **165**–**169** (Fig. 71) [217] with azo groups at the upper rim of calix[4]- or calix[6]arene can be next examples of the structural consequences of attaching photoactive azo moieties. For preparation of azo derivatives, a method of direct diazocoupling of the diazonium salt with calixarene core was elaborated. The expected effect of the introduction of photoswitchable azo groups was to control the flexibility and accessibility of the inner spaces of the dendrimers by *trans*–*cis* isomerization. It was found that receptors **166**, **167**, and **169** (with the azo groups in the *trans* configuration) do not form complexes despite the increased accessibility of the inner spaces of the dendrimers.



Scheme 21 Synthesis of *O*-dialkyl azocalix[4]arenes **161–164** with two free phenolic groups by diquinone route [216]

Fig. 71 Dendricalixarenes **165–169** functionalized with photoswitchable azo moieties and model compound **170** [217]



Complex formation was not observed also upon photoinduced conversion into *cis* forms. However a model compound **170** forms a stable complex of high binding constant with 4-(4-dimethylaminostyryl)-*N*-methylpyridinium iodide (DASPMI) of rod-like geometry.

Photoactive azo residue was also used for joining two molecules of calix[4]arenes to form dimeric structures (Fig. 72) [218]. Azo linked calix[4]arene **171** exists in a locked *trans* form [219]. The modification of the structure by providing azobenzene linker between two calix[4]arene molecules results in more flexible structure, for which the reversible *trans*–*cis* isomerization can be achieved. UV illumination of **172** solution in chloroform leads to photo-stationary state with 35:65 *trans* : *cis* ratio (according to ^1H NMR measurements). The reverse *cis*–*trans* isomerization rate of **172** in the dark, is solvent polarity, viscosity and pH dependent. These properties of dimeric calix[4]arene make it useful as a light-responsive molecular container.

Calixarene skeletons with crown ether functionality form metal cation complexes with the involvement of the macrocyclic polyether unit. Complexation of metal cations by azobenzene crown ether *p*-*tert*-butylcalix[4]arene **173** (Fig. 73) was analyzed using DFT calculations [220].

The optimized structures of crown ether *p*-*tert*-butylcalix[4]arene showed that more stable is this form in which azo moiety has *trans* configuration. Proposed models also showed that for both *trans* and *cis* isomers better complexation efficiency for alkali metal cations characterizes crown-ether with benzene-rings (*exo*) pockets. The obtained models also proved the higher affinity of macrocycle for sodium over potassium ion. Preferential sodium cation complexation was found for the *trans*-complex in the *exo*-binding mode comparing the *cis*-*exo* analogue.

Calix[4]biscrown compound bearing azobenzene unit **174** was described as a colorimetric receptor for Hg^{2+} ions in acetonitrile (Fig. 74) [221]. According to UV–Vis analysis it was stated that complex of 1:1 stoichiometry is formed.

Fig. 72 Dimeric azocalix[4]arenes: **171**—not isomerisable, **172**—undergoing reversible *trans*–*cis* isomerization [218]

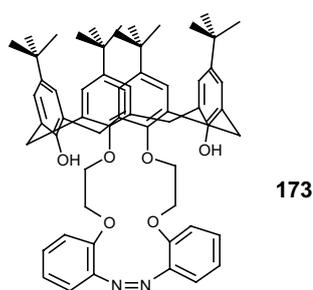
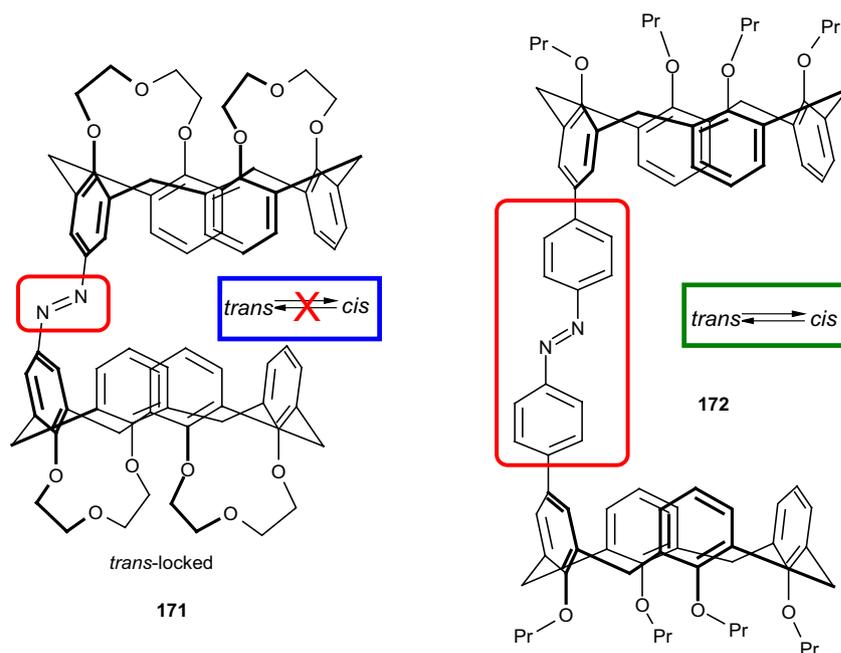


Fig. 73 Azobenzene crown ether *p*-*tert*-butylcalix[4]arene **173** studied analyzed by DFT calculations [220]

Complexation is manifested by change of color from red to pale yellow only when perchlorate anions are used as the counterion. X-ray analysis of $\text{Hg}(\text{NO}_3)_2$ -**174** complex shows crystal structure with the formula $[\text{Hg}(\text{174})(\text{NO}_3)_2]_n$ where each of *exo*-coordinated Hg^{2+} is in a distorted tetrahedral surroundings with its coordination sites occupied by two sulfur atoms from two different ligand molecules and two NO_3^- anions. Addition of potassium salt to Hg^{2+} -**174** solution in acetonitrile affects its UV–Vis spectrum indicating formation of heterodinuclear complex.

Supramolecular systems, where ionic species are transported across bilayer lipid membranes are used as artificial models for natural photo-excitable membranes. It was demonstrated that ether derivative of *p*-*tert*-butylcalix[4]arene bearing photoresponsive dimethylaminoazobenzene moiety **175** (Scheme 22) - acts as ion carrier in visible light (> 400 nm) driven transport of sodium cations across lipid bilayer (soybean phospholipids) [222].

Macrocyclic compounds with azo groups are interesting objects not only as molecular receptors or building blocks in molecular devices or machines. Some of them show anti-bacterial activity. An example can be azocalix[4]arene **176** (Fig. 75) obtained by diazo-coupling of tetradiazonium salt of calix [4] arene and 2,6-diaminopyridine, which shows strong activity against Gram-positive bacteria *Bacillus cereus*, while the activity against *Escherichia coli* is mild; the *Pseudomonas aeruginosa* shows resistance [223].

Cyclodextrins with anchored azo moiety

Cyclodextrins (CDs) known since the Villiers discovery [224] for more than 125 years, are cyclic oligosaccharides that consist of D-glucopyranose units linked by α -(1,4)-glycosidic bonds. Depending on the number of glucopyranose units CDs can be classified as α -, β - and γ -CD with six, seven and eight units, respectively (Fig. 76).

Hydrophilic outer surface and a hydrophobic cavity of the native CDs and their derivatives cause that these compounds form inclusion complexes with a number of guest molecules: organic, inorganic, biological molecules and ions [225]. The versatile career of CDs is due to their availability, biocompatibility, biodegradability and industrial scale production. That is why complexes of CDs with various molecules have found multitude applications in many branches of science and industry, for example: in supramolecular polymers [226–228], hydrogels [229], bioactive materials [230], drug delivery [231–234], dynamic materials [235], catalysis [236] and separation methods

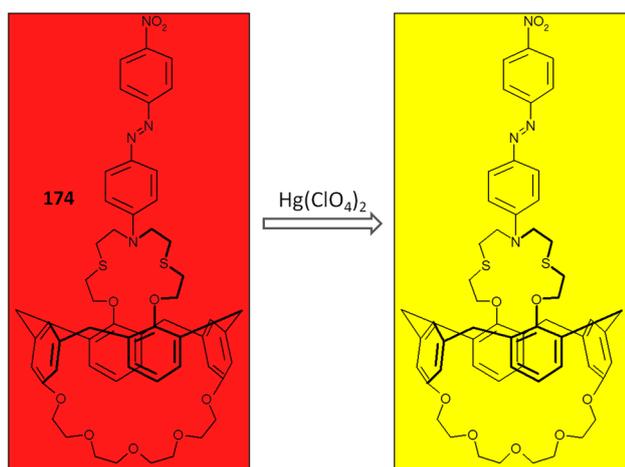


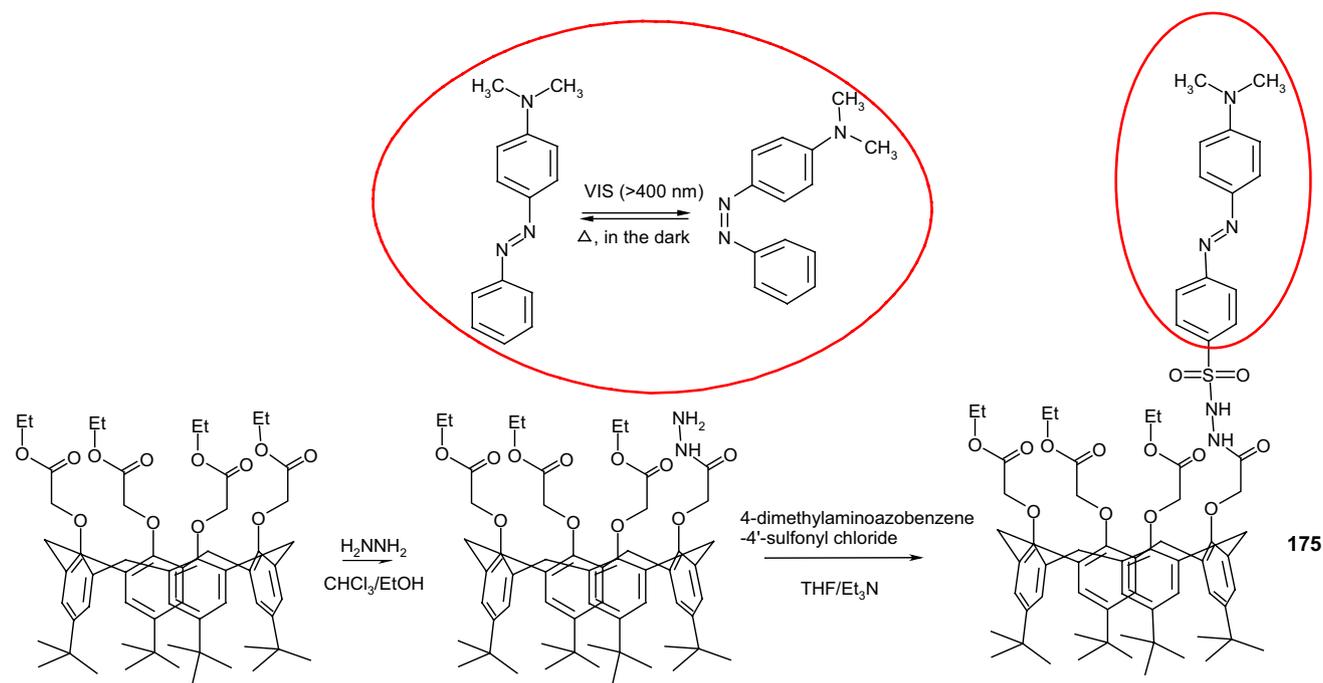
Fig. 74 Calix[4]biscrown **174** sensitive to mercury(II) perchlorate as chromogenic sensing molecule [221]

[237–239]. Substituted cyclodextrins give the possibility of the more specific modification of different type of substrates, molecules and surfaces [240], changing their physical and chemical properties, which opens the way to the new applications of such hopeful systems [241]. Among them, derivatives bearing azobenzene moiety at various locations of the cyclodextrin(s) containing architectures, are intensively studied due to their chromogenic

character and unique photo-responsive properties induced by the photochemical *trans*–*cis* isomerization.

