

## Deep eutectic solvents – A new platform in membrane fabrication and membrane-assisted technologies

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17 **Abstract**

18 Deep eutectic solvents (DESs) are a new class of solvents that can offset some of the primary drawbacks of typical solvents  
19 and ionic liquids (ILs). Since DESs fall into the guidelines of “Twelve Principles of Green Chemistry”, their implementation  
20 in several types of applications has exponentially increased over the last years. The usage of DESs has been directed to  
21 the designing, manufacture and purification of new materials and feedstocks. Very recently, great attention has been paid  
22 to new pioneering attempts aiming at DESs into the field of chemical engineering, including membrane science and  
23 technology. Even if just a few works have been currently reported in applying DESs in membranes, the consideration on  
24 this new type of solvents is continuously growing. This review compares and discusses the documented discoveries and  
25 breakthroughs carried out in applying DESs in membrane science. The scope of this review is emphasized in various  
26 scopes: i) new sustainable membrane preparation, ii) membrane-based technologies aided by DESs, iii) target molecules-  
27 DES interactions and iv) new membrane-DES structures providing enhanced physicochemical properties and thus  
28 separation performance. Here, besides the relevant insights in the field, we give the key hypotheses and strategies used  
29 by the scientists to reach a successful merging of both areas since the use of DESs in membranes is still challenging due  
30 to the compatibility issues of the DESs and membrane phase (either polymer and inorganic). The future directions and  
31 perspectives on using DESs in membranes are also given.

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34 **Keywords:** Green chemistry; deep eutectic solvents; membrane fabrication; membrane science and technology; separation  
35 techniques.

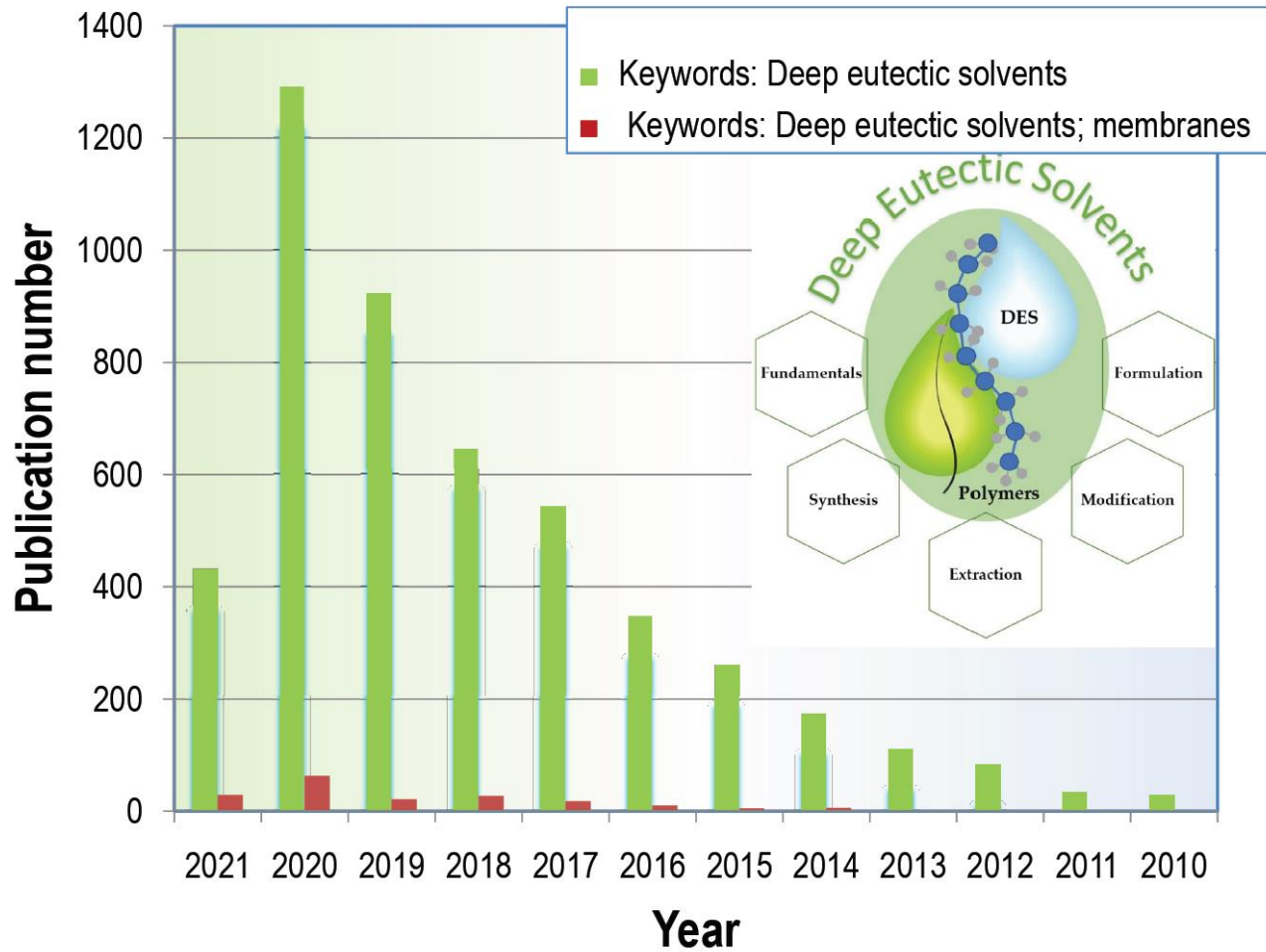
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## 37 1. Introduction

38 Deep eutectic solvents (DESs) can be considered as a new class of solvents which have been recognized as "green  
39 alternatives" to conventional solvents and ionic liquids (ILs) [1]. Their area of application is constantly growing and evolving,  
40 encompassing different applications including metal plating and coatings [2], sustainable media in organic reactions [3],  
41 extraction of biomolecules and chemical elements (Zn, Fe, Pb, etc.) from natural sources [4,5], analytical chemistry [6],  
42 removal of chemical contaminants from water and foods [7], CO<sub>2</sub> capture [8], desulfurization [9], formulation of stationary  
43 phases for chromatography [10], enzymatic biodiesel production [11], chemical catalysis [12], biotransformations [13],  
44 among others. Since the first introduction of choline chloride-based DES in 2003 by Abbott and co-workers [14], the  
45 implementation of DESs in modern chemistry has been exponentially explored as reflected in the growing number of  
46 publications on the topic (see **Figure 1**), demonstrating an increasing research interest in this field.

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**Figure 1.** Publications related to the applications of DES over past 10 years (until April 15<sup>th</sup>, 2021; source: Scopus).

Studies in red are those related to the application of DESs in membranes.

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52 Considering the “*Twelve Principles of Green Chemistry*” established by Anastas and Warner [15], DESs fully meet some of  
53 the proposed criteria (e.g. less hazardous synthesis, use of safer solvents and auxiliaries) fostering their use in various  
54 areas of research, along with their primary advantages compared with traditional solvents, including low toxicity and cost,  
55 easy handling, biodegradability, biocompatibility and reusability [16].

56 In addition to this, there is a current need of producing eco-friendly materials for either feedstock and technologies. To some  
57 extent, membrane technologies demand more environmentally friendly materials for the synthesis of membranes. To date,  
58 it is likely that chemically synthesized polymers are among the main materials used for membrane manufacture, which  
59 represents an environmental issue due to their low biodegradability and their fossil-based origin. Even if the use of DESs in  
60 membrane science and technology, as testified by the data reported in **Figure 1**, is still at a seminal stage, the number of  
61 works in the field is constantly growing. In principle, a solvent, as a dissolving phase in the dope polymer solution, plays a  
62 crucial role in membrane manufacturing, determining changes in membrane properties (mainly structure) and consequently  
63 influencing its separation performance [17,18]. The use of DESs in membrane preparation can be, therefore, considered as  
64 an emerging and valid option in order to impart specific physicochemical properties to the membranes. Besides analysing  
65 the ongoing progresses that have been made on the use of DESs as solvents or additives in membrane preparation [19],  
66 this review also focuses on the various uses of DES in the post-modification treatments of membranes, in DES-liquid  
67 supported membranes and in DES-assisted membrane processes [20]. Interestingly, DESs are able to improve the

68 membrane's separation performance thanks to a facilitated molecule transport and/or an adsorption mechanism occurring  
69 through the functional groups borne in their chemical structure [21–23]. However, the compatibility of DESs and polymer  
70 phases aimed to the preparation of homogeneous membranes with a dense morphology, is still a challenge. Therefore, this  
71 review also illustrates the guidelines of using DESs for dense membrane fabrication, together with the advantages and  
72 drawbacks of such membranes when attending specific selective gas and solvent separations. To finalize, brief feedbacks  
73 on concepts and synthesis of DESs are also given, along with the core interactions identified by researchers between target  
74 molecules and DES phases for enhanced target molecules separation.

