

# Graphene hydrogels with embedded metal nanoparticles as efficient catalysts in 4-nitrophenol reduction and methylene blue decolorization

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Synthesis and characterization of the graphene hydrogels with three different metallic nanoparticles, that is Au, Ag and Cu, respectively is presented. Synthesized in a one-pot approach graphene hydrogels with embedded metallic nanoparticles were tested as heterogeneous catalysts in a model reaction of 4-nitrophenol reduction. The highest activity was obtained for graphene hydrogel with Cu nanoparticles and additional reaction of methylene blue degradation was evaluated using this system. The obtained outstanding catalytic activity arises from the synergistic effect of graphene and metallic nanoparticles. The hydrogel form of the catalyst benefits in the easiness in separation from the reaction mixture (for example using tweezers) and reusability.

**Keywords:** graphene hydrogel, metallic nanoparticles, heterogeneous catalyst, 4-nitrophenol reduction.

## INTRODUCTION

Many of industrially important reactions i.e oil refining or bulk chemicals production require usage of catalysts. Heterogeneous catalysts are most frequently used, because of the easiness of their recovery and recycling and their conformity in continuous processing<sup>1</sup>. With the nanotechnology development, metallic nanoparticles gained much interest as potential catalysts in a number of reactions, including selective reduction of nitro compounds. Among others, gold nanoparticles (Au NPs) are the most frequently studied and immensity of papers can be found in the literature. The recent progress in state-of-the-art gold nanoparticle catalysts for chemoselective hydrogenations under liquid-phase conditions was reviewed in<sup>2</sup>. Silver and copper nanoparticles were also examined as a cheaper alternative to Au NPs in the nitro arenes reduction, however the literature in this topic is limited<sup>3–5</sup>. Metallic nanoparticles can be used in a colloidal form or they can be supported on different solids, either inorganic or organic<sup>3–16</sup>. Graphene and its derivatives (as graphene oxide or reduced graphene oxide) found its place in the catalysis as a support for nanoparticles, showing in some cases synergistic effect. The large surface area, electrical conductivity and mechanical strength make graphene the most promising substrate to be coupled with NPs catalysts to achieve high catalytic performance. Adsorption of reactants on graphene planes enhances their reduction by transferring free electrons from electron-rich graphene surface. Secondly, using graphene as a platform for metallic nanoparticles immobilization minimizes the NPs agglomeration and ensures the high dispersity of the catalyst. The reduction of 4-nitrophenol (4NP) to 4-aminophenol (4AP) has been performed using graphene/nanoparticles composites based on metals (e.g. Au, Pt, Pd, Ag, Ni etc.) or alloys (FeCo, ZnNi) with NaBH<sub>4</sub> as the reducing agent<sup>11–17</sup>. The mentioned 4-nitrophenol/NaBH<sub>4</sub> system is the preferred model for studying the activity of catalysts because the reaction can be easily followed by UV–vis spectroscopy.

Recent studies have shown that the self-assembly of graphene oxide (GO) under proper conditions leads to the creation of free-standing reduced graphene oxide

(rGO) hydrogels. Since the pioneer work of Shi's group, the self-assembled graphene hydrogels have gained much attention in different fields of science, including catalysis<sup>18–20</sup>. Such graphene hydrogel, comprising reduced graphene oxide and water is an interesting candidate for NPs support. Graphene hydrogel–NPs nanocomposites can be made by growing NPs directly on rGO surfaces in one-pot approach or by the simple mixing of pre-made NPs with GO, followed by its reduction. Au/graphene hydrogel synthesis, characterization and its use for catalytic reduction of 4-nitrophenol was described by Li et al.<sup>21</sup> The Au/graphene hydrogel was prepared through a hydrothermal process with graphene oxide as the graphene precursor, chloroauric acid as Au precursor and triethylenetetramine as a reductant. The Au nanoparticles synthesized in that one-pot synthesis were uniformly supported on the interconnected graphene sheets. More complex approach of Ag/graphene nanocomposite manufacturing was proposed by Dubey et al.<sup>22</sup> Graphene oxide and carbon spheres doped with silver nanoparticles were prepared separately, to be in the next step converted into hydrogel with the usage of cysteine. Such hybrid structures were found to be an efficient catalyst for the reduction of 4-NP into 4-AP using NaBH<sub>4</sub> as a reducing agent.

The Ag or Au 3D graphene-based porous monoliths were prepared by using an unidirectional freeze-drying method. The gold containing 3D porous monoliths showed superior catalytic activity for the reduction of 4-nitrophenol, while silver-based catalyst was tested in the Suzuki–Miyaura coupling reaction with unsatisfactory results<sup>23</sup>.

A simple one-step method for fabricating graphene-based hydrogels with interconnected 3D networks using Cu nanoparticles was developed. During this process, graphene oxide was reduced by Cu nanoparticles to form self-assembled graphene hydrogel with embedded nanoparticles. The unreacted Cu and Cu<sub>2</sub>O nanoparticles enwrapped in the graphene hydrogel structure showed good catalytic efficiency towards 4NP<sup>24</sup>. The usage of heterogeneous catalysts in a form of hydrogel composite is one of the most promising methodologies in developing green organic transformation processes because of

their intrinsic advantages such as easiness in separation from the reaction mixture (for example using tweezers) and reusability. Three dimensional graphene/metal nanocomposites have also unique properties, obtained through the cooperative action between the active metal species and carbonaceous support, providing unusual and efficient catalytic routes for organic synthesis. Moreover, harmful and hazardous to our environment compounds, as for example nitroarenes or dyes can be removed from wastewater *via* catalytic reduction in the presence of graphene composites.