Cyclodextrins are often utilized in a construction of the nanometer-scaled supramolecular architectures. The properties and the organization of self-assemblies can vary, inter alia, depending on reaction conditions. This allows to control the structure of supramolecular frameworks by tuning the conformation of the building blocks. The interesting example of systems of the same chemical composition, but different conformations was given by Liu and co-workers [242]. In Huisgen's cycloaddition, (Fig. 77) using the same reactants, but carrying out the reaction under different conditions two conformations: **177a**—self-locked and **177b**—self-unlocked were identified. The structures of obtained products were confirmed by X-ray and spectroscopic studies. In self-locked conformer—in solution and in solid state—azobenzene moiety is located in the own cavity. It presents a unique [1]rotaxane without a stopper part. Conformer **177b** was found to exist as a linear supramolecule.

[1]Rotaxanes **178** and **179** (Fig. 78) based on β -cyclodextrin skeleton, bearing azo moiety were also efficiently prepared via self-inclusion complexation and Suzuki-coupling capping in aqueous solution [243]. Obtained [1] rotaxanes **178** and **179** were characterized by absorption and induced circular dichroism spectra. [1]Rotaxane **179** undergoes photo- and thermal reversible *trans*–*cis* isomerization.



Scheme 22 Synthesis of a photoresponsive ion carrier **175** based on calix[4]arene [222]

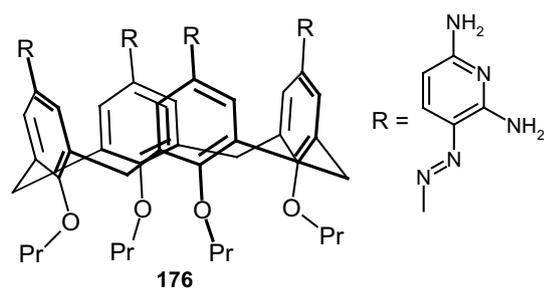


Fig. 75 Azocalix[4]arene showing antibacterial activity [223]

The goal of the introduction of azo moiety while modifying the cyclodextrins is the possibility of using them as photoswitchable systems. The efficient modification of the cyclodextrins' skeleton requires the elaboration of the

appropriate synthetic procedures. Huisgen's [3 + 2] cycloaddition was used for the efficient synthesis of azobenzene bridged β -cyclodextrin **180** (Fig. 79a) which was obtained in a click reaction in 72% yield [244].

An oxidative coupling in Sonogashira-type reaction has been used in synthesis of β -cyclodextrin derivative bearing an azobenzene group on the secondary face **180** (Fig. 79b) [245]. Optimized reaction conditions (degassing of the reaction mixture, the use of $[\text{Pd}(\text{PPh}_3)_4]$, pre-heated oil bath, diluted reductive H_2 atmosphere) allowed to diminish the formation of the dimeric side product. Such conditions resulted in 62% yield of the demanded azocompound.

Not only assembly/disassembly processes of supramolecular systems can be photocontrolled, but also morphological transformations of supramolecular assemblies can be light-induced. Liu and co-workers [246] described the nanotube–nanoparticle morphological conversion for the

Fig. 76 α -, β - and γ -CD and schematically their hydrophilic surface and a hydrophobic cavity

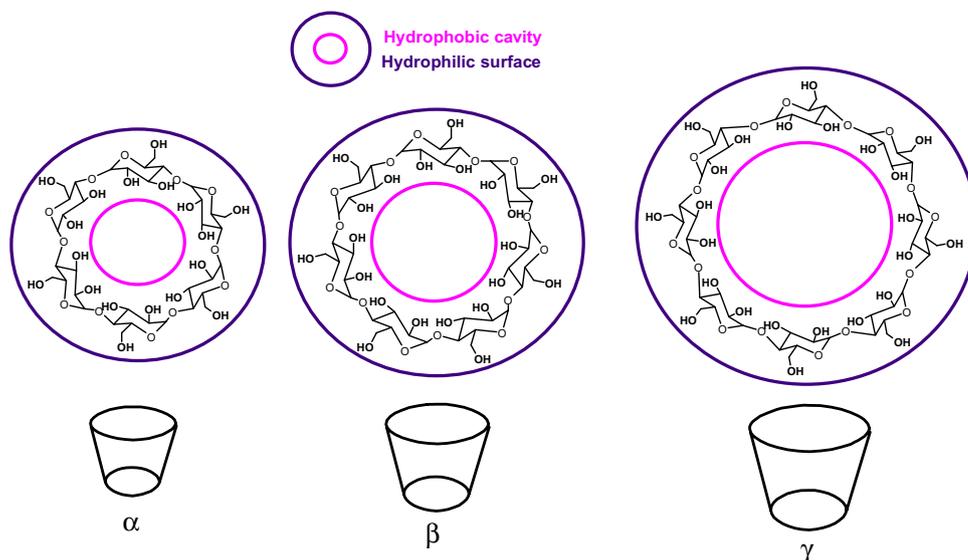


Fig. 77 The scheme of Huisgen's cycloadditions carried out under diverse conditions leading to different conformers **177a** and **177b**. Numbers of molecules in reproduced material correspond to follow-

ing numbers in this manuscript: 1 = **177a**; 2 = **177b**. Reprinted with permission from [242]. Copyright 2008 American Chemical Society

secondary assembly of amphiphilic porphyrin (**182**, guest molecule) mediated by azobenzene-bridged bis(permethyl- β -CD) (**183**, host molecule; Fig. 80). Azo-bridged β -CD, both as *trans* and *cis* isomer, forms complexes of 1:1 stoichiometry with guest molecule, porphyrin **182**. The estimated binding constant value for complex of *cis*-isomer is higher than for *trans* isomer. This can be a result of cooperative binding of the porphyrin derivative by two β -cyclodextrin cavities in *cis*-complex (the sandwich type complex). The interaction of *trans*-azobenzene-bridged bis(permethyl- β -cyclodextrin) **183** with amphiphilic porphyrin derivative **182** in aqueous solution (pH 7.2) leads to aggregate formation with hollow tubular structure. According to TEM images the average inner and outer diameters of obtained nanotubes were about 45 and 61 nm with a wall thickness of about 8 nm. It was deduced that interior and exterior surfaces of nanotubes are composed of units of *trans*-**183-182** complexes, whereas the alkyl chains of **182** interlace with each other in the middle of tubular walls. UV-light irradiation (at 365 nm) of *trans*-**183-182** solution induces morphological conversion being a result of *trans* to *cis* isomerization of the azobenzene moiety. TEM analysis confirmed that upon photoisomerization long *trans*-**183-182** nanotubes turn to nanospheres of average diameter of 180–220 nm. Subsequent irradiation of complex with the *cis* isomer enables nanotubes reconstruction. This photocontrolled process was found to be reversible and repeatable. TEM and DLS experiments confirmed that morphological switching can be repeated by irradiation for tens of times.

Photochemical properties of bis- β -cyclodextrin bearing azobenzene unit **184** was reported by Djedaini-Pilard and co-workers [247]. It was found that **184** can form two different inclusion complexes with adamantyl derivative **185** depending on **184** photoinduced isomers (Fig. 81). The *cis*-**184** complex of 1:1 stoichiometry is created when two cyclodextrin cavities bind simultaneously two adamantyl units of the guest. It was deduced that formation of supramolecular polymers with *n:n* stoichiometry for *trans*-**184** is the most probable.

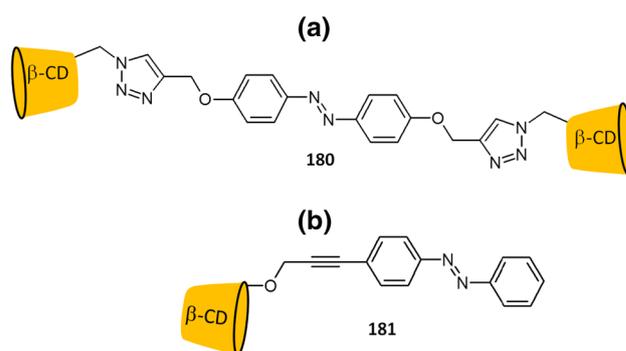


Fig. 79 a Azobenzene bridged β -cyclodextrins **180** [244]; b β -CD modified on a secondary faced with azobenzene **181** [245]

Azobenzene attached to the secondary face of β -CD was utilized as gated synthetic ion channel (Fig. 82) [248]. β -CD works as a channel for ionic species transport, whereas photoresponsible azobenzene unit acts as a gate, opening or closing upon irradiation. Dependent on *trans* or *cis* conformation of azobenzene moiety, cations or anions are preferentially transported.

Permethylated α -cyclodextrin-6^A-monoalcohol modified with azophenol **186** (Fig. 83) was used [249] for colorimetric detection and differentiation of primary, secondary and tertiary amines. Chloroform solutions of **186** change color in the presence of 1° and 2° amines as a consequence of spectral shift from 380 to 580 nm for primary and to 530 nm for secondary amines. Tertiary amines do not cause spectral changes. The values of the stability constants of **186** complexes with primary amines were found to be higher than with secondary amines. It was deduced that the number of possible hydrogen bonds formed between the oxygen atoms of crown ether and amine protons is crucial for the binding strength of the guest molecule.

Azobenzene bearing γ -cyclodextrin derivative **187** (Fig. 84) was tested as a probe for derivatives of phosphoric acid [250]. It was proved that the described host molecule reveals high selectivity towards ATP over other tested

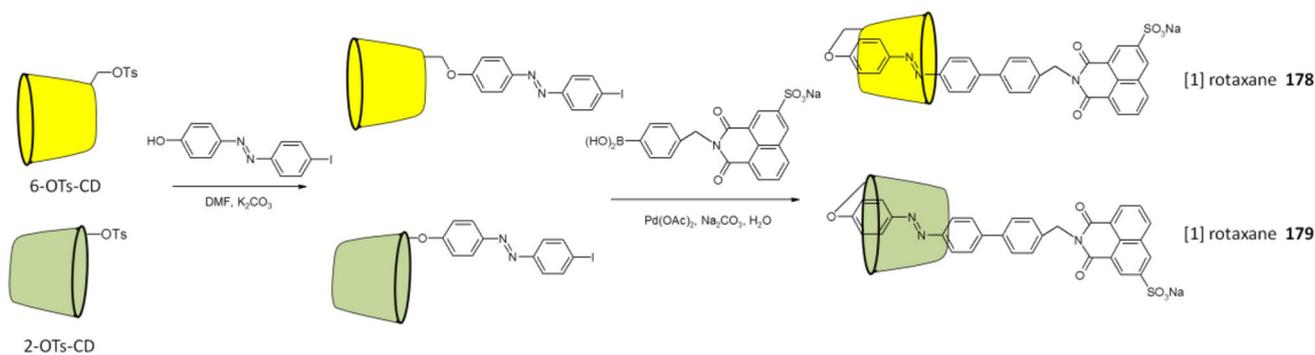


Fig. 78 Disparate oriented [1]rotaxanes **178** and **179** bearing azo moiety [243]