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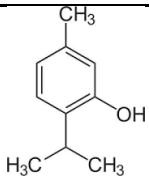
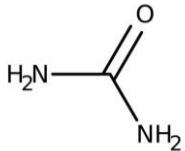
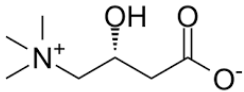
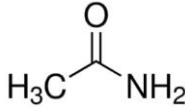
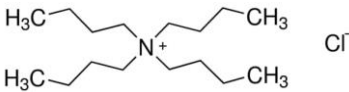
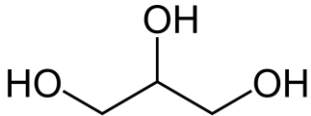
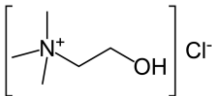

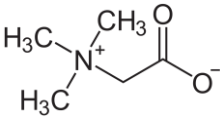
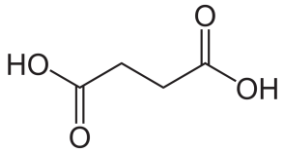
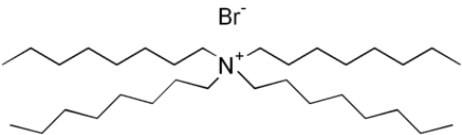
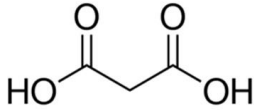
## 76 **2. Brief background on the properties of DESs**

77 DESs and ILs are considered as new types of green solvents. They share, in fact, several properties and advantages even  
78 if they are chemically quite different [24]. A typical DES is defined as a liquid system formed from a eutectic mixture of Lewis  
79 or Brønsted acids and bases. From a chemical perspective, DESs are synthesized by combining a hydrogen bond acceptor  
80 (HBA), like quaternary ammonium salts, with a hydrogen bond donor (HBD) compound (e.g., carboxylic acids, alcohols,  
81 sugars, amines, etc.) [25]. **Table 1** enlists a few typical examples of HBAs and HBDs agents used for the synthesis of DESs.

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**Table 1.** Typical HBDs and HBAs agents used for the synthesis of DESs [1,26,27].

HBA	Chemical structure	HBD	Chemical structure
Thymol		Urea	
L-Carnitine		Acetamide	
Tetrabutylammonium chloride		Glycerol	
Choline chloride		Ethylene glycol	
Trimethylglycine		Succinic acid	
Tetraoctylammonium bromide		Malonic acid	



85

86 DES classification is largely depending on the nature of the element compounds and they are generally categorized in four  
87 types, as detailed in **Table 2** [1]. For example, DESs based on  $MCl_x$  and quaternary ammonium salts (*type I*) are recognized  
88 as an analogous type to the metal halide/imidazolium salt systems. When DESs present hydrated metal halides and choline  
89 chloride, they are recognized as *type II*. The low cost of the hydrated metal salts together with their inherent air/moisture  
90 stability makes them a good option in large-scale applications. Regarding *type III* DESs, they are formed from choline  
91 chloride and HBDs, which make them attractive due to their ability to solvate a number of transition metal species, including  
92 chlorides and oxides. To date, *type III* DESs are likely to be the most studied, presenting low cost, non-toxicity and  
93 biodegradability [28]. If transition metals can be incorporated into ambient temperature eutectics, these DESs are termed  
94 *type IV*. Particularly, metal salts would not commonly ionize in non-aqueous media but  $ZnCl_2$ , for instance, has been  
95 demonstrated to form eutectic mixtures with urea, acetamide, ethylene glycol, among others.

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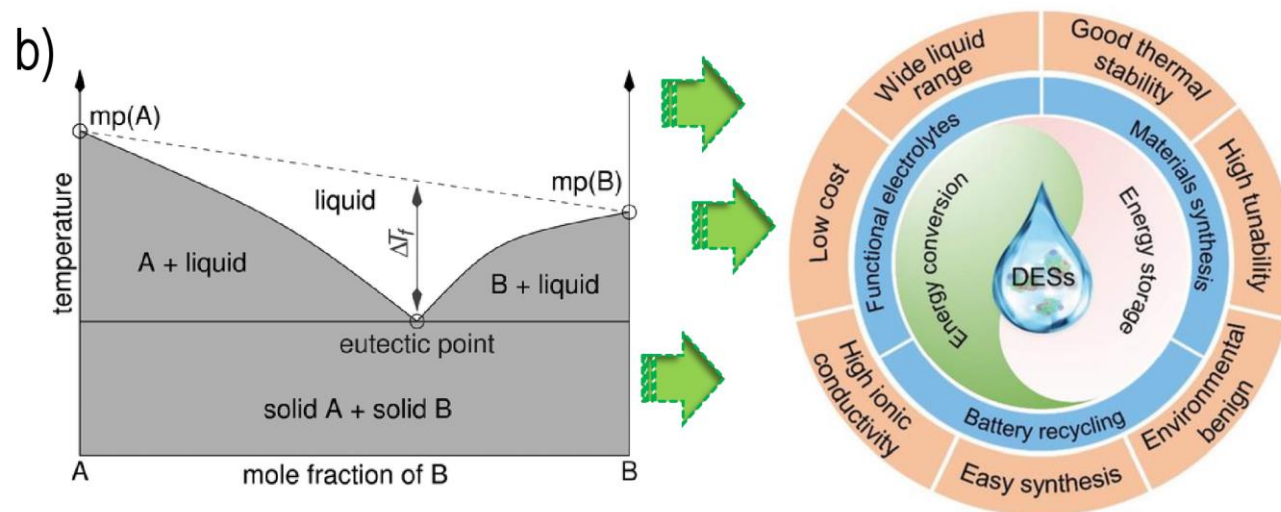
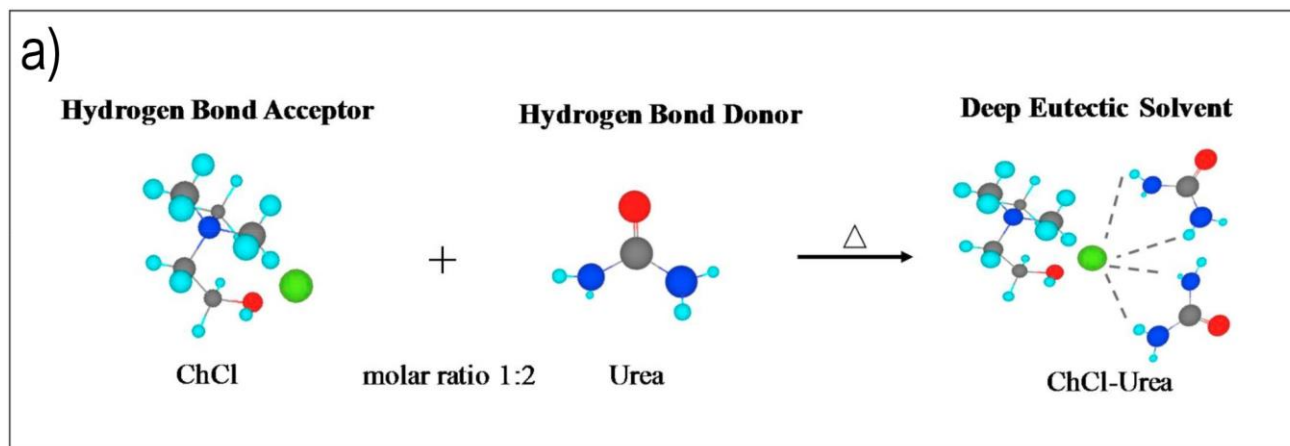