Here we present the synthesis and characterization of the graphene hydrogels with three different metallic nanoparticles, that is Au, Ag and Cu, respectively. The catalytic performance of all hybrid structures was evaluated using model reaction of 4-nitrophenol reduction with the excess of  $\text{NaBH}_4$ . The best results were obtained for graphene hydrogel containing Cu NPs, thus it was additionally tested towards decolorization of methylene blue solution.

## EXPERIMENTAL

Graphite flakes,  $\text{AgNO}_3$ ,  $\text{HAuCl}_4 \cdot x\text{H}_2\text{O}$  were purchased in Sigma-Aldrich. Potassium permanganate ( $\text{KMnO}_4$ ),  $\text{CuCl}_2 \cdot 2\text{H}_2\text{O}$ , L-ascorbic acid, 98% sulfuric acid ( $\text{H}_2\text{SO}_4$ ), 85% phosphoric acid ( $\text{H}_3\text{PO}_4$ ), 30% hydrogen peroxide ( $\text{H}_2\text{O}_2$ ), 35% hydrochloric acid ( $\text{HCl}$ ), and ethanol were purchased in POCh (Gliwice, Poland).

During the experiments the magnetic stirrer Heidolph MR hei-standard with hot plate was used and the centrifugation for all samples was done using Chemland model P3032 centrifuge with the speed 15 000 rpm for 10 min.

UV-Vis spectra were measured using Lambda 10, Perkin Elmer.

The morphology of reduced graphene oxide hydrogels with NPs was observed with a scanning electron microscopy (ESEM Quanta Feg 250, FEI). The elemental analysis was performed by energy-dispersive X-ray (EDX) spectroscopy was performed using the EDAX Genesis APEX 2i with Apollo X SDD spectrometer at 10 kV.

The presence of crystalline phases was checked by X-ray diffraction with the use of a Philips X'Pert diffractometer system. The XRD patterns of the samples were recorded at room temperature with  $\text{CuK}_\alpha$  radiation. Qualitative analysis of diffraction spectra was carried out with ICDD PDF database<sup>25</sup>.

Microphotographs were recorded using confocal microscope Olympus Lext 4000.

Raman spectra were recorded using Renishaw InVia spectroscope with argon ion laser operating at 514.5 nm focused through a 50x objective. Collected light was dispersed through a triple monochromator and detected with a charge-coupled device. The spectra were collected in the dark, with resolution of  $2\text{ cm}^{-1}$  in the range of  $100\text{--}3200\text{ cm}^{-1}$ .

The FTIR spectra were recorded using KBr pellet method on Perkin Elmer Frontier spectrophotometer.

### Synthesis of graphene oxide

Graphene oxide (GO) was prepared using method previously described by us<sup>26</sup>. Shortly, 0.2 g of graphite powder was mixed with 0.6 g of  $\text{KMnO}_4$  and then the

mixture of  $\text{H}_2\text{SO}_4/\text{H}_3\text{PO}_4$  (12 ml:3 ml) was slowly added. The suspension was left for 24 h with the magnetic stirring. Then 3% of hydrogen peroxide was added in order to stop the reaction. The suspension was centrifuged at 15 000 rpm for 10 min and washed several times with deionized water, 3% of  $\text{HCl}$  and with water once again. It was dried under vacuum overnight. To obtain graphene oxide suspension 100 mg of GO was dispersed in 100 ml of deionized water using ultrasonic bath for 4 h. The prepared 1 mg/1 ml suspension was used in further experiments.

The graphite flakes in a certain stage of their transformation from graphite to GO were withdrawn from the reaction mixture with a pipette and sandwiched between a microscope slide and a glass coverslip.

### Synthesis of graphene hydrogel/NPs composites

For the creation of graphene hydrogel, the hydrothermal method of reduction through self-assembly of reduced graphene oxide sheets was used for all samples. Ascorbic acid (AA) was chosen as a reductant of graphene oxide as well as metal ions to form nanoparticles. The previously prepared GO suspension (8 ml, 1 mg/ml) was mixed in a vial with metal precursor, that is  $\text{HAuCl}_4$ ,  $\text{AgNO}_3$  or  $\text{CuCl}_2$  and was sonicated for 1 minute. The  $\text{HAuCl}_4$ ,  $\text{AgNO}_3$  or  $\text{CuCl}_2$  were weighted in such amounts to keep the same metal content in all hydrogels, that is 1:1 w/w GO to metal. After addition of the excess of the reducing agent (32 mg AA) the vials were sealed and heated up in an oil bath up to  $90^\circ\text{C}$ . After 2–3 hours, the reduced graphene oxide hydrogels with embedded nanoparticles were formed. The created structures were thoroughly washed with deionized water in order to remove unreacted reductant and by-products. Purged hydrogels were stored in water at rt. To verify their catalytic activity the hydrogels were removed from the vials and put without drying into the reaction mixture.

Before microscopic and spectroscopic characterization the hydrogels should be dried, and therefore converted into aerogels. The drying protocol was as follows. Hydrogels were washed several times with deionized water and kept in ethanol for several days for the solvent exchange. All hydrogels were dried simultaneously under supercritical conditions for ethanol (63 bar,  $241^\circ\text{C}$ ). Teflon beakers containing hydrogels were placed in an autoclave with 100 ml of ethanol. Ethanol was added to the beakers and to the body of the autoclave. Firstly, the autoclave was flushed a few times with  $\text{N}_2$  gas and pre-pressurized to 50 bar. Samples were heated up to  $300^\circ\text{C}$  with a heating rate of  $5^\circ\text{C}/\text{min}$  and kept in this temperature for 2 h. Then the autoclave was evacuated at a constant temperature with a rate of about 0.7 bar/min. Finally, the autoclave was cooled down to room temperature. If necessary, the samples were disintegrated prior measurements.