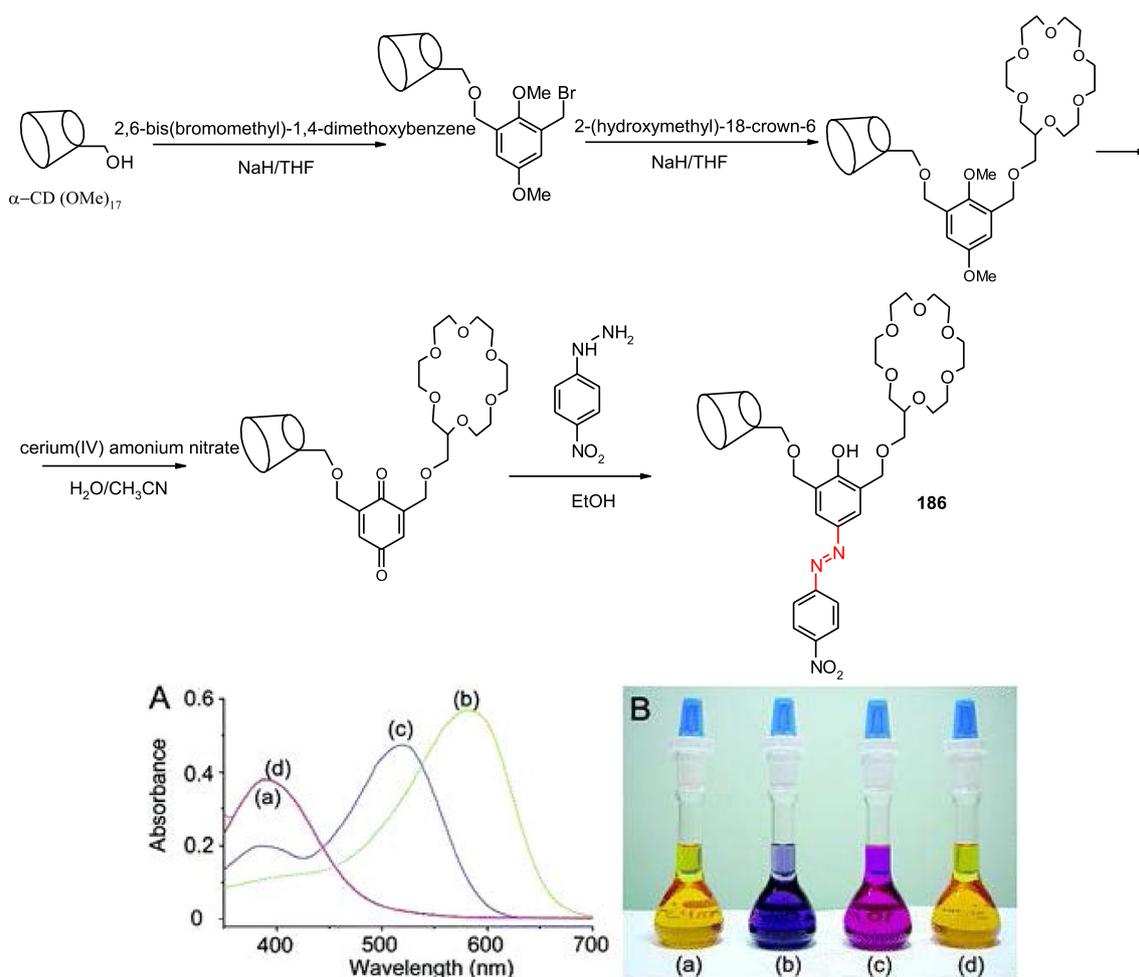


Fig. 83 Top: Synthetic procedure for **186**, bottom: **A** UV-Vis spectra and **B** photographs of **186** (0.03 mM) with amines (1000 equiv) in CHCl₃: **a** **186**, **b** **186** + *n*-octylamine, **c** **186** + di-*n*-butylamine, and **d**

186 + tri-*n*-butylamine. Reprinted with permission from [249]. Copyright 2006 American Chemical Society

delivery system [251]. The single particle consists of gold nanobipyramids (the core), mesoporous silica nanoparticles (the middle layer), and hyaluronic acid functionalized with α -cyclodextrin and azobenzene. Inside the silica pores anti-cancer drug—doxorubicin is loaded. Experiments carried out for human squamous carcinoma cells (representative cancer cells) and human keratinocyte cell (representative normal cells) revealed that these nanocomposites are able to specifically accumulate around the tumor tissue due to noncovalent interactions between hyaluronic acid and CD44 receptor overexpressed in cancer cells. Localized irradiation with near-infrared light (780 nm) converts *cis*-azobenzene to its *trans* isomer what leads to hydrogel formation due to noncovalent interactions between α -cyclodextrin and *trans*-azobenzene. Thanks to the presence of specific enzyme—hyaluronidase around the tumor cells the network in the hydrogel is degraded resulting in the anticancer drug release and its transport to the cancer cell nuclei.

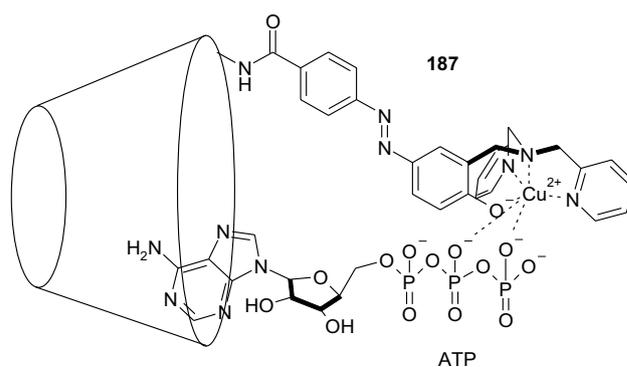


Fig. 84 Complex of cyclodextrin derivative **187** with ATP [250]

The association of artemisinin (ART) **188** with an azobenzene bridged bis(β -CD) derivative with an azobenzene 6–6' linker **189** in aqueous solution was investigated by circular dichroism (CD) spectroscopy (Fig. 85) [252].

It was shown that bis(β -CD) with *trans*-azobenzene unit binds artemisinin (1:1 complex) and this process can be light controlled. Upon irradiation at 363 nm *trans*-*cis* isomerization causes loss of the binding ability of artemisinin.

Polymers bearing macrocycle(s) and azo motif(s)

Azo derivatives found a vast range of applications in polymer science. Polymers containing azobenzene moiety have been intensively studied due to their photoresponsive properties ensuring the obtainment of functional materials. An azo group can be a part of a supramolecular system in polymer matrix (non-covalent interactions) or can be covalently bound within a polymer chain. Polymers, responding to light irradiation are widely investigated systems due to reversible (or irreversible) changes of physical properties [253]. This can be utilized in many branches of science.

The change of the polymer properties can be very often achieved by using molecules, which act as molecular containers. Elegant molecules of such properties are cyclodextrins, described earlier, which can be also used for design and synthesis of functional polymers.

A system utilizing noncovalent interactions between synthesized in a click reaction AZO- β -CD (Fig. 86), which interacts as a “dimer” with azo bearing polyester obtained in reaction of ϵ -caprolactone with *p*-aminoazobenzene (AZO-PCL) (Fig. 86a), was described by Ma et al. [254]. It was suggested that in aqueous solution micellar aggregates are formed due to host–guest interaction between (AZO- β -CD) and AZO-PCL (Fig. 86b). On the basis of ^1H NMR spectra it was suggested that the guest molecule in its *trans* form is included shallowly into cyclodextrin cavity from its wider site. After UV-light irradiation the transparent opalescence solution becomes turbid, what is a result of decomplexation followed by disaggregation. The uniform vesicles are reformed upon exposure of the solution to visible light. Authors propose possible use of the system in the control or release of drugs.

Photosensitive hydrogel based on α -CD, dodecyl-modified poly(acrylic acid), and a photoresponsive competitive guest [255] inspired further studies of self-assembling

systems with polymer side chains. Poly(acrylic acid)s (pAA) with $p3\alpha\text{CD}$ and $p6\alpha\text{CD}$ functionalities and pAA carrying azobenzene moieties (pC12Azo), were used for the construction of the photoresponsive system based on polymer–polymer interactions (Fig. 87) [256]. The properties of obtained systems were studied in details among others by steady-shear viscosity (η) measurements. The method was chosen because the interaction of the CD polymers with pC12Azo (formation of inclusion complexes of CD moieties in the CD polymers with side chains of guest polymers) may cause an increase of solution viscosity. The mixture of the $p3\alpha\text{CD}/p\text{C12Azo}$ and $p6\alpha\text{CD}/p\text{C12Azo}$ has shown contrast η changes upon photoirradiation: decrease in the case of the $p3\alpha\text{CD}/p\text{C12Azo}$ mixture, and increase of η for $p6\alpha\text{CD}/p\text{C12Azo}$ mixture. Irradiation with visible light causes the reverse process in the above cases, i.e. η values became similar to those before the UV exposure. The differences in η values were explained by the fact that UV light causes dissociation of inclusion complexes for the $p3\alpha\text{CD}/p\text{C12Azo}$ mixture, and the formation of interlocked complexes for the $p6\alpha\text{CD}/p\text{C12Azo}$ mixture (Fig. 87, bottom).

PEG-substituted CD with an azobenzene residue at the end of the PEG chain (6-Az-PEG600-HyCiO- β -CD) was obtained by Harada and co-workers (Fig. 88) [257]. The photochemically and thermally induced conformational changes in aqueous solutions were studied by 1D and 2D NMR analyses. It was found that at low concentration, 6-*trans*-Az-PEG600-HyCiO- β -CD forms different types of intermolecular, self-inclusion complexes or exists in an uncomplexed form depending on the temperature. An intermolecular complex is formed at high concentration. Regardless of the concentration, irradiation by UV light promotes complexation with the CD including the azobenzene part.

The attaching of azobenzene groups to side-chains of liquid crystalline polymers results in light-controllable polymer materials. Such films and coatings can be applied for example as optical molecular devices. One group of such materials are derivatives of crown ethers bearing residues able to form liquid crystalline (LC) phases. Complex formation by crown ether moiety can lead to the appearance or disruption of supramolecular structures. A series of photochromic azobenzene-crown-containing compounds forming crystalline and nematic phases were described by Shinkai and co-workers [258]. Photochromic crown ether-containing LC homopolymers and copolymers based on azobenzenes were later described also by Bobrovsky and co-workers [259]. The complexation of metal ions by these compounds cause the decrease of clearing temperature and sometimes the transition into the amorphous state. The investigation of the relationship between molecular architecture of this type polymers and their photo-optical properties and phase behavior was the main scope of studies. Bobrovsky and co-workers [260] described among others the synthesis and properties

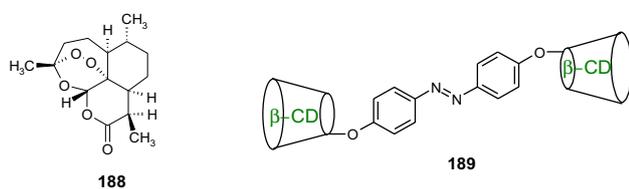


Fig. 85 Azobenzene bridged bis(β -CD) **189** and the formula of artemisinin **188** [252]