**Table 2.** Classification, general formula and terms of DESs [1].

Type	Formula	Terms
Type I	$\text{Cat}^+ \text{X}^- z\text{MCl}_x$	M = Zn, Sn, Fe, Al, Ga, In
Type II	$\text{Cat}^+ \text{X}^- z\text{MCl}_x \cdot y\text{H}_2\text{O}$	M = Cr, Co, Cu, Ni, Fe
Type III	$\text{Cat}^+ \text{X}^- z\text{RZ}$	Z = CONH <sub>2</sub> , COOH, OH
Type IV	$\text{MCl}_x + \text{RZ} = \text{MCl}_{x-1}^+ \cdot \text{RZ} + \text{MCl}_{x+1}^-$	M = Al, Zn Z = CONH <sub>2</sub> , OH

103 In DESs synthesis, hydrogen bonding is the main responsible to chemically join a HBA and a HBD [29], as graphically  
 104 represented in **Figure 2a**. For instance, the new self-associated mixture displays a eutectic phase characterized by melting  
 105 point value ( $T_m$ ) lower than that of each forming molecule (see **Figure 2b**). This is a quite well-known phenomenon as the  
 106 eutectic point represents the lowest temperature among all the individual compositions. Experimentally, most of the DESs

107 reported in literature display a  $T_m$  lower than 100°C [30]. Such a phenomenon occurs when two compounds with different  
108 molecular sizes are merged. Apart from this property, **Figure 2b** also describes some other interesting features of the new  
109 eutectic mixture, including good thermal stability, high tunability, high dissolubility, good chemical stability, among others.  
110 The physicochemical properties (such as acidity, viscosity, gas solubility, density, hydrophilicity/hydrophobicity,  $T_m$ , volatility)  
111 of DESs are greatly dictated by various factors, such as the type of HBA/HBD, the synthesis protocol and the operation  
112 parameters including the molar ratio, the temperature and the mixing time.



**Figure 2.** a) Graphical drawing of typical DES formation [30], and b) main features and properties of DESs including eutectic point on a two-component phase diagram [1].

117 In a polymer membrane containing DESs, both components (i.e., polymer and DES) can interact resulting in the following  
118 scenarios: i) the DES could act as a solvent only, ii) any of the DES elements (HBA or HBD) can be part of the  
119 polymerization, and iii) the DES can provoke significant changes in polymer properties, such as surface modification,  
120 morphological and structural changes [31]. At this point, the nature of DESs (i.e., hydrophilicity and hydrophobicity) is  
121 relevant when merged in membranes since specific membrane technologies (such as pervaporation, gas separation) require  
122 membrane surfaces with a hydrophilic or hydrophobic nature to differentiate polar and non-polar molecules [32,33]. As an  
123 example, hydrophilic pervaporation membranes preferentially transport polar solvent molecules (including water) hindering  
124 the permeation of less polar (or non-polar) molecules. On the contrary, hydrophobic pervaporation membranes preferentially  
125 transport less polar (or non-polar) molecules, retaining polar molecules. Here, the application of DESs and their nature  
126 should be smartly selected in order to accurately tune membrane properties facilitating the permeation of the target  
127 molecules.

128 Therefore, the next section is devoted to the different strategies so far employed in applying DESs in different scenarios of  
129 membrane technologies.

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### 131 **3. Breakthroughs on DES applications in membrane fabrication and DES-assisted membrane technologies.**

#### 132 *3.1. DESs in membrane preparation and DES-polymer membranes*



133 As in any polymer preparation protocol, a solvent is primarily required to create the dope solution and then to form a  
134 membrane using specific techniques, such as phase inversion (including solvent-induced, temperature-induced or vapour-  
135 induced phase separation), dense-film casting method, electrospinning and track etching [34,35]. For the preparation of  
136 ultrafiltration (UF) membranes, Jiang et al. [36] introduced various imidazole-based DESs, based on organic Cl, Br salts  
137 and organic imidazole (IM) molecules, as pore-former additives in polyethersulfone (PES) membranes using the phase  
138 inversion technique. In principle, a pore-former is used in membrane fabrication to promote the porosity and to enhance the  
139 membrane pore size resulting in more permeable membranes. Since water is typically used as a non-solvent during phase  
140 inversion, water-soluble pore-formers (like polyethylene glycol) are generally employed. They exhibit, in fact, hydrophilic  
141 groups which make them easily extractable by the formed membrane thanks to their solubility in the aqueous coagulation  
142 medium [37]. The same principle has been also explored in the case of hydrophilic DESs. Due to the affinity of the selected  
143 DES with water, the exchange rate between the solvent (e.g., NMP) and non-solvent (water) in the coagulation bath can be  
144 accelerated favouring the fabrication of membranes with a more porous architecture [19,36]. As shown in **Figure 3**, Jiang  
145 et al. [36] demonstrated that the incorporation of DESs, especially tetrabutylphosphonium bromide- imidazole (P<sub>4444</sub>Br/IM),  
146 into the membrane matrix, can turn the pores into macrovoids in PES membranes; particularly, this was more evidenced  
147 when the DES concentration was increased from 0 to 2 wt.%. During filtration tests, the PES UF membranes exhibited a  
148 water permeability of 781 L m<sup>-2</sup> h<sup>-1</sup>, which was up to 6 times higher respect to the reference PES membrane not containing

149 any DES. At the same time, rejection tests revealed that bovine serum albumin (BSA) was minimally affected by DES  
150 incorporation since all the membranes presented a BSA rejection as high as 97.7%. According to the author's findings, such  
151 a high BSA rejection rate was the result of the narrowly distributed pore diameters, along with the reduction in the effective  
152 pore size. Moreover, by filtrating humic acid solutions, the water flux decline was found to be less pronounced and  
153 characterised by more stable performance. This suggested that DES based membranes can benefit of improved antifouling  
154 properties as a result of their lower surface roughness. These findings are in agreement with the ones documented by  
155 Maalige et al. [38], who treated film composite polyamide membranes with different DESs based on HBA (choline chloride)  
156 and HBD (ethylene glycol, urea and glycerol). The chemical surface modification considerably increased the permeation  
157 rate by 2-5 fold compared with the non-treated membrane with negligible changes in rejection. The enhanced measured  
158 flux was the consequence of an improved surface wettability obtained after the DES treatment; as observed by Jiang et al.  
159 [36], which also reported an increased surface smoothness for the membranes prepared with DESs. This result was  
160 attributed to the existence of hydrogen bonding between the DES and the polyamide moiety, which was confirmed by zeta-  
161 potential analysis [38]. By following a similar strategy, Vatanpour et al. [39] recently synthesised DESs based on ethylene  
162 glycol and choline chloride, the so-called ethaline, which was later proposed as a hydrophilic pore-former to fabricate  
163 PES/polyvinyl pyrrolidone (PVP) nanofiltration (NF) membranes. It was generally concluded that the use of ethaline led to  
164 a series of benefits in the resulting membranes, such as: i) the formation of uniform pores on the membrane surface, ii) the