### Study on the catalytic activity of GO/NPs hydrogels

In order to investigate the catalytic activity of the as-prepared graphene-metallic hydrogels at room temperature, the reduction reaction of 4-nitrophenol by  $\text{NaBH}_4$  was carried out. The aqueous solutions of 4NP ( $10^{-3}\text{ M}$ ) and  $\text{NaBH}_4$  (0.1 M) were prepared directly before experiments. The 3 ml of 4NP solution and 2

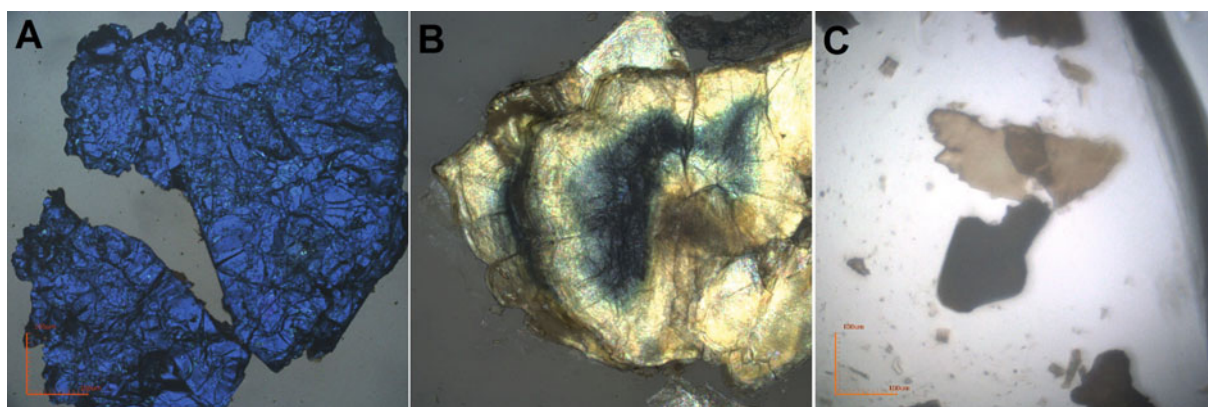
ml of  $\text{NaBH}_4$  solution were added to the vial containing catalyst and the mixture was diluted to 15 ml.

The methylene blue aqueous solution (5 ml,  $10^{-5}$  M) was mixed with  $\text{NaBH}_4$  (10  $\mu\text{l}$ , 0.1 M). The UV-Vis spectra were recorded in the range of 200–900 nm.

## RESULTS AND DISCUSSION

Synthesis and characterization of graphene hydrogels.

The GO was successfully prepared by method previously described by us<sup>26</sup>. The material was characterized using Raman, UV-Vis and FTIR spectroscopy, and scanning electron microscopy. The details can be found in our previous papers<sup>26, 27</sup> and in the section below. Additionally, the progress of the oxidation of graphite during the synthesis was monitored by optical microscopy. In Figure 1. graphene flake at different stages of oxidation is presented. Figure 1A shows blue graphene flake at the preliminary stage of the reaction. The blue color is characteristic for intercalation product of graphite with sulfuric acid. In the next photos (Fig. 1B and B) not-fully and fully-oxidized graphite flakes can be seen. Oxidized areas along the flake's perimeter appear semitransparent light-brown, whereas the unoxidized areas in the center of the flake appear blue-colored (Fig. 1A). In the case of fully oxidized flakes they whole surface is semitransparent and brownish.

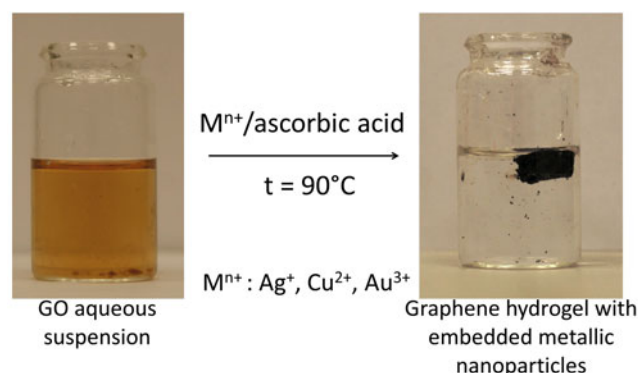


**Figure 1.** Optical images of graphite flakes at different stages of oxidation: A) intercalation product of graphite with sulfuric acid formed in the beginning of the reaction; B) partially oxidized and C) fully oxidized graphite

This simple and low-cost method can be used to control the reaction progress, instead of the expensive and not always available electron microscopy. Such approach can be included into the quality management in the scaled-up production of GO.

The graphene hydrogel formation was carried out by known hydrothermal method<sup>18, 26, 27</sup>. The cylindrically shaped hydrogel/NPs composites were obtained in all cases. The representative image, presenting hydrogel formation can be seen in Figure 2. The obtained hydrogels were denoted as GH\_Au, GH\_Ag and GH\_Cu, for reduced graphene oxide hydrogel with gold, silver and copper nanoparticles, respectively.

After the reduction, the structural and chemical properties of GO change dramatically and different spectroscopic methods can be used to characterize GO and reduced GO (rGO). The conversion of GO into hydrogel was observed with naked eye and the changing properties were monitored by FTIR, UV-VIS, Raman