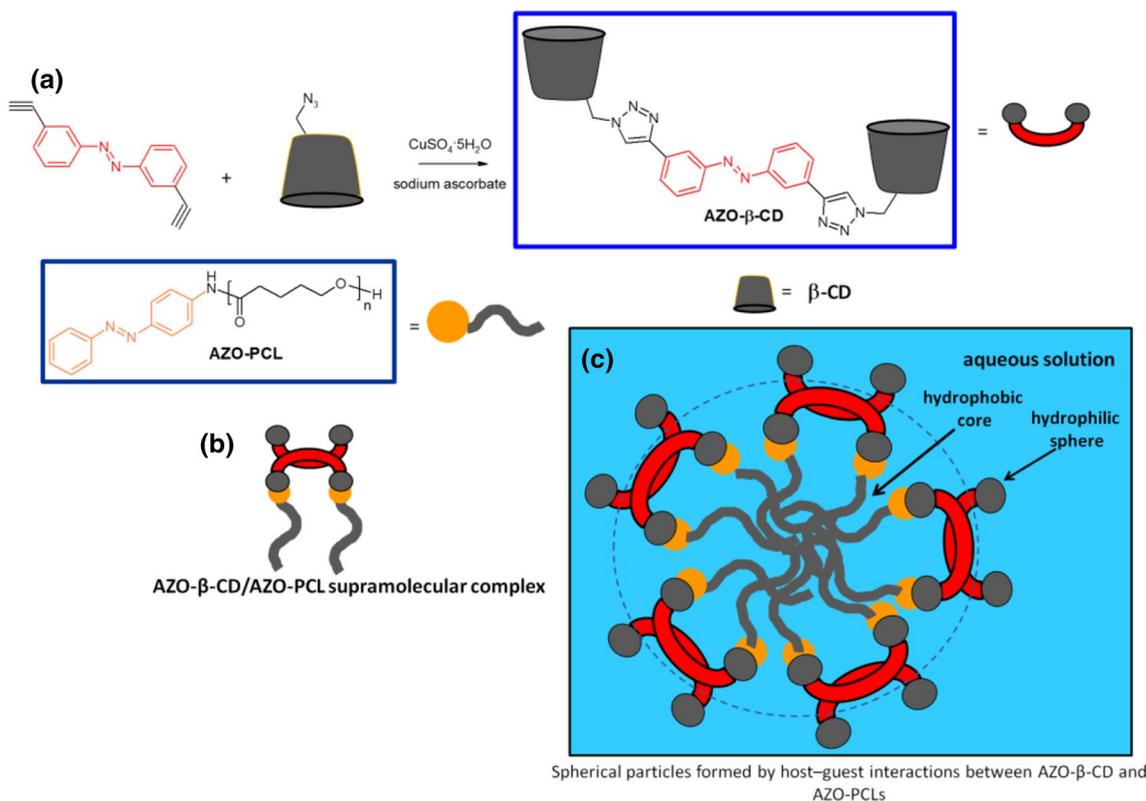


Fig. 86 **a** The synthesis of AZO-β-CD by click reaction, **b** AZO-β-CD/AZO-PCL supramolecular complex, and **c** the possible aggregation mode: spherical particles are formed by host-guest interactions between AZO-β-CD and AZO-PCLs [254]

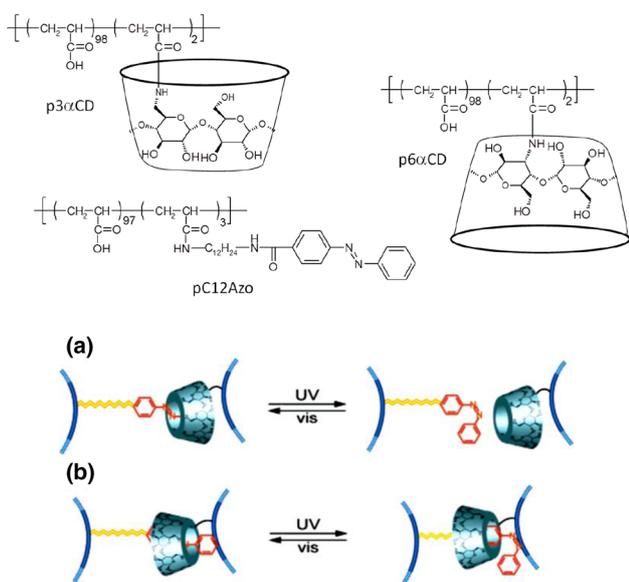


Fig. 87 Top: p3αCD and p6αCD and pAA carrying azobenzene moieties (pC₁₂Azo) used for studies of self-assembly. Bottom: schematic representation for interactions of CD and azo moieties upon irradiation with UV and visible light for **(a)** p3αCD/pC₁₂Azo, **(b)** p6αCD/pC₁₂Azo. Reprinted with permission from [256]. Copyright 2006 American Chemical Society

of two types of polymers differing in the position of the crown ether in relation to the photoresponsive azobenzene residues (Fig. 89). Macrocylic moiety was linked directly to chromogenic residue (Fig. 89, left) or via carboxymethylene spacer (Fig. 89, right). Phase behavior, spectral properties and kinetics of photo-orientation processes inside thin films of polymers shown in Fig. 89 were found to be dependent on the location of crown ether with respect to the residue bearing azo group. In the case where a crown ether was introduced as separated non photochromic side group the decrease of the degree of photoinduced orientational order was found. Complex formation with potassium ions by compound shown in Fig. 89 (left) results in the decrease in degree of the photoinduced order. Possible application in the creation of new sensing materials was suggested.

Zhu and co-workers [261] described linear **190** and **191** and cyclic **192** and **193** (Scheme 23) amphiphilic polymers containing azobenzene moieties. Macrocylic polymers were obtained in Cu(I)-catalyzed azide-alkyne cycloaddition to achieve intramolecular macroring closure process, one of the most popular and powerful “click” synthetic reaction [262].

According to the obtained results, azomacrocycles exhibit increased glass transition temperatures, faster *trans-cis-trans* photoisomerization, and enhanced

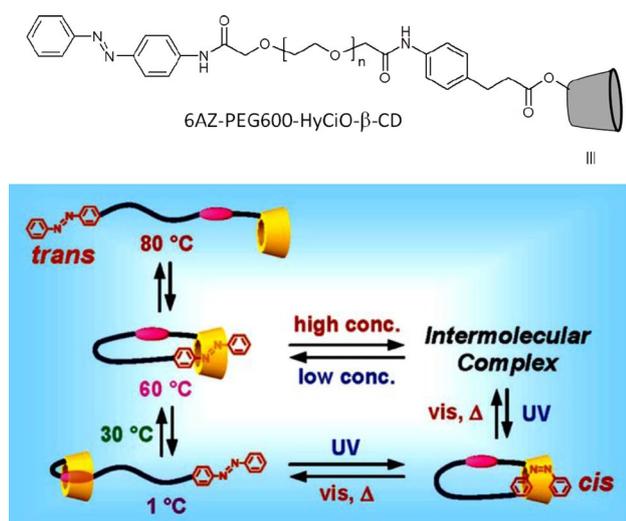


Fig. 88 Top: 6-Az-PEG600-HyCiO-β-CD, Bottom: proposed conformational changes of 6-Az-PEG600-HyCiO-β-CD in aqueous solutions by external stimuli. Reprinted with permission from [257]. Copyright 2007 American Chemical Society

fluorescence intensity in comparison to their acyclic analogues. In water:THF mixture (1:1, v/v) both macrocyclic and linear polymers self-assemble into spherical nanoparticles. The size of aggregates formed by cyclic compounds are significantly smaller than those of corresponding linear analogues due to more dense and compact packing of macrocycle-bearing particles. Alternating irradiation of nanospheres with ultraviolet (365 nm) and visible (435 nm) light causes isomerization of the azo group located in polymer main chain. This induces reversible shift of the hydrophilic-hydrophobic balance of macromolecules and leads to the dissociation and reaggregation of the particles. The photoresponsive behavior is slower for nanospheres containing cyclic polymers than in the case of particles with materials of linear structure.

The skeleton of the photoresponsive polymers also can be enriched with other functionalities that for example are able to form complexes with metal cations.

Wiktorowicz et al. [263] prepared polymers comprising dibenzo-18-crown-6 moieties joined by azo bridges **194**, **195** using reductive coupling procedure (Fig. 90). Spectrophotometric measurements showed that the polymers are pH-sensitive and exhibit solvatochromic properties. Alternating irradiation of the polymers with UV and visible light induces reversible *trans*–*cis*–*trans* photoisomerization. Due to the presence of crown ether cavity, the described polymers interact with Ba²⁺ ions and also with low molar mass pyridinium type guests, leading to complex-induced phase separation in solvents of lower polarity. In alcohols the polymers reveal thermo-responsive behavior exhibiting the upper critical

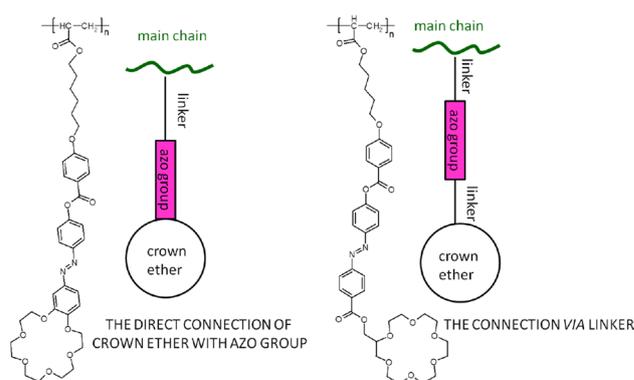
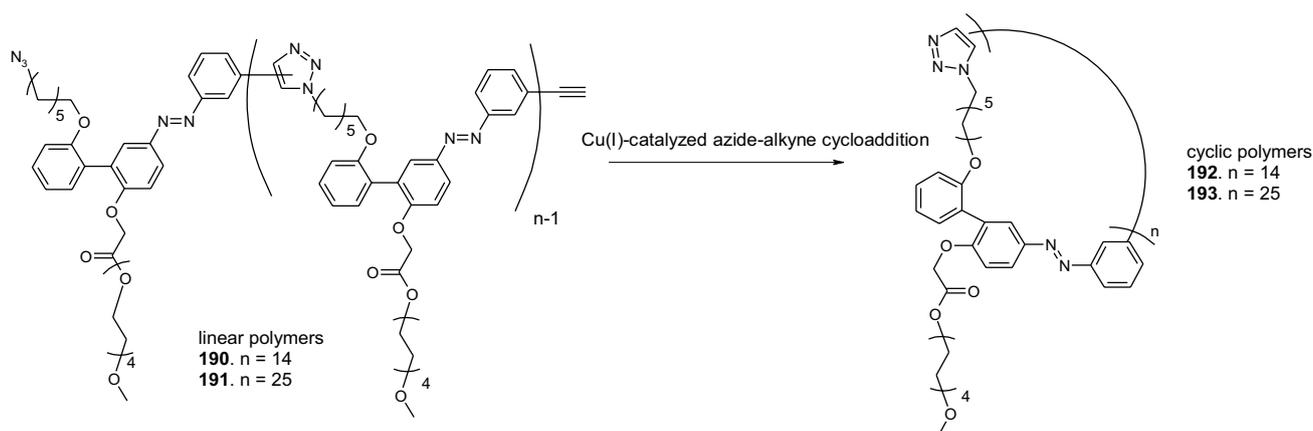


Fig. 89 Photochromic crown ethers-containing LC polymers and their schematic representation [260]

solution temperature type transitions. This effect depends on the polymer concentration and the degree of polymerization.

Cozan and co-workers [264] described the preparation of copoly(ether sulfone)s with azocrown ether and fluorene fragments. The polymers **196**–**200** (Fig. 91) showed good solubility in solvents of different polarity. Thermogravimetric analysis showed the lowest thermal stability of the copolymer **196** among all investigated polymers as it contains only azo-crown ether units that are sensitive to thermal degradation. The insertion of fluorene moieties into a polymer chain significantly enhances thermal stability. The *trans* to *cis* isomerization of the polymers in DMSO occurs after irradiation with UV light (at 375 nm). The rate constant of the first order photoisomerization increases with decreasing the number of azobenzene units. It was also found that complexation of K⁺ inside the macrocyclic cavity increases *trans* to *cis* isomerization rate.