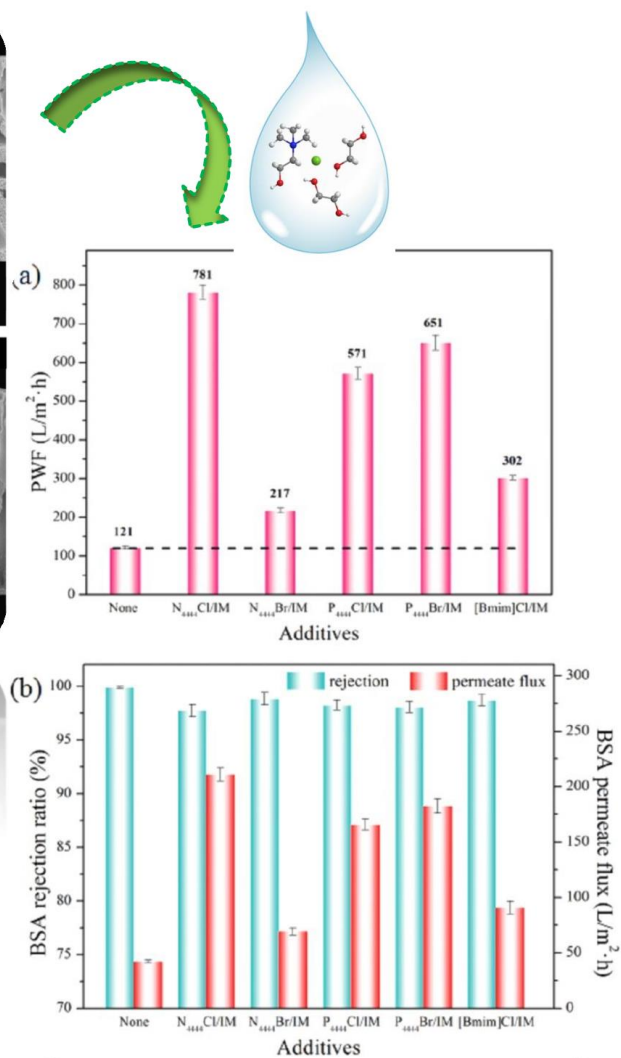
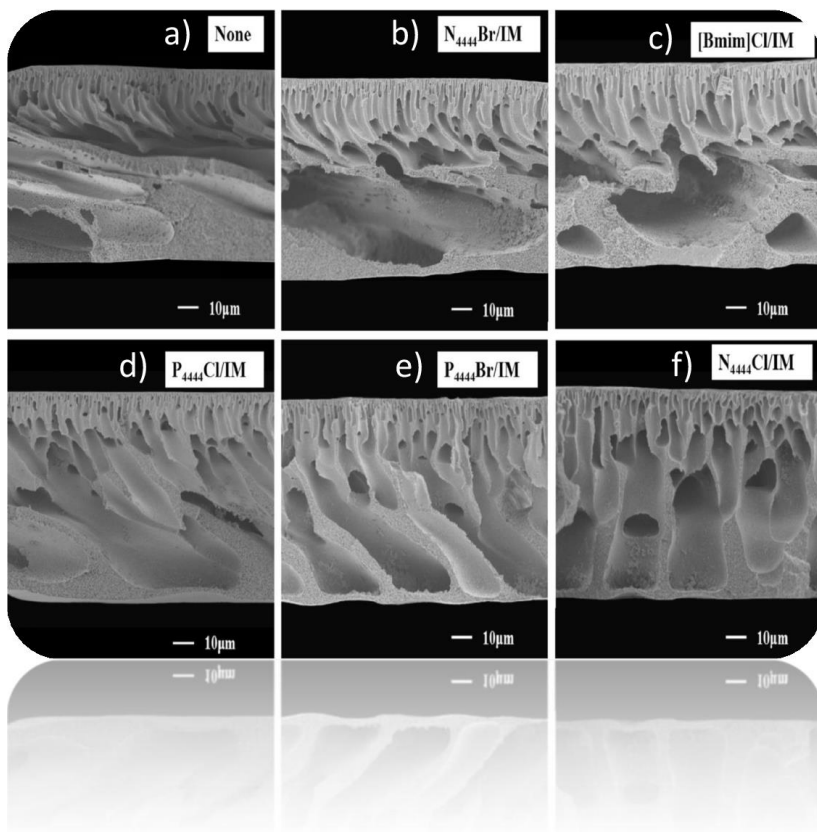
165 creation of large macro-voids in the finger-like layer, iii) an enhanced surface smoothness, iv) an improvement in  
166 hydrophilicity, v) a high protein rejection and dye removal, and vi) a high permeation rate.

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**Figure 3.** Morphological structure of PES UF membranes tailored by DESs as pore-formers; a) none, b) PES-N<sub>444</sub>Br/IM, c) PES-[Bmim]Cl/IM, d) PES-P<sub>444</sub>Cl/IM, e) PES-P<sub>444</sub>Br/IM, f) PES-N<sub>444</sub>Cl/IM [36].

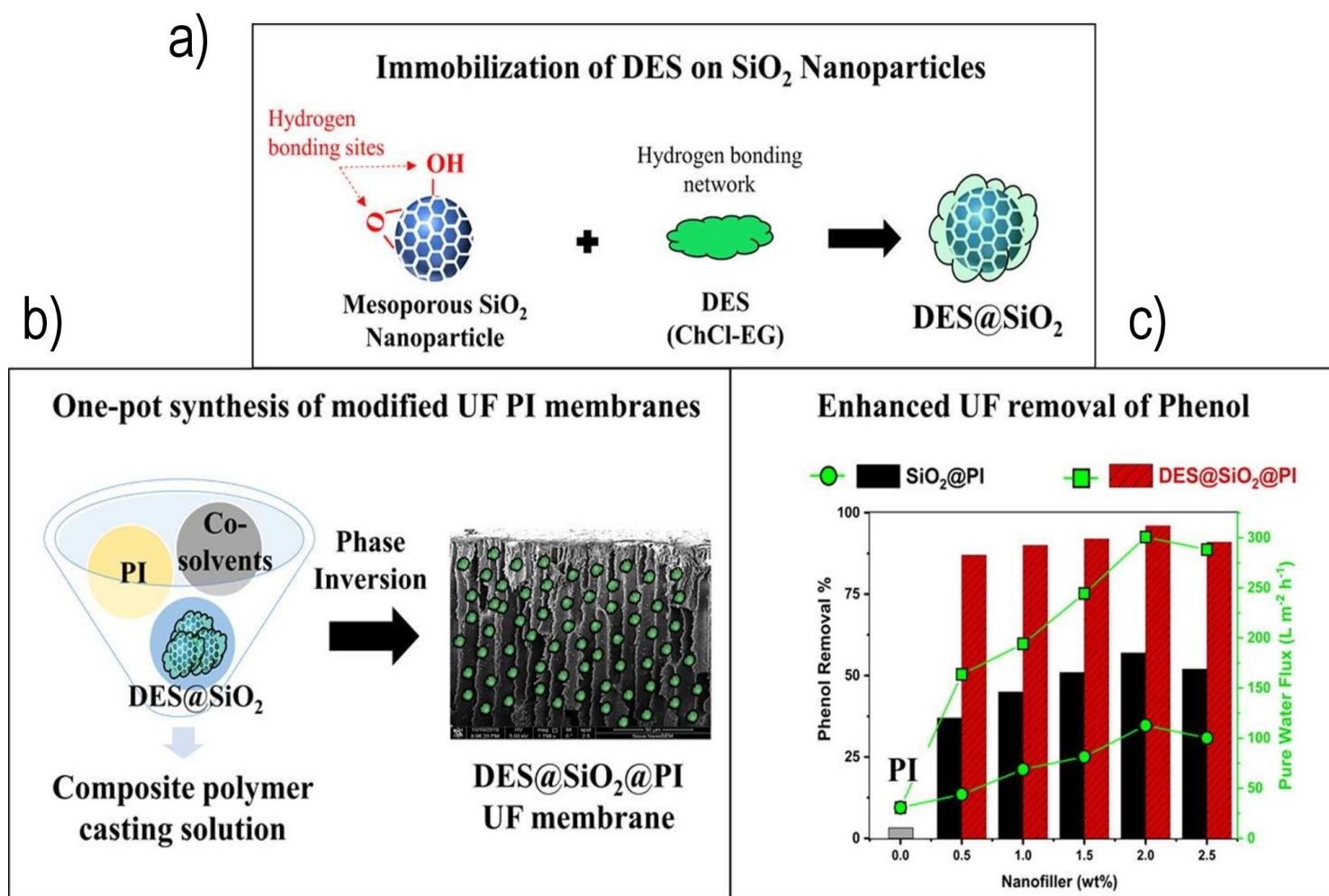


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175 An outstanding application in preparing composite membranes was done by Kuttiani Ali [40], who utilized choline chloride-  
176 ethylene glycol (ChCl-EG) DES for the post-impregnation of silica nanoparticles that were subsequently embedded into  
177 polyimide UF membranes, as illustrated in **Figure 4a&b**. The impregnation of the DES on the nanoparticles was found to  
178 have a negligible effect on their morphology. The resulting DES-modified silica revealed an inter-spacing of 3.3 nm, while  
179 TEM analysis evidenced the existence of a porous shell comprising nanoclusters over the surface of the nanoparticles.  
180 According to the authors' speculation, such nanoclusters could be related to the DES in solid-state. A content of 2 wt.% of  
181 DES modified nanoparticles was found to be the optimal loading in the polymer matrix resulting in membranes with the best  
182 mechanical properties. In vacuum filtration experiments (at 85 kPa), the UF membranes doped with the nanofiller displayed  
183 high water permeation as high as  $300 \text{ L m}^{-2} \text{ h}^{-1}$  and a rejection efficiency of 96% when applied for the treatment of a  $30 \text{ mg}$   
184  $\text{L}^{-1}$  of aqueous phenol solution. **Figure 4c** clearly describes the substantial effect of DES-modified silica into UF membrane  
185 compared with non-modified silica. Regarding the exceptional performance in terms of phenol removal, it was evidenced  
186 that the significant adsorption of phenol on the hydrated surface of the silica-polyimide membranes was due to hydrogen  
187 bonding and carboxylic moieties of polyimide. Additionally, hydrogen bonding between phenol and silanol and siloxane  
188 groups in silica surfaces could contribute to the efficient phenol removal.