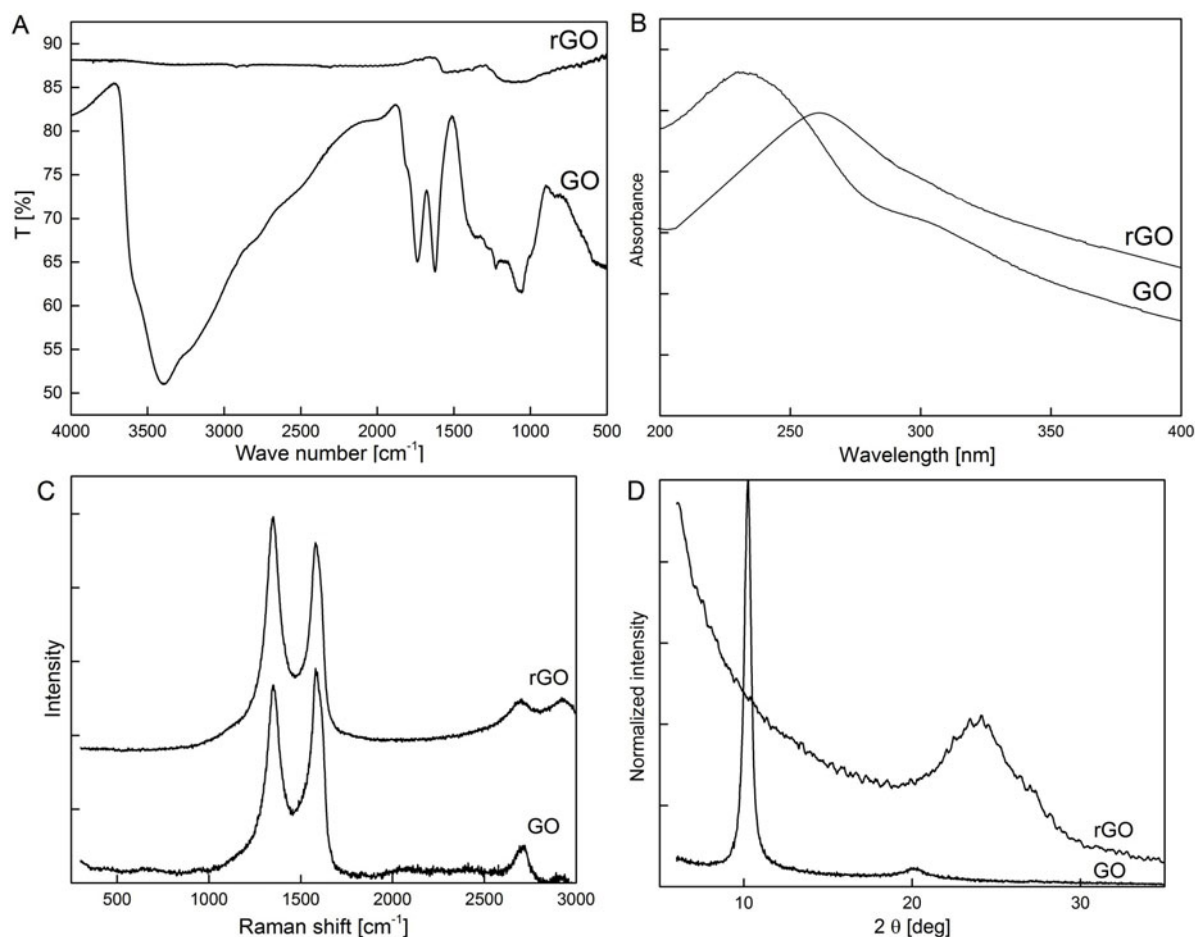
**Figure 2.** Scheme of hydrogel/NPs composites synthesis

and XRD spectroscopy. In Figure 3 the spectra recorded for GO and rGO can be seen. As for the spectroscopic measurements the dried hydrogel should be disintegrated, it was denoted in Figure 3 as rGO as it was not in the form of hydrogel. The presented results are focused only on changes concerning the carbonaceous structure. The structural and spectroscopic features connected with embedded nanoparticles will be discussed later.

Figure 3A shows FTIR spectrum of GO and rGO. Distinct band in GO spectrum at  $1735\text{ cm}^{-1}$  confirms the presence of carbonyl groups. The small band at  $1590\text{ cm}^{-1}$  is assigned to the stretching vibrations of C=C conjugated with C=C or C=O bonds. The broad multiple band

between  $1420$  and  $1000\text{ cm}^{-1}$  are placed in the region characteristic for stretching of C-O bond in alcohols and epoxides. The broad, strong band centered at c.a.  $3300\text{ cm}^{-1}$  and second strong band at  $1640\text{ cm}^{-1}$  refer to the water molecules bonded to GO due to its hydrophilic nature. After reduction, all strong bands disappeared proving the successful removal of oxygen containing groups. The vibrational bands originating from water molecules are also not visible as hydrophobic nature of rGO do not facilitate the water adsorption. The UV-VIS spectra of GO and rGO aqueous suspensions are presented in Figure 3B. The main maximum absorption band at GO spectrum, observed at 230 nm, can be ascribed to  $\pi \rightarrow \pi^*$  electron transition in conjugated carbon-carbon double bonds in the GO plane. The second, smaller band at 310 nm is connected with  $n \rightarrow \pi^*$  transition of nonbonding electrons in oxygen atoms connected with double C=C bonds. Upon reduction, a red shift of the first peak and the disappearance of the second one was observed





**Figure 3.** A) FTIR; B) UV-VIS; C) Raman and D) XRD spectra of GO and rGO

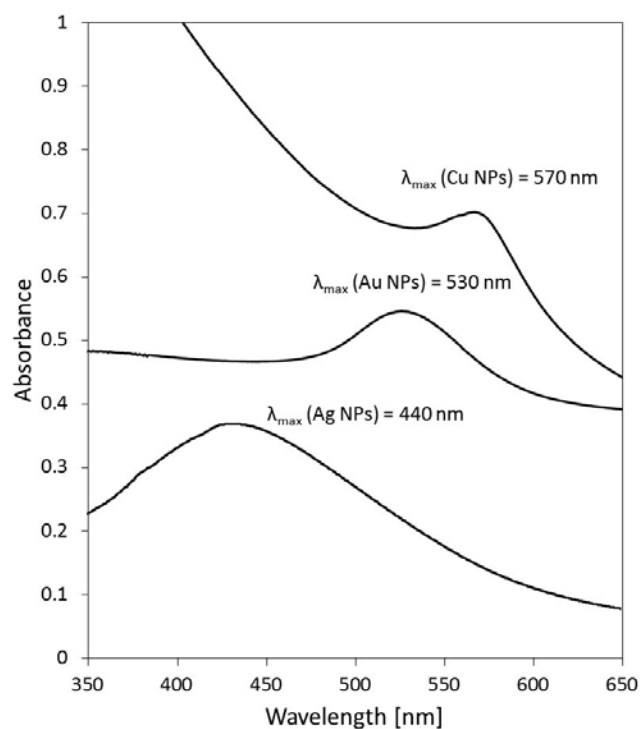
(Fig. 3 B, rGO), confirming again the successful removal of oxygen functionalities.