Photo-induced structural transitions of azo compound bearing dibenzo-24-crown-8 (DB24C8) moiety, dibenzylammonium salt (DBA), and 1,2,3-triazole groups were tested by Dong et al. [265]. Due to host–guest interactions between DB24C8 and DBA from separated molecules linear supramolecular polymers of 1:1 threaded structures (pseudorotaxanes) are formed. The presence of azobenzene moiety allows to control the complex formation, as *trans*-azobenzene-appended DBA interacts with DB24C8 stronger than its *cis* isomer. After addition of [PdCl₂(PhCN)₂], 1,2,3-triazole rings of different polymer chains are linked together by the metal coordination, what leads to the formation of cross-linked supramolecular polymers. In dichloromethane the cross-linked assemblies have a form of red gel. UV irradiation (365 nm) of linear and branched polymers induces *trans* to *cis* isomerization, resulting in weaker host–guest interactions and, in a consequence, dissociation of supramolecular polymers. The structural change of cross-linked supramolecular polymer is manifested by naked eye observable decrease



Scheme 23 Linear **190** and **191** and cyclic **192** and **193** photoresponsive polymers described by Zhu and co-workers [261]

Fig. 90 Structure of poly(azodibenzo-18-crown-6-ether)s **194** and **195** [263]

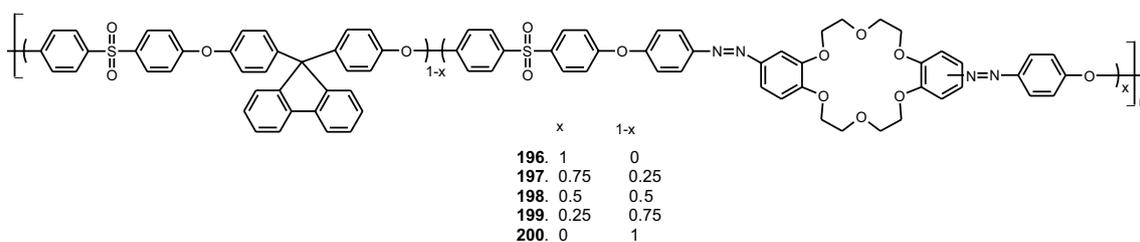
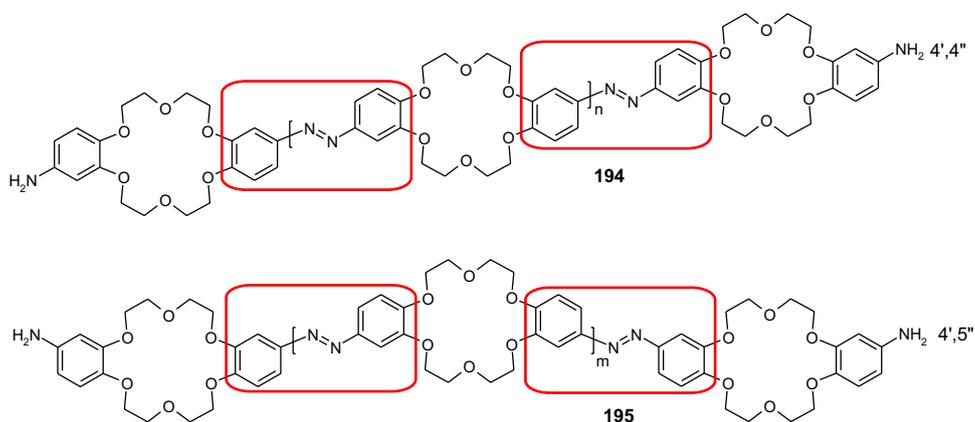


Fig. 91 Copoly(ether sulfone)s with azo dibenzo-18-crown-6 and fluorene fragments [264]

of viscosity. The reformation to the gel state is achieved by exposing the solution to visible light (430 nm).

The photoisomerization of a series of macrocyclic oligomers containing azobenzene moiety in the main chain and their linear analogs was studied by Zhu and co-workers [266] (Scheme 24). Tetraethylene glycol (TEG) was chosen as the building block for the preparation of amphiphilic polymers of good solubility. According to the

UV–Vis spectrophotometry it was shown that the *trans* to *cis* and reverse process are the first order reactions for both linear **201** and cyclic **202** compounds. The estimated values of rate constants for macrocyclic oligomers are distinctly higher (for *trans* to *cis* isomerization) and slightly higher (for *cis* to *trans* isomerization) in comparison with results for linear ones, especially for $n = 1$. This can be

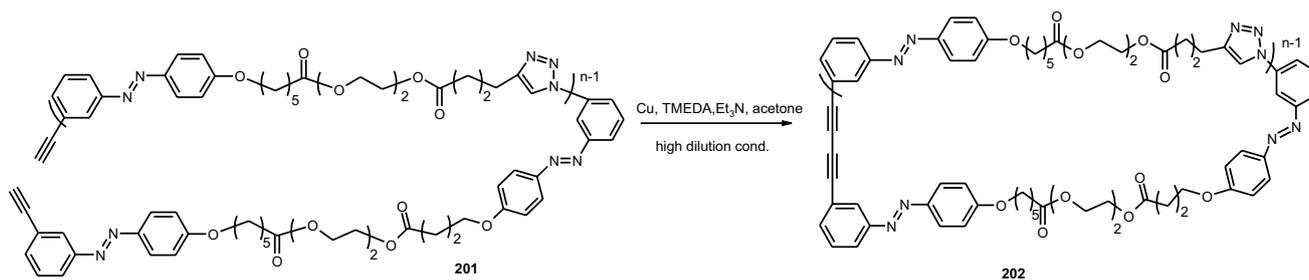
explained by the more stable conformation of cyclic *cis*-azobenzene than linear *trans* analog.

In turn, poly(ethylene glycol)methyl ether was used as building block for other amphiphilic copolymers with cyclic azobenzene unit **203** [267]. For comparative purposes a linear analog **204** was also obtained. The synthetic route is shown in Scheme 25. The obtained copolymers assemble in phosphate buffer solution (pH 7.4) into stable vesicles with hydrophobic blocks containing the azobenzene moieties aggregated in the membranes of the vesicles, and the hydrophilic PEG arrangements on the outer and inner surface of the vesicles. Due to presence of azo moiety the obtained polymers are not only photoresponsive, but also sensitive towards reducing reagents. These properties were used for the investigation of the encapsulation and release of Nile Red (NR- a model compound for drug delivery system) and anticancer drug doxorubicin (DOX). NR-loaded vesicles are fluorescent. The intensity of fluorescence can be controlled by illumination with UV light (365 nm). The reverse *cis*–*trans* process occurs upon irradiation with visible light at 435 nm.

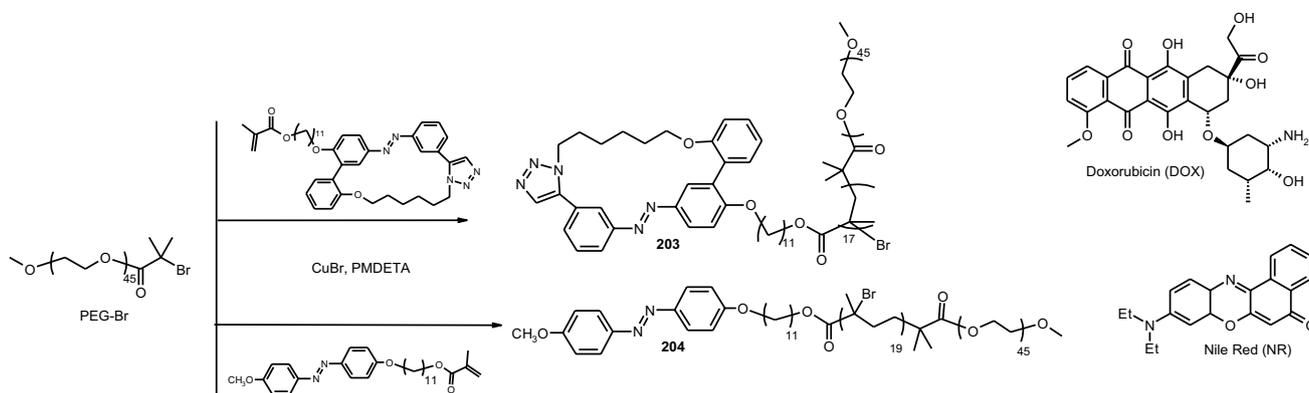
Azo compounds can be reduced by azoreductase or popular reducing reagents, such as for example sodium dithionite. The result is azo bond cleavage, which can be used in effective drug transport. DOX-loaded vesicles were

investigated in reductant-release of the encapsulated substance. The release rate of DOX from cyclic polymer **203** is higher compared with the linear analog **204**. This points out the importance of investigated copolymers—particularly cyclic compounds—as potential agents in the treatment of colon disease.

Combination of cyclodextrin and azobenzene units bearing polymer were used by Winnik and co-workers [268] to obtain molecular “charm bracelets”. The cyclic poly(*N*-isopropylacrylamide) with azobenzene inserted in the main chain **205** (Fig. 92) was synthesized by the “click” ring closure of α -azobenzene ω -azido poly(*N*-isopropylacrylamide) in the form of inclusion complex with α -cyclodextrin. UV-light irradiation of aqueous solution of **205** (at 365 nm) induces motion within the molecule as *cis*-azobenzene unit obtained upon photoisomerization, due to its size, is expelled from the α -cyclodextrin cavity pushing the host to the other sections of the polymeric ring. The *trans* to *cis* photoisomerization does not affect the temperature of phase transition of the polymer, whereas in the case of the analog without cyclodextrin the temperature increases by 1.7 °C. This may be explained assuming that the enhanced polarity due to *trans* to *cis* isomerization in polymer **205** is overshadowed by the strong hydrophilicity of α -cyclodextrin



Scheme 24 The synthesis of the molecularly-defined linear **201** and cyclic oligomers **202** (n = 1–6) [266]



Scheme 25 Poly(ethylene glycol)methyl ether based amphiphilic copolymers with cyclic azobenzene unit **203** and its linear analog **204** [267]. The chemical formulas of doxorubicin and Nile Red are also shown

moiety interlocked along the polymer ring and no change of the phase transition temperature is observed.

Miscellaneous macrocyclic systems bearing azo group(s)

Molecular containers such as pillararenes are promising blocks for building of photoswitchable assemblies. Ogoshi et al. [269] obtained supramolecular polymers consisting of *trans*-azobenzene-bridged pillar[5]arene dimer **206** and bispyridinium cations linked by hexamethylene unit **207** (Fig. 93). On the basis of ^1H NMR spectroscopy it was stated that in dichloromethane at low concentration (2 mM) complexes of 1:1 stoichiometry are formed, in which pyridinium cation moiety **207** is included inside the cavity of **206**. At higher concentration (100 mM) supramolecular assemblies were detected according to DOSY ^1H NMR experiments. Irradiation of diluted solution with UV light induces *trans* to *cis* isomerization of component **206**. At the photostationary state the ratio of *trans* to *cis* isomer is 26:74. At high concentration nearly half of *trans*-**206** does not convert into the *cis* form. Efficient reverse process occurs after exposure to visible light (436 nm). Under equilibrium the ratio of *trans* to *cis* isomer is 93:7.

It was demonstrated that photoisomerization from *trans* to *cis* form weakens the host–guest interactions, probably due to the steric hindrance caused by the *cis* isomer of **206**. As a consequence, the created at high concentration, supramolecular polymers disassembly after UV-light irradiation. Photo-switching between assembly and disassembly of supramolecular system looks completely reversible by alternating irradiation between visible and UV light.

Cavitands **208** and **209** (Scheme 26) bearing azo moiety integrated with macrocyclic [270] structure undergo *trans*–*cis* photoisomerization upon illumination with UV light (365 nm). *Cis*–*trans* conversion proceeds by heating to 164 °C for 5 min or irradiating with 450 nm light for 20 min. *Trans*–*cis* and *cis*–*trans* cycles can be repeated 5 times without degradation of the system. Both the *trans* isomers of **208** and **209** have deep cavities able to bind guest molecules. In fact, **208** and **209** were found to form complexes with small molecules of adamantane series in d_{12} -mesitylene. The highest values of stability constants were found for 1-adamantanecarbonitrile and 2-adamantanone. It was explained assuming the possibility of stabilization of formed complexes by hydrogen bonding and polar interactions with the upper rim of the cavitands. The complexation of adamantane guests can be light controlled, namely irradiation controls uptake and release of guest for **208**.