189





**Figure 4.** a) Immobilization of choline chloride-ethylene glycol (ChCl-EG) DES on SiO<sub>2</sub>, b) embedding of DES impregnated silica nanoparticles into polyimide UF membranes, c) removal efficiency of phenol [40].

193

194 A DES based on urea and guanidine hydrochloride (molar ratio of 2:1) was used for the defined exfoliation of silk fibers  
195 (diameter between 20-100 nm and length between 0.3-10  $\mu\text{m}$ ) towards functional membrane preparation. In this work, Tan  
196 et al. [41] proved that the DES improved the exfoliation protocol of the fibers thanks to its ability to permeate into the silk  
197 fibers, loosening their structure and disrupting hydrogen bonds. This diminished the strength of the hydrophobic interactions  
198 in the silk. The resulting fibers showed better mechanical properties allowing to manufacture membranes for vacuum  
199 filtration testing. Here, the authors evaluated the ability of these membranes to remove ions, dyes, and protein considering  
200 the amphiphilic properties of the silk fibers. Membranes with a thickness of 18  $\mu\text{m}$  exhibited a rejection efficiency of over  
201 97% towards several dyes (e.g., Rhodamine B, congo red, methylene blue), along with exceptional protein uptake (over  
202 96%); unfortunately, such membranes were not able to retain  $\text{Cu}^{2+}$  ions. In a subsequent work, Tan et al. [42] determined  
203 that silk protein nanofibers produced using a DES-assisted extraction protocol can be a promising alternative in tissue  
204 engineering applications since the fibers demonstrated exceptional cyto-compatibility, flexibility and mechanical stability.  
205 An important contribution of DESs relies on the facilitated water transport once incorporated in membranes. According to  
206 Seyyed Shahabi et al. [43], composite thin-film polyamide reverse osmosis membranes modified with choline chloride-urea  
207 (1 wt.%) increased the water permeation (up to  $56 \text{ L m}^{-2} \text{ h}^{-1}$ ) and salt rejection (over 96.4%) by 27% and 3%, respectively,  
208 in comparison with a pristine polyamide membrane. The improvement was attributed to the ability of the DES of tuning



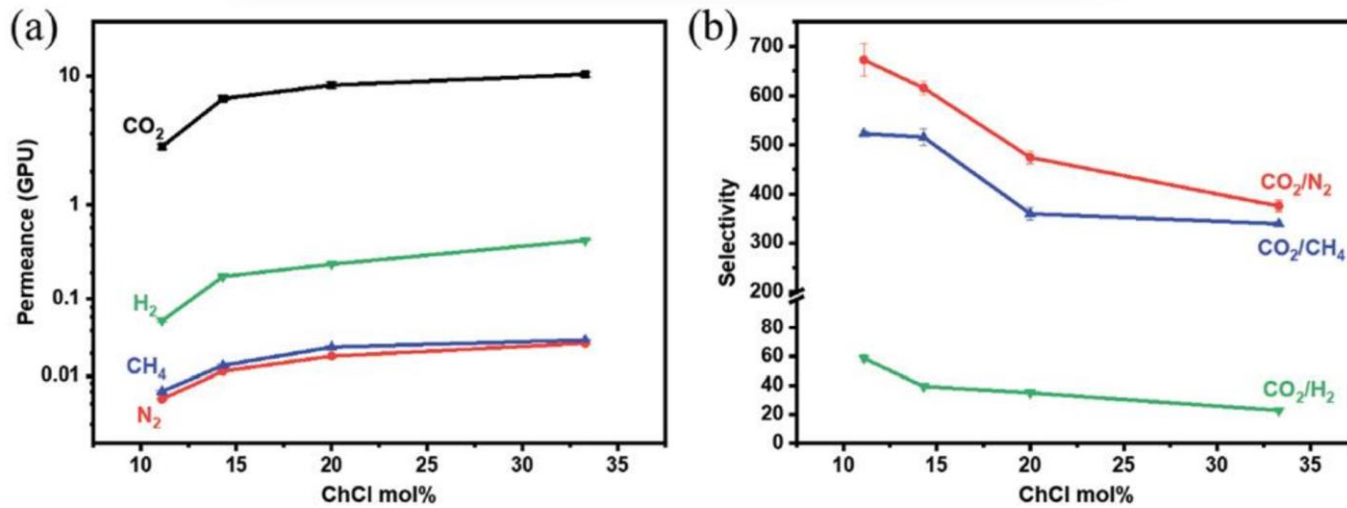
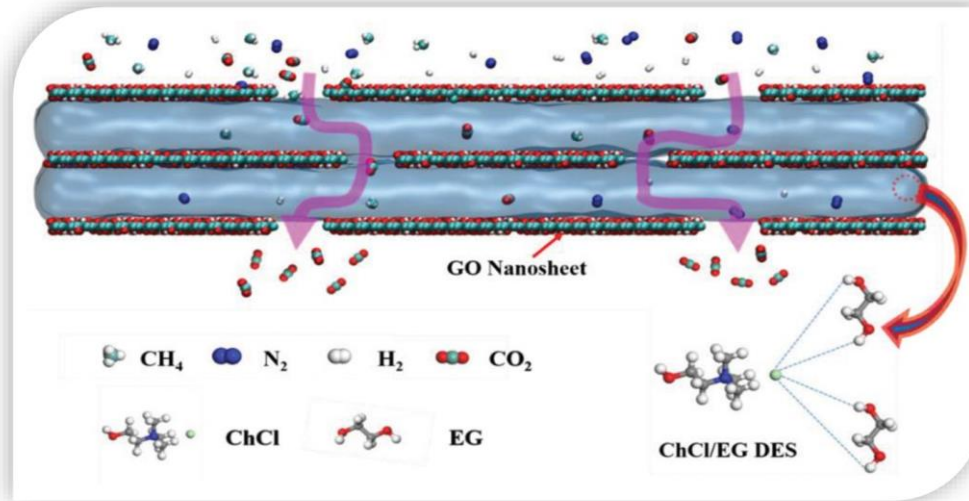
209 surface membrane properties by enhancing their surface hydrophilicity and smoothness due to the presence of hydroxyl  
210 functional groups. Lately, ChCl-EG was systematically applied for the chemical functionalization of graphene oxide (GO)  
211 NF membranes [44]. The developed membranes were able to display unprecedented water permeability of  $124 \text{ L m}^{-2} \text{ h}^{-1}$ ,  
212 which represents a 5-7 times higher permeability compared with permeability of  $22 \text{ L m}^{-2} \text{ h}^{-1}$  in unmodified GO membrane,  
213 and a high rejection toward salt and dyes (nearly 99%). Eventually, the DES provoked a great impact on the structural  
214 properties of GO in various scenarios: *i)* it shifted the d-spacing of GO, *ii)* it decreased their lateral size and *iii)* it decreased  
215 the wettability properties of the final membrane.