Raman spectroscopy is very frequently used to characterize carbonaceous materials, including GO and its derivatives. The Raman spectrum of GO (Fig. 3C, GO) consists of two main modes: a broad D mode at  $1350\text{ cm}^{-1}$  and G mode at  $1580\text{ cm}^{-1}$ . The prominent D peak arises from the structural imperfections created by the attachment of hydroxyl and epoxide groups on the carbon basal plane and G peak refers to the vibration of carbon-carbon bonds in a basal plane of GO. The overtone 2D-band can be seen in the range of  $2700$  and  $3000\text{ cm}^{-1}$ . The intensity of the overtone 2D-band with respect to the D and G peaks is small. Raman spectrum of rGO (Fig. 3C) is similar to the spectrum of its precursor, however showing higher D band intensity than G band. As reported elsewhere comparing the D to G ratios do not offer any insights about the chemical changes in the samples, in contrast to the changes of 2D to G ratios<sup>28</sup>. Moreover, after the reduction a small peak shift for the 2D-band and no obvious shift for the D-band or the G-band can be observed, showing that changes in the overtone region of the Raman spectrum is reduction indicative. The small 2D-peak shift, together with decrease of its intensity in respect to G band ( $I_{2D}/I_G = 0.12$  for GO and  $I_{2D}/I_G = 0.05$  for rGO) confirms the reduction of GO to rGO. The last method used to characterize the obtained materials was XRD spectroscopy. The XRD patterns of GO and rGO are shown in Figure 3D. The XRD pattern of GO reveals the diffraction peak at  $2\theta = 10.23^\circ$  which can be attributed

to the (002) diffraction line of graphite. The Bragg's equation was used to calculate the interlayer spacing (d-spacing) of GO which was equal to  $0.863\text{ nm}$ . The XRD pattern of rGO shows a diffraction peak at  $2\theta = 23.77^\circ$  which corresponds to the d-spacing  $0.374\text{ nm}$ . The decrease in the interlayer spacing is a proof of the chemical reduction and the removal of oxygen-containing functional groups from GO plane. The broad XRD signal of rGO indicated the poor ordering of sheets along their stacking direction. It also demonstrated that the structure of rGO was composed of the few-layer stacked sheets, which was consistent with the results obtained by SEM imaging.

In order to characterize the obtained hydrogel/NPs composites they were washed several times with deionized water and kept in ethanol for several days for the solvent exchange. All hydrogels were dried simultaneously under supercritical conditions for ethanol. Dried hydrogels were examined by scanning electron microscopy (SEM) with, UV-Vis spectroscopy and X-ray diffraction (XRD).

Due to the surface plasmon resonance phenomenon colloids of metallic nanoparticles can be characterized by UV-vis spectroscopy. The synthesized hydrogel/NPs composites were dispersed in water and absorption spectra were recorded. In Figure 4 the absorbance spectrum of three different samples can be seen. The characteristic bands at wavelengths ca.  $440\text{ nm}$ ,  $530\text{ nm}$  and  $570\text{ nm}$  refer to silver, gold and copper nanoparticles, respectively. The considerable broadening of the absorption band in the case of GH\_Au and GH\_Ag samples indicates the polydispersity of the analyzed nanoparticles, which is in

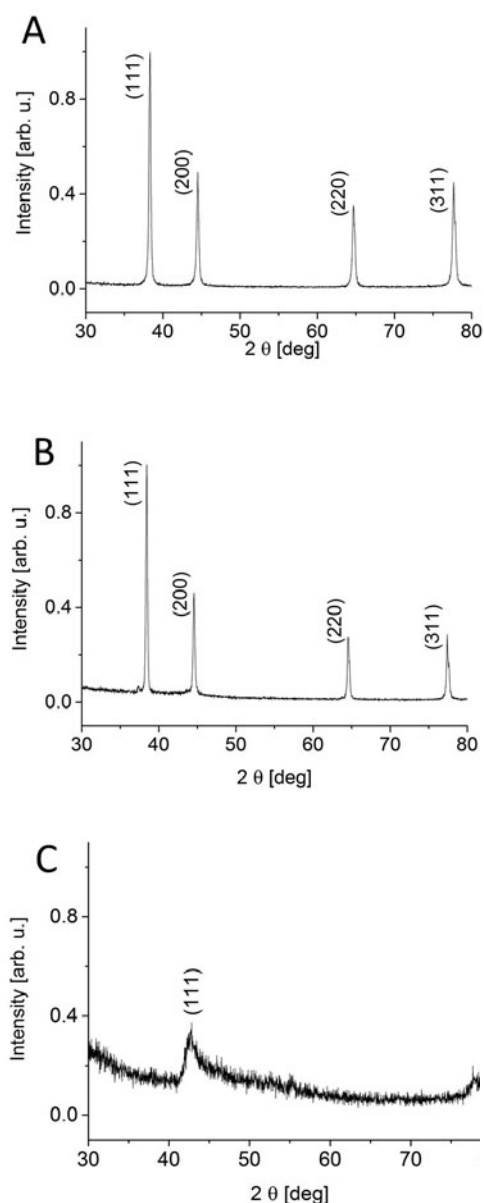


**Figure 4.** Absorption spectra of dispersed GH\_Au, GH\_Ag and GH\_Cu, showing characteristic bands for Au, Ag and Cu nanoparticles

agreement with SEM images. The absorption spectrum of GH\_Cu sample is in agreement with previously reported results for Cu nanoparticles of diameters below 10 nm<sup>29</sup>.