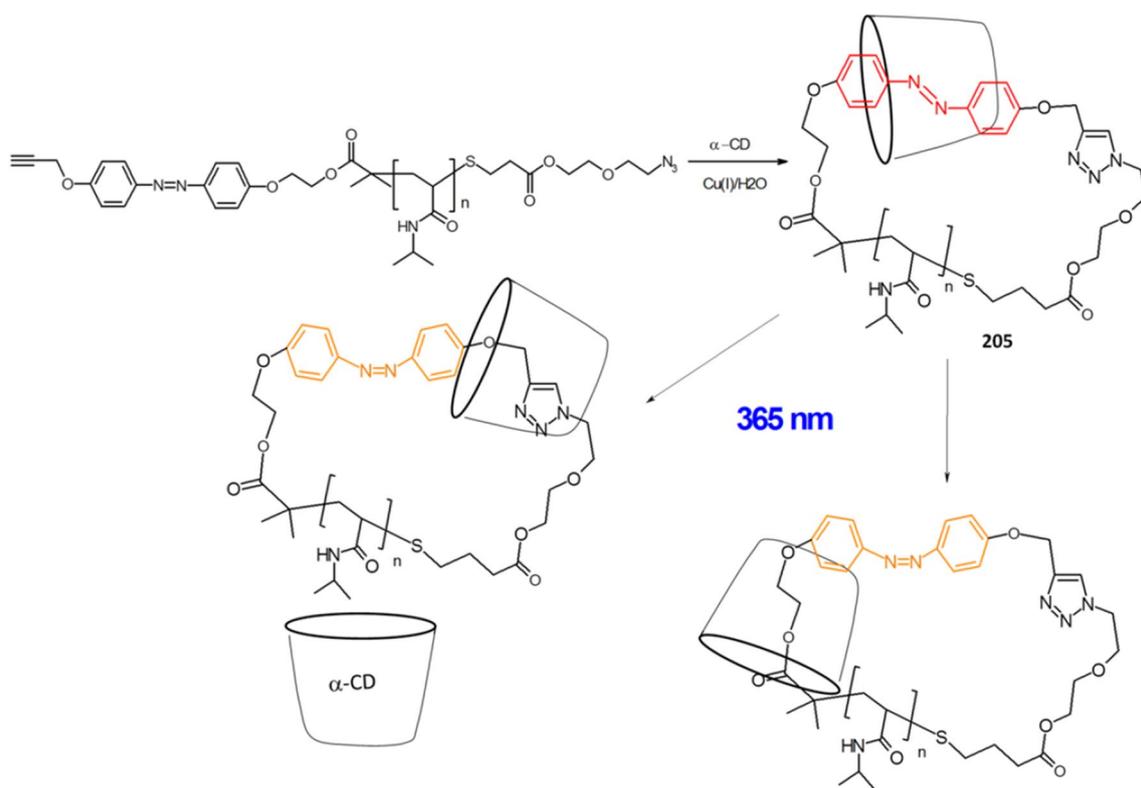
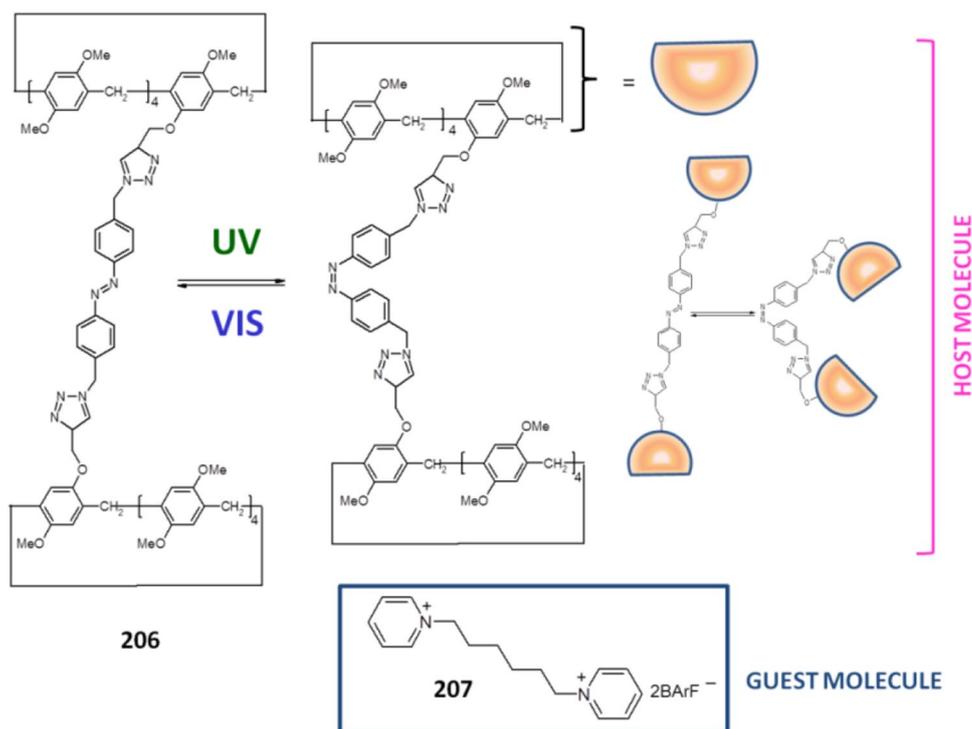


Fig. 92 Cyclic azo-poly(*N*-isopropylacrylamide) with interlocked α -cyclodextrin **205** and photoinduced molecular motion within the polymer [268]

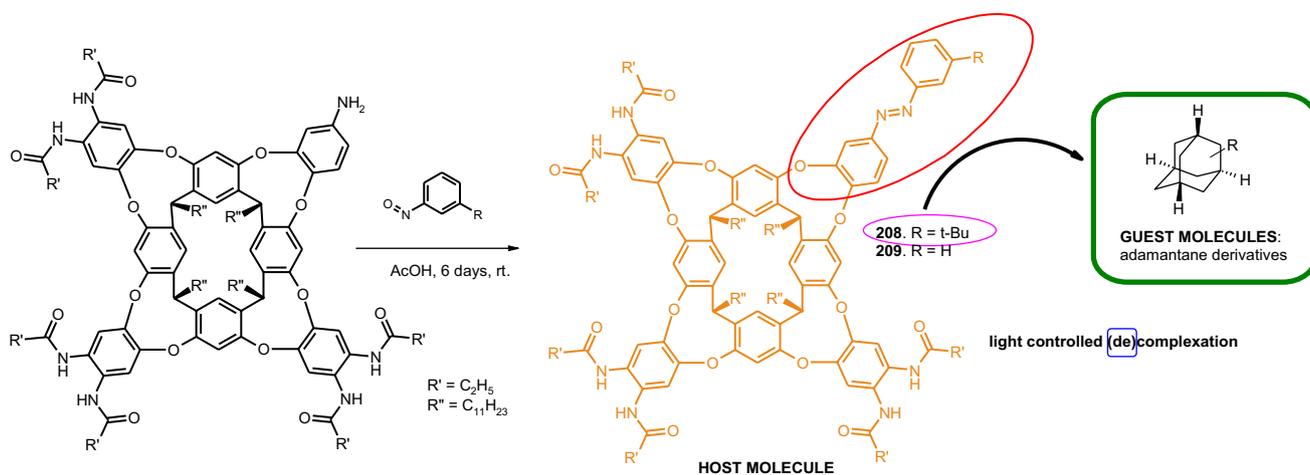
Fig. 93 Chemical structures of the azobenzene-bridged pillar[5]arene dimer **206** and the homoditopic pyridinium guest molecule **207** [269]



Azo moieties can constitute part of macrocyclic Schiff bases, as for example fluorescent product of [2 + 2] condensation of *N,N'*-bis-(2-hydroxybenzaldehyde-5-yl)-benzene-1,3-diazene and benzene-1,2-diamine **210** (Fig. 94) [271]. The stoichiometry of complexes of **210** with zinc(II), copper(II) and nickel(II) is 1:2 (L:M) as showed by elemental analyses and spectral studies. Fluorescence spectra registered in DMSO showed quenching of fluorescence of Schiff base upon metal binding.

Another example of azo derivative of Schiff base type can be chiral macrocycle **211** (Fig. 95), with three azobenzene

residues [272]. This compound was obtained by [3 + 3] condensation reaction of enantiomerically pure *trans* 1,2-diaminocyclohexane with azobenzene-4,4'-dicarbaldehyde in dichloromethane. The subsequent sodium borohydride reduction of **211** produces macrocyclic hexamine **212**, also with three azobenzene units. Irradiation of chloroform solution of (*R,R,R,R,R*)-**211** with 365 nm light for 30 min, causes the decrease of absorption peak at 348 nm and increase of bands intensity at 273 and 450 nm. The back process occurs upon leaving the solution at room temperature for 48 h. Similar observations were made for reduced analog



Scheme 26 Synthesis of azo derivatives of cavitands **208**, **209** and their complexation properties [270]

(*R,R,R,R,R,R*)-**212**. Interesting properties were found for **211** dissolved in benzene. In this solvent a translucent and orange colored gel was obtained. Scanning electron microscopy (SEM) measurements of the obtained material showed the presence of elongated fibers with diameters of around 1 μm in the dried gel. The illumination of the gel with UV light for several hours led to gel–sol transformation. The reverse process occurs upon heating the sol. Macrocycle **212** forms inclusion complexes with various aromatic organic guest molecules. Complexes of 1:1 stoichiometry were found for benzene and toluene as the guests and 2:1 (**212**:guest) when *o*-, *m*-, and *p*-xylenes were complexed. This can indicate better complementarity of the host and benzene or toluene than in the case of larger xylene molecules.

Not always photoinduced transformations are reversible. Red colored compound **213** (Fig. 96) obtained in reaction of 3,3'-dihydroxy-4,4'-bipyridine and

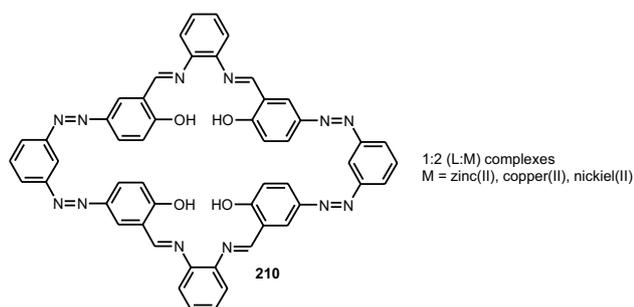


Fig. 94 Macrocyclic Schiff base **210** bearing inherent azo groups [271]

azobenzene-2,2'-dicarboxylic acid in dichloromethane is a highly-strained cyclophane **213** comprising azobenzene and methyl viologen units [273]. Cyclic voltamperometric measurements showed its unique irreversible electrochemical behavior. *Trans*–*cis* isomerisation upon visible light illumination of **213** is also irreversible.

The photo- and redox properties of azo compounds can be extended to more sophisticated systems due to incorporation of transition metal cations into their structure. In such cases both the photoisomerization of azo compounds and chemical and physical properties of transition metal cations (optical, redox, magnetic etc.) can be utilized for construction of functional systems. Among the others, tetranuclear macrocyclic gold(I) alkynyl phosphine complexes with two azobenzene moieties, were obtained (shown schematically in Fig. 97) and investigated as photoswitchable system [274]. It was found that the photo switching of gold(I) complex could be locked or unlocked with a second input: by the addition or removal of silver(I) ions.

The conformational change of the molecule which is a consequence of the reversible *trans*–*cis* isomerization and the red-ox properties of iron are good examples of construction block for multi-stimuli molecular devices. The interlocking of a ferrocene-based rotary module with a photochromic azo unit of molecular machines operating *via* power-conversion mechanisms can be constructed. Such systems resemble daily used devices such as pliers **214** shown in Fig. 98 [275].