216

### 217 3.2. DESs in membrane preparation for gas separation and pervaporation

218 Membrane gas separation and pervaporation are among the most performing processes for the separation of gas and liquid  
219 phases. Both technologies require a continuous non-porous membrane to carry out the selective separation in the gas (or  
220 vapour) state [45,46]. In practice, the presence of porous structures and defects in the membrane matrix leads to a lack of  
221 the separation efficiency. Therefore, the suitable application of DESs in gas separation and pervaporation membranes is  
222 quite challenging. Nevertheless, researchers have implemented intelligent strategies in applying DESs in such technologies.  
223 For instance, Lin et al. [47] designed a  $\text{CO}_2$ -selective membrane by depositing ChCl-EG into GO nanosheets, as  
224 schematically represented in **Figure 5**. In this case, the replacement of the traditional ILs with their green analogues DESs

225 was successfully achieved; where the liquid DES phase considerably facilitated the CO<sub>2</sub> diffusion. After confirming the CO<sub>2</sub>  
226 selective properties, the authors varied the molar ratio of DES and membrane thickness finding exceptional CO<sub>2</sub> separation  
227 against light gases (e.g., N<sub>2</sub>, CH<sub>4</sub> and H<sub>2</sub>). Such parameters (molar ratio and thickness) were deeply evaluated since the  
228 molar ratio of HBA and HBD species greatly determines the physicochemical properties (solubility, viscosity, etc.) of the  
229 resulting eutectic mixture [48] and may influence the membrane performance, while the second parameter dictates, to some  
230 extent, the permeable properties of the resulting membrane [49]. In Lin's work, it was noted that the permeance of four  
231 gases increased as the choline chloride mole fraction increased (see **Figure 5a**), while the CO<sub>2</sub>/light gas selectivity behaved  
232 oppositely (see **Figure 5b**). Even though the membrane containing a DES with molar ratio 1:2 offered the lowest selectivity,  
233 it still displayed a selectivity for CO<sub>2</sub>/N<sub>2</sub>, CO<sub>2</sub>/CH<sub>4</sub>, and CO<sub>2</sub>/H<sub>2</sub> of 370, 339, and 23, respectively, proving its preferential CO<sub>2</sub>  
234 transportation compared with light gases. In this membrane concept, it was clear that the selective solution-diffusion  
235 mechanism across the DES-filled laminated GO (see **Figure 5**) took place. The CO<sub>2</sub> solubility and diffusion in DESs were  
236 not well described but it was speculated that intermolecular hydrogen bond dominates the CO<sub>2</sub> absorption in choline-based  
237 DESs. It is important mentioning that DESs were also able to improve membranes durability (in terms of long-term operation)  
238 and to enhance their thermal stability. The latter property was the result of the hydrogen bonding between the GO sheets  
239 and DES; for example, the hydrogen bonding attraction of GO for Cl<sup>-</sup> and ethylene glycol induced the formation of Cl and  
240 ethylene glycol enriched layers which improved the final stability.



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**Figure 5.** Graphical drawing of ChCl-EG-deposited GO membrane; a) its CO<sub>2</sub> transport mechanism, and b) its separation performance [47].

244

245 Using the same concept as Lin et al., DES-supported liquid membranes were fabricated by Saeed et al. [50]. Basically,  
246 Saeed and co-workers used betaine-based natural DESs that were embedded into a porous PVDF support. In the  
247 permeation experiments, the CO<sub>2</sub> permeability raised from 25 to 29 Barrer ( $10^{-10}$  cm<sup>3</sup> (STP) cm cm<sup>-2</sup> s<sup>-1</sup> cmHg<sup>-1</sup>) by replacing  
248 different HBD from tartaric acid to malic acid; but betaine stood as the dominant element with high solubility and diffusivity  
249 according to COO<sup>-</sup>CO<sub>2</sub> interactions. Such carboxylate group (COO<sup>-</sup>) is, in fact, recognized for its high CO<sub>2</sub> affinity [51].  
250 Additionally, the ideal CO<sub>2</sub>/CH<sub>4</sub> selectivity varied from 51 to 56 when tartaric acid was substituted by malic acid as the HBD.  
251 Very recently, Craveiro et al. [52] introduced choline chloride and urea (1:2) into a polytetrafluoroethylene (PTFE) support  
252 for CO<sub>2</sub> separation. This DES was selected according to its higher CO<sub>2</sub> solubility, which was accompanied by high  
253 permeability values (up to 100 Barrer). Van der Waals's interactions were the predominant forces formed between the gas  
254 molecules and the DES. By using choline chloride and levulinic acid (molar ratio of 1:2), de Castro et al. [53] also evidenced  
255 acceptable CO<sub>2</sub> permeability between 36-77 Barrer with good CO<sub>2</sub>/N<sub>2</sub> selectivity (varying from 17 to 32).  
256 Poly-DESs constituted by choline chloride (as HBA) and two different HBDs (such as polyacrylic acid and polyacrylamide)  
257 were prepared by Ishaq et al. [54], who sodden them into porous PVDF membranes for the potential CO<sub>2</sub> capture. The  
258 doped membranes showed a CO<sub>2</sub>/CH<sub>4</sub> selectivity as high as 50 compared to pristine PVDF membranes; moreover, CO<sub>2</sub>/N<sub>2</sub>  
259 selectivity was approximately 60. Both acceptable performances were linked to the basicity, molar free volume and the

260 hydrogen bonding strength of the DESs. In a current attempt aiming at producing new CO<sub>2</sub>-selective membranes, Lin et al.  
261 [55] nanoconfined ChCl-EG (molar ratio of 1:4) into nanoslits of laminated MXene and later assessed their capacity of  
262 separating CO<sub>2</sub> from various lighter gases. Experimentally, the MXene supported DES membrane exhibited a CO<sub>2</sub>  
263 permeance of 26 gas permeation unit (GPU) (expressed as 10<sup>-6</sup> cm<sup>3</sup> (STP) cm<sup>-2</sup> s<sup>-1</sup> cmHg<sup>-1</sup>), with CO<sub>2</sub>/N<sub>2</sub>, CO<sub>2</sub>/CH<sub>4</sub> and  
264 CO<sub>2</sub>/H<sub>2</sub> selectivity values of 319, 249 and 12, respectively. In this case, the hydrogen bonding interactions guaranteed  
265 simultaneously the linking of DESs onto MXene and the CO<sub>2</sub> transport across the 2D material.

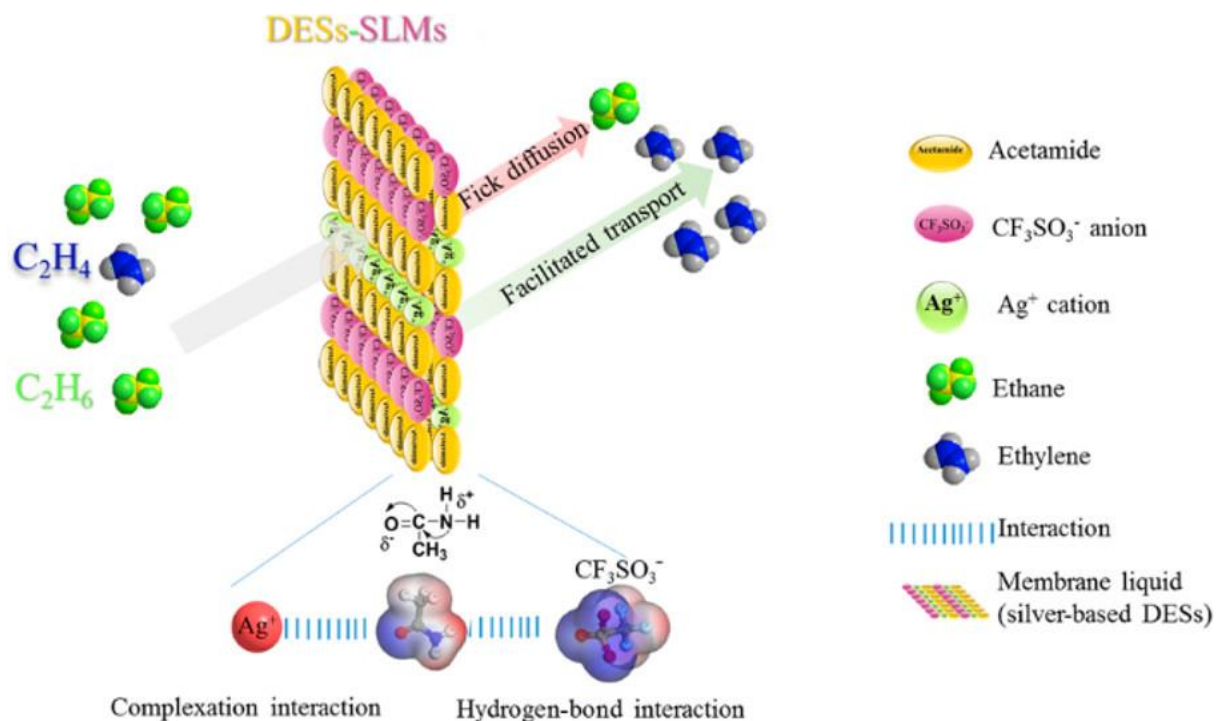
266 Towards different gas separation applications, metal chlorides into 1-butyl-3-methylimidazolium chloride- cuprous  
267 monochloride ([BMIM]Cl)/CuCl DESs were implemented to enhance the activity of the Cu<sup>+</sup> carrier for ethylene/ethane  
268 separation [56]. In fact, the authors have deeply analysed in a series of studies the complexation effect of several metals  
269 (such as Cu, Ag, Al) in gas transport properties of DESs [56–60]. In a first study, the initial impregnation of DESs into PVDF  
270 support membranes was performed, providing an ethylene/ethane selectivity of 17.8, which corresponded to higher values  
271 compared with the membranes without ZnCl<sub>2</sub> (displaying a selectivity of 10.7) [56]. The preferential permeation of ethylene  
272 through DES was the primary factor for the high membrane performance since ethylene gas molecules can positively  
273 interact with Cu<sup>+</sup> via  $\pi$ -bond complexation. According to the study, such DES-supported membranes presented stable long-  
274 term operation over 150 h. Concurrently, the same research group investigated a second DES, choline chloride/glycerol  
275 (molar ratio 1:2) dissolving CuCl, for the ethylene/ethane separation using liquid membranes [57]. At this time, the viscosity