Additional examination of the samples was carried out by X-ray diffraction method (XRD). The obtained XRD patterns confirmed the presence of metallic nanoparticles in the reduced graphene matrix (Fig 5). In the case of GH\_Au sample the reflections at  $2\theta$  angle 38.3°, 44.5°, 64.7° and 77.8° correspond to the (111), (200), (220) and (311) crystal faces of Au nanoparticles. Similarly, the face-centered cubic silver crystals were found in the GH\_Ag sample, with diffraction reflections at 38.4° (111), 44.6° (200), 64.6° (220) and 77.4° (311). In the results obtained for the GH\_Cu sample, relatively small but distinct reflection at 42.6° can be seen. It corresponds to the (111) planes of copper crystals. The other expected reflections at 50.8 and 74.4° are not observed, most probably because of too low intensity. The average crystallite sizes calculated from the widening of the Au (111), Ag (111) and Cu(111) maxima by the Scherrer's formula were 61 nm, 84 and 7 nm, respectively. Intensity of XRD reflections observed for gold- and silver containing samples is high whereas this of Cu samples is low. This difference may be caused by higher atom scattering factors of gold and silver in comparison to copper as well as by highly nanocrystalline nature of copper particles. These results are in good agreement with UV-Vis spectra (Fig. 4) and SEM images, presented in Figure 6.

As can be seen in Figure 6A and B, the obtained gold and silver nanoparticles are polydispersed, but with diameters not exceeding 100 nm. The EDX analysis confirmed the presence of these metals at reasonable level, which was in accordance with metal-to-carbon ratio used in the reaction (Figure 6D). The SEM imaging of GH\_Cu (Figure 6C) sample displayed very small copper particles. As can be seen in the inset, the size of nanoparticles do not exceed 10 nm, which is almost

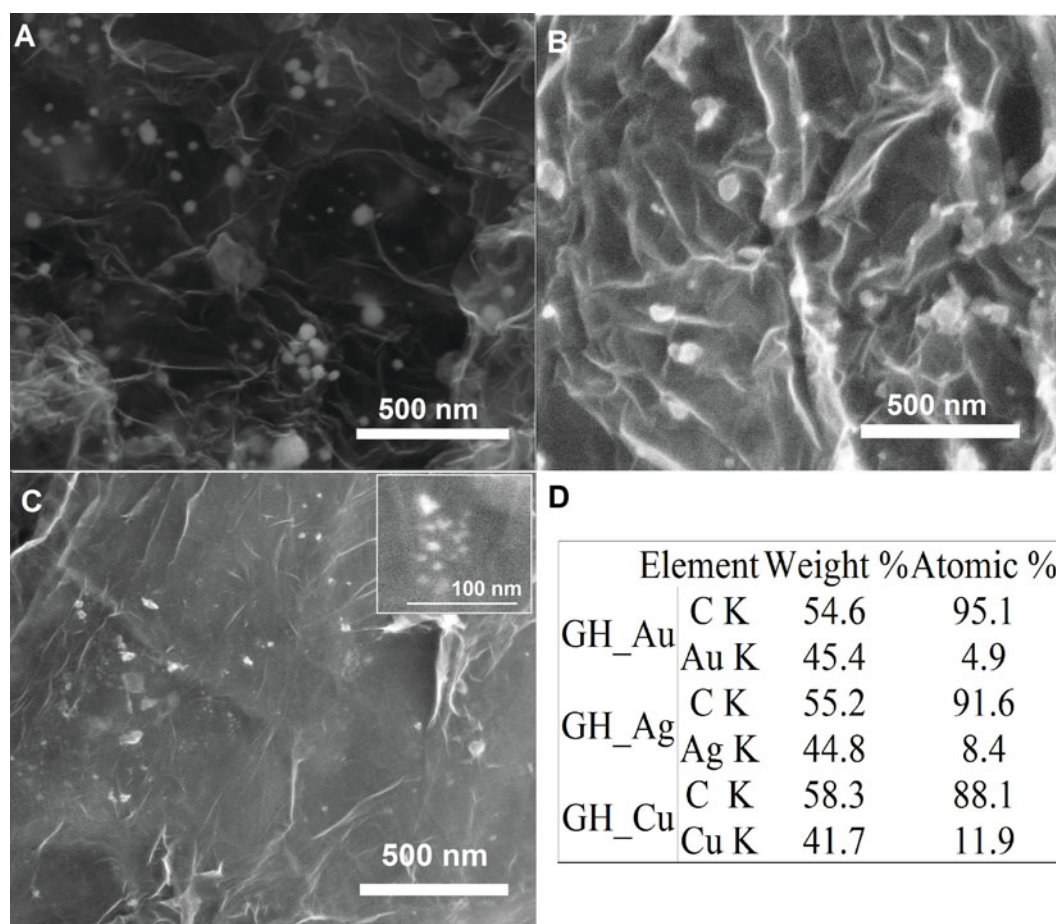


**Figure 5.** XRD patterns of A) GH\_Au, B) GH\_Ag and C) GH\_Cu samples

beyond the microscope resolution. The EDX analysis confirmed the presence of copper atoms (Figure 6 D). Taking into account the data obtained by UV-Vis and XRD spectroscopy, together with microscopic analysis it can be concluded, that Cu NPs below 10 nm were synthesized, which are much smaller as compared to other two metallic NPs, that is Ag and Au.