Azov et al. [276] investigated macrocyclization of tetrathiafulvalene dithiolates with bis-bromomethylazobenzenes

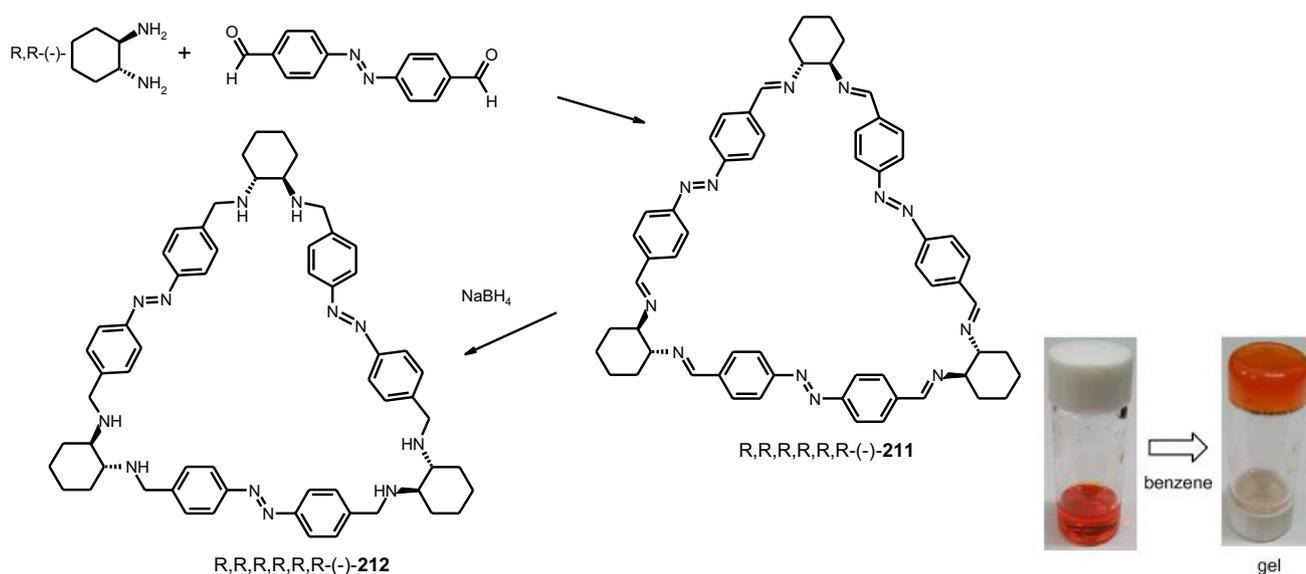


Fig. 95 Left: the synthesis of chiral macrocyclic Schiff base **211** and its reduced derivative **212**. Right: sol–gel transformation of **211**. Reprinted from [272]. Copyright 2010 with permission from Elsevier

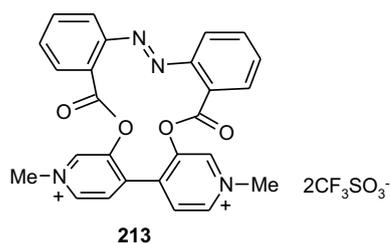


Fig. 96 Highly-strained cyclophane **213** comprising azobenzene and methyl viologen units [273]

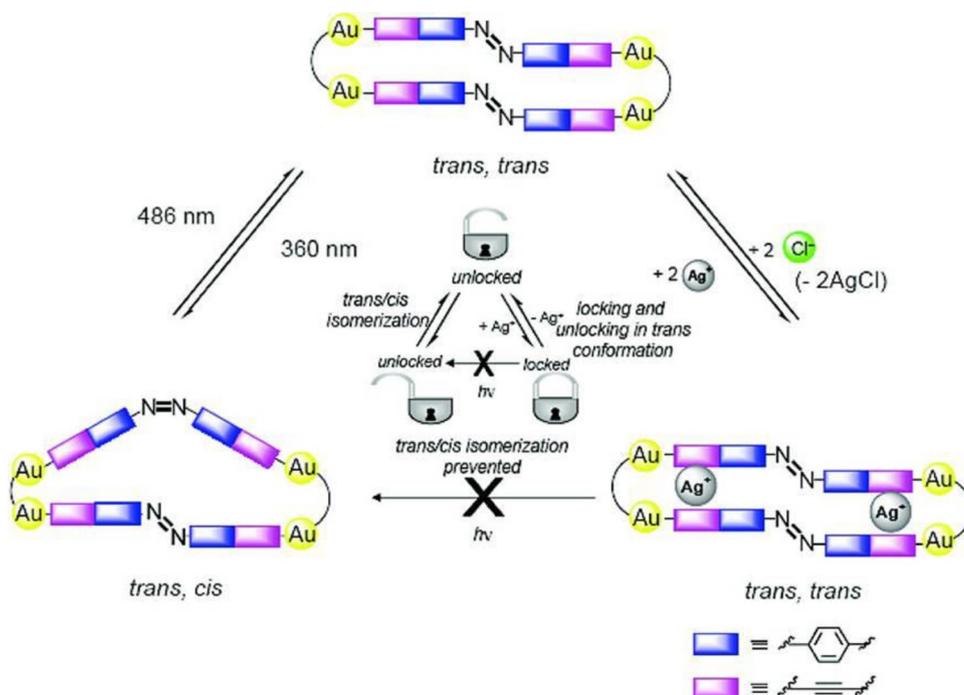
(Ab) under high dilution conditions (Fig. 99). The reaction afforded [1 + 1] cyclization product **215** with *m*-Ab and [2 + 2] cyclization product **216** with *p*-Ab in good yields (above 66%). Irradiation of *p*-Ab with UV light (365 nm, 0 °C) before reaction results in obtainment of *cis*-azobenzene bearing product **217** (1 + 1 cyclization type). Analysis of cyclic voltammograms registered in dichloromethane/0.1 M Bu₄NClO₄ showed that the electrochemical properties of tetrathiafulvalene moiety strongly depend on configuration (*trans* or *cis*) of azobenzene unit.

Banerjee and co-workers [277] synthesized and compared properties of two covalent organic frameworks (COFs) (Fig. 100) being derivatives of triformylphloroglucinol and 4,4'-azodianiline (**Tp-Azo**) or 4,4'-diaminostilbene (**Tp-Stb**). Azo-functionalized COF **Tp-Azo** exhibits better stability, porosity and crystallinity than stilbene-bearing analogue **Tp-Stb**. The analysis of N₂ absorption isotherm of **Tp-Azo** treated with 9 M HCl indicates the retention of

intrinsic porosity of the azo-functionalized COF, whereas in the case of **Tp-Stb** decrease of porosity after the acid treatment was observed. According to TGA experiments it was stated that **Tp-Azo** possesses higher acid loading (5.4 wt%) than **Tp-Stb** (2.8 wt%). Doping of H₃PO₄ to the azo-functionalized COF leads to immobilization of the acid inside the framework pores, what enables proton transfer in both the anhydrous ($\sigma = 6.7 \times 10^{-5} \text{ S cm}^{-1}$ at 340 K) and hydrated state ($\sigma = 9.9 \times 10^{-4} \text{ S cm}^{-1}$ at 332 K under 98% relative humidity). Stilbene-bearing COF shows almost zero proton conductivity in anhydrous milieu and a poor proton conductivity value ($\sigma = 2.3 \times 10^{-5} \text{ S cm}^{-1}$) at 332 K under 98% relative humidity.

Azobenzene isomerization has been also utilized to drive functional changes in biomolecules such as: peptides, proteins, lipids, nucleic acids and carbohydrates. Comprehensive review of such applications can be found in the work of Beharry and Woolley from 2011 [72]. To apply azobenzene to direct protein conformational change in biological systems several requirements need to be fulfilled such as: (i) substantial structural change of azo bearing unit upon isomerization that can be coupled to protein conformational change, (ii) stability of the azo unit in a cellular environment, (iii) a suitable for cells and tissues irradiation wavelength and rate of thermal relaxation. Photocontrol of cyclic peptides was investigated among the others by Schutt et al. [278]. The authors described cyclization of a heptapeptide containing the Arg-Gly-Asp (RGD) sequence with 4-aminomethylphenylazobenzoic acid (AMPB). Studies of the cyclic peptide affinity to the cell surface receptor $\alpha\text{V}\beta\text{3}$ integrin

Fig. 97 Schematic diagram demonstrating the “locking” and “unlocking” mechanism brought about by the addition and removal of Ag⁺ ions in preventing and facilitating *trans*–*cis* isomerization of [Au₄(PⁿP)₂(C≡C–L–C≡C)₂]. Reprinted with permission from [274]. Copyright 2007 American Chemical Society



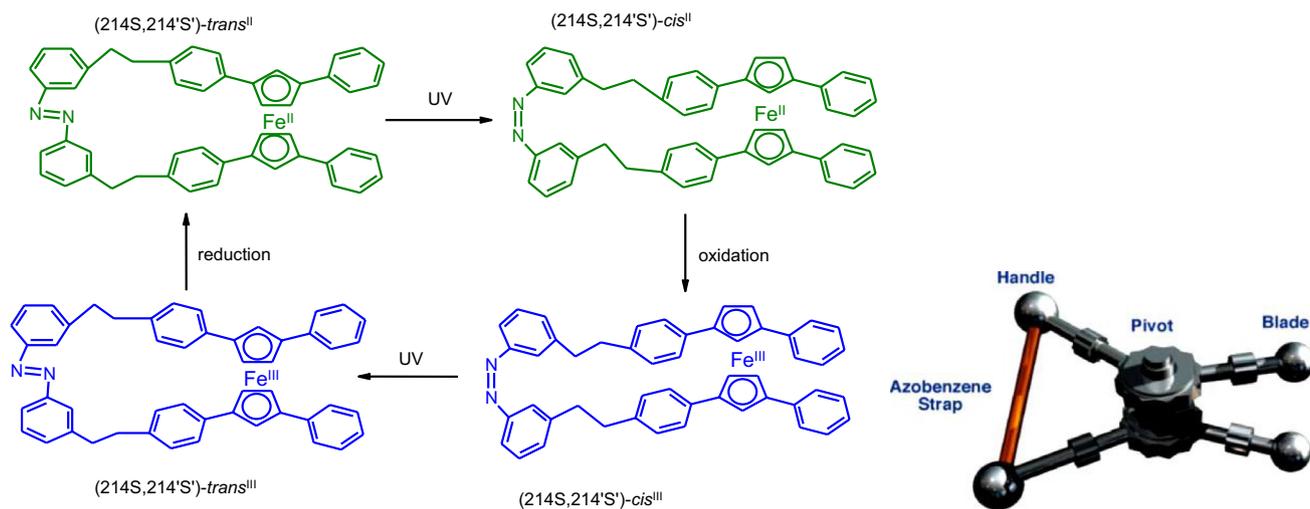


Fig. 98 Left: the operation of molecular pliers **214** by light and redox stimuli. Right: schematic illustration of molecular pliers. Reprinted from [275]. Copyright 2008 with permission from Royal Society of Chemistry