276 and the strength of hydrogen bond networks were ascribed as the responsible variables for the facilitated transport of  
277 ethylene in respect to the ethane molecules. In a more recent work, Jiang et al. [58] documented that ethylene/ethane  
278 separation can be reached by confining a ternary DES (based on silver nitrate with the aluminium nitrate  
279 nonahydrate/methylacetamide mixture with a molar ratio of 1:0.4:4) in a PVDF support. Here, ethylene selectivity was  
280 tentatively fostered due to the complexing interaction between  $\text{Ag}^+$  and methylacetamide, together with the complexing  
281 interaction between  $\text{Al}^{3+}$  and  $\text{NO}_3^-$  that weakened the electrostatic interaction between  $\text{Ag}^+$  and  $\text{NO}_3^-$ , apparently enhancing  
282 the interaction energy between ethylene molecule and  $\text{Ag}^+$  carrier (see **Figure 6**).





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**Figure 6.** Facilitated transport mechanism in DES-supported membranes for gas separation [59].

This phenomenon was confirmed for olefin/paraffin separation as well [59]. When compared with the available literature, an outperforming ethylene/ethane performance was stated due to the synergistic effect taking place in this system. It is important to point out that the so-called complexation effect has been speculated between hydrophobic L-menthol/salicylic acid and  $\text{Ag}^+$  ions, forming a new complex transport species driven by the difference of concentration [61].

290 Until now, most research in gas separation has been directed towards DES liquid membranes; however, the chemistry  
291 related to the DESs may allow their complete merging into polymer phases. This is the case of Pulyalina et al. [62], who  
292 recently prepared polyamide-imide/DES composites using zinc chloride and acetamide (molar ratio of 1:3). In this regard,  
293 the proposed DESs were able to be blended in the polymer phase by hydrogen bonds. The membranes did not show any  
294 visible defect, interfacial voids or clusters in membrane morphology, offering the possibility for separating water-isopropanol  
295 (IPA) mixtures via pervaporation. Experimentally, the membranes, containing 10 wt.% DES, could dehydrate IPA showing  
296 total fluxes of  $82 \text{ g m}^{-2} \text{ h}^{-1}$  and a separation factor of 216. On the other hand, the membranes with 5 wt.% of DES also  
297 demonstrated their ability in splitting gas pairs, such as  $\text{O}_2/\text{N}_2$  and  $\text{He}/\text{O}_2$ , with separation factors of 15.9 and 19.7,  
298 respectively. The addition of the DES significantly increased the permeation and separation efficiency with respect to the  
299 pristine polymer membrane. In an atypical solvent separation, Dietz et al. [63] separated furfural and hydroxymethylfurfural  
300 from diluted aqueous solutions. Hydrophobic DESs composed of various HBDs (decanoic acid, thymol) and HBAs  
301 (lidocaine, thymol, n-tetraoctylammonium bromide, thymol) were immobilized in commercial membrane supports. The DES  
302 containing thymol/lidocaine (molar ratio of 2:1) and then supported in polyethylene was found to be the best candidate for  
303 the separation of furfural and hydroxymethylfurfural from water.

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306 3.3. *DESs in membrane preparation for electromembrane processes*

307 An electromembrane process consists of a unit operation employing electric potential gradient as the driving force for ions  
308 transport. Currently, such membrane technique is proposed for the removal of charged components from solutions, e.g., for  
309 producing drinking water from brackish water. The pioneering applications of DESs in this field were carried out with the  
310 aim of improving the electrical conductivity of the membranes. In 2016, Rahman et al. [64] fabricated a polymer electrolyte  
311 based on poly(vinyl alcohol) (PVA) via electrospinning. The resulting electrospun membranes were immersed in N,N-  
312 diethylethanolammonium chloride/ethylene glycol DESs. This was systematically explored since there was evidence that  
313 the electrical conductivity of the DESs varied depending on the molar ratio. After soaking in DESs with a molar ratio of 1:2,  
314 the membranes demonstrated interesting electrical conductivity from  $2.78 \times 10^{-6}$  to  $2.27 \times 10^{-2} \text{ S cm}^{-1}$ . Based on this finding,  
315 the authors suggested that such membranes could be good candidates for specific applications in batteries, sensors, energy  
316 storage, etc. In the line of enhanced proton conduction membranes, Wong et al. [65] found a significant improvement in the  
317 proton conductivity of chitosan/carboxymethyl cellulose blend membranes when containing choline chloride/urea (1:2).  
318 Particularly, the membranes (composed of 50 wt.% chitosan and DES) registered the highest proton conductivity of  
319  $1.57 \times 10^{-2} \text{ S cm}^{-1}$ , which was comparable with the commercial Nafion-117 membrane (ca.  $8.6 \times 10^{-2} \text{ S cm}^{-1}$ ). When dealing  
320 with electrolyte applications, this membrane could also be a good candidate since it showed low water uptake (ca. 49%),  
321 inferring a lower structural damage. Herein, DES contributed to preserve the chitosan membrane structure that could be

322 weakened by merging with carboxymethyl cellulose. The DES contribution was ascribed to the strong ionic interaction of its  
323 oppositely charged ions with the polymeric chains, increasing the segmental mobility of the entire membrane [66].

324 DESs have not only been proposed to tune the properties of membranes but also as a liquid interface in electromembrane  
325 technologies. For instance, Hansen et al. [67] reported the electromembrane extraction of six polar drug analytes (tyramine,  
326 metaraminol, sotalol, ephedrine, atenolol and metoprolol) from drug-free plasma using non-ionic DES elements (such as  
327 camphor, coumarin, DL-menthol, and thymol) as liquid interface. For this application, commercial PVDF filter membranes  
328 of 0.45  $\mu\text{m}$  pore size (96-well MultiScreen-IP filter) were used as support for the DESs. Depending on the type of DESs, the  
329 electromembrane system offered a recovery efficiency that varied from 47 to 93% towards the drugs. In principle, the authors  
330 defined that the obtained DESs were highly aromatic acting as strong HBs and moderate HBAs. Interestingly, aromatic ( $\pi$   
331 type) interactions were identified for the transfer of bases, while hydrogen bonding was dominant for acids.

332 Karimi et al. [68] studied the effect of halogens on proton conductivity and fuel cell performance of Nafion membranes.  
333 Bromide ( $\text{Br}^-$ ) and iodide ( $\text{I}^-$ ) based DESs were thus prepared and used as immersion phases for the membranes.  
334 Experimentally, bromide-based DESs were likely to display the best conductivity of  $337 \text{ mS cm}^{-1}$ , which corresponded about  
335 120% higher value than the iodide-based Nafion/DES membranes; they also showed a 1200 % higher conductivity than  
336 other Nafion/IL polymer polyelectrolyte membranes at anhydrous conditions. The interaction between such DESs and  
337 Nafion membrane is not well explained but it is believed that in Nafion/DES composite membranes, in presence of ethylene



338 glycol, sodium iodide (NaI) can interact with a sulfonic acid group of Nafion (R–SO<sub>3</sub>H) to yield R–SO<sub>3</sub>Na. Thus, the existence  
339 of I<sup>-</sup> in Nafion/DES solution can be a consequence of a hydroiodic acid reaction with available oxygen (O<sub>2</sub>). Nevertheless,  
340 when any DES is attempted to be used as proton conductive media in polymer membranes, some criteria should be carefully  
341 taken into account [69], such as:

- 342 • DES may present the ability to dissociate protons from sulfonic acid groups of sulfonated polymers (like Nafion);
- 343 • Low DES viscosity;
- 344 • DES proton conductivity;
- 345 • Outstanding DES thermal stability;
- 346 • A feasible mechanism for DES addition to a polymer matrix.