## STUDY ON THE CATALYTIC PERFORMANCE OF GRAPHENE HYDROGELS

The catalytic activity of freshly prepared hydrogel/NPs composites was evaluated using model reaction of 4NP reduction. The change occurring in the solution was monitored visually as well as by UV-vis spectroscopy. Typical absorption band of 4-nitrophenol is observed at c.a. 320 nm and undergoes a bathochromic shift to 400 nm, after adding an aqueous solution of NaBH<sub>4</sub> due to the phenolate ion formation. For each sample the same amount of 4-NP solution was poured into the vials with different hydrogels. The solution of reductant was added directly before recording the UV-vis spectrum. In all cases



**Figure 6.** SEM images of A) GH\_Au; B) GH\_Ag and C) GH\_Cu; D) table with elemental analysis of the samples

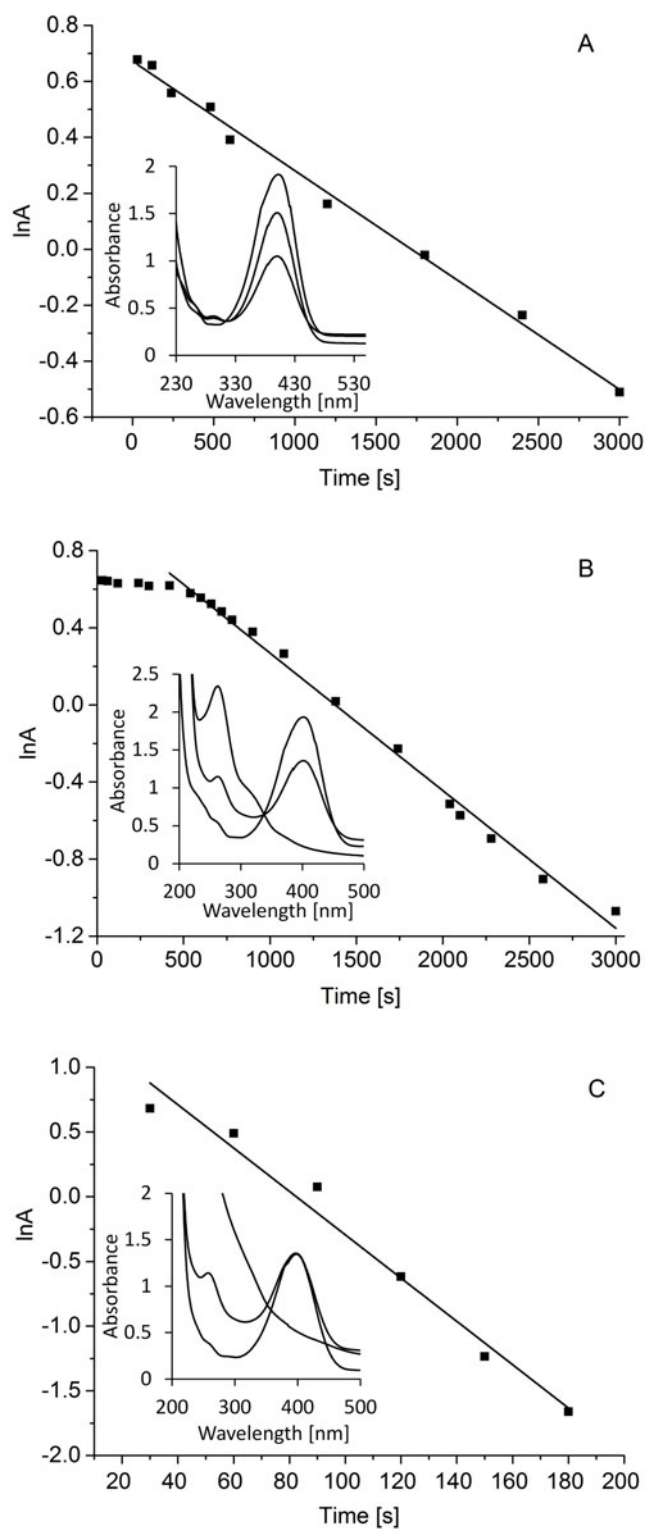
constant decrement of the 400 nm peak intensity was observed. In the case of GH\_Au sample the appearance of new peak at c.a. 290 nm referring to 4-aminophenol can be seen (inset in Fig. 7A). Similarly, if GH\_Ag sample was used as catalyst new peak at 290 and 230 nm were detected. Such observations are in agreement with the literature, proving the successful reduction of nitro to amino compound<sup>30–32</sup>.

It was also observed visually that the light yellow color of the 4-nitrophenol turned to yellowish green rapidly after adding the aqueous solution of NaBH<sub>4</sub> (formation of 4-nitrophenolate ion) and faded with time to become almost colorless. The situation was different when GH\_Cu hydrogel was implemented. In the case of GH\_Cu sample, the disappearance of 400 nm peak was observed immediately after addition of the reducing agent. Neither absorption maximum due to 4-NP nor due to 4-AP was observed in the UV-vis spectrum (inset in Fig. 7C).