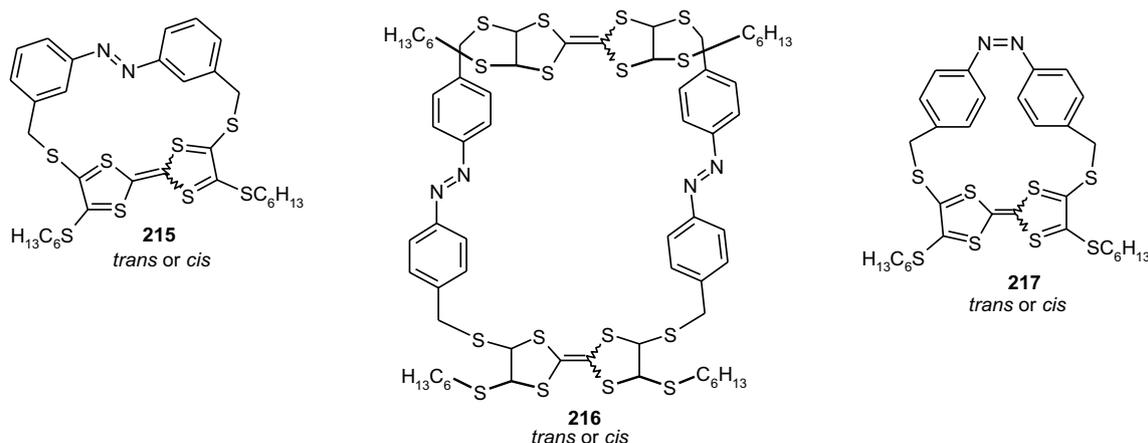


Fig. 99 Macroyclic azocompounds bearing tetrathiafulvalene units **215–217** [276]

show that RGD binds target protein stronger when azo unit is in its *trans* form. The cell adhesion can be also controlled by tethering of RGD peptide to a surface via azobenzene linker. 3-((4'-aminomethyl)phenylazo)benzoic acid was used to control the conformation of a cyclic peptide based on nNOS β -finger [279]. The *trans* isomer shows binding affinity towards target protein— α -1-syntrophin. Irradiation of the system with light at 330 nm enables the protein recognition. According to FTIR and NMR experiments, the isomerization azobenzene unit induces the formation of secondary, antiparallel β -type structure of the peptide ensuring the efficient interactions with α -1-syntrophin. Incorporation of azobenzene unit into protein disulfide isomerase *via* bis-cysteinyll active site was used to the obtainment of a simple model for allosteric conformational rearrangements [280]. It was stated, that the geometric changes accompanying

isomerization of the azo group induce a rearrangement of peptide sequence changing energy landscape of the peptide and both isomers *trans* and *cis* exist in defined conformational states stabilized by disulfide bridge. Derda et al. [281] proposed bis(allenamide) functionalized azobenzene reagents for conversion of cysteine containing peptides to light responsive macrocycles. In comparison with typically used bis-alkyl halides containing azobenzene the allenyl amide derivatives ensure 2–3 order of magnitude faster macrocyclization by cysteine ligation in model peptide and those displayed on M13 phage. Woolley and co-workers [282] incorporated a thiol reactive azobenzene cross linker **218** into peptide backbone receiving cyclic azopeptides **219–221** (Fig. 101). Upon irradiation of the peptides in aqueous solution with blue light at 400–450 nm *trans* to *cis* isomerization occurs. Obtained *cis* isomers relax thermally with a half-life

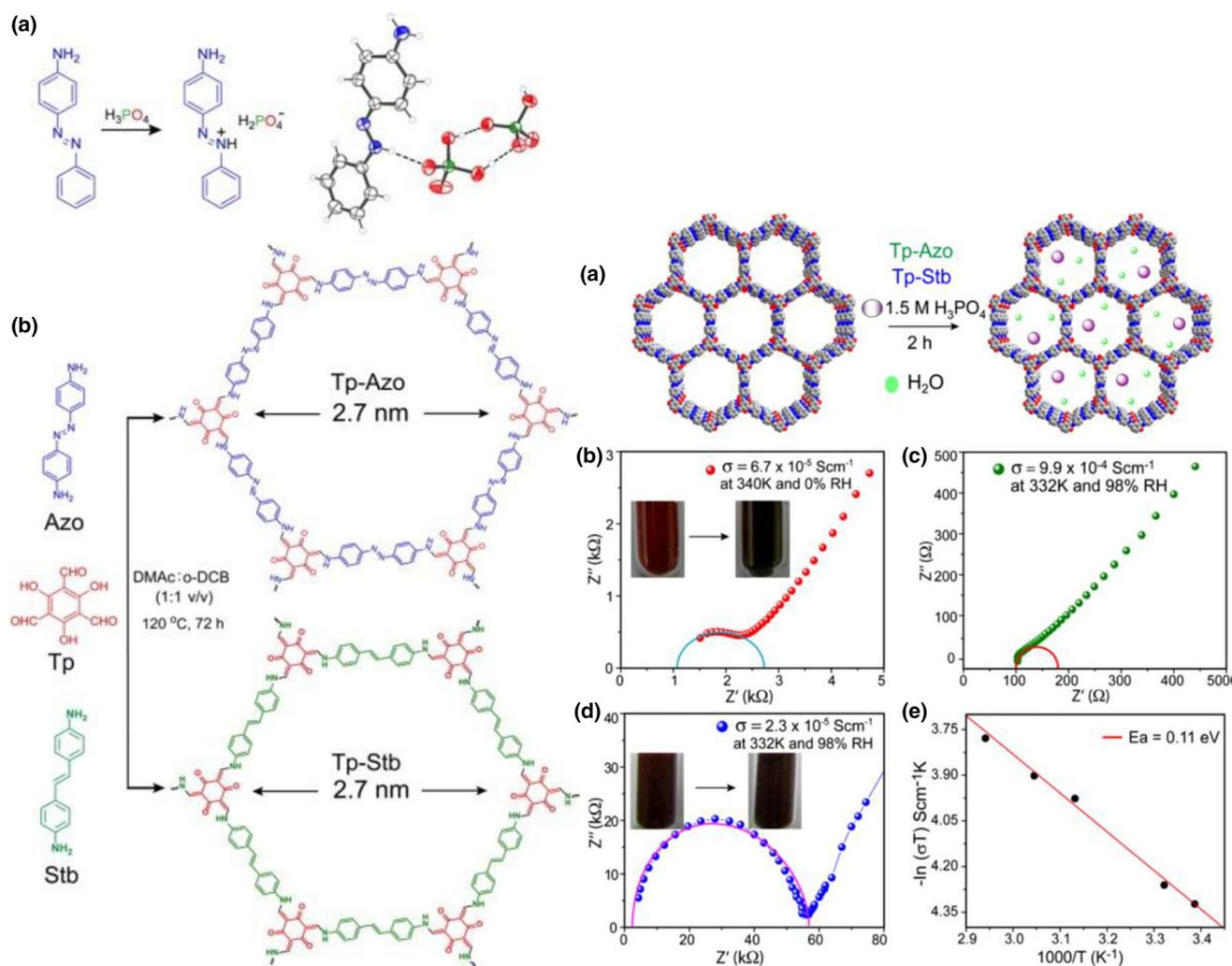


Fig. 100 Left: **a** crystal structure of 4-[(*E*)-phenyl-diazenyl]anilinium dihydrogen phosphate. **b** schematic of Tp-Azo and Tp-Stb synthesis. Right: **a** schematic of H₃PO₄ doping of COFs. Proton conductivity of PA@Tp-Azo in **b** anhydrous and **c** wet conditions. **d** Proton con-

ductivity of PA@Tp-Stb in wet conditions. **e** Arrhenius plot for PA@Tp-Azo in hydrous conditions. Reprinted with permission from [277]. Copyright 2014 American Chemical Society

of about 1 s. It was stated, that azobenene linker **218** can be used to control of helical content of attached peptide, as in its *trans* form the linker bridges Cys residues spaced $i, i + 15$ (peptide **219**) in an α -helix. Switching **218** to its *cis* isomer causes the decrease of the helix content of **219** and the increase of the helix content of **221**. After photoisomerization no helix content of **220** is detected.

Jaeschke and co-workers [283] described carbohydrate-based macrocycles obtained from isothiocyanate-armed bis-azobenzene glycosides and piperazine. Isomerization of glycoazobenzene precursor molecules before the reaction ensured more efficient macrocyclization (yields: 48–65%). Obtained *trans* macrocycles isomerize into their *cis* forms upon UV-light irradiation what results in tremendous change of chirality with a strong helical induction in the *cis* state. The isomerization process is fully reversible by thermal

relaxation, whereas upon irradiation with blue light only partially recovery of *trans* isomer is obtained.

Summary

The above review article is a subjective point of view on the current state of art in the synthesis and properties of selected azomacrocyclic compounds. It covers mainly the last 10 years, however, many of the previous works were also cited, to give more comprehensive background of the subject. Our intention was to underline the importance of very simple, seemingly tiny, functional –N=N– group, which can be incorporated into almost any molecule (material) giving extraordinary properties, especially when macrocyclic compounds are regarded. The presence of macrocyclic scaffold

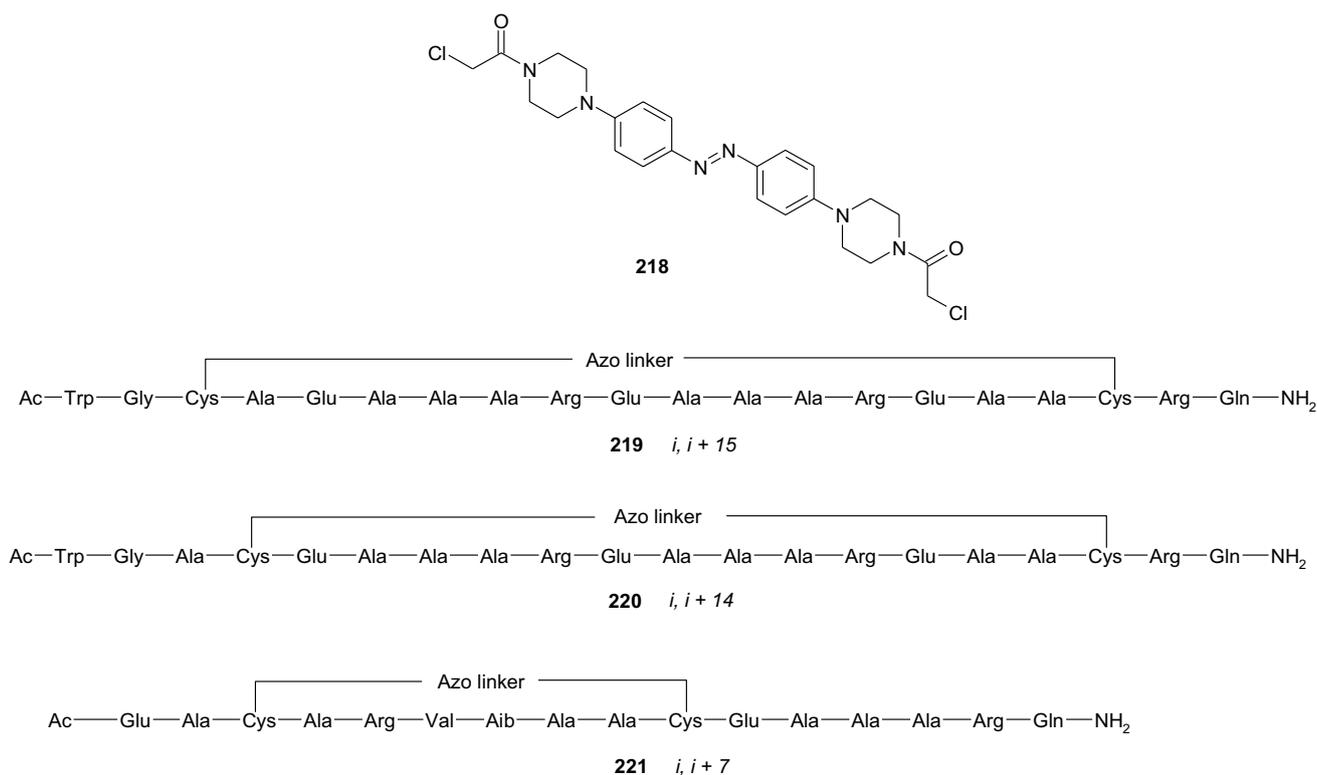


Fig. 101 Azobenzene cross linker **218** and primary sequence of the cross linked peptides **219–221** described by Woolley and co-workers [282]

can have an enormous influence on switching properties of azo group due to ring strain and substituent effects. Photochemical characteristic of cyclic azobenzenes depends also on other factors, such as the number of azo units in the macrocycle, the symmetry of total molecule and the degree of conjugation, what makes the design of azomacrocyclic compounds a challenging task. Reversible *trans*–*cis* isomerization gives an opportunity to control the macrocycles structures at the molecular level what can be utilized for instance in the development of light-induced assembly/disassembly processes of supramolecular systems or in morphological transformation of assemblies. Binding properties of macrocyclic hosts e.g. crown ethers or cyclodextrins can be regulated by photoswitching of azo moiety, what finds applications among the others in ion transport through membranes and controlled drug release systems. Chromogenic and electroactive properties of azo group enable effective macrocycle use in optical and electrochemical sensors development. In the above manuscript we wanted to signalize the multifarious areas of science, technology and medicine where macrocyclic azo compounds can find applications. We believe the review will be helpful for readers interested in organic, analytical and practical aspects of supramolecular chemistry.

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