347 Such criteria were suggested by Karimi et al., continuously exploring the improvement of properties of Nafion membranes.  
348 These membranes are probably the most prominent commercial cation exchange membranes for electrochemical  
349 applications [70]. In comparison to bromide-based Nafion/DES membranes that exhibited a conductivity of 337 mS cm<sup>-1</sup>  
350 [68], iodide-based Nafion/DES membranes tend to offer lower proton conductivity (ca. 153 mS cm<sup>-1</sup>) [69]. Although  
351 Nafion/DES membranes owned superconducting properties as a consequence of a Grotthuss-like mechanism for proton  
352 conduction, Karimi et al. [71] also observed that the presence of water can either have a constructive or destructive effect  
353 on the DES and Nafion/DES membranes, pointing out that such factor could be determinant in the performance of a polymer

354 electrolyte membrane. It is worth mentioning that the Grotthuss mechanism (or proton jumping) is the process in which an  
355 'excess' proton or proton defect diffuses through the hydrogen bond network of water molecules or other hydrogen-bonded  
356 liquids [72].

357

#### 358 3.4. *DES assisting membrane processes and extraction techniques*

359 Since DESs have proven their exceptional ability for extracting various types of solutes [6,73], they have been also involved  
360 in assisting extraction protocols. A successful approach based on the extraction of amino acids, by filling DES into the pores  
361 of commercial membranes, was obtained by Li et al. [74]. In this work, the process consisted of two independent diffusion  
362 cells, in which the membranes had direct contact. After evaluating several DESs, choline chloride/p-toluene sulfonic acid  
363 was the most effective (efficiency of 86.1%) in extracting tryptophan, pointing out the hydrogen bonding (between the DES  
364 and the amino acids) as the driving mechanism. In another work, Matsumoto et al. [75] developed lactic acid permeable  
365 polymer inclusion membranes (PIMs). At this point, DESs were prepared based on various HBDs (such as urea, glucose  
366 and several organic acids) and HBA (such as choline chloride, lidocaine, DL-menthol). The investigated DESs were  
367 combined with poly (vinylidene fluoride-co-hexafluoropropylene) (PVDF-HFP) and polyvinyl chloride (PVC). Interestingly,  
368 some of the DESs were water-soluble suggesting an important role as transport carriers, contributing to possible polar  
369 interactions (e.g., hydroxyl groups) between the membrane and the lactic acid for its permeation.



370 A new DES membrane-based microextraction protocol was developed by Shishov et al. [76], who strategically extracted  
371 analytes due to *in situ* DES formation between analytes and choline chloride (as HBA) supported in a hydrophilic poly  
372 (vinylidene fluoride-co-tetrafluoroethylene) membrane. The HBA promoted the extraction of phenolic compounds (such as  
373 guaiacol, eugenol, isoeugenol, among others) from food samples and concurrently formed DESs, which resulted in an  
374 extraction recovery rate ranging from 70 to 80%.

375

#### 376 **4. Concluding remarks and future directions in the field**

377 This review evidenced that the major application of DESs in membrane preparation deals with their use as pore former  
378 additives for porous asymmetric membranes. The resulting membranes have displayed different properties, tuned pore size,  
379 enhanced surface wettability and hydrophilicity, leading to outperforming permeation rates with no decline in membrane  
380 rejection. However, specific nanomaterials (e.g., silica nanoparticles) can also be impregnated by DESs tuning the  
381 performance properties once embedded in polymer composites. Furthermore, DESs have been also used in the preparation  
382 of DES-supported polymer membranes, proving greater effect than the conventional IL membranes in gas separation (CO<sub>2</sub>,  
383 ethylene, etc.). At this point, the complexation effect between metals (such as Cu, Ag Al) and gas molecules (ethylene,  
384 olefin) has been identified as the main mechanism for enhancing gas transport properties of DESs [56–59]. A few reports,  
385 but relevant, have demonstrated that a smart selection of DESs, according to their properties, can also be the key for the



386 preparation of defect-free dense membranes suitable for gas separation and pervaporation applications [47,62]. Regarding  
387 the preparation of membranes for electromembrane processes, DESs demonstrated to enhance the proton conductivity of  
388 commercial Nafion membranes by soaking them into the DES matrix. Also, DESs have assisted specific extraction and  
389 purification processes based on membrane processes thanks to their selectivity and exceptional analyte transport. By  
390 analyzing the current state of the art of DES application in membranes and membrane processes, the future directions and  
391 suggestions for the new researchers in the field are given as follows:

- 392 • Innovative DESs, e.g., based on protonated L-proline (as HBA) and glucose/xylitol (as HBD) [77], have been  
393 ultimately prepared by the research community. According to their natural components, they present a great potential  
394 in multiple applications (e.g., food, pharmaceutical industry, etc.) including membrane fabrication which have never  
395 been explored so far. It is quite possible that novel DESs could be applied due to their success in other chemical  
396 engineering applications [6,10]. Importantly, the molar ratio of HBA and HBD elements will inherently be studied since  
397 such a relationship strongly influences the resulting features of the eutectic mixture.
- 398 • In specific water separations using membranes, DESs have shown great outcomes thanks to the facilitated water  
399 transport related to the DES's polarity. However, the latter property is difficult to be defined in a DES and challenging  
400 to express quantitatively [78]. Here, it is likely that membranes will profit more from hydrophilic DESs associated with  
401 their polarity. On the contrary, when using hydrophobic DESs, the strength of hydrogen bonding will be determinant

402 in selective water transport [79].

403 • When DESs are attempted to assist specific extraction and purification processes using membranes, particular  
404 emphasis should be paid to the possible interactions that water content may display on DES and target biochemical  
405 molecules [80].

406 • To date, DESs have been successfully utilized for the preparation of porous membranes for pressure-driven  
407 membrane processes. At this point, there are no reports of utilizing DESs as pore formers for the fabrication of porous  
408 membranes for other membrane operations, such as membrane distillation [81], which requires porous membranes  
409 with improved anti-wettability properties.

410 • As a suggestion for the new researchers in the field, it is recommended to be focused on synthesizing new CO<sub>2</sub>-  
411 selective DESs since the CO<sub>2</sub> capture is one the most investigated and attractive process in membrane gas  
412 separation. Here, specific DESs (e.g., choline chloride/glycerol) have been identified as the new generation of CO<sub>2</sub>  
413 sorbents [82].

414 • As a current research gap, the study of the complexation effect of new DESs (such as L-menthol/salicylic acid) in the  
415 transport of specific molecules should be further analyzed. Even though there is evidence that DESs can selectively  
416 promote the transport of metal ions (Ag<sup>+</sup> ions) [61], it is needed to unveil and establish the real transport mechanism  
417 of species through DES to open up the field to new applications.

- 418 • Thanks to the application provided by Mehrabi et al. [44], DESs have opened up the possibility to chemically  
419 functionalize inorganic nanomaterials that are currently sought for their outperforming water transport. Here, DESs  
420 can tune and define the interlayer spacing (i.e., *d*-spacing) and lateral size of 2D materials (such as GO, MXene) for  
421 superior performing membranes in water purification [83], solvent dehydration [84], seawater desalination, among  
422 others.
- 423 • The biggest interest is related to application of DESs at their eutectic point composition – optimisation of membrane  
424 performance should include studies also for other molar ratios which could offer more advantageous properties.
- 425 • DESs are formed by weak (hydrogen bonding) interactions. Their components can be soluble in the feedstocks  
426 subjected to desired separation process. Although high selectivity and throughput of the membranes, the risk of  
427 cross-contamination (DESs components leaching issues) into the purified feed in every study should be evaluated.
- 428 • Obtained DES based membranes should be always studied for minimum 5 consecutive cycles for a real assessment  
429 of performance stability.
- 430 • Especially in the case of natural DESs used for membranes preparation, the aspects related to their biodegradability  
431 should be taken into consideration. They relate to a positive disposal of the membranes after their usage, but also to  
432 the risks related to their performance instability in case of treatment of aqueous solutions that can contain bacteria  
433 able to degrade DES components.

434

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442

443 **Conflict of Interest**

444 The authors declare no conflict of interest.

445

446 **References**

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