Figure 7 shows a linear correlation between  $\ln(A)$  ( $A$  is the absorbance at 400 nm) vs. reduction time, indicating that the reaction is a pseudo-first order for all three catalysts. However, in the case of GH\_Cu it was not possible to determine the rate constant of the reaction as the reduction proceeded faster than the measurement time. For GH\_Cu catalyst presented dependence was observed after 3 cycles of the reduction without regeneration of the catalyst (Fig. 7C). The rate constants ( $k$ ) determined from the slopes are  $0.4 \cdot 10^{-3} \text{ s}^{-1}$ ,  $0.7 \cdot 10^{-3} \text{ s}^{-1}$  and  $19.4 \cdot 10^{-3} \text{ s}^{-1}$  for GH\_Au, GH\_Ag and GH\_Cu, respectively. However, in the case of GH\_Cu the calculated rate constant refers to the degradation of

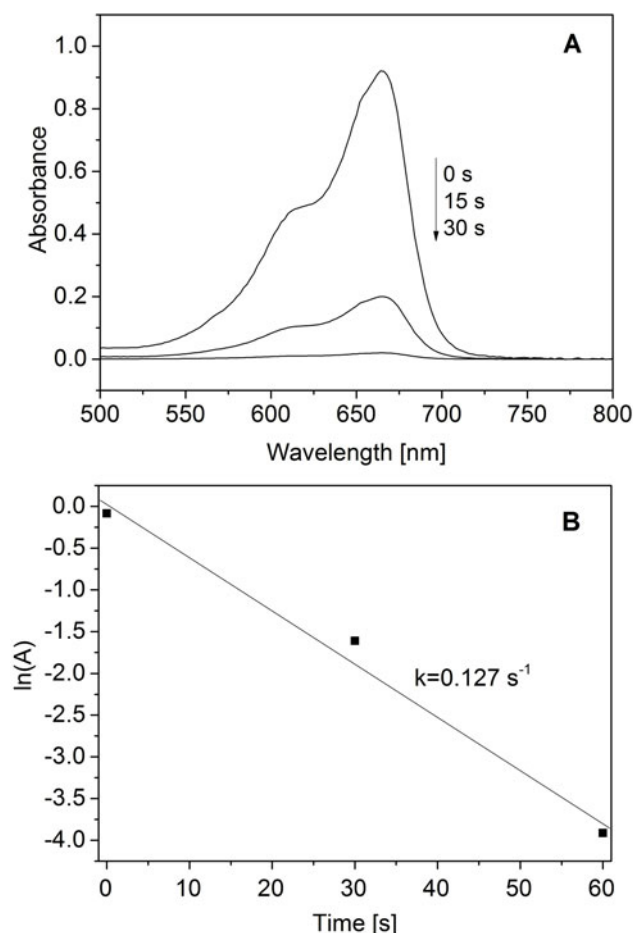
4NP, not to its conversion to 4-aminophenol. According to the literature, the smaller the metallic NPs are, the higher catalytic activity they show<sup>2,3</sup>. Here, the average size obtained from XRD results is c.a. 60 nm for Au NPs and 80 nm for Ag NPs, and from SEM images it can be seen, that the size distribution is wide. Thus, the observed catalytic activity is not as high as reported elsewhere<sup>21–23</sup>. In contrast, very small Cu NPs were obtained and they revealed outstanding properties. Such result is in particular beneficial, as the precursors for Cu NPs are freely available, more stable and less expensive as compared to silver or gold NPs precursors. One of the main problems with Cu NPs usage is their tendency to form oxides in the presence of air. The stability of the catalyst was monitored by XRD and UV-Vis spectroscopy and no changes were observed. The structure of graphene hydrogel/metallic NPs was not changed after the catalytic reaction. The metallic nanoparticles were present on the graphene sheets forming the hydrogel, as was controlled by SEM imaging. The leakage of metallic nanoparticles was also monitored by UV-Vis spectroscopy. As no plasmonic bands were present in the spectra after several cycles, it can be concluded, that no leakage occurred. Moreover no additional bands, for example due to metallic oxide formation were observed after weeks. The XRD pattern showed, that the metallic state was kept after the catalytic run for all analyzed samples, including Cu containing composite. In order to additionally verify the catalytic performance of GH\_Cu, the reduction of methylene blue (MB) with NaBH<sub>4</sub> in the presence of GH\_Cu was carried out. The MB is a water soluble dye used in medicine and pharmacy,





**Figure 7.** Plot of the logarithm of the absorbance at 400 nm vs. the reduction time for A) GH\_Au; B) GH\_Ag and C) GH\_Cu. Insets present changes in the absorption spectra of 4NP with time

causing however cardiovascular, dermatological and hematological problems. The methods of its removal from wastewaters are sought. The reaction progress was monitored by UV-Vis spectroscopy. The obtained results are presented in Figure 8. The decolorization of methylene blue solution occurred immediately after addition of reductant, when catalyst was present. The rate constant determined from the slope was equal to  $0.127 \text{ s}^{-1}$ . The reaction was repeated three times without catalyst regeneration, showing no increase of the reaction



**Figure 8.** A) Changes in the absorption spectra of MB with time during the reduction with  $\text{NaBH}_4$  in the presence of GH\_Cu catalyst; B) plot of the logarithm of the absorbance at 664 nm vs. the reduction time of MB

time. Without the catalyst, the fading of MB was not observed for days, when it was kept in the dark.

Sherazi et al. used colloidal copper stabilized with sodium dodecylsulfate as a catalyst in MB removal<sup>33</sup>. Very recently, copper nanoparticles wrapped with reduced graphene oxide were used as a catalyst in 4NP reduction<sup>34</sup>. More complex approach was presented by Atarod et al., who additionally implemented magnetic nanoparticles into the Cu/rGO system and used such composite as a magnetically separable and reusable catalyst for the reduction of 4-nitrophenol<sup>12</sup>. However, our approach benefits from the easiness of the catalyst preparation, removal from the reaction mixture and regeneration.

## CONCLUSIONS

Simple method of graphene hydrogels with embedded metal nanoparticles was presented. Graphene/Cu NPs hydrogels were found to be the most active towards degradation of two model compounds, that is 4-nitrophenol and methylene blue dye. The reaction was completed within few seconds. In the case of Au and Ag NPs supported on graphene hydrogel the reduction of 4-NP lead to the 4-aminophenol formation, with satisfactory rate constants. High adsorption ability of graphene and effective electron transfer from graphene to metallic

nanoparticles make the proposed system very efficient in the chemical reduction of different compounds. The usage of graphene hydrogel/metallic NPs as a reusable heterogeneous catalyst in various reaction will be further studied.

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